

Bis(dimethylamine- κ N)bis[4-(1,2,4-triazol-1-yl)benzoato- κ O]copper(II)

Lin Liu and Zheng-Bo Han*

College of Chemistry, Liaoning University, Shenyang, 110036, People's Republic of China. *Correspondence e-mail: ceshzb@lnu.edu.cn

Received 22 December 2021

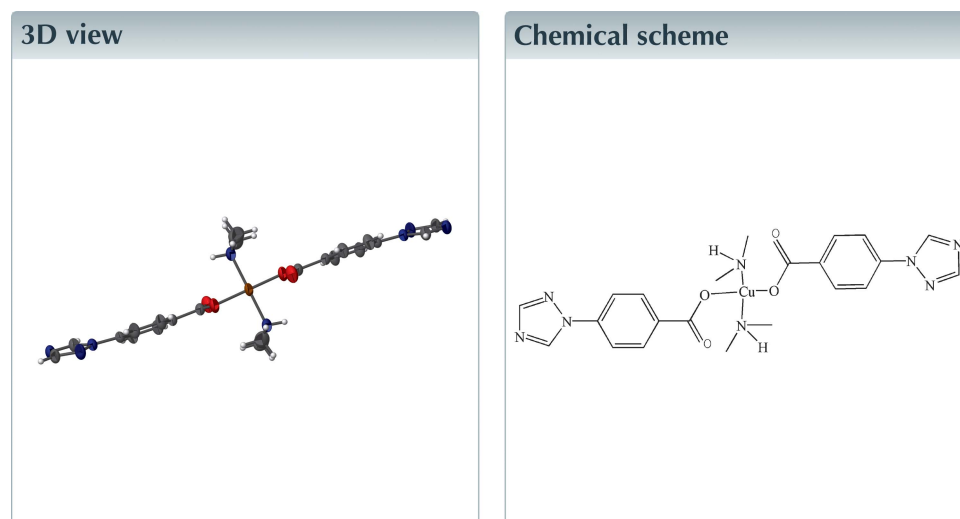
Accepted 12 January 2022

Edited by M. Weil, Vienna University of Technology, Austria

Keywords: crystal structure; coordination polymer; hydrogen bond.

Structural data: full structural data are available from iucrdata.iucr.org

In the title compound, $[\text{Cu}(\text{C}_9\text{H}_6\text{N}_3\text{O}_2)_2(\text{C}_2\text{H}_7\text{N})_2]$, the Cu^{2+} cation is situated on an inversion center and is coordinated by the N atoms of two dimethylamine ligands and the carboxylate O atoms of two 4-(1,2,4-triazol-1-yl)benzoate anions, leading to a slightly distorted square-planar N_2O_2 coordination environment. In the crystal, intermolecular $\text{N}-\text{H}\cdots\text{N}$ hydrogen bonds between the amine function and the central N atom of the triazole ring lead to the formation of ribbons parallel to $[1\bar{1}1]$. Weak intermolecular $\text{C}-\text{H}\cdots\text{O}$ hydrogen-bonding interactions are also observed that consolidate the crystal packing.



Structure description

The rational design of coordination polymers is based on the combination of metal ions and versatile organic ligands, resulting in various supramolecular assemblies. The resulting crystal structures determine the potential applications of the coordination polymers. Different polymers based on 4-(1,2,4-triazol-1-yl)benzoic acid complexes have been reported (Du *et al.*, 2014). They not only feature structural varieties, but also can be applied in gas storage (Wang *et al.*, 2012). In this context we have investigated crystals formed from a copper(II) solution and 4-(1,2,4-triazol-1-yl)benzoic acid under solvothermal conditions.

As shown in Fig. 1, the asymmetric unit of the title compound comprises one Cu^{II} atom, one 4-(1,2,4-triazol-1-yl)benzoate ligand, and one dimethylamine molecule generated *in situ* from the decomposition of the solvent dimethyl formamide. The complete molecule is generated by inversion symmetry. The Cu^{II} atom has a distorted square-planar coordination environment, being coordinated by two symmetry-related benzoato O atoms [$\text{Cu}-\text{O}1 = 1.9611(14) \text{ \AA}$] and two symmetry-related N atoms [$\text{Cu}-\text{N}4 = 2.0096(19) \text{ \AA}$ of the amine ligands]. The second carboxylate O atom of the anion seems to be too far away [$\text{Cu}-\text{O}2 = 2.80136(19) \text{ \AA}$] to contribute to a significant bonding. Nevertheless, the

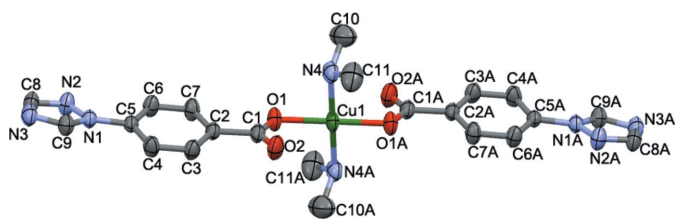


Figure 1

The molecular structure of the title compound, with atom labelling and displacement ellipsoids drawn at the 30% probability level. H atoms have been omitted for clarity. [Symmetry code: (A) $-x + 2, -y, -z + 1$.]

non-bonding O2 atom is involved as an acceptor in weak C—H···O hydrogen-bonding interactions (Table 1, Fig. 2). Stronger N—H···N hydrogen bonds between the amine NH group and the central N atom of the triazole ring are also observed.

Synthesis and crystallization

A mixture of $\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$ (0.0725 mg, 0.3 mmol), 4-(1,2,4-triazol-1-yl)benzoic acid (0.057 g, 0.3 mmol), dimethylformamide (5 ml), ethanol (5 ml) and water (5 ml) was placed in a Teflon reactor with a 23 ml capacity, which was heated at 433 K for 3 days and then cooled to room temperature at a rate of 10 K h^{-1} . Blue block-shaped crystals of the title compound were obtained in 52% yield after being washed with dimethylformamide and dried in air.

Refinement

Crystal data, data collection and structure refinement details are summarized in Table 2.

Funding information

Funding for this research was provided by: Scientific Research Foundation of the Education Department of Liaoning Province (grant No. LJC202004).

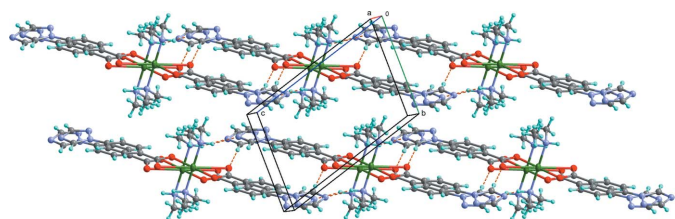


Figure 2

The crystal packing of the complex molecules. Hydrogen-bonding interactions are shown as dashed lines.

Table 1

Hydrogen-bond geometry ($\text{\AA}, ^\circ$).

$D-H \cdots A$	$D-H$	$H \cdots A$	$D \cdots A$	$D-H \cdots A$
$\text{N4}-\text{H4} \cdots \text{N3}^{\text{i}}$	0.91	2.17	3.034 (3)	160
$\text{C9}-\text{H9A} \cdots \text{O2}^{\text{ii}}$	0.93	2.50	3.428 (3)	173

Symmetry codes: (i) $-x + 1, -y + 1, -z$; (ii) $-x + 2, -y + 1, -z$.

Table 2

Experimental details.

Crystal data	
Chemical formula	$[\text{Cu}(\text{C}_9\text{H}_6\text{N}_3\text{O}_2)_2(\text{C}_2\text{H}_7\text{N})_2]$
M_r	530.06
Crystal system, space group	Triclinic, $P\bar{1}$
Temperature (K)	293
a, b, c (\AA)	6.3657 (5), 8.1428 (7), 12.1896 (11)
α, β, γ ($^\circ$)	72.595 (2), 89.376 (2), 87.805 (2)
V (\AA^3)	602.47 (9)
Z	1
Radiation type	Mo $K\alpha$
μ (mm^{-1})	0.95
Crystal size (mm)	$0.36 \times 0.32 \times 0.27$
Data collection	
Diffractometer	Bruker SMART CCD
Absorption correction	Multi-scan (SADABS; Bruker, 2002)
$T_{\text{min}}, T_{\text{max}}$	0.727, 0.785
No. of measured, independent and observed [$I > 2\sigma(I)$] reflections	3971, 2727, 2345
R_{int}	0.017
$(\sin \theta/\lambda)_{\text{max}}$ (\AA^{-1})	0.650
Refinement	
$R[F^2 > 2\sigma(F^2)], wR(F^2), S$	0.038, 0.093, 1.03
No. of reflections	2727
No. of parameters	160
H-atom treatment	H-atom parameters constrained
$\Delta\rho_{\text{max}}, \Delta\rho_{\text{min}}$ (e \AA^{-3})	0.28, -0.20

Computer programs: SMART and SAINT (Bruker, 2002), SHELXS97 (Sheldrick, 2008), SHELXL (Sheldrick, 2015), Mercury (Macrae *et al.*, 2020) and publCIF (Westrip, 2010).

References

- Bruker (2002). SADABS, SAINT and SMART. Bruker AXS Inc., Madison, Wisconsin, USA.
- Du, M., Li, C.-P., Chen, M., Ge, Z.-W., Wang, X., Wang, L. & Liu, C.-S. (2014). *J. Am. Chem. Soc.* **136**, 10906–10909.
- Macrae, C. F., Sovago, I., Cottrell, S. J., Galek, P. T. A., McCabe, P., Pidcock, E., Platings, M., Shields, G. P., Stevens, J. S., Towler, M. & Wood, P. A. (2020). *J. Appl. Cryst.* **53**, 226–235.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Sheldrick, G. M. (2015). *Acta Cryst.* **C71**, 3–8.
- Wang, Y.-L., Fu, J.-H., Wei, J.-J., Xu, X., Li, X.-F. & Liu, Q.-Y. (2012). *Cryst. Growth Des.* **12**, 4663–4668.
- Westrip, S. P. (2010). *J. Appl. Cryst.* **43**, 920–925.

full crystallographic data

IUCrData (2022). 7, x220046 [https://doi.org/10.1107/S2414314622000463]

Bis(dimethylamine- κ N)bis[4-(1,2,4-triazol-1-yl)benzoato- κ O]copper(II)

Lin Liu and Zheng-Bo Han

Bis(dimethylamine- κ N)bis[4-(1,2,4-triazol-1-yl)benzoato- κ O]copper(II)*Crystal data*

[Cu(C₉H₆N₃O₂)₂(C₂H₇N)₂]

$M_r = 530.06$

Triclinic, $P\bar{1}$

$a = 6.3657$ (5) Å

$b = 8.1428$ (7) Å

$c = 12.1896$ (11) Å

$\alpha = 72.595$ (2)°

$\beta = 89.376$ (2)°

$\gamma = 87.805$ (2)°

$V = 602.47$ (9) Å³

$Z = 1$

$F(000) = 275$

$D_x = 1.461$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 1477 reflections

$\theta = 2.6$ – 26.4 °

$\mu = 0.95$ mm⁻¹

$T = 293$ K

Block, blue

$0.36 \times 0.32 \times 0.27$ mm

Data collection

Bruker SMART CCD

diffractometer

Radiation source: fine-focus sealed tube

Detector resolution: 10.0 pixels mm⁻¹

ω scan

Absorption correction: multi-scan

(*SADABS*; Bruker, 2002)

$T_{\min} = 0.727$, $T_{\max} = 0.785$

3971 measured reflections

2727 independent reflections

2345 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.017$

$\theta_{\max} = 27.5$ °, $\theta_{\min} = 2.6$ °

$h = -8 \rightarrow 8$

$k = -10 \rightarrow 10$

$l = -10 \rightarrow 15$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.038$

$wR(F^2) = 0.093$

$S = 1.03$

2727 reflections

160 parameters

0 restraints

Hydrogen site location: mixed

H-atom parameters constrained

$w = 1/[\sigma^2(F_o^2) + (0.0482P)^2 + 0.1332P]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.28$ e Å⁻³

$\Delta\rho_{\min} = -0.20$ e Å⁻³

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cu1	1.000000	0.000000	0.500000	0.03668 (13)
O1	0.8278 (3)	0.19597 (19)	0.40688 (13)	0.0456 (4)
N4	0.8929 (3)	-0.1440 (2)	0.40592 (16)	0.0478 (5)
H4	0.834029	-0.069234	0.342060	0.072*
O2	1.0899 (3)	0.2218 (2)	0.28158 (16)	0.0564 (5)
N3	0.3345 (3)	0.9752 (2)	-0.18059 (16)	0.0470 (5)
C8	0.1967 (4)	0.9638 (3)	-0.0944 (2)	0.0475 (5)
H8A	0.069359	1.026329	-0.105518	0.057*
N1	0.4455 (3)	0.7973 (2)	-0.01740 (15)	0.0357 (4)
C1	0.9149 (3)	0.2671 (3)	0.31047 (18)	0.0397 (5)
C2	0.7914 (3)	0.4135 (2)	0.22845 (17)	0.0347 (4)
C3	0.8841 (3)	0.5111 (3)	0.1285 (2)	0.0443 (5)
H3A	1.024977	0.489576	0.114867	0.053*
C4	0.7728 (3)	0.6395 (3)	0.0487 (2)	0.0462 (5)
H4A	0.838250	0.704495	-0.017864	0.055*
C5	0.5630 (3)	0.6712 (2)	0.06815 (17)	0.0332 (4)
C6	0.4692 (3)	0.5793 (3)	0.16901 (19)	0.0429 (5)
H6A	0.329651	0.603747	0.183663	0.052*
C7	0.5831 (3)	0.4506 (3)	0.24845 (18)	0.0438 (5)
H7A	0.518902	0.388216	0.316208	0.053*
C9	0.4888 (4)	0.8692 (3)	-0.12927 (18)	0.0437 (5)
H9A	0.611277	0.846901	-0.165463	0.052*
N2	0.2529 (3)	0.8588 (2)	0.00657 (16)	0.0467 (5)
C11	1.0630 (5)	-0.2323 (4)	0.3597 (3)	0.0766 (9)
H11A	1.164757	-0.150450	0.321505	0.115*
H11B	1.129585	-0.320275	0.421504	0.115*
H11C	1.005285	-0.283743	0.305845	0.115*
C10	0.7338 (6)	-0.2657 (4)	0.4651 (3)	0.0853 (10)
H10A	0.624985	-0.205376	0.494376	0.128*
H10B	0.674023	-0.317305	0.411925	0.128*
H10C	0.798323	-0.353837	0.527585	0.128*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cu1	0.0456 (2)	0.0340 (2)	0.02439 (18)	0.00871 (14)	-0.00500 (14)	-0.00062 (13)
O1	0.0569 (9)	0.0417 (8)	0.0312 (7)	0.0137 (7)	-0.0080 (7)	-0.0019 (6)
N4	0.0643 (12)	0.0411 (10)	0.0328 (9)	0.0055 (9)	-0.0129 (9)	-0.0034 (8)
O2	0.0430 (9)	0.0580 (10)	0.0580 (11)	0.0168 (8)	-0.0055 (8)	-0.0038 (8)
N3	0.0571 (11)	0.0418 (10)	0.0366 (10)	0.0088 (8)	-0.0123 (9)	-0.0043 (8)
C8	0.0512 (13)	0.0448 (12)	0.0427 (12)	0.0162 (10)	-0.0140 (10)	-0.0089 (10)
N1	0.0387 (9)	0.0317 (8)	0.0330 (8)	0.0042 (7)	-0.0060 (7)	-0.0045 (7)
C1	0.0453 (12)	0.0350 (10)	0.0360 (11)	0.0061 (9)	-0.0121 (9)	-0.0069 (8)
C2	0.0388 (10)	0.0325 (9)	0.0305 (10)	0.0036 (8)	-0.0075 (8)	-0.0060 (8)
C3	0.0333 (10)	0.0467 (12)	0.0446 (12)	0.0055 (9)	-0.0017 (9)	-0.0018 (10)

C4	0.0398 (11)	0.0451 (12)	0.0412 (12)	0.0019 (9)	0.0036 (9)	0.0055 (9)
C5	0.0365 (10)	0.0288 (9)	0.0315 (10)	0.0025 (8)	-0.0064 (8)	-0.0051 (7)
C6	0.0350 (10)	0.0496 (12)	0.0366 (11)	0.0087 (9)	0.0009 (9)	-0.0027 (9)
C7	0.0449 (12)	0.0460 (12)	0.0311 (10)	0.0067 (9)	0.0029 (9)	0.0017 (9)
C9	0.0482 (12)	0.0427 (11)	0.0349 (11)	0.0050 (9)	-0.0043 (9)	-0.0040 (9)
N2	0.0425 (10)	0.0498 (11)	0.0407 (10)	0.0145 (8)	-0.0048 (8)	-0.0049 (8)
C11	0.100 (2)	0.0749 (19)	0.0628 (18)	0.0288 (17)	-0.0177 (17)	-0.0357 (16)
C10	0.103 (3)	0.082 (2)	0.065 (2)	-0.0332 (19)	-0.0135 (18)	-0.0088 (17)

Geometric parameters (Å, °)

Cu1—O1	1.9611 (14)	C2—C3	1.380 (3)
Cu1—O1 ⁱ	1.9612 (14)	C2—C7	1.384 (3)
Cu1—N4	2.0096 (19)	C3—C4	1.375 (3)
Cu1—N4 ⁱ	2.0096 (19)	C3—H3A	0.9300
O1—C1	1.276 (3)	C4—C5	1.382 (3)
N4—C10	1.470 (4)	C4—H4A	0.9300
N4—C11	1.475 (4)	C5—C6	1.377 (3)
N4—H4	0.9071	C6—C7	1.382 (3)
O2—C1	1.240 (3)	C6—H6A	0.9300
N3—C9	1.314 (3)	C7—H7A	0.9300
N3—C8	1.345 (3)	C9—H9A	0.9300
C8—N2	1.315 (3)	C11—H11A	0.9600
C8—H8A	0.9300	C11—H11B	0.9600
N1—C9	1.343 (3)	C11—H11C	0.9600
N1—N2	1.368 (2)	C10—H10A	0.9600
N1—C5	1.421 (2)	C10—H10B	0.9600
C1—C2	1.507 (3)	C10—H10C	0.9600
O1—Cu1—O1 ⁱ	180.00 (8)	C3—C4—C5	119.5 (2)
O1—Cu1—N4	89.16 (7)	C3—C4—H4A	120.2
O1 ⁱ —Cu1—N4	90.84 (7)	C5—C4—H4A	120.2
O1—Cu1—N4 ⁱ	90.84 (7)	C6—C5—C4	119.96 (18)
O1 ⁱ —Cu1—N4 ⁱ	89.16 (7)	C6—C5—N1	120.69 (18)
N4—Cu1—N4 ⁱ	180.00 (7)	C4—C5—N1	119.34 (18)
C1—O1—Cu1	111.32 (13)	C5—C6—C7	119.81 (19)
C10—N4—C11	111.0 (2)	C5—C6—H6A	120.1
C10—N4—Cu1	113.77 (18)	C7—C6—H6A	120.1
C11—N4—Cu1	112.97 (17)	C6—C7—C2	120.9 (2)
C10—N4—H4	108.7	C6—C7—H7A	119.6
C11—N4—H4	103.5	C2—C7—H7A	119.6
Cu1—N4—H4	106.2	N3—C9—N1	110.8 (2)
C9—N3—C8	102.46 (18)	N3—C9—H9A	124.6
N2—C8—N3	116.0 (2)	N1—C9—H9A	124.6
N2—C8—H8A	122.0	C8—N2—N1	101.72 (18)
N3—C8—H8A	122.0	N4—C11—H11A	109.5
C9—N1—N2	109.03 (17)	N4—C11—H11B	109.5
C9—N1—C5	129.63 (18)	H11A—C11—H11B	109.5

N2—N1—C5	121.27 (17)	N4—C11—H11C	109.5
O2—C1—O1	124.09 (19)	H11A—C11—H11C	109.5
O2—C1—C2	119.7 (2)	H11B—C11—H11C	109.5
O1—C1—C2	116.21 (18)	N4—C10—H10A	109.5
C3—C2—C7	118.30 (18)	N4—C10—H10B	109.5
C3—C2—C1	119.88 (19)	H10A—C10—H10B	109.5
C7—C2—C1	121.79 (19)	N4—C10—H10C	109.5
C4—C3—C2	121.5 (2)	H10A—C10—H10C	109.5
C4—C3—H3A	119.3	H10B—C10—H10C	109.5
C2—C3—H3A	119.3		
C9—N3—C8—N2	-0.1 (3)	C9—N1—C5—C4	17.1 (3)
Cu1—O1—C1—O2	-2.6 (3)	N2—N1—C5—C4	-166.2 (2)
Cu1—O1—C1—C2	176.67 (13)	C4—C5—C6—C7	-2.7 (3)
O2—C1—C2—C3	-8.7 (3)	N1—C5—C6—C7	176.47 (19)
O1—C1—C2—C3	172.0 (2)	C5—C6—C7—C2	0.5 (4)
O2—C1—C2—C7	169.3 (2)	C3—C2—C7—C6	1.6 (3)
O1—C1—C2—C7	-10.0 (3)	C1—C2—C7—C6	-176.4 (2)
C7—C2—C3—C4	-1.6 (4)	C8—N3—C9—N1	0.0 (3)
C1—C2—C3—C4	176.4 (2)	N2—N1—C9—N3	0.0 (3)
C2—C3—C4—C5	-0.5 (4)	C5—N1—C9—N3	177.02 (19)
C3—C4—C5—C6	2.7 (3)	N3—C8—N2—N1	0.1 (3)
C3—C4—C5—N1	-176.5 (2)	C9—N1—N2—C8	0.0 (2)
C9—N1—C5—C6	-162.0 (2)	C5—N1—N2—C8	-177.35 (18)
N2—N1—C5—C6	14.7 (3)		

Symmetry code: (i) $-x+2, -y, -z+1$.

Hydrogen-bond geometry ($\text{\AA}, ^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N4—H4 \cdots N3 ⁱⁱ	0.91	2.17	3.034 (3)	160
C9—H9A \cdots O2 ⁱⁱⁱ	0.93	2.50	3.428 (3)	173

Symmetry codes: (ii) $-x+1, -y+1, -z$; (iii) $-x+2, -y+1, -z$.