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μ -1,6,7,12-Tetraazaperylene- $\kappa^4 N^1, N^{12}: N^6, N^7$ -bis[chlorido(η^6 -*p*-cymene)ruthenium(II)] bis(hexafluoridophosphate) acetone disolvate

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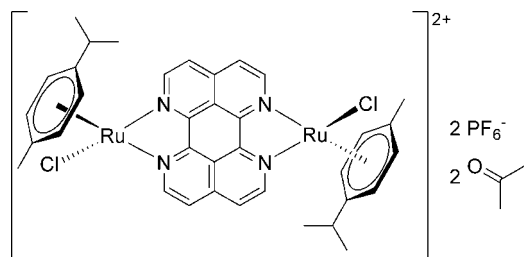
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; R factor = 0.030; wR factor = 0.080; data-to-parameter ratio = 13.7.

In the title compound, $[\text{Ru}_2(\text{C}_{10}\text{H}_{14})_2\text{Cl}_2(\text{C}_{16}\text{H}_8\text{N}_4)](\text{PF}_6)_2 \cdot 2\text{C}_3\text{H}_6\text{O}$, the binuclear Ru^{II} complex dication, $[\{\text{RuCl}(\eta^6\text{-cym})\}_2(\mu\text{-tape})]^{2+}$, built up by a planar 1,6,7,12-tetraazaperylene (tape) bridge, two η^6 -bound cymene (cym) ligands and two chloride ligands, includes an inversion center. The Ru^{II} atom shows the typical piano-stool motif for arene coordination. The counter-charge is provided by a hexafluoridophosphate anion and the asymmetric unit is completed by an acetone molecule of crystallization. The components of the structure are connected into a three-dimensional architecture by $\text{C}-\text{H} \cdots \text{O}/\text{F}/\text{Cl}$ interactions.

Related literature

For related Ru^{II} -arene complexes, see: Bennett & Smith (1974); Robertson *et al.* (1980); Govindaswamy *et al.* (2007); Betanzos-Lara *et al.* (2012). For tetraazaperylene-bridged Ru^{II} complexes, see: Brietzke, Mickler, Kelling & Holdt (2012); Brietzke, Mickler, Kelling, Schilde *et al.* (2012).



Experimental

Crystal data

$[\text{Ru}_2(\text{C}_{10}\text{H}_{14})_2\text{Cl}_2(\text{C}_{16}\text{H}_8\text{N}_4)] \cdot (\text{PF}_6)_2 \cdot 2\text{C}_3\text{H}_6\text{O}$
 $M_r = 1203.82$
 Triclinic, $P\bar{1}$
 $a = 8.6289$ (5) Å
 $b = 11.9346$ (7) Å
 $c = 12.7785$ (7) Å
 $\alpha = 66.099$ (4)°
 $\beta = 83.536$ (4)°
 $\gamma = 77.572$ (4)°
 $V = 1174.45$ (12) Å³
 $Z = 1$
 Mo $K\alpha$ radiation
 $\mu = 0.91$ mm⁻¹
 $T = 293$ K
 $1.50 \times 0.62 \times 0.17$ mm

Data collection

Stoe IPDS-2 diffractometer
 Absorption correction: integration (*X-RED*; Stoe & Cie, 2011)
 $T_{\text{min}} = 0.614$, $T_{\text{max}} = 0.860$
 15353 measured reflections
 4133 independent reflections
 3971 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.047$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.030$
 $wR(F^2) = 0.080$
 $S = 1.01$
 4133 reflections
 302 parameters
 H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 1.17$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.67$ e Å⁻³

Table 1

Selected bond lengths (Å).

C9—Ru1	2.214 (2)	C14—Ru1	2.200 (3)
C10—Ru1	2.188 (2)	Cl1—Ru1	2.3844 (7)
C11—Ru1	2.204 (3)	N1—Ru1	2.105 (2)
C12—Ru1	2.212 (3)	N2—Ru1	2.105 (2)
C13—Ru1	2.185 (3)		

Table 2

Hydrogen-bond geometry (Å, °).

$D-\text{H} \cdots A$	$D-\text{H}$	$\text{H} \cdots A$	$D \cdots A$	$D-\text{H} \cdots A$
C1—H1 \cdots O1	0.93	2.39	3.194 (4)	145
C3—H3 \cdots Cl1 ⁱ	0.93	2.84	3.637 (3)	145
C3—H3 \cdots F6 ⁱⁱ	0.93	2.58	3.205 (4)	125
C4—H4 \cdots F6 ⁱⁱ	0.93	2.60	3.223 (4)	125
C19—H19B \cdots F4 ⁱⁱⁱ	0.96	2.46	3.321 (5)	149

Symmetry codes: (i) $-x, -y + 1, -z$; (ii) $x - 1, y, z$; (iii) $x, y + 1, z$.

Data collection: *X-AREA* (Stoe & Cie, 2011); cell refinement: *X-AREA*; data reduction: *X-RED* (Stoe & Cie, 2011); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL2013* (Sheldrick, 2008); molecular graphics: *DIAMOND* (Brandenburg, 2012) and *ORTEP-3 for Windows* (Farrugia, 2012); software used to prepare material for publication: *SHELXL2013*.

Supporting information for this paper is available from the IUCr electronic archives (Reference: TK5277).

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supporting information

Acta Cryst. (2014). E70, m39–m40 [doi:10.1107/S160053681400035X]

μ -1,6,7,12-Tetraazaperylene- $\kappa^4 N^1, N^{12}: N^6, N^7$ -bis[chlorido(η^6 -*p*-cymene)ruthenium(II)] bis(hexafluoridophosphate) acetone disolvate

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S1. Experimental

S1.1. Synthesis and crystallization

[RuCl $_2$ (η^6 -cym)] $_2$ was prepared as described in literature (Bennett & Smith, 1974). [$\{\text{RuCl}(\eta^6\text{-cym})\}_2(\mu\text{-tape})\}(\text{PF}_6)_2$ was prepared in the same way as described in literature for [RuCl(η^6 -cym)phen]PF $_6$ (Robertson *et al.*, 1980) using only one equivalent tape instead of an excess of phenanthroline (phen). Crystals suitable for X-ray structure analysis were obtained by vapor diffusion of diethyl ether into a saturated acetone solution of [$\{\text{RuCl}(\eta^6\text{-cym})\}_2(\mu\text{-tape})\}(\text{PF}_6)_2$. Therefore, the solution was placed into a flask, connected with another diethylether containing flask and a slight vacuum was applied. Dark crystals began to form at ambient temperature within one night.

S1.2. Refinement

All hydrogen atoms were calculated in their expected positions and refined as riding atoms with $U_{\text{iso}}(\text{H})=1.2U_{\text{eq}}(\text{C})$ with the exception of methyl hydrogen atoms, which were refined with $U_{\text{iso}}(\text{H})=1.5U_{\text{eq}}(\text{C})$. The maximum residual electron density peak of 1.17 eÅ $^{-3}$ was located 1.45 Å from the Ru atom.

S2. Results and discussion

1,6,7,12-Tetraazaperylene (tape) is a D_{2h} -symmetric bis(α, α' -diimine)-type bridging ligand containing an extended π -heteroaromatic system. As for 2,2'-bipyrimidine (bpym), tape can be used to construct a series of dinuclear metal complexes. These dinuclear Ru^{II} complexes of tape display intense low-energy $d\pi(\text{Ru}) \rightarrow \pi^*(\text{tape})$ MLCT absorption bands. Separated metal oxidations for the bimetallic tape complexes and the intervalence-charge-transfer (IVCT) transition, measured for [$\{\text{Ru}(\text{L}-\text{N}_4\text{Me}_2)\}_2(\mu\text{-tape})\}^{5+}$, where L-N $_4$ Me $_2$ is N, N' -dimethyl-2,11-diaza[3.3](2,6)-pyridinophane, at 2472 nm indicate a high degree of electronic interaction between the two ruthenium centres, mediated through the tape bridging ligand (Brietzke, Mickler, Kelling & Holdt, 2012; Brietzke, Mickler, Kelling, Schilde *et al.*, 2012). Here, we report the crystal structure of μ -1,6,7,12-tetraazaperylene-bis[chlorido-(η^6 -*p*-cymene)-ruthenium(II)] bis(hexafluorophosphate) acetone solvate, Fig. 1 & Table 1. The *p*-cymene (cym) ligands are coordinated to ruthenium in the typical piano-stool like mode. Comparison of the published [$\{\text{RuCl}(\eta^6\text{-cym})\}_2(\mu\text{-bpym})\}(\text{PF}_6)_2$ structure (Govindaswamy *et al.*, 2007) and [$\{\text{RuCl}(\eta^6\text{-cym})\}_2(\mu\text{-tape})\}(\text{PF}_6)_2$ show slightly shorter Ru—N distances for the tape complex (Ru—N1, Ru—N2 = 2.117 (3) Å, 2.122 (3) Å and 2.105 (2) Å, 2.105 (2) Å for the bpym and tape complexes, respectively). These shorter bonds should be mainly lead back to the slightly higher π -acceptor strength of tape. Furthermore, only very small differences in the distance between ruthenium(II) and the *p*-cymene centroid (1.6886 (3) Å and 1.6847 (2) Å for bpym and tape complex, respectively) and exactly the same Ru—Cl bond length (2.384 (1) Å, 2.384 (7) Å) were observed. The Ru—Ru distance in [$\{\text{RuCl}(\eta^6\text{-cym})\}_2(\mu\text{-tape})\}(\text{PF}_6)_2$ is 8.0715 (5) Å, and therefore 0.088 Å longer than in [$\{\text{Ru}(\text{L}-$

$\text{N}_4\text{Me}_2\}_2(\mu\text{-tape})(\text{PF}_6)_4$ (Brietzke, Mickler, Kelling, Schilde *et al.*, 2012), what is affected by longer Ru—N bonds in the arene complex. However, the Ru—N bond lengths (were N is part of a bpy type ligand) of $[\{\text{RuCl}(\eta^6\text{-cym})\}_2(\mu\text{-tape})](\text{PF}_6)_2$ are closely related (Ru—N(bpy type) 2.121–2.080 Å) with earlier reported ruthenium arene complexes (Betanzos-Lara *et al.*, 2012; Govindaswamy *et al.*, 2007). The three-dimensional structure of the title compound is build from staggered layers, spaced by *p*-cymene moieties, hexafluorophosphate anions and the acetone solvent molecules. In the crystal structure the tape moieties interact *via* C3 with chlorine and hexafluorophosphate (F6), which is also linked to the adjacent C4, forming bifurcated non classical hydrogen bonds, Table 2. Further weak hydrogen bonds can be found between tape (C1) and acetone (O1), with the result that acetone build a bridge *via* C19 to hexafluorophosphate (F4), Fig. 3.

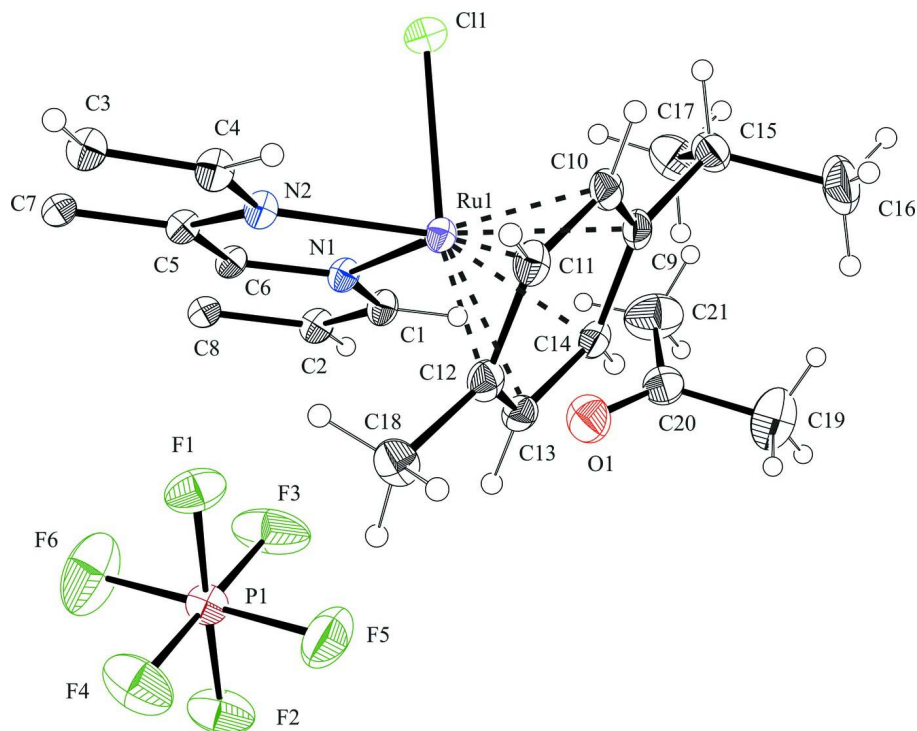


Figure 1

ORTEP drawing of asymmetric unit with the atomic numbering scheme and 30% probability displacement ellipsoids for non-hydrogen atoms.

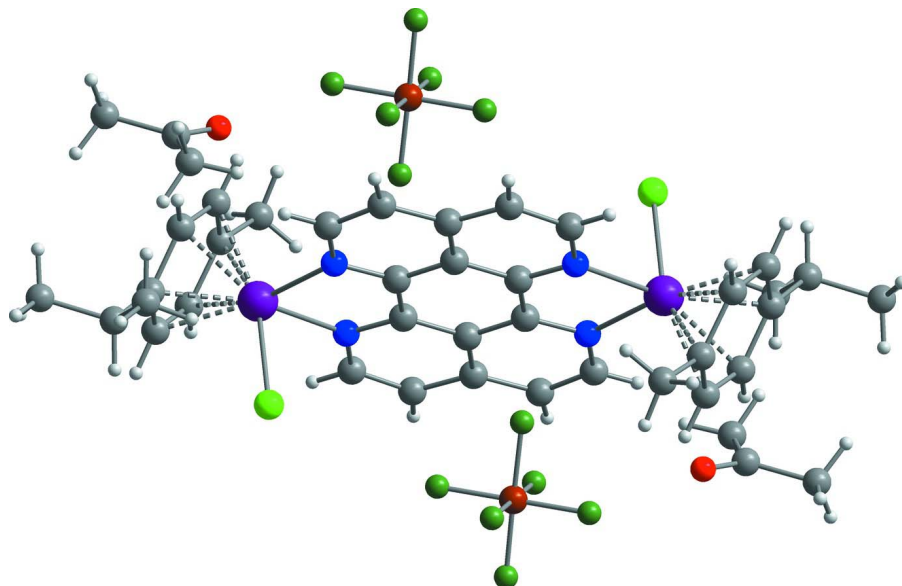


Figure 2
The molecular structure of the title compound.

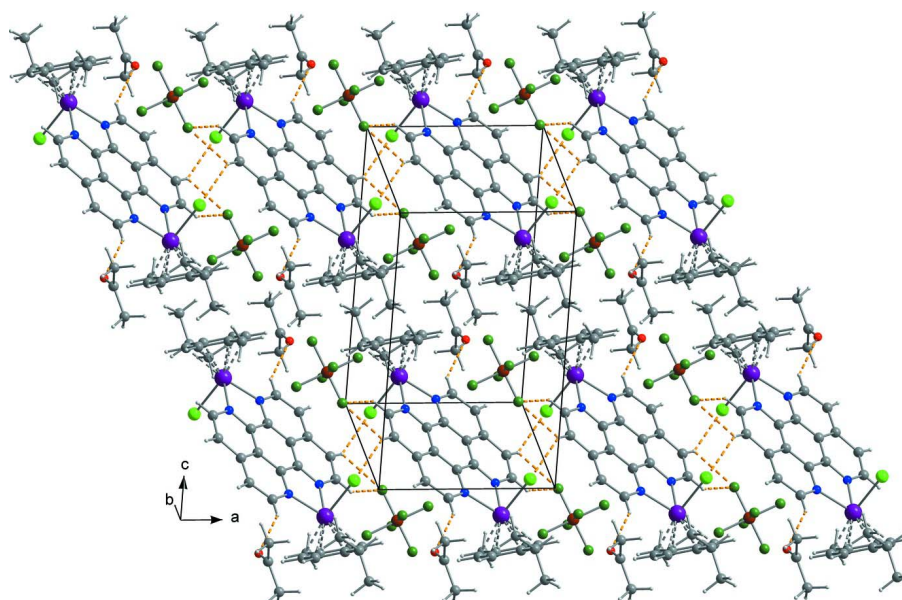


Figure 3
A packing diagram of the title compound is displayed approximately along the *b* axis. Hydrogen bonds are shown as orange dashed lines. One C—H···F hydrogen bond running along *b* was omitted for clarity.

**μ -1,6,7,12-Tetraazaperylene- $\kappa^4 N^1, N^{12}; N^6, N^7$ -bis[chlorido(η^6 -*p*-cymene)ruthenium(II)]
bis(hexafluoridophosphate) acetone disolvate**

Crystal data

[Ru₂(C₁₀H₁₄)₂Cl₂(C₁₆H₈N₄)](PF₆)₂·2C₃H₆O
M_r = 1203.82
Triclinic, *P* $\bar{1}$

Hall symbol: -P 1
a = 8.6289 (5) Å
b = 11.9346 (7) Å

$c = 12.7785 (7) \text{ \AA}$
 $\alpha = 66.099 (4)^\circ$
 $\beta = 83.536 (4)^\circ$
 $\gamma = 77.572 (4)^\circ$
 $V = 1174.45 (12) \text{ \AA}^3$
 $Z = 1$
 $F(000) = 604$
 $D_x = 1.702 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$
 Cell parameters from 20121 reflections
 $\theta = 1.7\text{--}29.8^\circ$
 $\mu = 0.91 \text{ mm}^{-1}$
 $T = 293 \text{ K}$
 Needle, black
 $1.50 \times 0.62 \times 0.17 \text{ mm}$

Data collection

Stoe IPDS-2
 diffractometer
 Radiation source: sealed X-ray tube
 Detector resolution: 6.67 pixels mm^{-1}
 ω scan
 Absorption correction: integration
 (*X-RED*; Stoe & Cie, 2011)
 $T_{\min} = 0.614$, $T_{\max} = 0.860$

15353 measured reflections
 4133 independent reflections
 3971 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.047$
 $\theta_{\max} = 25.0^\circ$, $\theta_{\min} = 1.7^\circ$
 $h = -9 \rightarrow 10$
 $k = -14 \rightarrow 14$
 $l = -15 \rightarrow 15$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.030$
 $wR(F^2) = 0.080$
 $S = 1.01$
 4133 reflections
 302 parameters
 0 restraints
 Primary atom site location: structure-invariant
 direct methods
 Secondary atom site location: difference Fourier
 map

Hydrogen site location: inferred from
 neighbouring sites
 H-atom parameters constrained
 $w = 1/[\sigma^2(F_o^2) + (0.045P)^2 + 1.286P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.003$
 $\Delta\rho_{\max} = 1.17 \text{ e \AA}^{-3}$
 $\Delta\rho_{\min} = -0.67 \text{ e \AA}^{-3}$
 Extinction correction: *SHELXL2013* (Sheldrick,
 2008), $F_c^* = kF_c[1 + 0.001x F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$
 Extinction coefficient: 0.0088 (8)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on all data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.5372 (3)	0.7198 (2)	0.1198 (2)	0.0278 (5)
H1	0.5127	0.7695	0.1619	0.033*
C2	0.6772 (3)	0.7223 (2)	0.0576 (2)	0.0282 (5)
H2	0.7457	0.7720	0.0592	0.034*
C3	0.1439 (3)	0.3571 (2)	0.0791 (2)	0.0288 (5)
H3	0.0679	0.3099	0.0853	0.035*
C4	0.1221 (3)	0.4332 (2)	0.1380 (2)	0.0293 (5)

H4	0.0302	0.4364	0.1831	0.035*
C5	0.3591 (3)	0.5000 (2)	0.0674 (2)	0.0237 (5)
C6	0.4691 (3)	0.5778 (2)	0.0619 (2)	0.0231 (5)
C7	0.3911 (3)	0.4242 (2)	0.0050 (2)	0.0231 (5)
C8	0.7183 (3)	0.6500 (2)	−0.0090 (2)	0.0250 (5)
C9	0.1106 (3)	0.7601 (3)	0.2988 (2)	0.0302 (6)
C10	−0.0022 (3)	0.6859 (3)	0.3042 (2)	0.0313 (6)
H10	−0.1038	0.7252	0.2778	0.038*
C11	0.0364 (3)	0.5555 (3)	0.3479 (2)	0.0312 (6)
H11	−0.0400	0.5097	0.3515	0.037*
C12	0.1928 (3)	0.4922 (2)	0.3873 (2)	0.0314 (6)
C13	0.3041 (3)	0.5644 (3)	0.3818 (2)	0.0307 (6)
H13	0.4067	0.5251	0.4062	0.037*
C14	0.2629 (3)	0.6972 (2)	0.3395 (2)	0.0292 (5)
H14	0.3380	0.7428	0.3389	0.035*
C15	0.0633 (4)	0.9007 (3)	0.2507 (3)	0.0383 (7)
H15	−0.0254	0.9237	0.1999	0.046*
C16	0.0016 (5)	0.9403 (3)	0.3501 (3)	0.0629 (11)
H16A	0.0839	0.9136	0.4044	0.094*
H16B	−0.0883	0.9027	0.3868	0.094*
H16C	−0.0294	1.0295	0.3213	0.094*
C17	0.1938 (4)	0.9683 (3)	0.1810 (3)	0.0527 (8)
H17A	0.2254	0.9444	0.1174	0.079*
H17B	0.2834	0.9467	0.2283	0.079*
H17C	0.1555	1.0567	0.1532	0.079*
C18	0.2359 (4)	0.3522 (3)	0.4299 (3)	0.0396 (7)
H18A	0.1643	0.3221	0.3997	0.059*
H18B	0.2284	0.3167	0.5120	0.059*
H18C	0.3426	0.3287	0.4051	0.059*
C19	0.5166 (6)	0.9785 (4)	0.3738 (4)	0.0748 (12)
H19A	0.4216	1.0407	0.3534	0.112*
H19B	0.5910	1.0067	0.4036	0.112*
H19C	0.4914	0.9021	0.4310	0.112*
C20	0.5881 (4)	0.9566 (3)	0.2701 (3)	0.0446 (7)
C21	0.6172 (6)	1.0667 (5)	0.1694 (4)	0.0792 (13)
H21A	0.6894	1.1064	0.1881	0.119*
H21B	0.5188	1.1241	0.1458	0.119*
H21C	0.6625	1.0419	0.1082	0.119*
Cl1	0.09547 (8)	0.78153 (6)	0.03483 (5)	0.03517 (17)
F1	0.5627 (2)	0.3410 (2)	0.26787 (17)	0.0589 (5)
F2	0.8898 (2)	0.3000 (2)	0.3836 (2)	0.0675 (6)
F3	0.7334 (4)	0.4617 (2)	0.2626 (3)	0.0938 (10)
F4	0.7241 (3)	0.1753 (2)	0.3893 (3)	0.0811 (8)
F5	0.6323 (3)	0.3400 (3)	0.4315 (2)	0.0837 (8)
F6	0.8209 (3)	0.2954 (4)	0.2213 (3)	0.1058 (11)
N1	0.4310 (2)	0.64737 (19)	0.12313 (17)	0.0247 (4)
N2	0.2288 (2)	0.50461 (19)	0.13364 (17)	0.0253 (4)
O1	0.6209 (3)	0.8519 (2)	0.2737 (2)	0.0611 (7)

P1	0.72700 (9)	0.31873 (7)	0.32587 (7)	0.03586 (18)
Ru1	0.20978 (2)	0.63314 (2)	0.21078 (2)	0.02344 (10)

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0287 (13)	0.0305 (12)	0.0306 (12)	-0.0114 (11)	0.0015 (11)	-0.0163 (10)
C2	0.0275 (13)	0.0300 (12)	0.0326 (13)	-0.0140 (11)	0.0003 (11)	-0.0139 (11)
C3	0.0250 (12)	0.0325 (13)	0.0336 (13)	-0.0145 (11)	0.0032 (11)	-0.0141 (11)
C4	0.0233 (12)	0.0356 (13)	0.0336 (13)	-0.0125 (11)	0.0061 (11)	-0.0166 (11)
C5	0.0224 (11)	0.0265 (11)	0.0231 (11)	-0.0090 (10)	0.0017 (9)	-0.0092 (9)
C6	0.0232 (12)	0.0268 (11)	0.0217 (11)	-0.0091 (10)	0.0007 (9)	-0.0103 (9)
C7	0.0225 (12)	0.0254 (11)	0.0231 (11)	-0.0091 (10)	-0.0010 (10)	-0.0088 (9)
C8	0.0241 (12)	0.0267 (12)	0.0258 (12)	-0.0098 (10)	-0.0007 (10)	-0.0094 (10)
C9	0.0299 (13)	0.0352 (14)	0.0296 (13)	-0.0034 (11)	0.0049 (11)	-0.0197 (11)
C10	0.0230 (12)	0.0410 (14)	0.0325 (13)	-0.0035 (11)	0.0058 (11)	-0.0199 (12)
C11	0.0277 (13)	0.0392 (14)	0.0308 (13)	-0.0126 (11)	0.0100 (11)	-0.0176 (11)
C12	0.0349 (14)	0.0333 (13)	0.0241 (12)	-0.0078 (12)	0.0068 (11)	-0.0107 (10)
C13	0.0303 (13)	0.0366 (14)	0.0239 (12)	-0.0040 (11)	0.0004 (10)	-0.0120 (11)
C14	0.0306 (13)	0.0370 (14)	0.0268 (12)	-0.0086 (11)	0.0011 (10)	-0.0186 (11)
C15	0.0356 (15)	0.0341 (14)	0.0459 (16)	-0.0002 (12)	-0.0046 (13)	-0.0186 (13)
C16	0.079 (3)	0.0439 (18)	0.068 (2)	0.0035 (18)	0.009 (2)	-0.0350 (18)
C17	0.0503 (19)	0.0348 (16)	0.067 (2)	-0.0084 (14)	-0.0035 (17)	-0.0128 (15)
C18	0.0425 (16)	0.0319 (14)	0.0383 (15)	-0.0088 (13)	0.0075 (13)	-0.0090 (12)
C19	0.088 (3)	0.084 (3)	0.073 (3)	-0.039 (3)	0.021 (2)	-0.046 (2)
C20	0.0364 (16)	0.0520 (19)	0.0522 (18)	-0.0107 (14)	-0.0066 (14)	-0.0248 (15)
C21	0.076 (3)	0.085 (3)	0.057 (2)	-0.024 (3)	-0.001 (2)	-0.005 (2)
Cl1	0.0354 (3)	0.0383 (3)	0.0310 (3)	-0.0088 (3)	-0.0050 (3)	-0.0108 (3)
F1	0.0433 (11)	0.0809 (14)	0.0505 (11)	-0.0207 (10)	-0.0098 (9)	-0.0170 (10)
F2	0.0412 (11)	0.0592 (12)	0.0888 (16)	-0.0176 (10)	-0.0221 (11)	-0.0062 (11)
F3	0.104 (2)	0.0417 (12)	0.116 (2)	-0.0260 (13)	-0.0552 (18)	0.0085 (13)
F4	0.0662 (15)	0.0428 (11)	0.124 (2)	-0.0188 (11)	-0.0134 (15)	-0.0148 (13)
F5	0.0653 (15)	0.138 (2)	0.0661 (14)	-0.0051 (15)	-0.0010 (12)	-0.0656 (16)
F6	0.0733 (17)	0.181 (3)	0.103 (2)	-0.053 (2)	0.0486 (16)	-0.094 (2)
N1	0.0239 (10)	0.0284 (10)	0.0254 (10)	-0.0091 (9)	0.0016 (8)	-0.0128 (8)
N2	0.0240 (10)	0.0289 (10)	0.0265 (10)	-0.0100 (9)	0.0037 (9)	-0.0128 (9)
O1	0.0586 (15)	0.0611 (16)	0.0792 (18)	-0.0093 (13)	-0.0124 (14)	-0.0416 (14)
P1	0.0298 (4)	0.0372 (4)	0.0405 (4)	-0.0110 (3)	0.0007 (3)	-0.0131 (3)
Ru1	0.02053 (13)	0.02803 (14)	0.02507 (14)	-0.00703 (9)	0.00287 (9)	-0.01342 (9)

Geometric parameters (Å, °)

C1—C2	1.368 (4)	C13—H13	0.9300
C1—N1	1.376 (3)	C14—Ru1	2.200 (3)
C1—H1	0.9300	C14—H14	0.9300
C2—C8	1.411 (4)	C15—C17	1.515 (5)
C2—H2	0.9300	C15—C16	1.531 (4)
C3—C4	1.369 (4)	C15—H15	0.9800

C3—C8 ⁱ	1.415 (4)	C16—H16A	0.9600
C3—H3	0.9300	C16—H16B	0.9600
C4—N2	1.366 (3)	C16—H16C	0.9600
C4—H4	0.9300	C17—H17A	0.9600
C5—N2	1.335 (3)	C17—H17B	0.9600
C5—C7	1.399 (3)	C17—H17C	0.9600
C5—C6	1.442 (3)	C18—H18A	0.9600
C6—N1	1.327 (3)	C18—H18B	0.9600
C6—C7 ⁱ	1.400 (3)	C18—H18C	0.9600
C7—C6 ⁱ	1.400 (3)	C19—C20	1.493 (5)
C7—C8 ⁱ	1.410 (3)	C19—H19A	0.9600
C8—C7 ⁱ	1.410 (3)	C19—H19B	0.9600
C8—C3 ⁱ	1.415 (4)	C19—H19C	0.9600
C9—C14	1.405 (4)	C20—O1	1.203 (4)
C9—C10	1.430 (4)	C20—C21	1.465 (5)
C9—C15	1.512 (4)	C21—H21A	0.9600
C9—Ru1	2.214 (2)	C21—H21B	0.9600
C10—C11	1.399 (4)	C21—H21C	0.9600
C10—Ru1	2.188 (2)	Cl1—Ru1	2.3844 (7)
C10—H10	0.9300	F1—P1	1.595 (2)
C11—C12	1.432 (4)	F2—P1	1.592 (2)
C11—Ru1	2.204 (3)	F3—P1	1.575 (2)
C11—H11	0.9300	F4—P1	1.574 (2)
C12—C13	1.401 (4)	F5—P1	1.579 (2)
C12—C18	1.506 (4)	F6—P1	1.574 (3)
C12—Ru1	2.212 (3)	N1—Ru1	2.105 (2)
C13—C14	1.427 (4)	N2—Ru1	2.105 (2)
C13—Ru1	2.185 (3)		
C2—C1—N1	123.2 (2)	C12—C18—H18A	109.5
C2—C1—H1	118.4	C12—C18—H18B	109.5
N1—C1—H1	118.4	H18A—C18—H18B	109.5
C1—C2—C8	120.4 (2)	C12—C18—H18C	109.5
C1—C2—H2	119.8	H18A—C18—H18C	109.5
C8—C2—H2	119.8	H18B—C18—H18C	109.5
C4—C3—C8 ⁱ	120.2 (2)	C20—C19—H19A	109.5
C4—C3—H3	119.9	C20—C19—H19B	109.5
C8 ⁱ —C3—H3	119.9	H19A—C19—H19B	109.5
N2—C4—C3	123.2 (2)	C20—C19—H19C	109.5
N2—C4—H4	118.4	H19A—C19—H19C	109.5
C3—C4—H4	118.4	H19B—C19—H19C	109.5
N2—C5—C7	122.9 (2)	O1—C20—C21	123.9 (4)
N2—C5—C6	117.0 (2)	O1—C20—C19	119.5 (4)
C7—C5—C6	120.2 (2)	C21—C20—C19	116.6 (4)
N1—C6—C7 ⁱ	123.5 (2)	C20—C21—H21A	109.5
N1—C6—C5	116.9 (2)	C20—C21—H21B	109.5
C7 ⁱ —C6—C5	119.7 (2)	H21A—C21—H21B	109.5
C6 ⁱ —C7—C5	120.2 (2)	C20—C21—H21C	109.5

C6 ⁱ —C7—C8 ⁱ	119.9 (2)	H21A—C21—H21C	109.5
C5—C7—C8 ⁱ	119.9 (2)	H21B—C21—H21C	109.5
C2—C8—C7 ⁱ	116.0 (2)	C6—N1—C1	117.0 (2)
C2—C8—C3 ⁱ	127.8 (2)	C6—N1—Ru1	114.24 (17)
C7 ⁱ —C8—C3 ⁱ	116.2 (2)	C1—N1—Ru1	128.77 (17)
C14—C9—C10	117.4 (2)	C5—N2—C4	117.6 (2)
C14—C9—C15	122.8 (3)	C5—N2—Ru1	113.90 (17)
C10—C9—C15	119.8 (2)	C4—N2—Ru1	128.38 (16)
C14—C9—Ru1	70.90 (14)	F6—P1—F4	90.27 (19)
C10—C9—Ru1	70.05 (14)	F6—P1—F3	89.3 (2)
C15—C9—Ru1	130.05 (19)	F4—P1—F3	178.93 (15)
C11—C10—C9	121.8 (3)	F6—P1—F5	179.13 (19)
C11—C10—Ru1	72.07 (14)	F4—P1—F5	88.87 (16)
C9—C10—Ru1	72.03 (14)	F3—P1—F5	91.57 (18)
C11—C10—H10	119.1	F6—P1—F2	90.33 (16)
C9—C10—H10	119.1	F4—P1—F2	90.70 (13)
Ru1—C10—H10	129.3	F3—P1—F2	88.33 (13)
C10—C11—C12	120.4 (3)	F5—P1—F2	89.85 (14)
C10—C11—Ru1	70.78 (15)	F6—P1—F1	90.38 (15)
C12—C11—Ru1	71.39 (15)	F4—P1—F1	90.59 (13)
C10—C11—H11	119.8	F3—P1—F1	90.39 (13)
C12—C11—H11	119.8	F5—P1—F1	89.46 (13)
Ru1—C11—H11	130.7	F2—P1—F1	178.53 (13)
C13—C12—C11	118.1 (2)	N1—Ru1—N2	77.91 (8)
C13—C12—C18	121.5 (3)	N1—Ru1—C13	95.62 (9)
C11—C12—C18	120.3 (3)	N2—Ru1—C13	117.75 (9)
C13—C12—Ru1	70.36 (15)	N1—Ru1—C10	157.21 (10)
C11—C12—Ru1	70.77 (15)	N2—Ru1—C10	124.11 (10)
C18—C12—Ru1	128.85 (18)	C13—Ru1—C10	79.68 (10)
C12—C13—C14	121.1 (3)	N1—Ru1—C14	95.53 (9)
C12—C13—Ru1	72.49 (15)	N2—Ru1—C14	154.82 (9)
C14—C13—Ru1	71.58 (14)	C13—Ru1—C14	37.97 (10)
C12—C13—H13	119.4	C10—Ru1—C14	67.02 (10)
C14—C13—H13	119.4	N1—Ru1—C11	158.10 (9)
Ru1—C13—H13	128.9	N2—Ru1—C11	97.57 (9)
C9—C14—C13	121.2 (3)	C13—Ru1—C11	67.23 (10)
C9—C14—Ru1	71.98 (15)	C10—Ru1—C11	37.15 (10)
C13—C14—Ru1	70.45 (15)	C14—Ru1—C11	79.41 (10)
C9—C14—H14	119.4	N1—Ru1—C12	120.53 (9)
C13—C14—H14	119.4	N2—Ru1—C12	94.32 (9)
Ru1—C14—H14	130.9	C13—Ru1—C12	37.15 (11)
C9—C15—C17	114.1 (2)	C10—Ru1—C12	67.90 (10)
C9—C15—C16	108.1 (2)	C14—Ru1—C12	67.84 (10)
C17—C15—C16	112.0 (3)	C11—Ru1—C12	37.84 (10)
C9—C15—H15	107.5	N1—Ru1—C9	119.65 (9)
C17—C15—H15	107.5	N2—Ru1—C9	161.77 (10)
C16—C15—H15	107.5	C13—Ru1—C9	68.20 (10)
C15—C16—H16A	109.5	C10—Ru1—C9	37.92 (11)

C15—C16—H16B	109.5	C14—Ru1—C9	37.11 (10)
H16A—C16—H16B	109.5	C11—Ru1—C9	68.03 (10)
C15—C16—H16C	109.5	C12—Ru1—C9	81.09 (10)
H16A—C16—H16C	109.5	N1—Ru1—Cl1	86.45 (6)
H16B—C16—H16C	109.5	N2—Ru1—Cl1	84.46 (6)
C15—C17—H17A	109.5	C13—Ru1—Cl1	157.67 (7)
C15—C17—H17B	109.5	C10—Ru1—Cl1	89.80 (7)
H17A—C17—H17B	109.5	C14—Ru1—Cl1	119.71 (7)
C15—C17—H17C	109.5	C11—Ru1—Cl1	114.67 (8)
H17A—C17—H17C	109.5	C12—Ru1—Cl1	152.21 (8)
H17B—C17—H17C	109.5	C9—Ru1—Cl1	91.43 (7)
N1—C1—C2—C8	1.0 (4)	C11—C12—C13—Ru1	-54.0 (2)
C8 ⁱ —C3—C4—N2	-0.3 (4)	C18—C12—C13—Ru1	124.2 (2)
N2—C5—C6—N1	-0.1 (3)	C10—C9—C14—C13	1.5 (4)
C7—C5—C6—N1	179.9 (2)	C15—C9—C14—C13	-178.5 (2)
N2—C5—C6—C7 ⁱ	179.2 (2)	Ru1—C9—C14—C13	-52.6 (2)
C7—C5—C6—C7 ⁱ	-0.8 (4)	C10—C9—C14—Ru1	54.1 (2)
N2—C5—C7—C6 ⁱ	-179.2 (2)	C15—C9—C14—Ru1	-125.9 (2)
C6—C5—C7—C6 ⁱ	0.8 (4)	C12—C13—C14—C9	-1.9 (4)
N2—C5—C7—C8 ⁱ	1.2 (4)	Ru1—C13—C14—C9	53.3 (2)
C6—C5—C7—C8 ⁱ	-178.8 (2)	C12—C13—C14—Ru1	-55.2 (2)
C1—C2—C8—C7 ⁱ	-1.0 (4)	C14—C9—C15—C17	37.2 (4)
C1—C2—C8—C3 ⁱ	179.1 (2)	C10—C9—C15—C17	-142.9 (3)
C14—C9—C10—C11	-0.1 (4)	Ru1—C9—C15—C17	-54.7 (4)
C15—C9—C10—C11	180.0 (2)	C14—C9—C15—C16	-88.1 (3)
Ru1—C9—C10—C11	54.5 (2)	C10—C9—C15—C16	91.9 (3)
C14—C9—C10—Ru1	-54.5 (2)	Ru1—C9—C15—C16	-180.0 (3)
C15—C9—C10—Ru1	125.5 (2)	C7 ⁱ —C6—N1—C1	-0.5 (4)
C9—C10—C11—C12	-1.1 (4)	C5—C6—N1—C1	178.8 (2)
Ru1—C10—C11—C12	53.4 (2)	C7 ⁱ —C6—N1—Ru1	178.40 (18)
C9—C10—C11—Ru1	-54.4 (2)	C5—C6—N1—Ru1	-2.4 (3)
C10—C11—C12—C13	0.8 (4)	C2—C1—N1—C6	-0.2 (4)
Ru1—C11—C12—C13	53.8 (2)	C2—C1—N1—Ru1	-178.90 (19)
C10—C11—C12—C18	-177.5 (2)	C7—C5—N2—C4	-1.2 (4)
Ru1—C11—C12—C18	-124.4 (2)	C6—C5—N2—C4	178.8 (2)
C10—C11—C12—Ru1	-53.1 (2)	C7—C5—N2—Ru1	-177.53 (18)
C11—C12—C13—C14	0.7 (4)	C6—C5—N2—Ru1	2.5 (3)
C18—C12—C13—C14	179.0 (2)	C3—C4—N2—C5	0.8 (4)
Ru1—C12—C13—C14	54.7 (2)	C3—C4—N2—Ru1	176.49 (19)

Symmetry code: (i) $-x+1, -y+1, -z$.

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C1—H1 \cdots O1	0.93	2.39	3.194 (4)	145
C3—H3 \cdots Cl1 ⁱⁱ	0.93	2.84	3.637 (3)	145

C3—H3···F6 ⁱⁱⁱ	0.93	2.58	3.205 (4)	125
C4—H4···F6 ⁱⁱⁱ	0.93	2.60	3.223 (4)	125
C19—H19B···F4 ^{iv}	0.96	2.46	3.321 (5)	149

Symmetry codes: (ii) $-x, -y+1, -z$; (iii) $x-1, y, z$; (iv) $x, y+1, z$.