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Diaquabis(3-chlorobenzoato-κO)-bis(nicotinamide-κN^1)cobalt(II)

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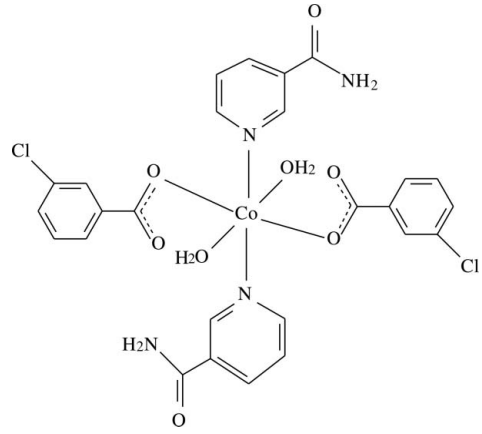
Key indicators: single-crystal X-ray study; T = 294 K; mean σ(C–C) = 0.002 Å; R factor = 0.027; wR factor = 0.074; data-to-parameter ratio = 13.7.

In the title complex, $[Co(C_7H_4ClO_2)_2(C_6H_6N_2O)_2(H_2O)_2]$, the Co^{II} atom is located on an inversion center and is coordinated by two 3-chlorobenzoate (CB) anions, two nicotinamide (NA) ligands and two water molecules. The four O atoms in the equatorial plane form a slightly distorted square-planar arrangement, while the slightly distorted octahedral coordination is completed by the two N atoms of the NA ligands in the axial positions. The dihedral angle between the carboxylate group and the adjacent benzene ring is 9.14 (9)°, while the pyridine and benzene rings are oriented at a dihedral angle of 82.18 (8)°. In the crystal, N—H···O and O—H···O hydrogen bonds link the molecules into a two-dimensional network lying parallel to (101). π–π stacking between parallel pyridine rings of adjacent molecules [centroid–centroid distance = 3.7765 (8) Å] further stabilizes the crystal structure.

Related literature

For literature on niacin, see: Krishnamachari (1974). For information on the nicotinic acid derivative *N*,*N*-diethylnicotinamide, see: Bigoli *et al.* (1972). For related structures, see: Aydın *et al.* (2012); Hökelek *et al.* (1996, 2009*a*,*b*); Hökelek & Necefoğlu (1998, 2007); Necefoğlu *et al.* (2011*a*,*b*); Sertçelik *et al.* (2012*a*,*b*,*c*). For bond-length data, see: Allen *et al.* (1987).



Experimental

Crystal data

$[Co(C_7H_4ClO_2)_2(C_6H_6N_2O)_2(H_2O)_2]$
M_r = 650.32
Monoclinic, $P2_1/n$
a = 11.5181 (3) Å
b = 8.8191 (2) Å
c = 13.5089 (3) Å
β = 90.546 (2)°
V = 1372.16 (6) Å^3
Z = 2
Mo $K\alpha$ radiation
μ = 0.88 mm^{-1}
T = 294 K
0.35 × 0.22 × 0.18 mm

Data collection

Bruker APEXII CCD area-detector diffractometer
Absorption correction: multi-scan (*SADABS*; Bruker, 2012)
T_{min} = 0.793, T_{max} = 0.854
18960 measured reflections
2797 independent reflections
2667 reflections with $I > 2\sigma(I)$
R_{int} = 0.027

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.027
$wR(F^2)$ = 0.074
S = 1.11
2797 reflections
204 parameters
52 restraints
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.32 e Å^{-3}
$\Delta\rho_{min}$ = −0.33 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
N2—H21···O3^{i}	0.84 (3)	2.24 (3)	2.876 (2)	133 (2)
N2—H22···O4^{ii}	0.87 (2)	2.27 (2)	3.012 (2)	143 (2)
O4—H41···O1^{iii}	0.92 (2)	1.68 (2)	2.5822 (16)	164 (2)
O4—H42···O3^{iv}	0.83 (3)	1.98 (3)	2.7892 (15)	166 (2)

Symmetry codes: (i) $-x+\frac{1}{2}, y+\frac{1}{2}, -z+\frac{3}{2}$; (ii) $-x+\frac{1}{2}, y-\frac{1}{2}, -z+\frac{3}{2}$; (iii) $-x, -y+2, -z+2$; (iv) $x, y+1, z$.

Data collection: *APEX2* (Bruker, 2012); cell refinement: *SAINT* (Bruker, 2012); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 2012); software used to prepare material for publication: *WinGX* (Farrugia, 2012) and *PLATON* (Spek, 2009).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SU2606).

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supporting information

Acta Cryst. (2013). E**69**, m349–m350 [doi:10.1107/S160053681301458X]

Diaquabis(3-chlorobenzoato-κO)bis(nicotinamide-κN^1)cobalt(II)

Nihat Bozkurt, Nefise Dilek, Nagihan Çaylak Delibaş, Hacali Necefoğlu and Tuncer Hökelek

S1. Comment

As a part of our ongoing investigations of transition metal complexes of nicotinamide (NA), one form of niacin (Krishnamachari, 1974), and/or the nicotinic acid derivative *N*,*N*-diethylnicotinamide (DENA), an important respiratory stimulant (Bigoli *et al.*, 1972), the title compound was synthesized and its crystal structure is reported herein.

In the title mononuclear complex, Co^{II} atom is located on an inversion center and is coordinated by two 3-chlorobenzoate (CB) anions, two nicotinamide (NA) ligands and two water molecules, all ligands coordinating in a monodentate manner (Fig. 1). The crystal structures of similar complexes of Cu^{II}, Co^{II}, Ni^{II}, Mn^{II} and Zn^{II} ions, $[Cu(C_7H_5O_2)_2(C_{10}H_{14}N_2O)_2]$ (Hökelek *et al.*, 1996), $[Cu(C_7H_4BrO_2)_2(C_6H_6N_2O)_2(H_2O)_2]$ (Necefoğlu *et al.*, 2011*a*), $[Co(C_6H_6N_2O)_2(C_7H_4NO_4)_2(H_2O)_2]$ (Hökelek & Necefoğlu, 1998), $[Co(C_9H_9O_2)_2(C_{10}H_{14}N_2O)_2(H_2O)_2]$ (Necefoğlu *et al.*, 2011*b*), $[Co(C_7H_4IO_2)_2(C_6H_6N_2O)_2(H_2O)_2]$ (Aydın *et al.*, 2012), $[Co(C_8H_5O_3)_2(C_6H_6N_2O)_2(H_2O)_2]$ (Sertçelik *et al.*, 2012*a*), $[Ni(C_7H_4ClO_2)_2(C_6H_6N_2O)_2(H_2O)_2]$ (Hökelek *et al.*, 2009*a*), $[Ni(C_5H_5O_3)_2(C_6H_6N_2O)_2(H_2O)_2]$ (Sertçelik *et al.*, 2012*b*), $[Mn(C_9H_{10}NO_2)_2(H_2O)_4].2H_2O$ (Hökelek & Necefoğlu, 2007), $[Zn(C_7H_4BrO_2)_2(C_6H_6N_2O)_2(H_2O)_2]$ (Hökelek *et al.*, 2009*b*) and $[Zn(C_8H_5O_3)_2(C_{10}H_{14}N_2O)_2(H_2O)_2]$ (Sertçelik *et al.*, 2012*c*) have been reported. In the copper(II) complex mentioned above the two benzoate ions coordinate to the Cu^{II} atom as bidentate ligands, while in the other structures all the ligands coordinate in a monodentate manner.

In the title complex, Fig. 1, the four symmetry related O atoms (O2, O2a, O4 and O4a) [symmetry code: (a) - *x*, - *y*, - *z*] in the equatorial plane around the Co^{II} ion form a slightly distorted square-planar arrangement, while the slightly distorted octahedral coordination is completed by the two symmetry related N atoms of the NA ligands (N1 and N1a) in the axial positions. The near equalities of the C1—O1 [1.2435 (19) Å] and C1—O2 [1.2677 (18) Å] bonds in the carboxylate group indicate delocalized bonding arrangement, rather than localized single and double bonds. The Co—O bond lengths are 2.0592 (10) Å (for benzoate oxygens) and 2.1385 (10) Å (for water oxygens), and the Co—N bond length is 2.1641 (11) Å, close to standard values (Allen *et al.*, 1987). The Co atom is displaced out of the mean-plane of the carboxylate group (O1/C1/O2) by -0.5077 (1) Å. The dihedral angle between the planar carboxylate group and the adjacent benzene ring (C2—C7) is 9.14 (9)°. The benzene (C2—C7) and the pyridine (N1/C8—C12) rings are oriented at a dihedral angle of 82.18 (8)°.

In the crystal, N—H···O and O—H···O hydrogen bonds (Table 1) link the molecules into a two-dimensional network lying parallel to (101). $\pi\cdots\pi$ stacking between the pyridine rings, $Cg\cdots Cg^{i}$ [symmetry code: (i) -x, -y+1, -z+2, where Cg is the centroid of ring N1/C8—C12] further stabilizes the crystal structure, with a centroid-centroid distance of 3.7765 (8) Å.

S2. Experimental

The title compound was prepared by the reaction of $CoSO_4.H_2O$ (0.865 g, 5 mmol) in H_2O (25 ml) and nicotinamide (1.22 g, 50 mmol) in H_2O (100 ml) with sodium 3-chlorobenzoate (1.79 g, 10 mmol) in H_2O (100 ml) at room

temperature. The mixture was filtered and set aside to crystallize at ambient temperature for several days, giving pink single crystals.

S3. Refinement

Atoms H21 and H22 (for NH_2) and H41 and H42 (for H_2O) were located in a difference Fourier map and were refined freely. The C-bound H-atoms were positioned geometrically with C—H = 0.93 Å for aromatic H-atoms, and constrained to ride on their parent atoms, with $U_{iso}(H) = 1.2 \times U_{eq}(C)$.

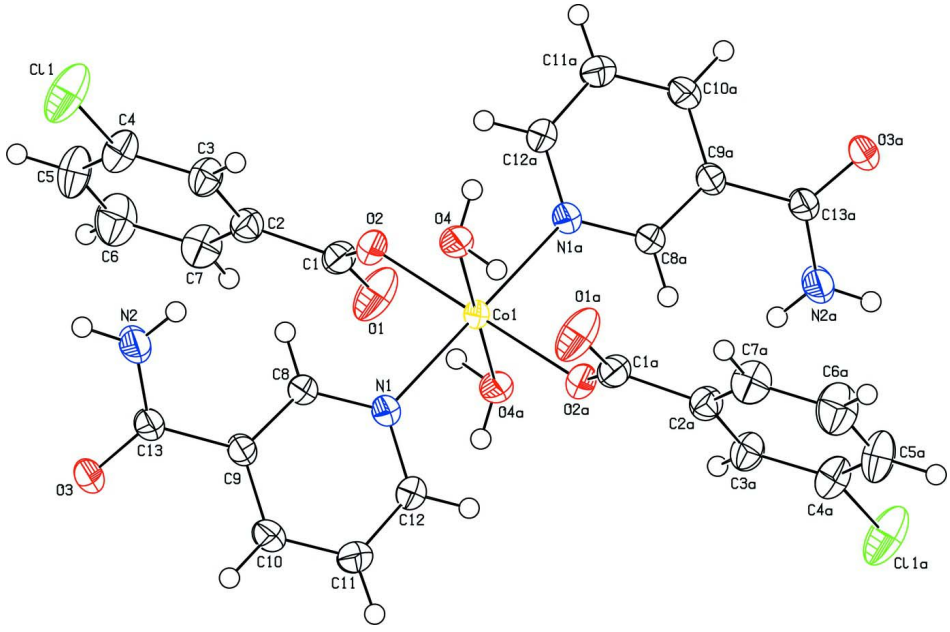

Figure 1

The molecular structure of the title molecule, with atom labelling. Displacement ellipsoids are drawn at the 50% probability level [symmetry code: (a) - *x*, - *y*, - *z*].

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Crystal data

$[Co(C_7H_4ClO_2)_2(C_6H_6N_2O)_2(H_2O)_2]$
$M_r = 650.32$
Monoclinic, $P2_1/n$
Hall symbol: -P 2yn
$a = 11.5181$ (3) Å
$b = 8.8191$ (2) Å
$c = 13.5089$ (3) Å
$\beta = 90.546$ (2)°
$V = 1372.16$ (6) Å^3
$Z = 2$
$F(000) = 666$
$D_x = 1.574$ Mg m^{-3}
Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Cell parameters from 4857 reflections
$\theta = 2.3$–24.4°
$\mu = 0.88$ mm^{-1}
$T = 294$ K
Block, pink
0.35 × 0.22 × 0.18 mm

Data collection

Bruker APEXII CCD area-detector diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
φ and ω scans
Absorption correction: multi-scan (*SADABS*; Bruker, 2012)
$T_{min} = 0.793$, $T_{max} = 0.854$
18960 measured reflections
2797 independent reflections
2667 reflections with $I > 2\sigma(I)$

R_{int} = 0.027
θ_{max} = 26.4°, θ_{min} = 2.3°
h = −14→14
k = −11→10
l = −16→16

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)]$ = 0.027
$wR(F^2)$ = 0.074
S = 1.11
2797 reflections
204 parameters
52 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H atoms treated by a mixture of independent and constrained refinement
$w = 1/[\sigma^2(F_o^2) + (0.0393P)^2 + 0.5258P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max}$ = 0.001
$\Delta\rho_{max}$ = 0.32 e Å^{-3}
$\Delta\rho_{min}$ = −0.33 e Å^{-3}
Extinction correction: *SHELXL97* (Sheldrick, 2008), $Fc^*=kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
Extinction coefficient: 0.0334 (17)

Special details

Geometry. Bond distances, angles etc. have been calculated using the rounded fractional coordinates. All su's are estimated from the variances of the (full) variance-covariance matrix. The cell esds are taken into account in the estimation of distances, angles and torsion angles

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2, conventional R-factors R are based on F, with F set to zero for negative F^2. The threshold expression of $F^2 > 2\text{sigma}(F^2)$ is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
Co1	0.00000	1.00000	1.00000	0.0221 (1)
Cl1	0.57400 (4)	0.77742 (8)	0.87425 (3)	0.0604 (2)
O1	0.17264 (10)	0.87980 (18)	1.18080 (9)	0.0518 (4)
O2	0.17291 (8)	0.94984 (12)	1.02212 (8)	0.0299 (3)
O3	0.16810 (10)	0.34420 (12)	0.81282 (10)	0.0415 (4)
O4	0.04558 (10)	1.07995 (12)	0.85606 (8)	0.0298 (3)
N1	−0.02319 (10)	0.77771 (13)	0.93492 (9)	0.0258 (3)
N2	0.23334 (16)	0.56716 (18)	0.76095 (16)	0.0592 (6)
C1	0.21833 (12)	0.88839 (17)	1.09796 (11)	0.0293 (4)
C2	0.33518 (12)	0.81535 (17)	1.08465 (11)	0.0286 (4)
C3	0.39495 (13)	0.83285 (18)	0.99657 (11)	0.0316 (4)
C4	0.49988 (13)	0.7599 (2)	0.98487 (12)	0.0368 (5)
C5	0.54679 (15)	0.6700 (2)	1.05879 (14)	0.0464 (6)
C6	0.48691 (17)	0.6533 (2)	1.14586 (14)	0.0496 (6)
C7	0.38168 (15)	0.7255 (2)	1.15907 (12)	0.0388 (5)
C8	0.06524 (12)	0.70512 (15)	0.89207 (11)	0.0268 (4)
C9	0.05526 (12)	0.56102 (15)	0.85229 (10)	0.0258 (4)
C10	−0.05151 (14)	0.48869 (16)	0.85705 (12)	0.0314 (4)
C11	−0.14353 (13)	0.56408 (18)	0.89921 (13)	0.0356 (5)
C12	−0.12591 (12)	0.70790 (16)	0.93721 (11)	0.0303 (4)
C13	0.15664 (14)	0.48213 (16)	0.80673 (12)	0.0304 (4)

H3	0.36460	0.89300	0.94610	0.0380*
H5	0.61770	0.62160	1.04990	0.0560*
H6	0.51750	0.59290	1.19620	0.0590*
H7	0.34200	0.71360	1.21820	0.0470*
H8	0.13670	0.75380	0.88890	0.0320*
H10	−0.06090	0.39100	0.83220	0.0380*
H11	−0.21640	0.51890	0.90200	0.0430*
H12	−0.18830	0.75810	0.96560	0.0360*
H21	0.221 (2)	0.659 (3)	0.7492 (19)	0.066 (7)*
H22	0.293 (2)	0.526 (3)	0.7323 (19)	0.063 (7)*
H41	−0.028 (2)	1.098 (3)	0.8311 (18)	0.069 (7)*
H42	0.079 (2)	1.163 (3)	0.8530 (17)	0.055 (6)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Co1	0.0201 (2)	0.0183 (2)	0.0280 (2)	0.0011 (1)	0.0066 (1)	−0.0026 (1)
Cl1	0.0389 (2)	0.1013 (4)	0.0412 (3)	0.0171 (2)	0.0140 (2)	0.0029 (2)
O1	0.0339 (6)	0.0875 (10)	0.0342 (6)	0.0108 (6)	0.0090 (5)	0.0038 (6)
O2	0.0230 (5)	0.0306 (5)	0.0362 (5)	0.0039 (4)	0.0044 (4)	0.0007 (4)
O3	0.0403 (6)	0.0194 (5)	0.0651 (8)	0.0024 (4)	0.0211 (5)	−0.0007 (5)
O4	0.0297 (5)	0.0267 (6)	0.0332 (5)	−0.0023 (4)	0.0098 (4)	−0.0001 (4)
N1	0.0256 (6)	0.0205 (5)	0.0313 (6)	0.0005 (4)	0.0049 (5)	−0.0027 (4)
N2	0.0569 (10)	0.0231 (7)	0.0985 (14)	0.0036 (7)	0.0507 (10)	0.0049 (8)
C1	0.0247 (7)	0.0308 (7)	0.0324 (7)	−0.0020 (5)	0.0034 (5)	−0.0042 (6)
C2	0.0261 (7)	0.0297 (7)	0.0301 (7)	0.0001 (6)	−0.0001 (5)	−0.0025 (6)
C3	0.0273 (7)	0.0369 (8)	0.0307 (7)	0.0051 (6)	0.0000 (6)	0.0031 (6)
C4	0.0286 (7)	0.0489 (10)	0.0331 (8)	0.0054 (6)	0.0038 (6)	−0.0028 (7)
C5	0.0344 (8)	0.0564 (11)	0.0482 (10)	0.0200 (8)	−0.0027 (7)	0.0004 (8)
C6	0.0479 (10)	0.0572 (11)	0.0435 (10)	0.0165 (9)	−0.0075 (8)	0.0127 (8)
C7	0.0401 (9)	0.0459 (9)	0.0303 (8)	0.0038 (7)	0.0005 (6)	0.0045 (7)
C8	0.0251 (6)	0.0214 (6)	0.0339 (7)	−0.0006 (5)	0.0068 (5)	−0.0010 (5)
C9	0.0292 (7)	0.0197 (6)	0.0287 (7)	0.0017 (5)	0.0063 (5)	0.0004 (5)
C10	0.0336 (8)	0.0220 (7)	0.0388 (8)	−0.0030 (5)	0.0051 (6)	−0.0062 (5)
C11	0.0270 (7)	0.0312 (8)	0.0486 (9)	−0.0058 (6)	0.0062 (6)	−0.0072 (7)
C12	0.0249 (7)	0.0276 (7)	0.0384 (8)	0.0010 (5)	0.0073 (6)	−0.0032 (6)
C13	0.0320 (8)	0.0211 (7)	0.0382 (8)	0.0000 (5)	0.0112 (6)	−0.0025 (6)

Geometric parameters (Å, º)

Co1—O2	2.0592 (9)	C2—C3	1.389 (2)
Co1—O4	2.1385 (11)	C3—C4	1.380 (2)
Co1—N1	2.1640 (12)	C4—C5	1.381 (2)
Co1—O2^{i}	2.0592 (9)	C5—C6	1.377 (3)
Co1—O4^{i}	2.1385 (11)	C6—C7	1.382 (3)
Co1—N1^{i}	2.1640 (12)	C8—C9	1.3841 (19)
Cl1—C4	1.7350 (16)	C9—C13	1.497 (2)
O1—C1	1.2435 (19)	C9—C10	1.387 (2)

O2—C1	1.2676 (18)	C10—C11	1.379 (2)
O3—C13	1.2262 (18)	C11—C12	1.383 (2)
O4—H41	0.92 (2)	C3—H3	0.9300
O4—H42	0.83 (3)	C5—H5	0.9300
N1—C12	1.3344 (18)	C6—H6	0.9300
N1—C8	1.3392 (18)	C7—H7	0.9300
N2—C13	1.317 (2)	C8—H8	0.9300
N2—H21	0.84 (3)	C10—H10	0.9300
N2—H22	0.87 (2)	C11—H11	0.9300
C1—C2	1.504 (2)	C12—H12	0.9300
C2—C7	1.384 (2)		

O2—Co1—O4	87.55 (4)	Cl1—C4—C3	119.73 (12)
O2—Co1—N1	88.84 (4)	C3—C4—C5	121.47 (15)
O2—Co1—O2^{i}	180.00	Cl1—C4—C5	118.79 (13)
O2—Co1—O4^{i}	92.45 (4)	C4—C5—C6	118.90 (16)
O2—Co1—N1^{i}	91.16 (4)	C5—C6—C7	120.53 (17)
O4—Co1—N1	87.69 (4)	C2—C7—C6	120.25 (15)
O2^{i}—Co1—O4	92.45 (4)	N1—C8—C9	123.09 (13)
O4—Co1—O4^{i}	180.00	C8—C9—C13	121.62 (13)
O4—Co1—N1^{i}	92.31 (4)	C8—C9—C10	118.33 (13)
O2^{i}—Co1—N1	91.16 (4)	C10—C9—C13	120.05 (12)
O4^{i}—Co1—N1	92.31 (4)	C9—C10—C11	118.86 (13)
N1—Co1—N1^{i}	180.00	C10—C11—C12	119.01 (14)
O2^{i}—Co1—O4^{i}	87.55 (4)	N1—C12—C11	122.80 (13)
O2^{i}—Co1—N1^{i}	88.84 (4)	N2—C13—C9	117.24 (13)
O4^{i}—Co1—N1^{i}	87.69 (4)	O3—C13—N2	121.59 (16)
Co1—O2—C1	126.89 (9)	O3—C13—C9	121.16 (14)
H41—O4—H42	105 (2)	C2—C3—H3	120.00
Co1—O4—H41	99.0 (15)	C4—C3—H3	120.00
Co1—O4—H42	117.1 (16)	C4—C5—H5	121.00
Co1—N1—C8	121.14 (9)	C6—C5—H5	121.00
Co1—N1—C12	120.97 (9)	C5—C6—H6	120.00
C8—N1—C12	117.88 (12)	C7—C6—H6	120.00
H21—N2—H22	117 (2)	C2—C7—H7	120.00
C13—N2—H21	121.8 (16)	C6—C7—H7	120.00
C13—N2—H22	120.5 (17)	N1—C8—H8	118.00
O1—C1—O2	125.34 (14)	C9—C8—H8	118.00
O1—C1—C2	117.92 (13)	C9—C10—H10	121.00
O2—C1—C2	116.70 (13)	C11—C10—H10	121.00
C1—C2—C3	120.41 (13)	C10—C11—H11	121.00
C1—C2—C7	119.92 (13)	C12—C11—H11	120.00
C3—C2—C7	119.63 (14)	N1—C12—H12	119.00
C2—C3—C4	119.23 (14)	C11—C12—H12	119.00

O4—Co1—O2—C1	−177.85 (12)	O2—C1—C2—C7	−169.65 (14)
N1—Co1—O2—C1	−90.11 (12)	C1—C2—C3—C4	−177.53 (14)
O4^{i}—Co1—O2—C1	2.15 (12)	C7—C2—C3—C4	0.1 (2)

N1^{i}—Co1—O2—C1	89.89 (12)	C1—C2—C7—C6	177.50 (15)
O2—Co1—N1—C8	−26.16 (11)	C3—C2—C7—C6	−0.1 (2)
O2—Co1—N1—C12	153.10 (11)	C2—C3—C4—Cl1	178.62 (12)
O4—Co1—N1—C8	61.44 (11)	C2—C3—C4—C5	0.0 (2)
O4—Co1—N1—C12	−119.31 (11)	Cl1—C4—C5—C6	−178.62 (14)
O2^{i}—Co1—N1—C8	153.84 (11)	C3—C4—C5—C6	0.1 (3)
O2^{i}—Co1—N1—C12	−26.90 (11)	C4—C5—C6—C7	−0.1 (3)
O4^{i}—Co1—N1—C8	−118.56 (11)	C5—C6—C7—C2	0.1 (3)
O4^{i}—Co1—N1—C12	60.69 (11)	N1—C8—C9—C10	0.1 (2)
Co1—O2—C1—O1	−18.0 (2)	N1—C8—C9—C13	−178.77 (13)
Co1—O2—C1—C2	159.69 (10)	C8—C9—C10—C11	1.3 (2)
Co1—N1—C8—C9	177.93 (11)	C13—C9—C10—C11	−179.86 (14)
C12—N1—C8—C9	−1.4 (2)	C8—C9—C13—O3	146.12 (16)
Co1—N1—C12—C11	−177.98 (12)	C8—C9—C13—N2	−33.2 (2)
C8—N1—C12—C11	1.3 (2)	C10—C9—C13—O3	−32.7 (2)
O1—C1—C2—C3	−174.24 (15)	C10—C9—C13—N2	148.03 (17)
O1—C1—C2—C7	8.2 (2)	C9—C10—C11—C12	−1.3 (2)
O2—C1—C2—C3	7.9 (2)	C10—C11—C12—N1	0.0 (2)

Symmetry code: (i) −*x*, −*y*+2, −*z*+2.

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N2—H21···O3^{ii}	0.84 (3)	2.24 (3)	2.876 (2)	133 (2)
N2—H22···O4^{iii}	0.87 (2)	2.27 (2)	3.012 (2)	143 (2)
O4—H41···O1^{i}	0.92 (2)	1.68 (2)	2.5822 (16)	164 (2)
O4—H42···O3^{iv}	0.83 (3)	1.98 (3)	2.7892 (15)	166 (2)

Symmetry codes: (i) −*x*, −*y*+2, −*z*+2; (ii) −*x*+1/2, *y*+1/2, −*z*+3/2; (iii) −*x*+1/2, *y*−1/2, −*z*+3/2; (iv) *x*, *y*+1, *z*.