

2-(4-Methylphenyl)-6-nitro-1,3-benzoxazole

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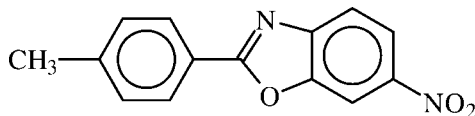
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 Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.006$ Å; R factor = 0.055; wR factor = 0.168; data-to-parameter ratio = 12.3.

The title compound, $\text{C}_{14}\text{H}_{10}\text{N}_2\text{O}_3$, is a π -conjugated molecule containing a benzoxazole aromatic fused heterobicycle. The benzoxazole ring system is planar within 0.01 Å. The molecule assumes an approximately flat conformation, the benzoxazole ring system forming dihedral angles of 6.52 (12) and 7.4 (3)° with the benzene ring and the nitro group, respectively. In the crystal, molecules are connected by very weak $\text{C}-\text{H}\cdots\text{O}$ hydrogen interactions, forming chains running parallel to the a or c axes. The methyl H atoms are disordered over two sets of sites of equal occupancy rotated by 60°.

Related literature

For general information on heterocycles in organic electronics and optoelectronics, see: Dalton (2002); Heeger (2010). For heterocycle-based semiconductors, optoelectronic and piezoelectric materials, see: Carella, Centore, Sirigu *et al.* (2004); Centore, Ricciotti *et al.* (2012); Centore, Concilio *et al.* (2012). For structural and theoretical analysis of conjugation in heterocycle-based organic molecules, see: Carella, Centore, Fort *et al.* (2004); Gainsford *et al.* (2008). For structural and theoretical analysis of conjugation in heterocycle-based metallorganic compounds, see: Takjoo *et al.* (2011); Takjoo & Centore (2013). For theoretical computations on similar compounds, see: Capobianco *et al.* (2012, 2013). For the synthesis of related heterocyclic compounds, see: Bruno *et al.* (2002); Centore *et al.* (2007); Piccialli *et al.* (2013); Centore, Fusco, Capobianco *et al.* (2013). For hydrogen bonding in crystals see: Desiraju & Steiner (1999); Centore, Fusco, Jazbinsek *et al.* (2013).



Experimental

Crystal data

$\text{C}_{14}\text{H}_{10}\text{N}_2\text{O}_3$	$V = 2432.8 (19) \text{ \AA}^3$
$M_r = 254.24$	$Z = 8$
Orthorhombic, $Pbca$	Mo $K\alpha$ radiation
$a = 27.251 (4) \text{ \AA}$	$\mu = 0.10 \text{ mm}^{-1}$
$b = 7.4457 (6) \text{ \AA}$	$T = 293 \text{ K}$
$c = 11.990 (9) \text{ \AA}$	$0.40 \times 0.20 \times 0.20 \text{ mm}$

Data collection

Enraf–Nonius MACH3 diffractometer	970 reflections with $I > 2\sigma(I)$
2968 measured reflections	$R_{\text{int}} = 0.020$
2140 independent reflections	1 standard reflections every 120 min
	intensity decay: none

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.055$	174 parameters
$wR(F^2) = 0.168$	H-atom parameters constrained
$S = 1.11$	$\Delta\rho_{\text{max}} = 0.14 \text{ e \AA}^{-3}$
2140 reflections	$\Delta\rho_{\text{min}} = -0.19 \text{ e \AA}^{-3}$

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{C10}-\text{H10}\cdots\text{O2}^i$	0.93	2.61	3.502 (6)	160
$\text{C1}-\text{H1B}\cdots\text{O2}^{ii}$	0.96	2.80	3.143 (5)	102

 Symmetry codes: (i) $-x + \frac{1}{2}, -y, z - \frac{1}{2}$; (ii) $x - \frac{1}{2}, -y + \frac{1}{2}, -z$.

Data collection: *MACH3/PC* Software (Nonius, 1996); cell refinement: *CELLFITW* (Centore, 2004); data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SIR97* (Altomare *et al.*, 1999); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 2012) and *Mercury* (Macrae *et al.*, 2006); software used to prepare material for publication: *WinGX* (Farrugia, 2012).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: RZ5053).

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supporting information

Acta Cryst. (2013). E69, o667–o668 [https://doi.org/10.1107/S1600536813008970]

2-(4-Methylphenyl)-6-nitro-1,3-benzoxazole**Roberto Centore, Vincenzo Piccialli and Angela Tuzi****S1. Comment**

Heterocycles are important compounds of synthetic chemistry. Besides their long standing and relevant application as drugs and bioactive compounds, aromatic heterocycles are playing a fundamental role in modern material chemistry as building blocks of conjugated active molecules in some emerging fields of organic electronics and optoelectronics: conducting polymers and organic solar cells (Heeger, 2010), organic field-effect transistors (Centore, Ricciotti *et al.*, 2012), nonlinear optically active and piezoelectric compounds (Dalton, 2002; Carella, Centore, Sirigu *et al.* (2004); Centore, Concilio *et al.*, 2012). The chemical investigation is mainly directed to the synthesis of new molecules or conjugated polymers containing heterocyclic moieties. However, also the structural investigation of the molecules is relevant, pointing towards the quantitative evaluation of the structural parameters related to the conjugation (Carella, Centore, Fort *et al.*, 2004; Gainsford *et al.*, 2008; Capobianco *et al.*, 2012; Capobianco *et al.*, 2013). Following our interest in the synthesis and characterization of new heterocyclic compounds, including metal containing heterocyclic compounds (Takjoo *et al.*, 2011; Takjoo & Centore, 2013) for applications as advanced materials and bioactive compounds, and in the analysis of crystal structures controlled by the formation of H bonds (Centore, Fusco, Jazbinsek *et al.*, 2013), we report, in the present paper, the structural investigation of the title compound, shown in Scheme 1. 2-(4-methyl)-phenyl-6-nitro-benzoxazole is an organic dye containing the 6-nitrobenzoxazole acceptor group conjugated with a 4-methylphenyl moiety. The 6-nitrobenzoxazole-2-yl moiety has been used in the synthesis of polymers showing quadratic NLO behaviour (Bruno *et al.*, 2002).

The molecular structure of the title compound is shown in Fig. 1. The phenyl and benzoxazole rings are nearly coplanar, the dihedral angle between the mean planes being $6.7(1)^\circ$. That structural feature is in accordance with the expected π conjugation of the compound.

Molecules in the crystal form rows through very weak hydrogen interactions between methyl or aromatic C–H donors and oxygen acceptors of the nitro group (Fig. 2 and Fig. 3; Table 1). The chains, which have graph set symbol $C^1_1(14)$ and $C^1_1(7)$ are generated, respectively, by the *b* and *a* glide planes.

S2. Experimental

The title compound was prepared by reaction of 2-amino-5-nitrophenol (5.00 g, 32.4 mmol) with toluic acid (4.41 g, 32.4 mmol) in polyphosphoric acid (150 g) at 150°C . The dehydration procedure is analogous to that we have already described for the synthesis of similar chromophores (Bruno *et al.*, 2002; Centore *et al.*, 2007). Purification of the title compound was obtained by recrystallization from ethanol. The final yield was 5.69 g (69%). M. p. 437 K. Single crystals were obtained by slow evaporation of an ethanol solution. $^1\text{H-NMR}$ (CDCl_3) δ 2.47 (s, 3H), 7.38 (d, 2H, $J = 7.9$ Hz), 7.82 (d, 1H, $J = 8.5$ Hz), 8.18 (d, 2H, $J = 8.3$ Hz), 8.33 (d, 1H, $J = 8.7$ Hz), 8.48 (d, 1H, $J = 1.8$ Hz).

S3. Refinement

All H atoms were generated stereochemically. In particular, the methyl group is disordered over two sets of sites of equal occupancy rotated from each other by 60°. All H atoms were refined by a riding model with C—H = 0.93–0.96 Å and with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ or $1.5U_{\text{eq}}(\text{C})$ for methyl hydrogen atoms.

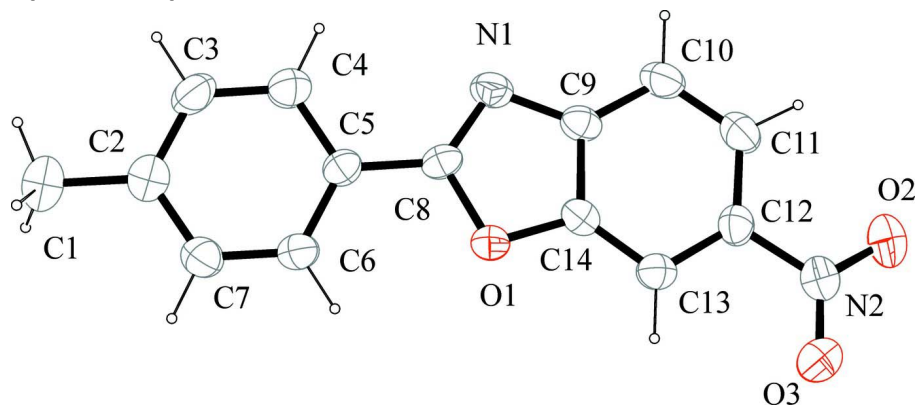


Figure 1

ORTEP view of the molecular structure of the title compound. Thermal ellipsoids are drawn at 50% probability level. Only one orientation of the disordered methyl group is shown.

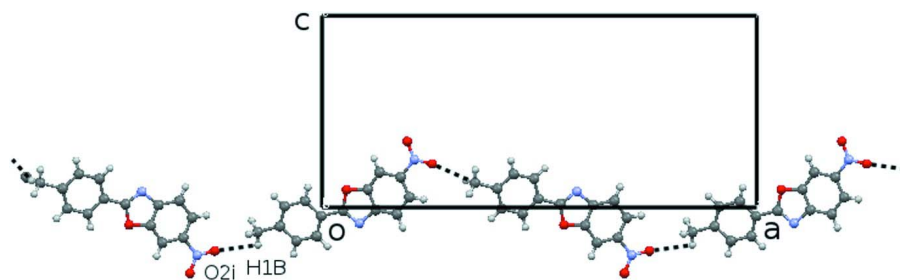


Figure 2

Row of molecules of the title compound running along the *a* axis. Only one orientation of the disordered methyl group is shown. Hydrogen bonds are shown as dashed lines.

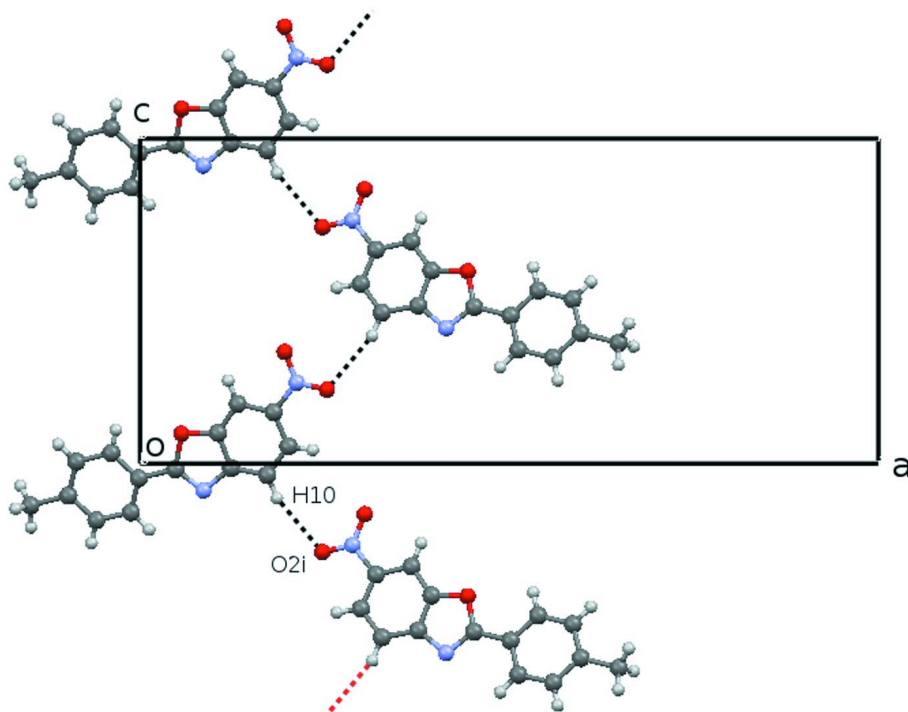


Figure 3

Row of molecules of the title compound running along the *c*. Only one orientation of the disordered methyl group is shown. Hydrogen bonds are shown as dashed lines.

2-(4-Methylphenyl)-6-nitro-1,3-benzoxazole

Crystal data

$C_{14}H_{10}N_2O_3$

$M_r = 254.24$

Orthorhombic, *Pbca*

Hall symbol: -P 2ac 2ab

$a = 27.251 (4) \text{ \AA}$

$b = 7.4457 (6) \text{ \AA}$

$c = 11.990 (9) \text{ \AA}$

$V = 2432.8 (19) \text{ \AA}^3$

$Z = 8$

$F(000) = 1056$

$D_x = 1.388 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 25 reflections

$\theta = 12.2\text{--}12.4^\circ$

$\mu = 0.10 \text{ mm}^{-1}$

$T = 293 \text{ K}$

Prism, brown

$0.40 \times 0.20 \times 0.20 \text{ mm}$

Data collection

Enraf–Nonius MACH3

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

non-profiled ω scans

2968 measured reflections

2140 independent reflections

970 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.020$

$\theta_{\text{max}} = 25.0^\circ$, $\theta_{\text{min}} = 1.5^\circ$

$h = -12 \rightarrow 32$

$k = -3 \rightarrow 8$

$l = -5 \rightarrow 14$

1 standard reflections every 120 min

intensity decay: none

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.055$
 $wR(F^2) = 0.168$
 $S = 1.11$
 2140 reflections
 174 parameters
 0 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: inferred from
 neighbouring sites
 H-atom parameters constrained
 $w = 1/[\sigma^2(F_o^2) + (0.0553P)^2 + 0.383P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} < 0.001$
 $\Delta\rho_{\max} = 0.14 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.19 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
O1	0.05625 (9)	0.1644 (3)	0.0918 (2)	0.0545 (7)	
O2	0.25311 (13)	-0.0885 (6)	0.2271 (3)	0.1125 (14)	
O3	0.19587 (12)	-0.0394 (6)	0.3439 (3)	0.1144 (14)	
N1	0.08562 (12)	0.1994 (4)	-0.0827 (3)	0.0614 (10)	
N2	0.21215 (15)	-0.0363 (6)	0.2493 (4)	0.0771 (11)	
C1	-0.14836 (14)	0.4375 (6)	-0.1283 (4)	0.0819 (15)	
H1A	-0.1684	0.3329	-0.1391	0.123*	0.57 (5)
H1B	-0.1477	0.5068	-0.1958	0.123*	0.57 (5)
H1C	-0.1618	0.5090	-0.0690	0.123*	0.57 (5)
H1D	-0.1502	0.5663	-0.1302	0.123*	0.43 (5)
H1E	-0.1709	0.3923	-0.0735	0.123*	0.43 (5)
H1F	-0.1568	0.3901	-0.2002	0.123*	0.43 (5)
C2	-0.09695 (16)	0.3808 (5)	-0.0983 (4)	0.0624 (12)	
C3	-0.05929 (17)	0.3930 (6)	-0.1742 (4)	0.0754 (13)	
H3	-0.0656	0.4411	-0.2443	0.091*	
C4	-0.01233 (16)	0.3358 (6)	-0.1496 (4)	0.0710 (13)	
H4	0.0122	0.3420	-0.2035	0.085*	
C5	-0.00180 (14)	0.2695 (5)	-0.0448 (3)	0.0516 (10)	
C6	-0.03907 (14)	0.2611 (5)	0.0325 (3)	0.0621 (11)	
H6	-0.0325	0.2182	0.1038	0.074*	
C7	-0.08606 (15)	0.3154 (6)	0.0058 (4)	0.0682 (13)	
H7	-0.1107	0.3075	0.0592	0.082*	
C8	0.04769 (15)	0.2121 (5)	-0.0175 (3)	0.0521 (10)	
C9	0.12336 (14)	0.1393 (5)	-0.0134 (3)	0.0548 (11)	
C10	0.17235 (15)	0.1012 (6)	-0.0351 (4)	0.0724 (13)	

H10	0.1855	0.1146	-0.1062	0.087*
C11	0.20086 (14)	0.0426 (6)	0.0532 (4)	0.0718 (13)
H11	0.2337	0.0139	0.0416	0.086*
C12	0.18069 (14)	0.0268 (6)	0.1583 (4)	0.0592 (11)
C13	0.13235 (14)	0.0639 (5)	0.1837 (3)	0.0571 (11)
H13	0.1194	0.0525	0.2551	0.069*
C14	0.10530 (14)	0.1191 (5)	0.0941 (3)	0.0509 (10)

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0553 (17)	0.0632 (18)	0.0449 (16)	0.0015 (13)	0.0038 (13)	0.0002 (14)
O2	0.057 (2)	0.171 (4)	0.110 (3)	0.020 (2)	-0.005 (2)	0.013 (3)
O3	0.086 (2)	0.190 (4)	0.068 (2)	0.022 (3)	-0.004 (2)	0.022 (3)
N1	0.064 (2)	0.073 (3)	0.0474 (19)	-0.0022 (18)	0.010 (2)	-0.0039 (19)
N2	0.058 (2)	0.096 (3)	0.077 (3)	-0.006 (2)	-0.004 (2)	0.006 (3)
C1	0.073 (3)	0.070 (3)	0.103 (4)	0.010 (3)	-0.024 (3)	-0.008 (3)
C2	0.071 (3)	0.051 (3)	0.066 (3)	0.000 (2)	-0.008 (3)	-0.011 (3)
C3	0.090 (3)	0.087 (3)	0.049 (3)	0.014 (3)	-0.011 (3)	0.007 (3)
C4	0.075 (3)	0.086 (3)	0.052 (3)	0.009 (3)	0.002 (2)	-0.001 (3)
C5	0.060 (2)	0.051 (3)	0.044 (2)	0.000 (2)	-0.002 (2)	-0.003 (2)
C6	0.064 (3)	0.073 (3)	0.050 (2)	-0.004 (3)	0.002 (2)	0.004 (2)
C7	0.060 (3)	0.078 (3)	0.067 (3)	-0.006 (2)	0.004 (2)	0.001 (3)
C8	0.067 (3)	0.051 (3)	0.038 (2)	0.000 (2)	-0.001 (2)	-0.003 (2)
C9	0.058 (3)	0.058 (3)	0.048 (2)	-0.009 (2)	0.008 (2)	-0.007 (2)
C10	0.065 (3)	0.093 (4)	0.058 (3)	-0.003 (3)	0.019 (3)	-0.001 (3)
C11	0.053 (3)	0.092 (3)	0.070 (3)	-0.001 (3)	0.010 (2)	-0.003 (3)
C12	0.050 (2)	0.066 (3)	0.062 (3)	-0.003 (2)	-0.005 (2)	-0.002 (2)
C13	0.059 (3)	0.060 (3)	0.052 (3)	-0.004 (2)	0.005 (2)	-0.001 (2)
C14	0.048 (2)	0.051 (2)	0.054 (3)	-0.005 (2)	0.006 (2)	-0.005 (2)

Geometric parameters (Å, °)

O1—C8	1.378 (4)	C3—H3	0.9300
O1—C14	1.379 (4)	C4—C5	1.380 (5)
O2—N2	1.212 (4)	C4—H4	0.9300
O3—N2	1.218 (5)	C5—C6	1.376 (5)
N1—C8	1.299 (4)	C5—C8	1.452 (5)
N1—C9	1.396 (5)	C6—C7	1.381 (5)
N2—C12	1.465 (5)	C6—H6	0.9300
C1—C2	1.506 (5)	C7—H7	0.9300
C1—H1A	0.9600	C9—C14	1.387 (5)
C1—H1B	0.9600	C9—C10	1.389 (5)
C1—H1C	0.9600	C10—C11	1.383 (6)
C1—H1D	0.9600	C10—H10	0.9300
C1—H1E	0.9600	C11—C12	1.380 (5)
C1—H1F	0.9600	C11—H11	0.9300
C2—C7	1.373 (6)	C12—C13	1.380 (5)

C2—C3	1.375 (6)	C13—C14	1.366 (5)
C3—C4	1.381 (6)	C13—H13	0.9300
C8—O1—C14	104.2 (3)	C5—C4—H4	120.1
C8—N1—C9	104.6 (3)	C3—C4—H4	120.1
O2—N2—O3	122.3 (4)	C6—C5—C4	118.4 (4)
O2—N2—C12	118.5 (4)	C6—C5—C8	121.4 (4)
O3—N2—C12	119.1 (4)	C4—C5—C8	120.2 (4)
C2—C1—H1A	109.5	C5—C6—C7	121.0 (4)
C2—C1—H1B	109.5	C5—C6—H6	119.5
H1A—C1—H1B	109.5	C7—C6—H6	119.5
C2—C1—H1C	109.5	C2—C7—C6	121.0 (4)
H1A—C1—H1C	109.5	C2—C7—H7	119.5
H1B—C1—H1C	109.5	C6—C7—H7	119.5
C2—C1—H1D	109.5	N1—C8—O1	114.8 (3)
H1A—C1—H1D	141.1	N1—C8—C5	128.7 (4)
H1B—C1—H1D	56.3	O1—C8—C5	116.6 (3)
H1C—C1—H1D	56.3	C14—C9—C10	119.5 (4)
C2—C1—H1E	109.5	C14—C9—N1	109.0 (3)
H1A—C1—H1E	56.3	C10—C9—N1	131.4 (4)
H1B—C1—H1E	141.1	C11—C10—C9	117.4 (4)
H1C—C1—H1E	56.3	C11—C10—H10	121.3
H1D—C1—H1E	109.5	C9—C10—H10	121.3
C2—C1—H1F	109.5	C12—C11—C10	120.1 (4)
H1A—C1—H1F	56.3	C12—C11—H11	119.9
H1B—C1—H1F	56.3	C10—C11—H11	119.9
H1C—C1—H1F	141.1	C13—C12—C11	124.4 (4)
H1D—C1—H1F	109.5	C13—C12—N2	117.3 (4)
H1E—C1—H1F	109.5	C11—C12—N2	118.3 (4)
C7—C2—C3	117.7 (4)	C14—C13—C12	113.7 (4)
C7—C2—C1	121.1 (4)	C14—C13—H13	123.1
C3—C2—C1	121.2 (4)	C12—C13—H13	123.1
C2—C3—C4	122.0 (4)	C13—C14—O1	127.8 (4)
C2—C3—H3	119.0	C13—C14—C9	124.8 (4)
C4—C3—H3	119.0	O1—C14—C9	107.4 (4)
C5—C4—C3	119.9 (4)		
C7—C2—C3—C4	2.4 (7)	N1—C9—C10—C11	-179.9 (4)
C1—C2—C3—C4	-177.7 (4)	C9—C10—C11—C12	-1.1 (7)
C2—C3—C4—C5	-2.1 (7)	C10—C11—C12—C13	0.8 (7)
C3—C4—C5—C6	0.5 (7)	C10—C11—C12—N2	179.8 (4)
C3—C4—C5—C8	-178.9 (4)	O2—N2—C12—C13	172.0 (4)
C4—C5—C6—C7	0.8 (6)	O3—N2—C12—C13	-5.9 (7)
C8—C5—C6—C7	-179.8 (4)	O2—N2—C12—C11	-7.0 (7)
C3—C2—C7—C6	-1.0 (6)	O3—N2—C12—C11	175.1 (5)
C1—C2—C7—C6	179.0 (4)	C11—C12—C13—C14	0.0 (6)
C5—C6—C7—C2	-0.6 (7)	N2—C12—C13—C14	-179.0 (4)
C9—N1—C8—O1	-0.3 (4)	C12—C13—C14—O1	179.4 (4)

C9—N1—C8—C5	179.6 (4)	C12—C13—C14—C9	-0.6 (6)
C14—O1—C8—N1	0.7 (4)	C8—O1—C14—C13	179.2 (4)
C14—O1—C8—C5	-179.2 (3)	C8—O1—C14—C9	-0.8 (4)
C6—C5—C8—N1	174.5 (4)	C10—C9—C14—C13	0.3 (6)
C4—C5—C8—N1	-6.1 (7)	N1—C9—C14—C13	-179.3 (3)
C6—C5—C8—O1	-5.6 (6)	C10—C9—C14—O1	-179.7 (3)
C4—C5—C8—O1	173.8 (4)	N1—C9—C14—O1	0.7 (4)
C8—N1—C9—C14	-0.2 (4)	C4—C5—C8—N1	-6.1 (7)
C8—N1—C9—C10	-179.8 (4)	C4—C5—C8—O1	173.8 (4)
C14—C9—C10—C11	0.5 (6)	C11—C12—N2—O3	175.1 (5)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
C10—H10 \cdots O2 ⁱ	0.93	2.61	3.502 (6)	160
C1—H1B \cdots O2 ⁱⁱ	0.96	2.80	3.143 (5)	102

Symmetry codes: (i) $-x+1/2, -y, z-1/2$; (ii) $x-1/2, -y+1/2, -z$.