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2-Amino-6-methylpyridinium 4-hydroxybenzoate

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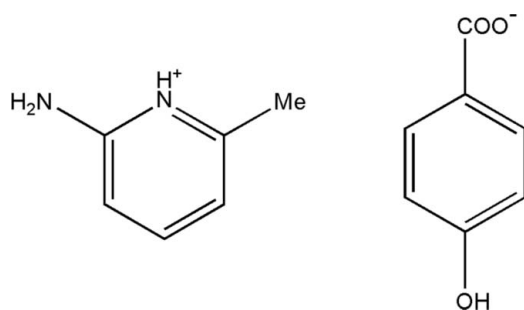
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.041; wR factor = 0.122; data-to-parameter ratio = 18.4.

In the title molecular salt, $\text{C}_6\text{H}_9\text{N}_2^+\cdot\text{C}_7\text{H}_5\text{O}_3^-$, the dihedral angle between the benzene ring and the CO_2 group in the anion is $6.1(2)^\circ$. In the crystal, the cation and anion are linked by $\text{N}-\text{H}\cdots\text{O}$ and $\text{C}-\text{H}\cdots\text{O}$ hydrogen bonds, and the anions are connected by $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds, forming a three-dimensional network.

Related literature

For general background to methylpyridinium derivatives, see: Blessing (1986); Brahadeeswaran *et al.* (2006); Brown (1976); Kvenvolden *et al.* (1971); Tomaru *et al.* (1991).



Experimental

Crystal data

$\text{C}_6\text{H}_9\text{N}_2^+\cdot\text{C}_7\text{H}_5\text{O}_3^-$
 $M_r = 246.26$
Monoclinic, $P2_1/c$
 $a = 11.9488(3)$ Å
 $b = 9.2952(3)$ Å

$c = 12.4067(3)$ Å
 $\beta = 117.116(2)^\circ$
 $V = 1226.51(6)$ Å³
 $Z = 4$
Mo $K\alpha$ radiation

$\mu = 0.10$ mm⁻¹
 $T = 293$ K

0.20 × 0.18 × 0.17 mm

Data collection

Bruker SMART APEXII CCD diffractometer
Absorption correction: multi-scan (SADABS; Bruker, 2008)
 $T_{\min} = 0.981$, $T_{\max} = 0.984$

11403 measured reflections
3084 independent reflections
2471 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.023$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.041$
 $wR(F^2) = 0.122$
 $S = 1.04$
3084 reflections
168 parameters

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 0.25$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.20$ e Å⁻³

Table 1
Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{N1}-\text{H1}\cdots\text{O1}^{\text{i}}$	0.86	2.00	2.8499 (13)	169
$\text{N2}-\text{H2A}\cdots\text{O2}^{\text{j}}$	0.86	1.94	2.7879 (14)	168
$\text{N2}-\text{H2B}\cdots\text{O1}^{\text{ii}}$	0.86	2.18	2.9902 (14)	157
$\text{O3}-\text{H3A}\cdots\text{O2}^{\text{iii}}$	0.97 (2)	1.67 (2)	2.6281 (14)	168.5 (19)
$\text{C4}-\text{H4}\cdots\text{O3}^{\text{iv}}$	0.93	2.51	3.4134 (17)	163

Symmetry codes: (i) $-x + 1, -y + 2, -z + 1$; (ii) $-x + 1, y - \frac{1}{2}, -z + \frac{1}{2}$; (iii) $-x + 2, y - \frac{1}{2}, -z + \frac{3}{2}$; (iv) $-x + 2, -y + 1, -z + 1$.

Data collection: APEX2 (Bruker, 2008); cell refinement: SAINT (Bruker, 2008); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 2012) and DIAMOND (Brandenburg, 1998); software used to prepare material for publication: SHELXL97 and PLATON (Spek, 2009).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: LX2273).

References

- Blessing, R. H. (1986). *Acta Cryst.* **B42**, 613–621.
Brahadeeswaran, S., Onduka, S., Takagi, M., Takahashi, Y., Adachi, H., Yoshimura, M., Mori, Y. & Sasaki, T. (2006). *J. Cryst. Growth*, **292**, 441–444.
Brandenburg, K. (1998). *DIAMOND*. Crystal Impact GbR, Bonn, Germany.
Brown, I. D. (1976). *Acta Cryst.* **A32**, 24–31.
Bruker (2008). APEX2, SADABS and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.
Farrugia, L. J. (2012). *J. Appl. Cryst.* **45**, 849–854.
Kvenvolden, K. A., Lawless, J. G. & Ponnampuram, C. (1971). *Proc. Natl Acad. Sci. USA*, **68**, 486–490.
Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.
Tomaru, S., Matsumoto, S., Kurihara, T., Suzuki, H., Oobara, N. & Kaino, T. (1991). *Appl. Phys. Lett.* **58**, 2583–2585.

supporting information

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2-Amino-6-methylpyridinium 4-hydroxybenzoate

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S1. Comment

Pyridine heterocycles and their derivatives are present in many large molecules having photo chemical, electro chemical and catalytic applications. Pyridine derivatives possess nonlinear optical (NLO) properties (Tomaru *et al.*, 1991). 4-*N,N*-dimethylamino-4'-*N'*-methyl stilbazolium tosylate (DAST) is used in generating and detecting terahertz (THz) frequencies (Brahadeeswaran *et al.*, 2006). Carboxylic acids are believed to have existed in the prebiotic earth (Kvenvolden *et al.*, 1971) and form aggregation patterns. An attempt is made to solve the pyridine based crystal structures to explore the NLO behaviour.

The crystal structure of the title compound (Fig. 1) consists of aminomethylpyridinium cation and hydroxybenzoate anion connected *via* N—H \cdots O & C—H \cdots O hydrogen bonds (Blessing, 1986; Brown, 1976). The pyridinium ring is essentially planar, with a maximum deviation of -0.005 (1) Å for atom N1. The dihedral angle between the pyridinium ring in the cation and the benzene ring in the anion is 78.32 (7)°.

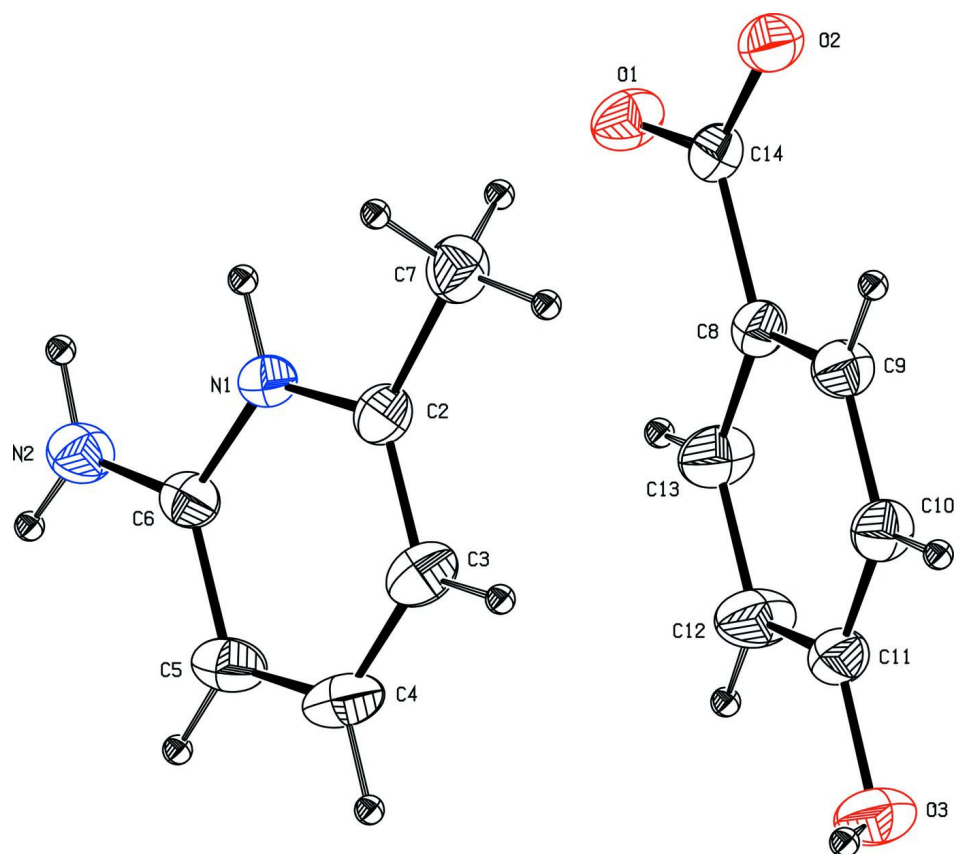
In the crystal structure (Fig. 2), the cation and anion are linked by N—H \cdots O and C—H \cdots O hydrogen bonds (Table 1), and the anions are connected by O—H \cdots O hydrogen bonds (Table 1), forming a three-dimensional network.

S2. Experimental

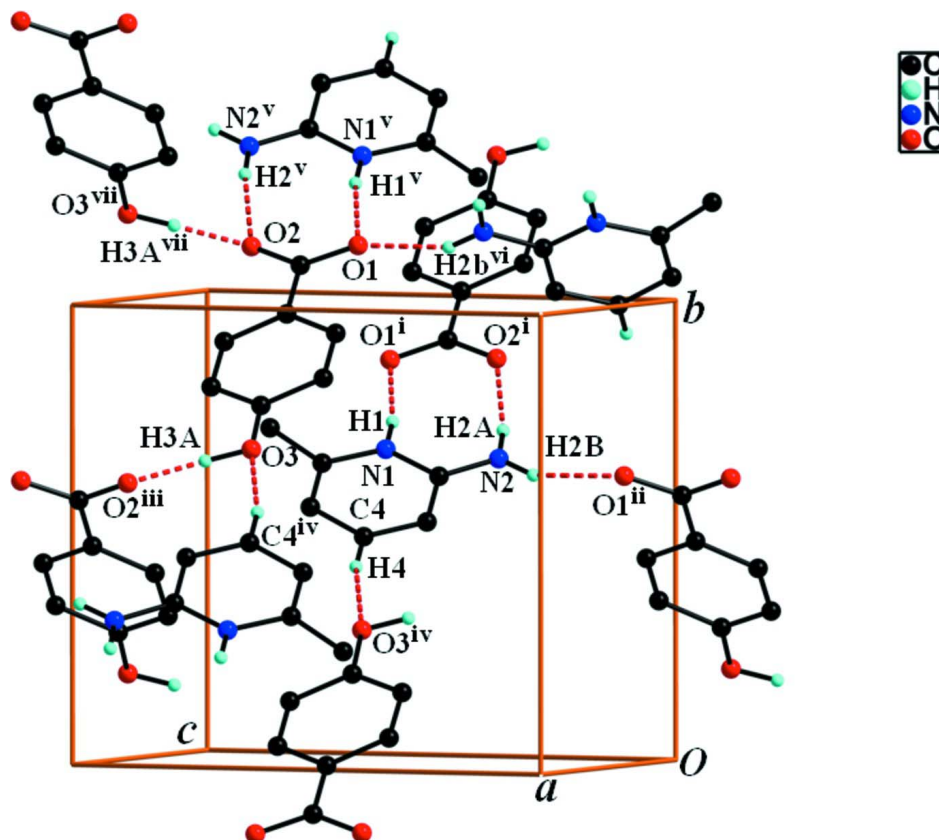
Methanol solutions of 2-amino-6-methylpyridine (54 mg) and 4-hydroxybenzoic acid (69 mg) were mixed together and stirred for about 1 h to get a homogeneous mixture. The resulting solution was allowed to evaporate at 303 K slowly in a water bath which has a temperature accuracy of $\pm 0.01^\circ$ C at ambient atmosphere. Brownish crystals with developed morphology of title compound were obtained after 15 days.

S3. Refinement

Especially about N—H & O—H according to RES. H atoms were positioned geometrically (N—H = 0.85–0.90 Å, O—H = 0.95–0.97 and C—H = 0.93–0.98 Å) and allowed to ride on their parent atoms, with $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C})$ for methyl H $1.2U_{\text{eq}}(\text{C})$ for other H atoms.

**Figure 1**

The molecular structure of the title compound with the atom numbering scheme. Displacement ellipsoids are drawn at the 30% probability level. H atoms are presented as small spheres of arbitrary radius.

**Figure 2**

A view of the N—H...O, C—H...O and O—H...O hydrogen bonds (dotted lines) in the crystal structure of the title compound. H atoms not participating in hydrogen-bonding were omitted for clarity. [Symmetry code: (i) $-x + 1, -y + 2, -z + 1$; (ii) $-x + 1, y - 1/2, -z + 1/2$; (iii) $-x + 2, y - 1/2, -z + 3/2$; (iv) $-x + 2, -y + 1, -z + 1$; (v) $-x + 1, -y + 2, -z + 1$; (vi) $-x + 1, y + 1/2, -z + 1/2$; (vii) $-x + 2, y + 1/2, -z + 3/2$.]

2-Amino-6-methylpyridinium 4-hydroxybenzoate

Crystal data

$C_6H_9N_2^+ \cdot C_7H_5O_3^-$

$M_r = 246.26$

Monoclinic, $P2_1/c$

Hall symbol: $-P 2_1/bc$

$a = 11.9488(3) \text{ \AA}$

$b = 9.2952(3) \text{ \AA}$

$c = 12.4067(3) \text{ \AA}$

$\beta = 117.116(2)^\circ$

$V = 1226.51(6) \text{ \AA}^3$

$Z = 4$

$F(000) = 520$

$D_x = 1.334 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 2671 reflections

$\theta = 1.9\text{--}28.4^\circ$

$\mu = 0.10 \text{ mm}^{-1}$

$T = 293 \text{ K}$

Block, white crystalline

$0.20 \times 0.18 \times 0.17 \text{ mm}$

Data collection

Bruker SMART APEXII CCD
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

ω and φ scans

Absorption correction: multi-scan
(*SADABS*; Bruker, 2008)

$T_{\min} = 0.981, T_{\max} = 0.984$

11403 measured reflections

3084 independent reflections

2471 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.023$
 $\theta_{\text{max}} = 28.4^\circ$, $\theta_{\text{min}} = 1.9^\circ$

$h = -16 \rightarrow 15$
 $k = -10 \rightarrow 12$
 $l = -16 \rightarrow 15$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.041$
 $wR(F^2) = 0.122$
 $S = 1.04$
 3084 reflections
 168 parameters
 0 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: inferred from
 neighbouring sites
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.0582P)^2 + 0.2532P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\text{max}} < 0.001$
 $\Delta\rho_{\text{max}} = 0.25 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\text{min}} = -0.20 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2 , conventional R-factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F , and R-factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C2	0.63520 (11)	0.65353 (14)	0.57200 (11)	0.0399 (3)
C3	0.72343 (13)	0.56067 (17)	0.57177 (14)	0.0528 (3)
H3	0.7947	0.5393	0.6434	0.063*
C4	0.70557 (14)	0.49803 (16)	0.46291 (15)	0.0562 (4)
H4	0.7652	0.4342	0.4623	0.067*
C5	0.60208 (14)	0.52912 (14)	0.35780 (13)	0.0486 (3)
H5	0.5912	0.4868	0.2857	0.058*
C6	0.51139 (12)	0.62547 (13)	0.35806 (11)	0.0384 (3)
C7	0.64299 (12)	0.72809 (17)	0.68095 (12)	0.0490 (3)
H7A	0.5776	0.6934	0.6988	0.074*
H7B	0.6331	0.8298	0.6660	0.074*
H7C	0.7234	0.7094	0.7485	0.074*
C8	0.81847 (10)	0.98339 (13)	0.59471 (10)	0.0370 (3)
C9	0.90757 (11)	0.92379 (15)	0.70239 (11)	0.0415 (3)
H9	0.9070	0.9489	0.7747	0.050*
C10	0.99711 (12)	0.82801 (15)	0.70462 (11)	0.0433 (3)
H10	1.0556	0.7891	0.7778	0.052*
C11	0.99954 (11)	0.79001 (15)	0.59772 (11)	0.0435 (3)
C12	0.91149 (14)	0.84879 (19)	0.48968 (12)	0.0592 (4)
H12	0.9124	0.8242	0.4175	0.071*
C13	0.82247 (13)	0.94363 (18)	0.48875 (12)	0.0542 (4)

H13	0.7637	0.9819	0.4154	0.065*
C14	0.72117 (10)	1.08629 (13)	0.59372 (10)	0.0369 (3)
N1	0.53173 (9)	0.68267 (11)	0.46589 (8)	0.0359 (2)
H1	0.4763	0.7405	0.4674	0.043*
N2	0.40790 (11)	0.66254 (13)	0.26020 (9)	0.0482 (3)
H2A	0.3555	0.7216	0.2661	0.058*
H2B	0.3930	0.6276	0.1908	0.058*
O1	0.63458 (8)	1.12882 (11)	0.49535 (8)	0.0503 (3)
O2	0.73063 (8)	1.12516 (11)	0.69528 (8)	0.0474 (2)
O3	1.08494 (10)	0.69729 (13)	0.59355 (10)	0.0598 (3)
H3A	1.145 (2)	0.672 (2)	0.675 (2)	0.089 (6)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C2	0.0376 (6)	0.0431 (7)	0.0414 (6)	−0.0007 (5)	0.0200 (5)	0.0035 (5)
C3	0.0437 (7)	0.0584 (9)	0.0572 (8)	0.0125 (6)	0.0239 (6)	0.0110 (7)
C4	0.0607 (8)	0.0492 (8)	0.0751 (10)	0.0178 (7)	0.0453 (8)	0.0095 (7)
C5	0.0657 (8)	0.0392 (7)	0.0571 (8)	0.0039 (6)	0.0422 (7)	−0.0004 (6)
C6	0.0463 (6)	0.0351 (6)	0.0416 (6)	−0.0035 (5)	0.0269 (5)	−0.0008 (5)
C7	0.0444 (7)	0.0597 (8)	0.0395 (6)	−0.0010 (6)	0.0162 (5)	−0.0015 (6)
C8	0.0327 (5)	0.0395 (6)	0.0367 (6)	0.0008 (5)	0.0140 (5)	−0.0002 (5)
C9	0.0417 (6)	0.0474 (7)	0.0345 (6)	0.0047 (5)	0.0165 (5)	0.0002 (5)
C10	0.0390 (6)	0.0483 (7)	0.0358 (6)	0.0072 (5)	0.0110 (5)	0.0041 (5)
C11	0.0372 (6)	0.0472 (7)	0.0430 (6)	0.0062 (5)	0.0156 (5)	−0.0027 (5)
C12	0.0563 (8)	0.0815 (11)	0.0350 (6)	0.0240 (8)	0.0167 (6)	−0.0031 (7)
C13	0.0485 (7)	0.0720 (10)	0.0339 (6)	0.0205 (7)	0.0117 (5)	0.0031 (6)
C14	0.0328 (5)	0.0392 (6)	0.0391 (6)	−0.0015 (5)	0.0168 (5)	0.0015 (5)
N1	0.0373 (5)	0.0366 (5)	0.0379 (5)	0.0027 (4)	0.0208 (4)	0.0000 (4)
N2	0.0516 (6)	0.0566 (7)	0.0380 (5)	0.0035 (5)	0.0217 (5)	−0.0052 (5)
O1	0.0431 (5)	0.0630 (6)	0.0413 (5)	0.0166 (4)	0.0162 (4)	0.0076 (4)
O2	0.0415 (5)	0.0576 (6)	0.0425 (5)	0.0060 (4)	0.0188 (4)	−0.0047 (4)
O3	0.0539 (6)	0.0729 (7)	0.0498 (6)	0.0261 (5)	0.0212 (5)	0.0006 (5)

Geometric parameters (Å, °)

C2—N1	1.3606 (15)	C8—C14	1.5013 (16)
C2—C3	1.3634 (18)	C9—C10	1.3825 (17)
C2—C7	1.4837 (18)	C9—H9	0.9300
C3—C4	1.395 (2)	C10—C11	1.3855 (17)
C3—H3	0.9300	C10—H10	0.9300
C4—C5	1.358 (2)	C11—O3	1.3546 (15)
C4—H4	0.9300	C11—C12	1.3844 (18)
C5—C6	1.4070 (17)	C12—C13	1.3775 (19)
C5—H5	0.9300	C12—H12	0.9300
C6—N2	1.3241 (16)	C13—H13	0.9300
C6—N1	1.3547 (15)	C14—O1	1.2511 (14)
C7—H7A	0.9600	C14—O2	1.2653 (14)

C7—H7B	0.9600	N1—H1	0.8600
C7—H7C	0.9600	N2—H2A	0.8600
C8—C13	1.3873 (17)	N2—H2B	0.8600
C8—C9	1.3887 (16)	O3—H3A	0.97 (2)
N1—C2—C3	119.11 (12)	C10—C9—H9	119.2
N1—C2—C7	116.13 (11)	C8—C9—H9	119.2
C3—C2—C7	124.75 (12)	C9—C10—C11	119.92 (11)
C2—C3—C4	119.27 (13)	C9—C10—H10	120.0
C2—C3—H3	120.4	C11—C10—H10	120.0
C4—C3—H3	120.4	O3—C11—C12	117.87 (12)
C5—C4—C3	120.77 (12)	O3—C11—C10	122.92 (11)
C5—C4—H4	119.6	C12—C11—C10	119.21 (12)
C3—C4—H4	119.6	C13—C12—C11	120.25 (12)
C4—C5—C6	119.91 (12)	C13—C12—H12	119.9
C4—C5—H5	120.0	C11—C12—H12	119.9
C6—C5—H5	120.0	C12—C13—C8	121.50 (12)
N2—C6—N1	118.40 (11)	C12—C13—H13	119.2
N2—C6—C5	124.19 (12)	C8—C13—H13	119.2
N1—C6—C5	117.41 (12)	O1—C14—O2	122.68 (11)
C2—C7—H7A	109.5	O1—C14—C8	120.15 (11)
C2—C7—H7B	109.5	O2—C14—C8	117.16 (10)
H7A—C7—H7B	109.5	C6—N1—C2	123.52 (10)
C2—C7—H7C	109.5	C6—N1—H1	118.2
H7A—C7—H7C	109.5	C2—N1—H1	118.2
H7B—C7—H7C	109.5	C6—N2—H2A	120.0
C13—C8—C9	117.58 (11)	C6—N2—H2B	120.0
C13—C8—C14	121.54 (11)	H2A—N2—H2B	120.0
C9—C8—C14	120.88 (11)	C11—O3—H3A	109.1 (12)
C10—C9—C8	121.54 (11)		
N1—C2—C3—C4	0.1 (2)	C10—C11—C12—C13	-0.1 (3)
C7—C2—C3—C4	-179.57 (13)	C11—C12—C13—C8	0.3 (3)
C2—C3—C4—C5	0.3 (2)	C9—C8—C13—C12	-0.1 (2)
C3—C4—C5—C6	0.0 (2)	C14—C8—C13—C12	-179.76 (14)
C4—C5—C6—N2	179.79 (13)	C13—C8—C14—O1	6.19 (19)
C4—C5—C6—N1	-0.64 (19)	C9—C8—C14—O1	-173.41 (12)
C13—C8—C9—C10	-0.1 (2)	C13—C8—C14—O2	-174.60 (12)
C14—C8—C9—C10	179.48 (12)	C9—C8—C14—O2	5.80 (18)
C8—C9—C10—C11	0.3 (2)	N2—C6—N1—C2	-179.31 (11)
C9—C10—C11—O3	179.75 (13)	C5—C6—N1—C2	1.10 (17)
C9—C10—C11—C12	-0.1 (2)	C3—C2—N1—C6	-0.82 (18)
O3—C11—C12—C13	179.97 (15)	C7—C2—N1—C6	178.85 (11)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1...O1 ⁱ	0.86	2.00	2.8499 (13)	169

N2—H2A···O2 ⁱ	0.86	1.94	2.7879 (14)	168
N2—H2B···O1 ⁱⁱ	0.86	2.18	2.9902 (14)	157
O3—H3A···O2 ⁱⁱⁱ	0.97 (2)	1.67 (2)	2.6281 (14)	168.5 (19)
C4—H4···O3 ^{iv}	0.93	2.51	3.4134 (17)	163

Symmetry codes: (i) $-x+1, -y+2, -z+1$; (ii) $-x+1, y-1/2, -z+1/2$; (iii) $-x+2, y-1/2, -z+3/2$; (iv) $-x+2, -y+1, -z+1$.