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Trimethylammonium 2,6-dioxo-5-(2,4,6-trinitrophenyl)-1,2,3,6-tetrahydropyrimidin-4-olate

 D. Kalaivani,^{a*} M. Buvaneswari^a and S. Rajeswari^b

^aPG and Research Department of Chemistry, Seethalakshmi Ramaswami College, Tiruchirappalli 620 002, Tamil Nadu, India, and ^bDepartment of Chemistry, Faculty of Engineering and Technology, SRM University, Kattankulathur 603 203, Tamil Nadu, India

Correspondence e-mail: kalaivalaj@yahoo.co.in

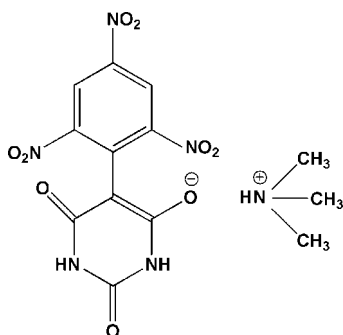
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Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.046; wR factor = 0.123; data-to-parameter ratio = 12.3.

In the title barbiturate salt (trivial name: trimethylammonium 2,4,6-trinitrophenylbarbiturate), $\text{C}_3\text{H}_{10}\text{N}^+\cdot\text{C}_{10}\text{H}_4\text{N}_5\text{O}_9^-$, the asymmetric unit contains two sets of anion–cation moieties. The dihedral angle between the rings in the anions are 44.0 (3) and 45.7 (3)°. Adjacent anions are connected into ribbons along [100] through $R_2^2(8)$ ring motifs formed by $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds involving the barbiturate rings. Attached to both sides of these ribbons *via* $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds are the trimethylammonium cations. $\text{C}-\text{H}\cdots\text{O}$ hydrogen bonds are also observed.

Related literature

For the biological activity of barbiturates, see: Nogrady (1988); Gitto *et al.* (2006). For side effects of barbiturates, see: Hardman & Limbird (2001); Rana *et al.* (2007). For barbiturates related to the title compound, see: Kalaivani *et al.* (2008); Kalaivani & Malarvizhi (2009).



Experimental

Crystal data

$\text{C}_3\text{H}_{10}\text{N}^+\cdot\text{C}_{10}\text{H}_4\text{N}_5\text{O}_9^-$
 $M_r = 398.30$
 Monoclinic, $P2_1/c$
 $a = 11.9828$ (12) Å
 $b = 30.802$ (3) Å
 $c = 9.5516$ (11) Å
 $\beta = 105.895$ (6)°

$V = 3390.6$ (6) Å³
 $Z = 8$
 Mo $K\alpha$ radiation
 $\mu = 0.13$ mm⁻¹
 $T = 296$ K
 $0.30 \times 0.20 \times 0.20$ mm

Data collection

Bruker Kappa APEX II CCD diffractometer
 Absorption correction: multi-scan (*SADABS*; Bruker, 2004)
 $T_{\min} = 0.90$, $T_{\max} = 0.97$

31796 measured reflections
 6580 independent reflections
 4977 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.030$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.046$
 $wR(F^2) = 0.123$
 $S = 1.02$
 6580 reflections
 536 parameters

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 0.34$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.30$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

<i>D</i> — <i>H</i> ··· <i>A</i>	<i>D</i> — <i>H</i>	<i>H</i> ··· <i>A</i>	<i>D</i> ··· <i>A</i>	<i>D</i> — <i>H</i> ··· <i>A</i>
N4—H4···O17 ⁱ	0.85 (2)	1.94 (2)	2.780 (2)	170 (2)
N5—H5···O16	0.86 (3)	1.93 (3)	2.778 (2)	170 (2)
N9—H9···O8	0.85 (2)	1.95 (2)	2.790 (2)	172 (2)
N10—H10···O7 ⁱⁱ	0.83 (2)	1.96 (2)	2.786 (2)	173 (2)
N11—H11A···O9 ⁱⁱⁱ	0.94 (3)	1.82 (3)	2.730 (2)	160 (3)
N12—H12A···O18	0.94 (3)	2.02 (3)	2.803 (3)	139 (3)
N12—H12A···O5 ^{iv}	0.94 (3)	2.16 (3)	2.856 (3)	130 (2)
C2—H2···O4 ^v	0.93	2.58	3.265 (3)	131
C16—H16···O2 ^{vi}	0.93	2.27	3.146 (3)	156
C24—H24B···O5	0.96	2.51	3.448 (4)	164
C24—H24C···O10 ⁱⁱⁱ	0.96	2.57	3.457 (4)	154
C26—H26B···O18 ^{iv}	0.96	2.50	3.400 (5)	156

Symmetry codes: (i) $x - 1, y, z$; (ii) $x + 1, y, z$; (iii) $x, -y + \frac{1}{2}, z + \frac{1}{2}$; (iv) $-x + 1, -y, -z + 1$; (v) $-x, -y, -z + 2$; (vi) $-x + 1, y + \frac{1}{2}, -z + \frac{3}{2}$.

Data collection: *APEX2* (Bruker, 2004); cell refinement: *SAINT-Plus* (Bruker, 2004); data reduction: *SAINT-Plus*; program(s) used to solve structure: *SIR92* (Altomare *et al.*, 1993); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3* (Farrugia, 1997) and *Mercury* (Macrae *et al.*, 2008); software used to prepare material for publication: *SHELXL97*.

The authors are thankful to the SAIF, IIT Madras, for the X-ray data collection.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: QK2022).

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supporting information

Acta Cryst. (2012). E68, o29–o30 [doi:10.1107/S1600536811051610]

Trimethylammonium 2,6-dioxo-5-(2,4,6-trinitrophenyl)-1,2,3,6-tetrahydro-pyrimidin-4-olate

D. Kalaivani, M. Buvaneswari and S. Rajeswari

S1. Comment

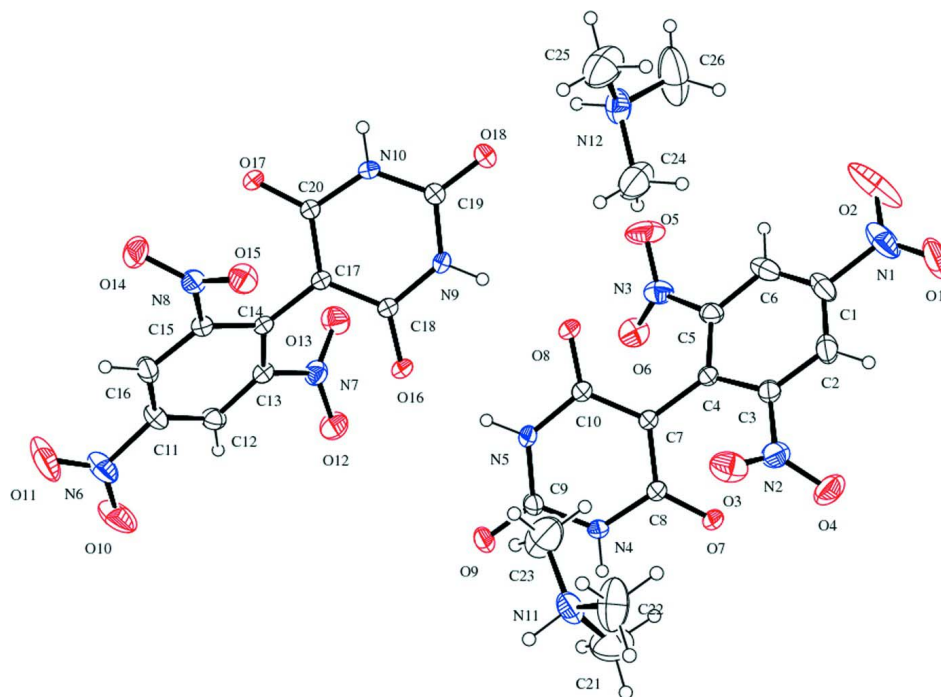
Barbiturates are biologically active molecules (Nogrady, 1988; Gitto *et al.*, 2006). Many reported barbiturates have noticeable side effects (Hardman & Limbird, 2001; Rana *et al.*, 2007). The growing interest in preparing new barbiturates prompted the present work. A barbiturate previously synthesized in our laboratory from 1-chloro-2,4-dinitrobenzene, barbituric acid and triethylamine has only one anion-cation pair in the asymmetric unit (Kalaivani *et al.*, 2008; Kalaivani & Malarvizhi, 2009). However, the asymmetric unit of the title barbiturate (Scheme), prepared from 1-chloro-2,4,6-trinitrobenzene, barbituric acid and trimethylamine, comprises two sets of anion-cation moieties (Fig. 1). Two sets of $R_2^2(8)$ hydrogen bond motifs are noticed between the adjacent anions. The hydrogen bonds N9—H9 \cdots O8 and N5—H5 \cdots O16 form the first set of $R_2^2(8)$ motif and the hydrogen bonds N10—H10 \cdots O7 and N4—H4 \cdots O17 form the second set of $R_2^2(8)$ motif. The aggregations of these motifs form an infinite one dimensional anionic ribbon extending along [100] direction. To either side of the anionic ribbon the trimethylammonium cations are hydrogen bonded through N11—H11A \cdots O9 and N12—H12A \cdots O18 hydrogen bonds in the terminal positions as depicted in Fig. 2. Furthermore a number of weak C—H \cdots O interactions (cf. Table 1) contribute to the three-dimensional coherence of the crystal structure.

S2. Experimental

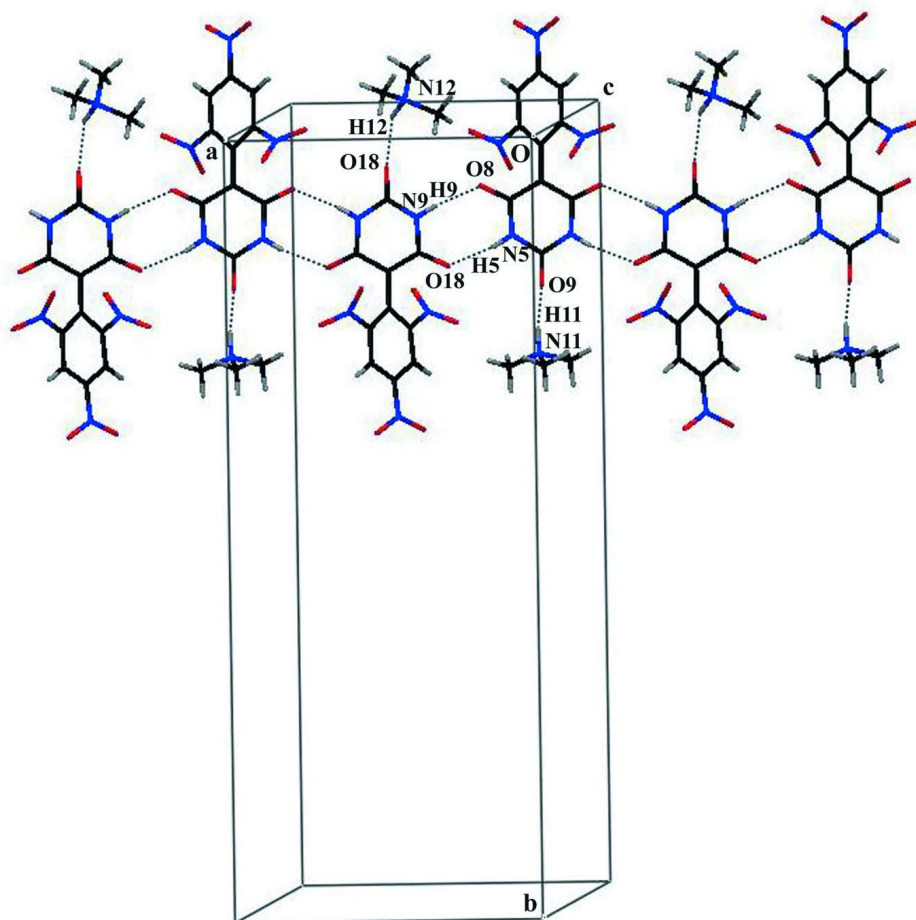
1-Chloro-2,4,6-trinitrobenzene (2.5 g, 0.01 mol) in 20 ml of absolute ethanol was mixed with barbituric acid (1.3 g, 0.01 mol) in 30 ml of absolute ethanol and heated to 40°C. Trimethylamine (3 ml ~0.05 mol) was then added and the mixture was shaken well for 2 hrs. Dark maroon red crystals obtained were filtered. The filtered crystals were powdered well and washed with 50 ml of dry ether and recrystallized from absolute ethanol (yield 80%; m.p. 245 °C). Good quality crystals, suitable for single-crystal X-ray studies were obtained by slow evaporation of the ethanolic solution of title compound at room temperature.

S3. Refinement

The six N-bound H atoms were located in a difference map and were freely refined without constraints. The C-bound hydrogen atoms were placed in calculated positions (C—H = 0.93–0.96 Å) and refined as riding atoms with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ for non-methyl and $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{C})$ for methyl H.

**Figure 1**

ORTEP diagram of (I) showing 30% probability displacement ellipsoids and the atomic numbering scheme.

**Figure 2**

View of a hydrogen bonded ribbon of barbiturate anions with the attached trimethylammonium cations.

Trimethylammonium 2,6-dioxo-5-(2,4,6-trinitrophenyl)-1,2,3,6-tetrahydropyrimidin-4-olate

Crystal data

$C_3H_{10}N^+ \cdot C_{10}H_4N_3O_9^-$

$M_r = 398.30$

Monoclinic, $P2_1/c$

Hall symbol: $-P\ 2_1/c$

$a = 11.9828\ (12)\ \text{\AA}$

$b = 30.802\ (3)\ \text{\AA}$

$c = 9.5516\ (11)\ \text{\AA}$

$\beta = 105.895\ (6)^\circ$

$V = 3390.6\ (6)\ \text{\AA}^3$

$Z = 8$

$F(000) = 1648$

$D_x = 1.561\ \text{Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 8488 reflections

$\theta = 2.2\text{--}29.2^\circ$

$\mu = 0.13\ \text{mm}^{-1}$

$T = 296\ \text{K}$

Block, brown

$0.30 \times 0.20 \times 0.20\ \text{mm}$

Data collection

Bruker Kappa APEX II CCD
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

ω and φ scan

Absorption correction: multi-scan
(*SADABS*; Bruker, 2004)

$T_{\min} = 0.90$, $T_{\max} = 0.97$

31796 measured reflections

6580 independent reflections

4977 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.030$
 $\theta_{\text{max}} = 26.0^\circ$, $\theta_{\text{min}} = 2.2^\circ$
 $h = -14 \rightarrow 14$

$k = -37 \rightarrow 37$
 $l = -11 \rightarrow 11$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.046$
 $wR(F^2) = 0.123$
 $S = 1.02$
 6580 reflections
 536 parameters
 0 restraints
 Primary atom site location: structure-invariant
 direct methods
 Secondary atom site location: difference Fourier
 map

Hydrogen site location: inferred from
 neighbouring sites
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.0538P)^2 + 1.671P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\text{max}} < 0.001$
 $\Delta\rho_{\text{max}} = 0.34 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\text{min}} = -0.30 \text{ e } \text{\AA}^{-3}$
 Extinction correction: *SHELXL97* (Sheldrick,
 2008), $F_c^* = kF_c[1 + 0.001x F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$
 Extinction coefficient: 0.0021 (4)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2 , conventional R-factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F , and R-factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.1280 (2)	-0.04978 (7)	0.6975 (3)	0.0520 (6)
C2	0.07387 (19)	-0.02594 (7)	0.7794 (3)	0.0507 (6)
H2	0.0396	-0.0393	0.8444	0.061*
C3	0.07113 (16)	0.01807 (6)	0.7636 (2)	0.0379 (5)
C4	0.11767 (14)	0.04061 (6)	0.6661 (2)	0.0301 (4)
C5	0.16732 (16)	0.01315 (6)	0.5837 (2)	0.0371 (4)
C6	0.17509 (19)	-0.03096 (7)	0.5989 (3)	0.0492 (6)
H6	0.2116	-0.0476	0.5431	0.059*
C7	0.11546 (15)	0.08708 (6)	0.6514 (2)	0.0294 (4)
C8	0.01360 (15)	0.11016 (6)	0.6479 (2)	0.0284 (4)
C9	0.10703 (16)	0.17708 (6)	0.6088 (2)	0.0330 (4)
C10	0.21401 (15)	0.10922 (6)	0.6360 (2)	0.0314 (4)
C11	0.57418 (19)	0.33262 (6)	0.5258 (2)	0.0418 (5)
C12	0.49039 (18)	0.30675 (6)	0.4401 (2)	0.0386 (5)
H12	0.4260	0.3186	0.3735	0.046*
C13	0.50449 (16)	0.26296 (6)	0.4557 (2)	0.0322 (4)
C14	0.59422 (15)	0.24254 (6)	0.5597 (2)	0.0288 (4)
C15	0.67477 (16)	0.27170 (6)	0.6415 (2)	0.0321 (4)
C16	0.66866 (19)	0.31572 (6)	0.6246 (2)	0.0417 (5)
H16	0.7272	0.3336	0.6788	0.050*

C17	0.60067 (15)	0.19623 (6)	0.5822 (2)	0.0303 (4)
C18	0.50129 (15)	0.17370 (6)	0.5898 (2)	0.0314 (4)
C19	0.61320 (16)	0.10691 (6)	0.6373 (2)	0.0372 (5)
C20	0.70684 (16)	0.17435 (6)	0.6027 (2)	0.0326 (4)
C21	-0.0029 (3)	0.19000 (12)	1.0012 (4)	0.0901 (10)
H21A	-0.0174	0.1597	0.9804	0.135*
H21B	-0.0129	0.2057	0.9117	0.135*
H21C	-0.0563	0.2008	1.0515	0.135*
C22	0.1314 (3)	0.17431 (12)	1.2299 (4)	0.0930 (11)
H22A	0.1229	0.1435	1.2147	0.140*
H22B	0.0744	0.1845	1.2757	0.140*
H22C	0.2077	0.1805	1.2913	0.140*
C23	0.1971 (3)	0.18289 (12)	1.0133 (4)	0.0934 (11)
H23A	0.2747	0.1878	1.0727	0.140*
H23B	0.1837	0.1997	0.9255	0.140*
H23C	0.1871	0.1526	0.9892	0.140*
C24	0.4800 (3)	0.00605 (12)	0.8010 (4)	0.0884 (10)
H24A	0.4493	-0.0136	0.8593	0.133*
H24B	0.4214	0.0122	0.7123	0.133*
H24C	0.5033	0.0325	0.8538	0.133*
C25	0.6768 (3)	-0.01717 (14)	0.8925 (4)	0.0988 (12)
H25A	0.6566	-0.0353	0.9635	0.148*
H25B	0.6975	0.0112	0.9329	0.148*
H25C	0.7414	-0.0297	0.8656	0.148*
C26	0.5432 (5)	-0.05449 (12)	0.6919 (5)	0.137 (2)
H26A	0.6101	-0.0716	0.6914	0.205*
H26B	0.4992	-0.0487	0.5935	0.205*
H26C	0.4959	-0.0702	0.7411	0.205*
N1	0.1341 (2)	-0.09686 (8)	0.7141 (3)	0.0774 (8)
N2	0.02277 (18)	0.04092 (7)	0.8676 (2)	0.0520 (5)
N3	0.20994 (17)	0.02933 (6)	0.4651 (2)	0.0493 (5)
N4	0.01520 (13)	0.15431 (5)	0.62608 (18)	0.0327 (4)
N5	0.20410 (14)	0.15362 (5)	0.61741 (19)	0.0352 (4)
N6	0.5592 (2)	0.37934 (6)	0.5148 (2)	0.0626 (6)
N7	0.42027 (15)	0.23778 (6)	0.34613 (19)	0.0409 (4)
N8	0.76942 (14)	0.25697 (5)	0.7650 (2)	0.0396 (4)
N9	0.51370 (14)	0.12959 (5)	0.6177 (2)	0.0383 (4)
N10	0.70604 (14)	0.13030 (5)	0.6297 (2)	0.0380 (4)
N11	0.11542 (19)	0.19574 (6)	1.0921 (2)	0.0538 (5)
N12	0.5802 (2)	-0.01380 (7)	0.7670 (2)	0.0569 (5)
O1	0.0869 (2)	-0.11310 (7)	0.7939 (5)	0.1399 (14)
O2	0.1908 (3)	-0.11639 (7)	0.6489 (2)	0.1265 (13)
O3	0.08047 (18)	0.06930 (6)	0.9405 (2)	0.0659 (5)
O4	-0.06914 (17)	0.02796 (7)	0.8806 (2)	0.0788 (6)
O5	0.29880 (18)	0.01269 (7)	0.4523 (2)	0.0780 (6)
O6	0.15519 (15)	0.05664 (6)	0.38352 (18)	0.0575 (4)
O7	-0.07663 (10)	0.09318 (4)	0.65984 (16)	0.0365 (3)
O8	0.30776 (11)	0.09178 (4)	0.64047 (17)	0.0423 (4)

O9	0.10310 (12)	0.21575 (4)	0.58780 (19)	0.0506 (4)
O10	0.4647 (2)	0.39359 (6)	0.4567 (3)	0.0874 (7)
O11	0.6406 (2)	0.40236 (6)	0.5663 (3)	0.1089 (9)
O12	0.32075 (13)	0.25037 (6)	0.3120 (2)	0.0589 (5)
O13	0.45480 (14)	0.20692 (5)	0.29116 (18)	0.0547 (4)
O14	0.86373 (13)	0.27312 (5)	0.7787 (2)	0.0605 (5)
O15	0.74674 (14)	0.23106 (5)	0.84766 (17)	0.0493 (4)
O16	0.40506 (11)	0.19019 (4)	0.57728 (17)	0.0403 (3)
O17	0.79831 (11)	0.19122 (4)	0.59685 (18)	0.0438 (4)
O18	0.61782 (13)	0.06785 (5)	0.6586 (2)	0.0563 (5)
H4	-0.046 (2)	0.1686 (7)	0.619 (2)	0.037 (6)*
H5	0.263 (2)	0.1680 (8)	0.607 (3)	0.050 (7)*
H9	0.455 (2)	0.1158 (7)	0.627 (2)	0.042 (6)*
H10	0.769 (2)	0.1172 (7)	0.640 (2)	0.046 (6)*
H11A	0.123 (3)	0.2258 (11)	1.111 (3)	0.091 (10)*
H12A	0.605 (3)	0.0039 (10)	0.701 (3)	0.082 (9)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0435 (12)	0.0251 (11)	0.0709 (16)	-0.0001 (9)	-0.0121 (11)	0.0031 (10)
C2	0.0354 (11)	0.0391 (12)	0.0711 (16)	-0.0048 (9)	0.0038 (11)	0.0197 (11)
C3	0.0274 (9)	0.0334 (10)	0.0516 (12)	0.0023 (8)	0.0089 (9)	0.0086 (9)
C4	0.0186 (8)	0.0273 (9)	0.0423 (11)	-0.0002 (7)	0.0046 (7)	0.0010 (8)
C5	0.0285 (10)	0.0337 (10)	0.0459 (12)	0.0032 (8)	0.0049 (8)	-0.0036 (9)
C6	0.0440 (12)	0.0348 (12)	0.0580 (14)	0.0105 (9)	-0.0043 (10)	-0.0105 (10)
C7	0.0226 (9)	0.0255 (9)	0.0406 (11)	-0.0002 (7)	0.0095 (7)	0.0016 (8)
C8	0.0240 (9)	0.0277 (9)	0.0336 (10)	-0.0004 (7)	0.0081 (7)	0.0008 (7)
C9	0.0278 (9)	0.0263 (10)	0.0427 (11)	-0.0017 (7)	0.0059 (8)	-0.0004 (8)
C10	0.0245 (9)	0.0288 (10)	0.0408 (11)	-0.0003 (7)	0.0089 (8)	0.0009 (8)
C11	0.0509 (12)	0.0252 (10)	0.0488 (13)	0.0038 (9)	0.0130 (10)	0.0035 (9)
C12	0.0382 (11)	0.0354 (11)	0.0426 (12)	0.0077 (8)	0.0118 (9)	0.0050 (9)
C13	0.0265 (9)	0.0338 (10)	0.0384 (11)	-0.0005 (8)	0.0123 (8)	-0.0006 (8)
C14	0.0233 (9)	0.0283 (9)	0.0389 (10)	0.0007 (7)	0.0153 (8)	-0.0004 (8)
C15	0.0274 (9)	0.0295 (10)	0.0400 (11)	0.0016 (7)	0.0103 (8)	0.0011 (8)
C16	0.0427 (11)	0.0287 (10)	0.0512 (13)	-0.0056 (9)	0.0086 (10)	-0.0027 (9)
C17	0.0236 (9)	0.0265 (9)	0.0432 (11)	0.0007 (7)	0.0133 (8)	-0.0007 (8)
C18	0.0248 (9)	0.0280 (9)	0.0436 (11)	0.0008 (7)	0.0131 (8)	-0.0007 (8)
C19	0.0317 (10)	0.0266 (10)	0.0558 (13)	-0.0001 (8)	0.0163 (9)	-0.0003 (9)
C20	0.0269 (9)	0.0252 (9)	0.0493 (12)	-0.0001 (7)	0.0166 (8)	-0.0011 (8)
C21	0.077 (2)	0.102 (3)	0.090 (2)	0.0316 (19)	0.0213 (18)	0.024 (2)
C22	0.103 (3)	0.100 (3)	0.068 (2)	-0.043 (2)	0.0111 (18)	0.0178 (18)
C23	0.082 (2)	0.116 (3)	0.091 (2)	-0.029 (2)	0.0376 (19)	-0.039 (2)
C24	0.0619 (18)	0.112 (3)	0.096 (2)	-0.0006 (18)	0.0298 (17)	0.028 (2)
C25	0.0619 (19)	0.140 (3)	0.087 (2)	0.003 (2)	0.0083 (17)	0.040 (2)
C26	0.227 (6)	0.081 (3)	0.130 (4)	-0.076 (3)	0.098 (4)	-0.039 (2)
N1	0.0761 (17)	0.0366 (13)	0.0909 (19)	0.0052 (12)	-0.0255 (14)	0.0109 (13)
N2	0.0495 (12)	0.0541 (12)	0.0587 (13)	0.0159 (10)	0.0256 (10)	0.0248 (10)

N3	0.0479 (11)	0.0517 (11)	0.0497 (11)	0.0081 (9)	0.0159 (9)	-0.0082 (9)
N4	0.0220 (8)	0.0269 (8)	0.0493 (10)	0.0040 (6)	0.0099 (7)	0.0017 (7)
N5	0.0226 (8)	0.0275 (8)	0.0561 (11)	-0.0036 (6)	0.0118 (7)	0.0042 (7)
N6	0.0797 (16)	0.0325 (11)	0.0665 (14)	0.0076 (11)	0.0046 (12)	0.0024 (10)
N7	0.0344 (9)	0.0435 (10)	0.0429 (10)	-0.0040 (8)	0.0071 (8)	0.0002 (8)
N8	0.0327 (9)	0.0342 (9)	0.0491 (11)	0.0004 (7)	0.0064 (8)	-0.0017 (8)
N9	0.0260 (8)	0.0278 (9)	0.0655 (12)	-0.0035 (7)	0.0201 (8)	0.0033 (8)
N10	0.0246 (8)	0.0279 (9)	0.0656 (12)	0.0049 (7)	0.0190 (8)	0.0027 (8)
N11	0.0733 (14)	0.0327 (10)	0.0583 (13)	-0.0046 (9)	0.0231 (11)	-0.0020 (9)
N12	0.0711 (14)	0.0527 (12)	0.0529 (12)	-0.0156 (10)	0.0272 (11)	0.0002 (10)
O1	0.0936 (19)	0.0485 (14)	0.275 (4)	-0.0012 (12)	0.047 (2)	0.064 (2)
O2	0.250 (4)	0.0466 (12)	0.0588 (13)	0.0597 (18)	0.0027 (17)	0.0002 (10)
O3	0.0789 (13)	0.0684 (12)	0.0515 (11)	0.0184 (11)	0.0197 (9)	0.0022 (9)
O4	0.0603 (11)	0.0869 (14)	0.1081 (16)	0.0190 (10)	0.0548 (11)	0.0474 (12)
O5	0.0752 (13)	0.0907 (14)	0.0825 (14)	0.0325 (11)	0.0459 (11)	-0.0048 (11)
O6	0.0557 (10)	0.0671 (11)	0.0479 (10)	0.0038 (9)	0.0113 (8)	0.0038 (8)
O7	0.0227 (6)	0.0326 (7)	0.0565 (9)	0.0023 (5)	0.0148 (6)	0.0072 (6)
O8	0.0217 (7)	0.0344 (8)	0.0731 (10)	0.0015 (5)	0.0169 (6)	0.0040 (7)
O9	0.0389 (8)	0.0248 (7)	0.0857 (12)	-0.0013 (6)	0.0131 (8)	0.0057 (7)
O10	0.1006 (17)	0.0439 (11)	0.0996 (16)	0.0302 (11)	-0.0035 (13)	0.0065 (10)
O11	0.1089 (18)	0.0312 (10)	0.151 (2)	-0.0119 (11)	-0.0244 (16)	0.0059 (12)
O12	0.0320 (8)	0.0625 (11)	0.0721 (11)	0.0003 (7)	-0.0030 (8)	0.0016 (9)
O13	0.0582 (10)	0.0518 (10)	0.0515 (10)	-0.0026 (8)	0.0108 (8)	-0.0162 (8)
O14	0.0315 (8)	0.0539 (10)	0.0856 (13)	-0.0096 (7)	-0.0015 (8)	0.0053 (9)
O15	0.0519 (9)	0.0494 (9)	0.0450 (9)	0.0027 (7)	0.0102 (7)	0.0097 (7)
O16	0.0236 (7)	0.0332 (7)	0.0684 (10)	0.0024 (5)	0.0197 (6)	0.0052 (7)
O17	0.0264 (7)	0.0306 (7)	0.0810 (11)	0.0005 (6)	0.0261 (7)	0.0031 (7)
O18	0.0430 (9)	0.0255 (8)	0.1043 (14)	0.0015 (6)	0.0269 (9)	0.0071 (8)

Geometric parameters (Å, °)

C1—C6	1.354 (4)	C21—N11	1.456 (4)
C1—C2	1.360 (4)	C21—H21A	0.9600
C1—N1	1.458 (3)	C21—H21B	0.9600
C2—C3	1.363 (3)	C21—H21C	0.9600
C2—H2	0.9300	C22—N11	1.438 (4)
C3—C4	1.395 (3)	C22—H22A	0.9600
C3—N2	1.460 (3)	C22—H22B	0.9600
C4—C5	1.395 (3)	C22—H22C	0.9600
C4—C7	1.438 (3)	C23—N11	1.443 (4)
C5—C6	1.367 (3)	C23—H23A	0.9600
C5—N3	1.452 (3)	C23—H23B	0.9600
C6—H6	0.9300	C23—H23C	0.9600
C7—C8	1.405 (2)	C24—N12	1.461 (4)
C7—C10	1.405 (2)	C24—H24A	0.9600
C8—O7	1.234 (2)	C24—H24B	0.9600
C8—N4	1.377 (2)	C24—H24C	0.9600
C9—O9	1.207 (2)	C25—N12	1.424 (4)

C9—N4	1.352 (2)	C25—H25A	0.9600
C9—N5	1.352 (2)	C25—H25B	0.9600
C10—O8	1.235 (2)	C25—H25C	0.9600
C10—N5	1.380 (2)	C26—N12	1.452 (4)
C11—C16	1.363 (3)	C26—H26A	0.9600
C11—C12	1.365 (3)	C26—H26B	0.9600
C11—N6	1.451 (3)	C26—H26C	0.9600
C12—C13	1.362 (3)	N1—O1	1.178 (4)
C12—H12	0.9300	N1—O2	1.201 (4)
C13—C14	1.398 (3)	N2—O4	1.209 (3)
C13—N7	1.462 (3)	N2—O3	1.210 (3)
C14—C15	1.391 (3)	N3—O6	1.210 (2)
C14—C17	1.441 (2)	N3—O5	1.218 (2)
C15—C16	1.365 (3)	N4—H4	0.85 (2)
C15—N8	1.466 (3)	N5—H5	0.86 (3)
C16—H16	0.9300	N6—O11	1.198 (3)
C17—C18	1.398 (2)	N6—O10	1.199 (3)
C17—C20	1.405 (2)	N7—O12	1.211 (2)
C18—O16	1.235 (2)	N7—O13	1.212 (2)
C18—N9	1.384 (2)	N8—O15	1.205 (2)
C19—O18	1.219 (2)	N8—O14	1.208 (2)
C19—N10	1.344 (2)	N9—H9	0.85 (2)
C19—N9	1.350 (2)	N10—H10	0.83 (2)
C20—O17	1.228 (2)	N11—H11A	0.94 (3)
C20—N10	1.381 (2)	N12—H12A	0.94 (3)
C6—C1—C2	121.6 (2)	N11—C22—H22C	109.5
C6—C1—N1	118.9 (3)	H22A—C22—H22C	109.5
C2—C1—N1	119.5 (3)	H22B—C22—H22C	109.5
C1—C2—C3	118.3 (2)	N11—C23—H23A	109.5
C1—C2—H2	120.8	N11—C23—H23B	109.5
C3—C2—H2	120.8	H23A—C23—H23B	109.5
C2—C3—C4	124.5 (2)	N11—C23—H23C	109.5
C2—C3—N2	114.0 (2)	H23A—C23—H23C	109.5
C4—C3—N2	121.33 (18)	H23B—C23—H23C	109.5
C5—C4—C3	112.67 (17)	N12—C24—H24A	109.5
C5—C4—C7	123.22 (17)	N12—C24—H24B	109.5
C3—C4—C7	124.11 (17)	H24A—C24—H24B	109.5
C6—C5—C4	124.8 (2)	N12—C24—H24C	109.5
C6—C5—N3	113.30 (19)	H24A—C24—H24C	109.5
C4—C5—N3	121.83 (18)	H24B—C24—H24C	109.5
C1—C6—C5	118.0 (2)	N12—C25—H25A	109.5
C1—C6—H6	121.0	N12—C25—H25B	109.5
C5—C6—H6	121.0	H25A—C25—H25B	109.5
C8—C7—C10	120.10 (16)	N12—C25—H25C	109.5
C8—C7—C4	119.90 (15)	H25A—C25—H25C	109.5
C10—C7—C4	119.97 (16)	H25B—C25—H25C	109.5
O7—C8—N4	118.79 (16)	N12—C26—H26A	109.5

O7—C8—C7	124.11 (16)	N12—C26—H26B	109.5
N4—C8—C7	117.08 (15)	H26A—C26—H26B	109.5
O9—C9—N4	122.42 (17)	N12—C26—H26C	109.5
O9—C9—N5	122.07 (17)	H26A—C26—H26C	109.5
N4—C9—N5	115.51 (16)	H26B—C26—H26C	109.5
O8—C10—N5	118.72 (16)	O1—N1—O2	124.3 (3)
O8—C10—C7	124.59 (17)	O1—N1—C1	118.2 (3)
N5—C10—C7	116.68 (16)	O2—N1—C1	117.4 (3)
C16—C11—C12	121.82 (18)	O4—N2—O3	124.9 (2)
C16—C11—N6	119.5 (2)	O4—N2—C3	116.9 (2)
C12—C11—N6	118.7 (2)	O3—N2—C3	117.97 (19)
C13—C12—C11	117.63 (19)	O6—N3—O5	124.2 (2)
C13—C12—H12	121.2	O6—N3—C5	119.76 (17)
C11—C12—H12	121.2	O5—N3—C5	116.0 (2)
C12—C13—C14	124.84 (18)	C9—N4—C8	125.18 (16)
C12—C13—N7	113.99 (18)	C9—N4—H4	116.2 (14)
C14—C13—N7	121.06 (16)	C8—N4—H4	118.5 (14)
C15—C14—C13	112.90 (16)	C9—N5—C10	125.41 (16)
C15—C14—C17	123.66 (17)	C9—N5—H5	115.7 (16)
C13—C14—C17	123.42 (17)	C10—N5—H5	118.8 (16)
C16—C15—C14	124.62 (18)	O11—N6—O10	122.2 (2)
C16—C15—N8	114.02 (17)	O11—N6—C11	119.1 (2)
C14—C15—N8	121.13 (16)	O10—N6—C11	118.6 (2)
C11—C16—C15	117.95 (19)	O12—N7—O13	124.23 (18)
C11—C16—H16	121.0	O12—N7—C13	117.30 (17)
C15—C16—H16	121.0	O13—N7—C13	118.40 (17)
C18—C17—C20	120.41 (16)	O15—N8—O14	124.99 (18)
C18—C17—C14	119.19 (15)	O15—N8—C15	117.93 (16)
C20—C17—C14	120.34 (15)	O14—N8—C15	117.02 (17)
O16—C18—N9	118.05 (16)	C19—N9—C18	125.36 (16)
O16—C18—C17	125.23 (17)	C19—N9—H9	117.0 (15)
N9—C18—C17	116.70 (16)	C18—N9—H9	117.6 (15)
O18—C19—N10	122.54 (18)	C19—N10—C20	125.79 (16)
O18—C19—N9	122.14 (18)	C19—N10—H10	117.8 (16)
N10—C19—N9	115.32 (17)	C20—N10—H10	116.4 (16)
O17—C20—N10	118.43 (16)	C22—N11—C23	114.3 (3)
O17—C20—C17	125.15 (17)	C22—N11—C21	110.7 (2)
N10—C20—C17	116.42 (16)	C23—N11—C21	110.1 (2)
N11—C21—H21A	109.5	C22—N11—H11A	107 (2)
N11—C21—H21B	109.5	C23—N11—H11A	108.9 (19)
H21A—C21—H21B	109.5	C21—N11—H11A	105 (2)
N11—C21—H21C	109.5	C25—N12—C26	114.9 (3)
H21A—C21—H21C	109.5	C25—N12—C24	111.8 (3)
H21B—C21—H21C	109.5	C26—N12—C24	108.2 (3)
N11—C22—H22A	109.5	C25—N12—H12A	105.4 (18)
N11—C22—H22B	109.5	C26—N12—H12A	106.4 (19)
H22A—C22—H22B	109.5	C24—N12—H12A	109.8 (18)

C6—C1—C2—C3	-2.2 (3)	C20—C17—C18—O16	177.9 (2)
N1—C1—C2—C3	178.9 (2)	C14—C17—C18—O16	0.8 (3)
C1—C2—C3—C4	1.7 (3)	C20—C17—C18—N9	-0.2 (3)
C1—C2—C3—N2	-173.5 (2)	C14—C17—C18—N9	-177.35 (18)
C2—C3—C4—C5	0.6 (3)	C18—C17—C20—O17	179.4 (2)
N2—C3—C4—C5	175.41 (18)	C14—C17—C20—O17	-3.5 (3)
C2—C3—C4—C7	-179.3 (2)	C18—C17—C20—N10	0.2 (3)
N2—C3—C4—C7	-4.5 (3)	C14—C17—C20—N10	177.31 (18)
C3—C4—C5—C6	-2.6 (3)	C6—C1—N1—O1	-175.5 (3)
C7—C4—C5—C6	177.25 (19)	C2—C1—N1—O1	3.5 (4)
C3—C4—C5—N3	173.45 (18)	C6—C1—N1—O2	7.0 (4)
C7—C4—C5—N3	-6.7 (3)	C2—C1—N1—O2	-174.0 (3)
C2—C1—C6—C5	0.3 (3)	C2—C3—N2—O4	-49.6 (3)
N1—C1—C6—C5	179.3 (2)	C4—C3—N2—O4	135.1 (2)
C4—C5—C6—C1	2.3 (3)	C2—C3—N2—O3	125.8 (2)
N3—C5—C6—C1	-174.08 (19)	C4—C3—N2—O3	-49.6 (3)
C5—C4—C7—C8	136.11 (19)	C6—C5—N3—O6	135.8 (2)
C3—C4—C7—C8	-44.0 (3)	C4—C5—N3—O6	-40.7 (3)
C5—C4—C7—C10	-42.0 (3)	C6—C5—N3—O5	-41.6 (3)
C3—C4—C7—C10	137.9 (2)	C4—C5—N3—O5	141.9 (2)
C10—C7—C8—O7	179.62 (18)	O9—C9—N4—C8	178.88 (19)
C4—C7—C8—O7	1.5 (3)	N5—C9—N4—C8	-1.4 (3)
C10—C7—C8—N4	1.4 (3)	O7—C8—N4—C9	-178.87 (18)
C4—C7—C8—N4	-176.73 (17)	C7—C8—N4—C9	-0.5 (3)
C8—C7—C10—O8	178.54 (19)	O9—C9—N5—C10	-177.6 (2)
C4—C7—C10—O8	-3.4 (3)	N4—C9—N5—C10	2.7 (3)
C8—C7—C10—N5	-0.3 (3)	O8—C10—N5—C9	179.20 (19)
C4—C7—C10—N5	177.84 (17)	C7—C10—N5—C9	-1.9 (3)
C16—C11—C12—C13	-0.5 (3)	C16—C11—N6—O11	17.1 (4)
N6—C11—C12—C13	-177.93 (19)	C12—C11—N6—O11	-165.3 (3)
C11—C12—C13—C14	4.6 (3)	C16—C11—N6—O10	-161.4 (2)
C11—C12—C13—N7	-171.62 (18)	C12—C11—N6—O10	16.2 (3)
C12—C13—C14—C15	-4.4 (3)	C12—C13—N7—O12	-40.6 (2)
N7—C13—C14—C15	171.59 (16)	C14—C13—N7—O12	143.01 (19)
C12—C13—C14—C17	174.00 (18)	C12—C13—N7—O13	136.46 (19)
N7—C13—C14—C17	-10.0 (3)	C14—C13—N7—O13	-39.9 (3)
C13—C14—C15—C16	0.2 (3)	C16—C15—N8—O15	131.1 (2)
C17—C14—C15—C16	-178.20 (19)	C14—C15—N8—O15	-43.6 (3)
C13—C14—C15—N8	174.30 (16)	C16—C15—N8—O14	-46.3 (3)
C17—C14—C15—N8	-4.1 (3)	C14—C15—N8—O14	139.05 (19)
C12—C11—C16—C15	-3.3 (3)	O18—C19—N9—C18	-178.5 (2)
N6—C11—C16—C15	174.1 (2)	N10—C19—N9—C18	0.7 (3)
C14—C15—C16—C11	3.5 (3)	O16—C18—N9—C19	-178.6 (2)
N8—C15—C16—C11	-170.99 (18)	C17—C18—N9—C19	-0.3 (3)
C15—C14—C17—C18	132.5 (2)	O18—C19—N10—C20	178.5 (2)
C13—C14—C17—C18	-45.7 (3)	N9—C19—N10—C20	-0.7 (3)
C15—C14—C17—C20	-44.6 (3)	O17—C20—N10—C19	-178.9 (2)
C13—C14—C17—C20	137.12 (19)	C17—C20—N10—C19	0.3 (3)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
N4—H4 \cdots O17 ⁱ	0.85 (2)	1.94 (2)	2.780 (2)	170 (2)
N5—H5 \cdots O16	0.86 (3)	1.93 (3)	2.778 (2)	170 (2)
N9—H9 \cdots O8	0.85 (2)	1.95 (2)	2.790 (2)	172 (2)
N10—H10 \cdots O7 ⁱⁱ	0.83 (2)	1.96 (2)	2.786 (2)	173 (2)
N11—H11 <i>A</i> \cdots O9 ⁱⁱⁱ	0.94 (3)	1.82 (3)	2.730 (2)	160 (3)
N12—H12 <i>A</i> \cdots O18	0.94 (3)	2.02 (3)	2.803 (3)	139 (3)
N12—H12 <i>A</i> \cdots O5 ^{iv}	0.94 (3)	2.16 (3)	2.856 (3)	130 (2)
C2—H2 \cdots O4 ^v	0.93	2.58	3.265 (3)	131
C16—H16 \cdots O2 ^{vi}	0.93	2.27	3.146 (3)	156
C24—H24 <i>B</i> \cdots O5	0.96	2.51	3.448 (4)	164
C24—H24 <i>C</i> \cdots O10 ⁱⁱⁱ	0.96	2.57	3.457 (4)	154
C26—H26 <i>B</i> \cdots O18 ^{iv}	0.96	2.50	3.400 (5)	156

Symmetry codes: (i) $x-1, y, z$; (ii) $x+1, y, z$; (iii) $x, -y+1/2, z+1/2$; (iv) $-x+1, -y, -z+1$; (v) $-x, -y, -z+2$; (vi) $-x+1, y+1/2, -z+3/2$.