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(*Z*)-Methyl 3-(2,4-dichlorophenyl)-3-hydroxyacrylate

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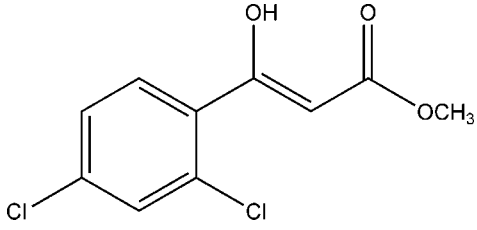
Key indicators: single-crystal X-ray study; T = 294 K; mean σ(C–C) = 0.003 Å; R factor = 0.034; wR factor = 0.086; data-to-parameter ratio = 17.1.

The molecular structure of the title compound, $C_{10}H_8Cl_2O_3$, exists in a *cis*-enol form, which is stabilized by a strong intramolecular O—H···O hydrogen bond. In the crystal, C—H···O interactions generate zigzag chains along the *c* axis which are, in turn, linked by further C—H···O interactions into sheets parallel to (100).

Related literature

For the synthesis of the title compound, see: Wu *et al.* (1997). For related structures, see: Mei & Huang (2007); Zheng, Fan *et al.* (2007); Zheng, Zheng *et al.* (2007). For the coordination properties of similar compounds, see: Nakamoto *et al.* (1970); Ma *et al.* (1999); Yoshida *et al.*(2005).



Experimental

Crystal data

$C_{10}H_8Cl_2O_3$
M_r = 247.06
Monoclinic, *Cc*
a = 15.889 (3) Å
b = 3.8242 (8) Å
c = 18.204 (4) Å
β = 108.18 (3)°
V = 1050.9 (4) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.60 mm^{-1}
T = 294 K
0.25 × 0.20 × 0.15 mm

Data collection

Rigaku SCXmini diffractometer
5011 measured reflections
2371 independent reflections
2170 reflections with $I > 2\sigma(I)$
R_{int} = 0.030

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.034
$wR(F^2)$ = 0.086
S = 1.07
2371 reflections
139 parameters
2 restraints
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.17 e Å^{-3}
$\Delta\rho_{min}$ = −0.18 e Å^{-3}
Absolute structure: Flack (1983), 1177 Friedel pairs
Flack parameter: 0.07 (6)

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O1—H1···O2	0.82	1.87	2.592 (3)	146
C3—H3···O2^{i}	0.93	2.48	3.356 (3)	157
C10—H10*B*···O1^{ii}	0.96	2.57	3.492 (3)	162

Symmetry codes: (i) $x, -y+1, z-\frac{1}{2}$; (ii) $x-\frac{1}{2}, y+\frac{1}{2}, z$.

Data collection: *CrystalClear* (Rigaku, 2005); cell refinement: *CrystalClear*; data reduction: *CrystalClear*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL/PC* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL/PC*.

This work was supported by the National Natural Science Foundation (81072577).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: LR2037).

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supporting information

Acta Cryst. (2012). E**68**, o4 [doi:10.1107/S1600536811051014]

(*Z*)-Methyl 3-(2,4-dichlorophenyl)-3-hydroxyacrylate

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S1. Comment

1,3-Diketones are versatile intermediates for the synthesis of some palladium(II) and platinum(II) compounds (Nakamoto *et al.*, 1970) and other coordination compounds (Ma *et al.*, 1999; Yoshida *et al.*, 2005). We present here the structure characterization of (*Z*)-methyl 3-(2,4-dichlorophenyl)-3-hydroxyacrylate.

The molecular structure (Fig.1) exists in a *cis*-enol form which is stabilized by a strong intramolecular O1—H1···O2 hydrogen bond. The crystal structure (Fig.2) is stabilized by intermolecular C—H···O interactions (Table 1) . The C3—H3···O2 interactions generated zigzag chains along the *c* axis which in turn are linked by C10—H10B···O1 interactions giving sheets parallel to (100).

S2. Experimental

The title compound was synthesized according to the literature procedure of Wu *et al.* (1997).Single crystals suitable for X-ray diffraction were obtained by slow evaporation of an ethanol solution at room temperature.

S3. Refinement

All H atoms were detected in a difference map, but all other H-atoms were placed in calculated positions and refined using a riding motion approximation, with C—H=0.93–0.96 Å, with U_{iso}(H)=1.2 or 1.5U_{eq}(C); O—H=0.82 Å, with U_{iso}(H)=1.5U_{eq}(O).

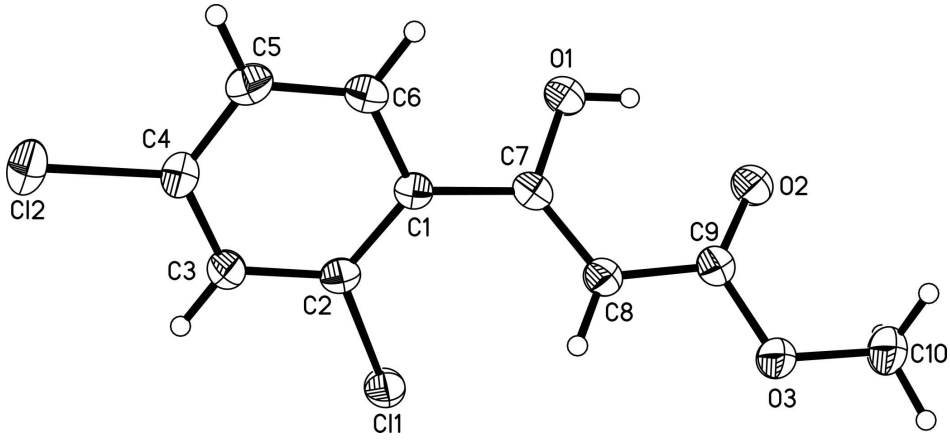

Figure 1

The molecular structure of the title compound. Displacement ellipsoids are drawn at the 30% probability level.

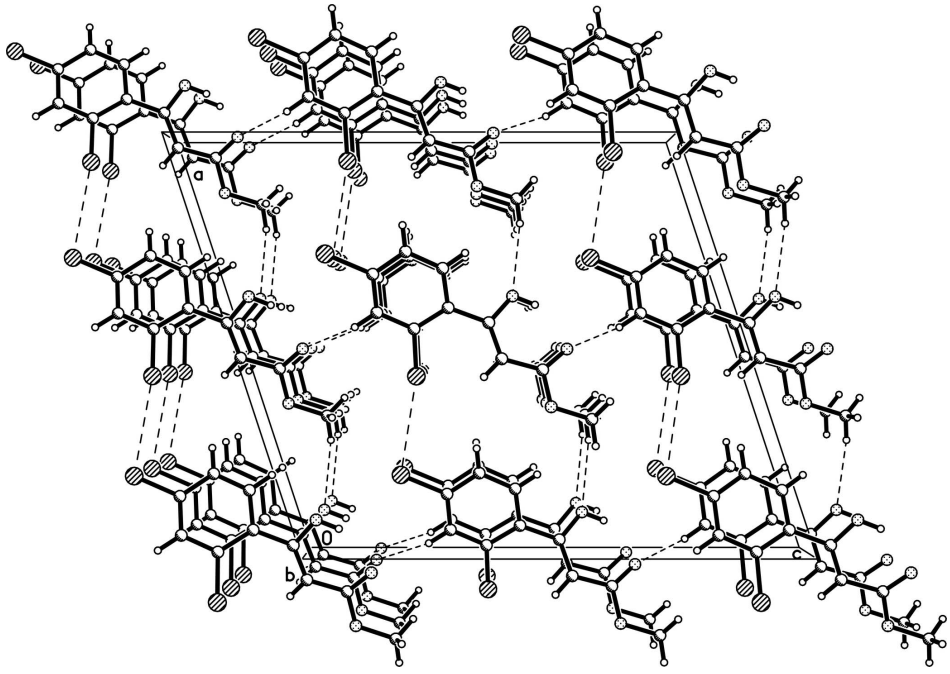

Figure 2

Packing diagram of the title compound viewed along the *b* axis. Hydrogen bonds are shown as dashed lines.

(*Z*)-Methyl 3-(2,4-dichlorophenyl)-3-hydroxyacrylate

Crystal data

$C_{10}H_8Cl_2O_3$
M_r = 247.06
Monoclinic, *Cc*
Hall symbol: C -2yc
a = 15.889 (3) Å
b = 3.8242 (8) Å
c = 18.204 (4) Å
β = 108.18 (3)°
V = 1050.9 (4) Å^3
Z = 4
F(000) = 504
D_x = 1.562 Mg m^{-3}
Mo *Kα* radiation, *λ* = 0.71073 Å
Cell parameters from 5046 reflections
θ = 3.3–27.0°
μ = 0.60 mm^{-1}
T = 294 K
Prism, colorless
0.25 × 0.20 × 0.15 mm

Data collection

Rigaku SCXmini diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
Detector resolution: 13.6612 pixels mm^{-1}
CCD_Profile_fitting scans
5011 measured reflections
2371 independent reflections
2170 reflections with $I > 2\sigma(I)$
R_{int} = 0.030
θ_{max} = 27.5°, θ_{min} = 4.1°
h = −20→20
k = −4→4
l = −23→23

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)]$ = 0.034
$wR(F^2)$ = 0.086
S = 1.07
2371 reflections
139 parameters
2 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites
H-atom parameters constrained
$w = 1/[\sigma^2(F_o^2) + (0.0413P)^2 + 0.0839P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} < 0.001$
$\Delta\rho_{max} = 0.17$ e Å^{-3}
$\Delta\rho_{min} = -0.18$ e Å^{-3}
Extinction correction: *SHELXL97* (Sheldrick, 2008), $Fc^* = kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
Extinction coefficient: 0.0193 (14)
Absolute structure: Flack (1983), 1177 Friedel pairs
Absolute structure parameter: 0.07 (6)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
C1	0.58573 (13)	0.7958 (6)	0.44452 (12)	0.0381 (5)
C2	0.53832 (13)	0.7161 (6)	0.36792 (12)	0.0392 (5)
C3	0.57482 (16)	0.7526 (7)	0.30884 (13)	0.0439 (5)
H3	0.5423	0.6944	0.2583	0.053*
C4	0.65998 (16)	0.8762 (6)	0.32574 (14)	0.0478 (5)
C5	0.70941 (16)	0.9602 (7)	0.40023 (15)	0.0520 (6)
H5	0.7669	1.0447	0.4109	0.062*
C6	0.67267 (16)	0.9175 (7)	0.45842 (14)	0.0495 (6)
H6	0.7065	0.9709	0.5088	0.059*
C7	0.55234 (15)	0.7569 (6)	0.51099 (12)	0.0424 (5)
C8	0.47153 (15)	0.8564 (6)	0.51265 (13)	0.0424 (5)
H8	0.4315	0.9548	0.4689	0.051*
C9	0.44659 (15)	0.8113 (7)	0.58190 (13)	0.0441 (5)
C10	0.3367 (2)	0.8937 (8)	0.64204 (16)	0.0612 (7)
H10A	0.3348	0.6501	0.6542	0.092*
H10B	0.2787	0.9934	0.6319	0.092*
H10C	0.3776	1.0126	0.6849	0.092*
Cl1	0.43104 (3)	0.55059 (16)	0.34217 (3)	0.05160 (19)
Cl2	0.70467 (4)	0.9249 (2)	0.25046 (4)	0.0757 (3)
O1	0.61264 (12)	0.6173 (6)	0.57216 (10)	0.0629 (5)
H1	0.5931	0.6120	0.6088	0.094*
O2	0.49301 (12)	0.6776 (6)	0.64092 (9)	0.0620 (5)
O3	0.36513 (11)	0.9310 (5)	0.57450 (9)	0.0521 (4)

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0333 (10)	0.0443 (11)	0.0353 (10)	−0.0018 (9)	0.0086 (8)	−0.0022 (8)
C2	0.0305 (9)	0.0449 (12)	0.0401 (11)	−0.0029 (8)	0.0081 (8)	−0.0011 (9)

C3	0.0384 (10)	0.0573 (13)	0.0350 (12)	−0.0029 (10)	0.0103 (10)	−0.0052 (10)
C4	0.0450 (12)	0.0598 (15)	0.0441 (12)	0.0013 (11)	0.0218 (10)	0.0021 (11)
C5	0.0355 (11)	0.0657 (15)	0.0541 (14)	−0.0083 (10)	0.0130 (11)	−0.0040 (12)
C6	0.0368 (12)	0.0687 (15)	0.0386 (12)	−0.0056 (11)	0.0056 (10)	−0.0073 (10)
C7	0.0412 (11)	0.0494 (12)	0.0329 (11)	−0.0003 (9)	0.0062 (9)	−0.0009 (9)
C8	0.0390 (11)	0.0525 (13)	0.0349 (11)	0.0029 (9)	0.0104 (9)	0.0063 (10)
C9	0.0423 (12)	0.0534 (13)	0.0366 (11)	0.0007 (10)	0.0124 (10)	0.0003 (10)
C10	0.0577 (15)	0.083 (2)	0.0505 (15)	0.0069 (14)	0.0281 (13)	0.0036 (13)
Cl1	0.0372 (3)	0.0734 (4)	0.0422 (3)	−0.0139 (3)	0.0093 (2)	−0.0054 (3)
Cl2	0.0592 (4)	0.1199 (7)	0.0596 (4)	−0.0099 (5)	0.0351 (4)	0.0000 (5)
O1	0.0466 (9)	0.1045 (15)	0.0347 (9)	0.0182 (10)	0.0085 (7)	0.0137 (9)
O2	0.0507 (10)	0.0981 (14)	0.0363 (9)	0.0163 (10)	0.0124 (8)	0.0170 (9)
O3	0.0434 (8)	0.0738 (12)	0.0412 (9)	0.0096 (8)	0.0164 (7)	0.0078 (8)

Geometric parameters (Å, º)

C1—C2	1.396 (3)	C7—O1	1.333 (3)
C1—C6	1.404 (3)	C7—C8	1.349 (3)
C1—C7	1.472 (3)	C8—C9	1.445 (3)
C2—C3	1.379 (3)	C8—H8	0.9300
C2—Cl1	1.740 (2)	C9—O2	1.210 (3)
C3—C4	1.375 (3)	C9—O3	1.340 (3)
C3—H3	0.9300	C10—O3	1.443 (3)
C4—C5	1.376 (4)	C10—H10A	0.9600
C4—Cl2	1.739 (2)	C10—H10B	0.9600
C5—C6	1.370 (3)	C10—H10C	0.9600
C5—H5	0.9300	O1—H1	0.8200
C6—H6	0.9300		
C2—C1—C6	116.5 (2)	O1—C7—C8	122.4 (2)
C2—C1—C7	125.31 (18)	O1—C7—C1	112.18 (19)
C6—C1—C7	118.17 (19)	C8—C7—C1	125.4 (2)
C3—C2—C1	121.98 (19)	C7—C8—C9	120.5 (2)
C3—C2—Cl1	116.21 (16)	C7—C8—H8	119.7
C1—C2—Cl1	121.77 (16)	C9—C8—H8	119.7
C4—C3—C2	119.0 (2)	O2—C9—O3	122.4 (2)
C4—C3—H3	120.5	O2—C9—C8	124.6 (2)
C2—C3—H3	120.5	O3—C9—C8	113.00 (19)
C3—C4—C5	121.3 (2)	O3—C10—H10A	109.5
C3—C4—Cl2	118.42 (19)	O3—C10—H10B	109.5
C5—C4—Cl2	120.24 (19)	H10A—C10—H10B	109.5
C6—C5—C4	119.0 (2)	O3—C10—H10C	109.5
C6—C5—H5	120.5	H10A—C10—H10C	109.5
C4—C5—H5	120.5	H10B—C10—H10C	109.5
C5—C6—C1	122.2 (2)	C7—O1—H1	109.5
C5—C6—H6	118.9	C9—O3—C10	115.47 (19)
C1—C6—H6	118.9		

C6—C1—C2—C3	0.4 (4)	C7—C1—C6—C5	179.7 (2)
C7—C1—C2—C3	−178.6 (2)	C2—C1—C7—O1	136.9 (2)
C6—C1—C2—Cl1	178.08 (18)	C6—C1—C7—O1	−42.0 (3)
C7—C1—C2—Cl1	−0.9 (3)	C2—C1—C7—C8	−44.8 (3)
C1—C2—C3—C4	−1.0 (4)	C6—C1—C7—C8	136.3 (3)
Cl1—C2—C3—C4	−178.9 (2)	O1—C7—C8—C9	−0.6 (4)
C2—C3—C4—C5	0.7 (4)	C1—C7—C8—C9	−178.8 (2)
C2—C3—C4—Cl2	−179.34 (19)	C7—C8—C9—O2	−2.1 (4)
C3—C4—C5—C6	0.3 (4)	C7—C8—C9—O3	178.5 (2)
Cl2—C4—C5—C6	−179.7 (2)	O2—C9—O3—C10	0.7 (4)
C4—C5—C6—C1	−1.0 (4)	C8—C9—O3—C10	−179.9 (2)
C2—C1—C6—C5	0.7 (4)		

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O1—H1···O2	0.82	1.87	2.592 (3)	146
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