

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

trans-Dichloridobis[(6-nicotinoyl-2-pyridyl- κ N⁶)(3-pyridyl- κ N)methanone]-copper(II)

Hong Qiang* and Fan Zhang

Department of Chemistry, Capital Normal University, Beijing 100048, People's Republic of China

Correspondence e-mail: qiang-hong@163.com

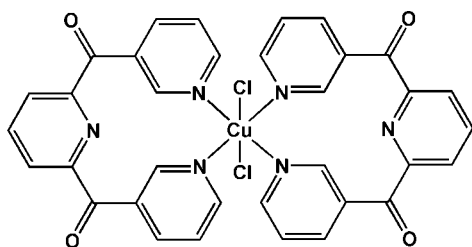
Received 6 May 2011; accepted 10 May 2011

Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.029; wR factor = 0.084; data-to-parameter ratio = 18.3.

In the title complex, $[\text{CuCl}_2(\text{C}_{17}\text{H}_{11}\text{N}_3\text{O}_2)_2]$, the Cu^{II} ion is located on an inversion center. It exhibits a distorted octahedral coordination geometry defined by two chloride anions at *trans* sites and four 3-pyridyl N atoms at equatorial sites from two (6-nicotinoyl-2-pyridyl)(3-pyridyl)methanone ligands. The (6-nicotinoyl-2-pyridyl)(3-pyridyl)methanone ligand can be viewed as having two pendant 3-pyridyl rings attached to a central pyridyl skeleton *via* separate carbonyl bridges, acting in a $\kappa^2\text{N},\text{N}'$ -chelating mode with its 3-pyridyl N atoms bound to the Cu^{II} ion. The pendant 3-pyridyl rings make a dihedral angle of $80.76(5)^\circ$. In the crystal, molecules are linked through intermolecular $\text{C}-\text{H}\cdots\pi$ and $\text{C}-\text{H}\cdots\text{O}$ interactions, forming a three-dimensional framework.

Related literature

For transition metal complexes with di-pyrid-2-yl ketone, see: Papaefstathiou & Perlepes (2002); Efthymiou *et al.* (2006). For the crystal structure of an analogous Cu^{II} complex, see: Wan *et al.* (2008). For $\text{C}-\text{H}\cdots\pi$ interactions, see: Umezawa *et al.* (1998).



Experimental

Crystal data

$[\text{CuCl}_2(\text{C}_{17}\text{H}_{11}\text{N}_3\text{O}_2)_2]$
 $M_r = 713.02$
 Monoclinic, $C2/c$
 $a = 18.728(3)$ Å
 $b = 11.8971(18)$ Å
 $c = 16.695(3)$ Å
 $\beta = 121.522(3)^\circ$

$V = 3170.9(8)$ Å³
 $Z = 4$
 Mo $K\alpha$ radiation
 $\mu = 0.91$ mm⁻¹
 $T = 293$ K
 $0.40 \times 0.30 \times 0.30$ mm

Data collection

Bruker APEXII CCD area-detector diffractometer
 Absorption correction: multi-scan (SADABS; Bruker, 2007)
 $T_{\text{min}} = 0.848$, $T_{\text{max}} = 1.000$
 11215 measured reflections
 3937 independent reflections
 3361 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.020$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.029$
 $wR(F^2) = 0.084$
 $S = 1.04$
 3937 reflections
 215 parameters
 H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.25$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.21$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

Cg1 is the centroid of the C13–C17,N3 ring.

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C13–H13A \cdots O1 ⁱ	0.93	2.61	3.418 (2)	146
C2–H2A \cdots Cg1 ⁱⁱ	0.93	2.73	3.621 (3)	162

Symmetry codes: (i) $x + \frac{1}{2}, y - \frac{1}{2}, z$; (ii) $-x + 1, y, -z + \frac{1}{2}$.

Data collection: APEX2 (Bruker, 2007); cell refinement: APEX2 and SAINT (Bruker, 2007); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXTL and PLATON (Spek, 2009).

The authors are grateful for financial support from the Science and Technology program, Beijing Municipal Education Commission.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: ZQ2102).

References

- Bruker (2007). APEX2, SAINT and SADABS. Bruker AXS Inc., Madison, Wisconsin, USA.
- Efthymiou, C. G., Raptopoulou, C. P., Terzis, A., Boca, R., Korabic, M., Mrozinski, J., Perlepes, S. P. & Bakalbassis, E. G. (2006). *Eur. J. Inorg. Chem.* **11**, 2236–2252.
- Papaefstathiou, G. S. & Perlepes, S. P. (2002). *Comments Inorg. Chem.* **23**, 249–274.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.
- Umezawa, Y., Tsuboyama, S., Honda, K., Uzawa, J. & Nishio, M. (1998). *Bull. Chem. Soc. Jpn.* **71**, 1202–1213.
- Wan, C. Q., Chen, X. D. & Mak, T. C. W. (2008). *CrystEngComm*, **10**, 475–478.

supporting information

Acta Cryst. (2011). E67, m758 [doi:10.1107/S1600536811017685]

***trans*-Dichloridobis[(6-nicotinoyl-2-pyridyl- κ N⁶)(3-pyridyl- κ N)methanone]copper(II)**

Hong Qiang and Fan Zhang

S1. Comment

Di-pyridin-2-yl-methanone (di-pyrid-2-yl ketone, DPK), (C₅H₄N)₂CO, is an extraordinarily versatile ligand among the thousands of basic building blocks that have been used in coordination chemistry and materials science (Papaefstathiou & Perlepes, 2002; Efthymiou *et al.*, 2006). Herein, we report the mononuclear Cu^{II} complex with the oligo-pyridyl ketone ligand (6-nicotinoyl-2-pyridyl)(3-pyridyl)methanone (abbreviated as L), a member of the pyridinylmethanone family.

In the crystal structure of the title complex, the center Cu^{II} adopts an octahedral coordination geometry with two chlorido depositing in *trans* to each other, and two 2,6-pyridinediylbis(3-pyridinyl)methanone ligands bound to the ion by four 3-pyridyl N atoms (Fig. 1). The Cu1—N3 and Cu1—Cl1 bond lengths equal 2.0412 (12) Å and 2.3087 (4) Å, respectively, while the Cu—N1 exhibits weak bonding with the Cu—N1 distance of 2.615 (1) Å. The latter Cu—N bonds are remarkably longer than that (about 2.03 Å) in the similar complex Cu(L)₂(BF₄)₂ (Wan *et al.* 2008). The pendant 3-pyridyl rings exhibit a dihedral angle of 80.76 (5)°. The mononuclear complex units link each other through the intermolecular C2—H2A $\cdots\pi$ and C13—H13A \cdots O1ⁱⁱ interactions to form a three-dimensional framework, as shown in Fig. 2. For the C—H $\cdots\pi$ interaction (Umezawa *et al.* 1998), the C2 \cdots Cg1 distance (where Cg1 is the centroid of the ring containing N3ⁱ; i: -x+1, -y, 0.5-z) is 3.621 (3) Å, and the C2—H2A \cdots Cg1 angle is 161.8°. For the C—H \cdots O interaction, the C13 \cdots O1ⁱⁱ distance is 3.418 (2) Å, and the C13—H13A \cdots O1ⁱⁱ angle is 146.2° (ii: x+0.5, y-0.5, z).

S2. Experimental

The (6-nicotinoyl-2-pyridyl)(3-pyridyl)methanone ligand was obtained following the reaction procedure as reported in literature (Wan *et al.*, 2008). Reaction of (6-nicotinoyl-2-pyridyl)(3-pyridyl)methanone (29 mg, 0.1 mmol) with CuCl₂ (7 mg, 0.05 mmol) in acetonitrile formed *trans*-[Cu(C₁₇H₁₁N₃O₂)₂Cl₂] as a blue solution, which was left stand in air for four days to obtain block-like crystals (yield 13.1mg, 61%).

S3. Refinement

The hydrogen atoms were placed in idealized positions and allowed to ride on the relevant carbon atoms, with C—H = 0.93 Å and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

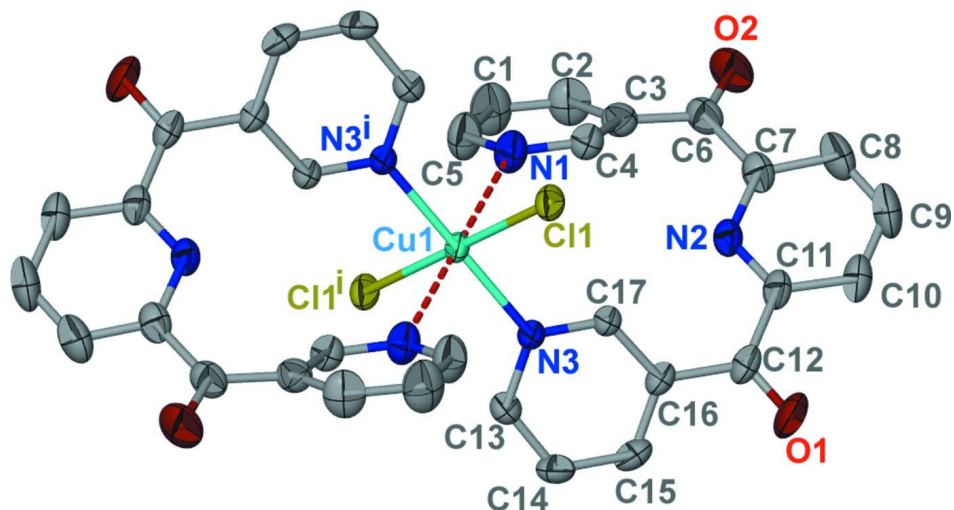


Figure 1

The atom-numbering scheme of the title complex (symmetry code: (i) $-x+1.5, -y+0.5, -z+1.5$) with red-dashed lines indicating weak Cu1-N1 bonding. Displacement ellipsoids are drawn at the 50% probability level and H atoms are omitted for clarity.

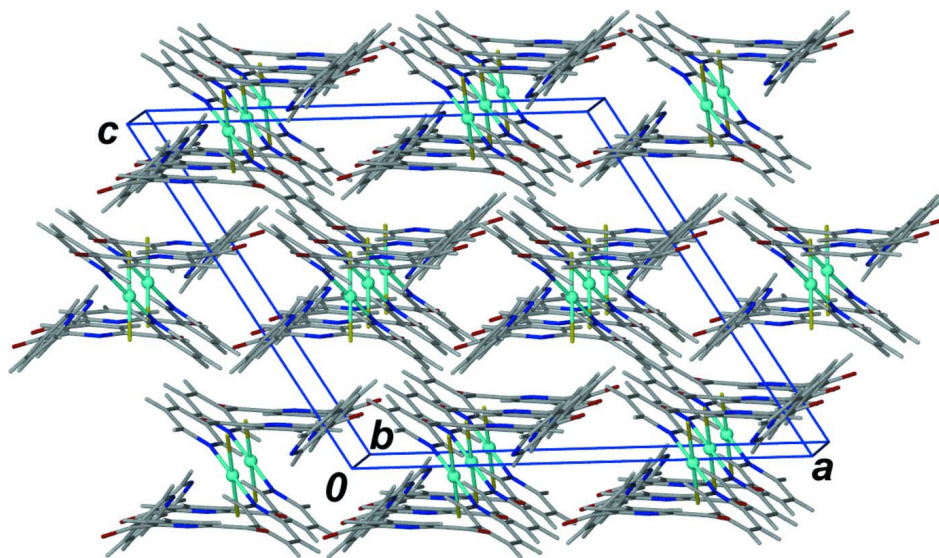


Figure 2

View of the crystal packing of the title compound.

***trans*-Dichloridobis[(6-nicotinoyl-2-pyridyl- κN^6)(3-pyridyl- κN)methanone]copper(II)**

Crystal data

$[\text{CuCl}_2(\text{C}_{17}\text{H}_{11}\text{N}_3\text{O}_2)_2]$

$M_r = 713.02$

Monoclinic, $C2/c$

Hall symbol: $-C\ 2yc$

$a = 18.728(3)\ \text{\AA}$

$b = 11.8971(18)\ \text{\AA}$

$c = 16.695(3)\ \text{\AA}$

$\beta = 121.522(3)^\circ$

$V = 3170.9(8)\ \text{\AA}^3$

$Z = 4$

$F(000) = 1452$

$D_x = 1.494\ \text{Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073\ \text{\AA}$

$\mu = 0.91\ \text{mm}^{-1}$

$T = 293$ K
Block, blue

$0.40 \times 0.30 \times 0.30$ mm

Data collection

Bruker APEXII CCD area-detector
diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
 ω scans
Absorption correction: multi-scan
(SADABS; Bruker, 2007)
 $T_{\min} = 0.848$, $T_{\max} = 1.000$

11215 measured reflections
3937 independent reflections
3361 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.020$
 $\theta_{\max} = 28.3^\circ$, $\theta_{\min} = 2.1^\circ$
 $h = -24 \rightarrow 24$
 $k = -15 \rightarrow 15$
 $l = -19 \rightarrow 22$

Refinement

Refinement on F^2
Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.029$
 $wR(F^2) = 0.084$
 $S = 1.04$
3937 reflections
215 parameters
0 restraints
Primary atom site location: structure-invariant
direct methods
Secondary atom site location: difference Fourier
map

Hydrogen site location: inferred from
neighbouring sites
H-atom parameters constrained
 $w = 1/[\sigma^2(F_o^2) + (0.0423P)^2 + 2.1803P]$
where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.001$
 $\Delta\rho_{\max} = 0.25 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.21 \text{ e } \text{\AA}^{-3}$
Extinction correction: SHELXL97 (Sheldrick,
2008), $F_c^* = kF_c[1 + 0.001x F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$
Extinction coefficient: 0.00079 (17)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2 , conventional R-factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F , and R-factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
Cu1	0.7500	0.2500	0.5000	0.03142 (9)
Cl1	0.69747 (3)	0.11938 (3)	0.38011 (3)	0.03905 (11)
O2	0.77486 (11)	0.74717 (10)	0.35782 (15)	0.0670 (5)
C12	0.73237 (11)	0.67642 (13)	0.36615 (13)	0.0423 (4)
C1	0.45389 (14)	0.2790 (2)	0.33217 (19)	0.0666 (6)
H1A	0.4120	0.2267	0.2971	0.080*
C2	0.43565 (12)	0.39132 (18)	0.32729 (16)	0.0588 (5)
H2A	0.3811	0.4166	0.2878	0.071*
C3	0.49909 (10)	0.46676 (15)	0.38157 (12)	0.0404 (4)
C4	0.57867 (10)	0.42349 (14)	0.44252 (12)	0.0386 (3)
H4A	0.6207	0.4729	0.4830	0.046*
N1	0.59789 (9)	0.31499 (12)	0.44589 (11)	0.0464 (3)

C5	0.53567 (13)	0.24547 (15)	0.39018 (18)	0.0571 (5)
H5A	0.5483	0.1698	0.3905	0.069*
C6	0.47982 (11)	0.58898 (16)	0.37172 (13)	0.0453 (4)
O1	0.41059 (9)	0.62239 (13)	0.35012 (13)	0.0676 (4)
C7	0.54433 (11)	0.67304 (14)	0.38234 (13)	0.0432 (4)
C8	0.53320 (14)	0.78598 (18)	0.39550 (17)	0.0602 (5)
H8A	0.4891	0.8084	0.4023	0.072*
C9	0.58907 (15)	0.86377 (16)	0.39825 (18)	0.0657 (6)
H9A	0.5826	0.9397	0.4062	0.079*
C10	0.65408 (13)	0.82833 (15)	0.38923 (14)	0.0520 (5)
H10A	0.6924	0.8795	0.3908	0.062*
C11	0.66154 (11)	0.71359 (13)	0.37759 (12)	0.0394 (4)
N2	0.60753 (9)	0.63678 (11)	0.37358 (10)	0.0378 (3)
N3	0.76337 (8)	0.36519 (10)	0.41813 (9)	0.0300 (3)
C14	0.74131 (10)	0.47290 (12)	0.41611 (11)	0.0314 (3)
H14A	0.7177	0.4935	0.4510	0.038*
C15	0.75214 (10)	0.55503 (12)	0.36437 (11)	0.0338 (3)
C16	0.78887 (11)	0.52392 (15)	0.31374 (12)	0.0419 (4)
H16A	0.7983	0.5770	0.2794	0.050*
C17	0.81099 (11)	0.41376 (15)	0.31523 (13)	0.0431 (4)
H17A	0.8353	0.3914	0.2815	0.052*
C13	0.79703 (10)	0.33587 (14)	0.36711 (11)	0.0368 (3)
H13A	0.8113	0.2611	0.3667	0.044*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cu1	0.04854 (17)	0.01975 (13)	0.03224 (15)	-0.00100 (10)	0.02548 (13)	0.00062 (9)
Cl1	0.0554 (2)	0.02824 (18)	0.0347 (2)	-0.00295 (15)	0.02439 (18)	-0.00262 (14)
O2	0.0755 (10)	0.0343 (7)	0.1049 (14)	-0.0040 (6)	0.0566 (10)	0.0174 (7)
C12	0.0487 (9)	0.0296 (7)	0.0459 (9)	-0.0008 (7)	0.0230 (8)	0.0109 (7)
C1	0.0446 (11)	0.0549 (12)	0.0815 (16)	-0.0138 (9)	0.0198 (11)	-0.0136 (11)
C2	0.0344 (9)	0.0595 (12)	0.0661 (13)	0.0007 (8)	0.0149 (9)	-0.0019 (10)
C3	0.0366 (8)	0.0426 (9)	0.0436 (9)	0.0033 (7)	0.0221 (7)	0.0027 (7)
C4	0.0371 (8)	0.0365 (8)	0.0390 (8)	-0.0007 (6)	0.0177 (7)	0.0016 (6)
N1	0.0412 (8)	0.0358 (7)	0.0534 (9)	0.0007 (6)	0.0186 (7)	0.0041 (6)
C5	0.0498 (11)	0.0376 (10)	0.0752 (15)	-0.0054 (8)	0.0266 (11)	-0.0042 (9)
C6	0.0433 (9)	0.0466 (9)	0.0468 (10)	0.0112 (8)	0.0242 (8)	0.0031 (8)
O1	0.0517 (8)	0.0637 (9)	0.0945 (12)	0.0184 (7)	0.0432 (9)	0.0034 (8)
C7	0.0444 (9)	0.0361 (8)	0.0428 (9)	0.0093 (7)	0.0185 (8)	0.0022 (7)
C8	0.0609 (13)	0.0440 (10)	0.0716 (14)	0.0169 (9)	0.0318 (11)	-0.0033 (10)
C9	0.0750 (15)	0.0305 (9)	0.0796 (15)	0.0102 (9)	0.0320 (13)	-0.0070 (9)
C10	0.0630 (12)	0.0273 (8)	0.0548 (11)	-0.0004 (8)	0.0233 (10)	0.0011 (7)
C11	0.0489 (9)	0.0264 (7)	0.0365 (8)	0.0033 (6)	0.0178 (7)	0.0052 (6)
N2	0.0430 (7)	0.0279 (6)	0.0397 (7)	0.0055 (5)	0.0196 (6)	0.0041 (5)
N3	0.0373 (7)	0.0248 (5)	0.0326 (6)	0.0011 (5)	0.0214 (6)	0.0023 (5)
C14	0.0374 (7)	0.0264 (7)	0.0343 (7)	0.0021 (6)	0.0213 (6)	0.0036 (6)
C15	0.0366 (8)	0.0284 (7)	0.0352 (8)	-0.0024 (6)	0.0179 (6)	0.0052 (6)

C16	0.0462 (9)	0.0437 (9)	0.0407 (9)	-0.0072 (7)	0.0262 (8)	0.0073 (7)
C17	0.0491 (10)	0.0495 (9)	0.0438 (9)	-0.0021 (8)	0.0335 (8)	-0.0011 (7)
C13	0.0428 (9)	0.0332 (7)	0.0394 (8)	0.0019 (6)	0.0251 (7)	-0.0011 (6)

Geometric parameters (Å, °)

Cu1—N3	2.0412 (12)	C7—N2	1.338 (2)
Cu1—N3 ⁱ	2.0412 (12)	C7—C8	1.395 (3)
Cu1—C11	2.3087 (4)	C8—C9	1.380 (3)
Cu1—C11 ⁱ	2.3087 (4)	C8—H8A	0.9300
O2—C12	1.215 (2)	C9—C10	1.369 (3)
C12—C15	1.495 (2)	C9—H9A	0.9300
C12—C11	1.502 (3)	C10—C11	1.396 (2)
C1—C2	1.371 (3)	C10—H10A	0.9300
C1—C5	1.376 (3)	C11—N2	1.339 (2)
C1—H1A	0.9300	N3—C14	1.3415 (18)
C2—C3	1.383 (3)	N3—C13	1.3438 (19)
C2—H2A	0.9300	C14—C15	1.3875 (19)
C3—C4	1.391 (2)	C14—H14A	0.9300
C3—C6	1.487 (2)	C15—C16	1.390 (2)
C4—N1	1.333 (2)	C16—C17	1.371 (2)
C4—H4A	0.9300	C16—H16A	0.9300
N1—C5	1.332 (2)	C17—C13	1.383 (2)
C5—H5A	0.9300	C17—H17A	0.9300
C6—O1	1.217 (2)	C13—H13A	0.9300
C6—C7	1.506 (3)		
N3—Cu1—N3 ⁱ	180.0	C9—C8—C7	118.57 (19)
N3—Cu1—C11	90.99 (4)	C9—C8—H8A	120.7
N3 ⁱ —Cu1—C11	89.01 (4)	C7—C8—H8A	120.7
N3—Cu1—C11 ⁱ	89.01 (4)	C10—C9—C8	119.50 (18)
N3 ⁱ —Cu1—C11 ⁱ	90.99 (4)	C10—C9—H9A	120.2
C11—Cu1—C11 ⁱ	180.0	C8—C9—H9A	120.2
O2—C12—C15	118.88 (17)	C9—C10—C11	118.42 (19)
O2—C12—C11	119.02 (16)	C9—C10—H10A	120.8
C15—C12—C11	122.10 (14)	C11—C10—H10A	120.8
C2—C1—C5	118.40 (19)	N2—C11—C10	123.12 (17)
C2—C1—H1A	120.8	N2—C11—C12	119.14 (14)
C5—C1—H1A	120.8	C10—C11—C12	117.71 (16)
C1—C2—C3	119.45 (18)	C7—N2—C11	117.58 (14)
C1—C2—H2A	120.3	C14—N3—C13	118.19 (13)
C3—C2—H2A	120.3	C14—N3—Cu1	120.94 (10)
C2—C3—C4	117.64 (17)	C13—N3—Cu1	120.84 (10)
C2—C3—C6	119.12 (16)	N3—C14—C15	123.02 (14)
C4—C3—C6	123.23 (16)	N3—C14—H14A	118.5
N1—C4—C3	123.46 (16)	C15—C14—H14A	118.5
N1—C4—H4A	118.3	C14—C15—C16	118.09 (14)
C3—C4—H4A	118.3	C14—C15—C12	123.33 (14)

C5—N1—C4	116.97 (16)	C16—C15—C12	118.41 (14)
N1—C5—C1	123.86 (18)	C17—C16—C15	118.97 (14)
N1—C5—H5A	118.1	C17—C16—H16A	120.5
C1—C5—H5A	118.1	C15—C16—H16A	120.5
O1—C6—C3	120.78 (18)	C16—C17—C13	119.84 (15)
O1—C6—C7	118.89 (17)	C16—C17—H17A	120.1
C3—C6—C7	120.22 (15)	C13—C17—H17A	120.1
N2—C7—C8	122.80 (18)	N3—C13—C17	121.86 (15)
N2—C7—C6	118.31 (15)	N3—C13—H13A	119.1
C8—C7—C6	118.79 (17)	C17—C13—H13A	119.1
C5—C1—C2—C3	-1.2 (4)	O2—C12—C11—C10	-7.4 (3)
C1—C2—C3—C4	-2.9 (3)	C15—C12—C11—C10	172.79 (16)
C1—C2—C3—C6	176.4 (2)	C8—C7—N2—C11	-0.2 (3)
C2—C3—C4—N1	5.2 (3)	C6—C7—N2—C11	176.08 (15)
C6—C3—C4—N1	-174.07 (17)	C10—C11—N2—C7	-0.8 (3)
C3—C4—N1—C5	-3.0 (3)	C12—C11—N2—C7	-178.81 (15)
C4—N1—C5—C1	-1.6 (3)	C11—Cu1—N3—C14	136.92 (12)
C2—C1—C5—N1	3.7 (4)	C11 ⁱ —Cu1—N3—C14	-43.08 (12)
C2—C3—C6—O1	30.2 (3)	C11—Cu1—N3—C13	-45.23 (12)
C4—C3—C6—O1	-150.5 (2)	C11 ⁱ —Cu1—N3—C13	134.77 (12)
C2—C3—C6—C7	-145.85 (19)	C13—N3—C14—C15	-0.4 (2)
C4—C3—C6—C7	33.5 (3)	Cu1—N3—C14—C15	177.50 (11)
O1—C6—C7—N2	-157.20 (19)	N3—C14—C15—C16	-1.1 (2)
C3—C6—C7—N2	18.9 (3)	N3—C14—C15—C12	-176.31 (15)
O1—C6—C7—C8	19.2 (3)	O2—C12—C15—C14	146.67 (19)
C3—C6—C7—C8	-164.64 (18)	C11—C12—C15—C14	-33.5 (2)
N2—C7—C8—C9	1.0 (3)	O2—C12—C15—C16	-28.5 (3)
C6—C7—C8—C9	-175.3 (2)	C11—C12—C15—C16	151.27 (16)
C7—C8—C9—C10	-0.8 (4)	C14—C15—C16—C17	1.5 (2)
C8—C9—C10—C11	-0.1 (3)	C12—C15—C16—C17	176.90 (16)
C9—C10—C11—N2	0.9 (3)	C15—C16—C17—C13	-0.4 (3)
C9—C10—C11—C12	178.98 (19)	C14—N3—C13—C17	1.6 (2)
O2—C12—C11—N2	170.73 (18)	Cu1—N3—C13—C17	-176.35 (13)
C15—C12—C11—N2	-9.1 (2)	C16—C17—C13—N3	-1.2 (3)

Symmetry code: (i) $-x+3/2, -y+1/2, -z+1$.

Hydrogen-bond geometry ($\text{\AA}, ^\circ$)

Cg1 is the centroid of the C13—C17,N3 ring.

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
C13—H13A \cdots O1 ⁱⁱ	0.93	2.61	3.418 (2)	146
C2—H2A \cdots Cg1 ⁱⁱⁱ	0.93	2.73	3.621 (3)	162

Symmetry codes: (ii) $x+1/2, y-1/2, z$; (iii) $-x+1, y, -z+1/2$.