

Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

# 1-(2-Oxo-3,4-dihydro-2H-1,3-benzoxazin-4-yl)urea monohydrate

Malahat M. Kurbanova,\* Aysel B. Novruzova, Atash V. Gurbanov, Sheyda R. Ismaylova and Abel M. Maharramov

Baku State University, Z. Khalilov Street 23, Baku, AZ-1148, Azerbaijan  
Correspondence e-mail: mkurbanova72@mail.ru

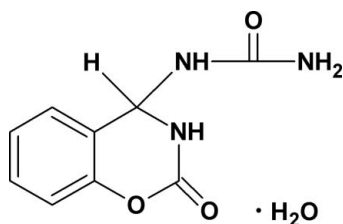
Received 2 August 2010; accepted 11 August 2010

Key indicators: single-crystal X-ray study;  $T = 296$  K; mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å;  $R$  factor = 0.043;  $wR$  factor = 0.117; data-to-parameter ratio = 16.9.

The organic molecule in the title hydrate,  $\text{C}_9\text{H}_9\text{N}_3\text{O}_3 \cdot \text{H}_2\text{O}$ , was obtained by the condensation of salicylic aldehyde with urea in acetonitrile. The oxazine ring adopts a slightly distorted sofa conformation, with the N atom deviating from the plane passing through the other atoms of the ring by 0.267 (2) Å. The crystal structure displays intermolecular  $\text{N}-\text{H} \cdots \text{O}$  and  $\text{O}-\text{H} \cdots \text{O}$  hydrogen bonding.

## Related literature

For details of the condensation of salicyl aldehyde with urea, see: Pandey *et al.* (2008); El-Hamouly *et al.* (2007); Bobowski & Shavel (1967).



## Experimental

### Crystal data

$\text{C}_9\text{H}_9\text{N}_3\text{O}_3 \cdot \text{H}_2\text{O}$   
 $M_r = 225.21$   
Monoclinic,  $P2_1/n$   
 $a = 5.3393$  (10) Å

$b = 8.5465$  (16) Å  
 $c = 21.846$  (4) Å  
 $\beta = 95.472$  (4)°  
 $V = 992.3$  (3) Å<sup>3</sup>

$Z = 4$   
Mo  $K\alpha$  radiation  
 $\mu = 0.12$  mm<sup>-1</sup>

$T = 296$  K  
 $0.30 \times 0.30 \times 0.30$  mm

### Data collection

Bruker APEXII CCD diffractometer  
Absorption correction: multi-scan (SADABS; Sheldrick, 2003)  
 $T_{\min} = 0.965$ ,  $T_{\max} = 0.965$

10468 measured reflections  
2448 independent reflections  
1942 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.058$

### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.043$   
 $wR(F^2) = 0.117$   
 $S = 1.00$   
2448 reflections

145 parameters  
H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.28$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.29$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H} \cdots A$	$D-\text{H}$	$\text{H} \cdots A$	$D \cdots A$	$D-\text{H} \cdots A$
$\text{N1}-\text{H1} \cdots \text{O2}^{\text{i}}$	0.90	2.05	2.946 (2)	174
$\text{N2}-\text{H2A} \cdots \text{O3}^{\text{ii}}$	0.90	2.41	3.088 (2)	132
$\text{N3}-\text{H3A} \cdots \text{O4}^{\text{iii}}$	0.93	2.14	3.054 (2)	169
$\text{N3}-\text{H3B} \cdots \text{O4}$	0.96	1.99	2.923 (2)	166
$\text{O4}-\text{H4A} \cdots \text{O2}^{\text{iv}}$	0.94	1.91	2.774 (2)	153
$\text{O4}-\text{H4B} \cdots \text{O3}^{\text{v}}$	0.94	1.94	2.842 (2)	161

Symmetry codes: (i)  $-x+1, -y+2, -z+1$ ; (ii)  $x-1, y, z$ ; (iii)  $x+1, y, z$ ; (iv)  $x-1, y-1, z$ ; (v)  $-x+1, -y+1, -z+1$ .

Data collection: APEX2 (Bruker, 2005); cell refinement: SAINT-Plus (Bruker, 2001); data reduction: SAINT-Plus; program(s) used to solve structure: SHELXTL (Sheldrick, 2008); program(s) used to refine structure: SHELXTL; molecular graphics: SHELXTL; software used to prepare material for publication: SHELXTL.

The authors thank Dr Victor N. Khrustalev for fruitful discussions and help with this work.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: RK2226).

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## supporting information

*Acta Cryst.* (2010). E66, o2348 [https://doi.org/10.1107/S1600536810032411]

**1-(2-Oxo-3,4-dihydro-2H-1,3-benzoxazin-4-yl)urea monohydrate**

**Malahat M. Kurbanova, Aysel B. Novruzova, Atash V. Gurbanov, Sheyda R. Ismaylova and Abel M. Maharramov**

**S1. Comment**

3,4-Dihydro-3-alkyl-2H-1,3-benzoxazin-2-ones, 4,4'-oxobis(3,4-dihydro-3-alkyl-2H-benzoxazin-2-ones), and 1-(3,4-dihydro-3-alkyl-2-oxo-2H-1,3-benzoxazin-4-yl)-1,3-dialkylureas were obtained by the condensation of *o*-hydroxyaromatic aldehydes with alkylisocyanates (Bobowski & Shavel, 1967). The formation 1,1'-[(2-hydroxyphenyl)methylene]diurea by the condensation of salicyl aldehyde with two molecules of carbamide is well known from the literature (El-Hamouly *et al.*, 2007; Pandey *et al.*, 2008).

The title molecule contains a fused bicycle system containing two six-membered rings (benzene and oxazin) (Fig. 1). The oxazin cycle adopts a slightly distorted *sofa* conformation, with the nitrogen N1 atom deviating from plane passed through the other atoms of the cycle by 0.267 (2) Å. Such disposition of the nitrogen atom is defined by the intermolecular N1–H1<sup>⋯</sup>O2<sup>i</sup> hydrogen bonding (Table 1). Symmetry code: (i) -x+1, -y+2, -z+1. The nitrogen N1 atom has a pyramidalized configuration (sum of bond angles at the nitrogen N1 atom is 352.7°, the C2–N1–O1–C1 torsion angle is 22.4 (2)°). The carbamide substituent of the oxazin cycle is in axial position (the C9–N2–C2–N1 torsion angle is -64.31 (15)°).

Compound **I** possesses one asymmetric center at the C2, but the crystal of **I** is racemate.

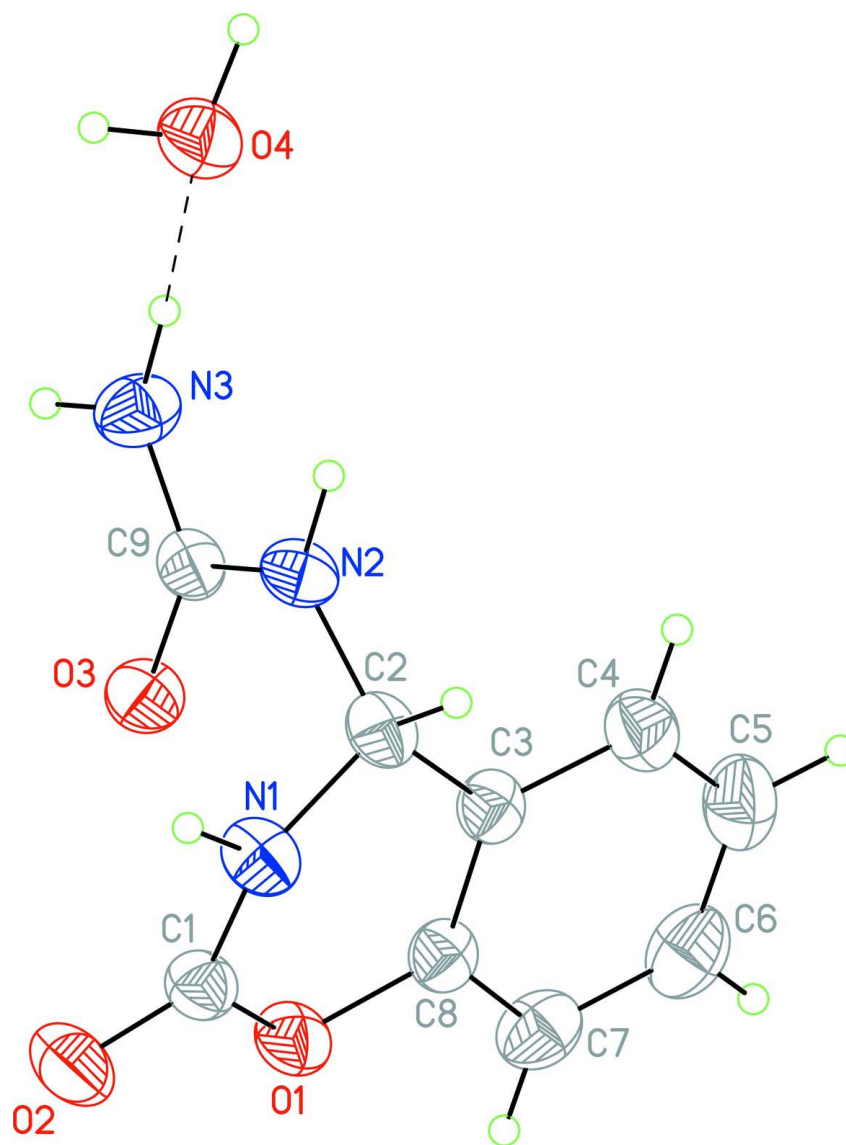
In crystal of **I**, the organic molecules and solvate water molecules are bound into complex two-tier layers parallel to (0 0 1) by the hydrogen bonds (Fig. 2, Table 1).

**S2. Experimental**

The solution of salicylic aldehyde (0.02 mol), urea (0.03 mol), 1 ml trichloroacetic acid, 3 ml *AcOH* in 15 ml acetonitril was mixed for 3-4 h (Fig. 3). Of a reaction course observed using thin layer chromatography (*TLC*) method. After cooling the product was filtered and washed out by ethanol. The target product was obtained by re-crystallization of a water solution of ethanol (1:3) as white crystals. *M.p.* = 498-500 K. IR,  $\nu/\text{cm}^{-1}$ : 1456-1541 (C=C), 1646 (NH–C(O)–NH<sub>2</sub>), 1710 (C(O)–O), 3349 (NH(lactam), associated), 3346 (C(O)–NH<sub>2</sub>). Mass spectrum, *m/z*: 230 [*M*+Na]<sup>+</sup>. <sup>1</sup>H NMR ((CD<sub>3</sub>)<sub>2</sub>SO, 293 K): 8.5 (s, 1H, NH), 7.3(m, 2H, H6, H7), 7.15 (m, 2H, H3, H5), 7.05 (d, 1H, H8, <sup>3</sup>J = 7.2), 6.2 (d, 1H, H4, <sup>3</sup>J = 7.5), 5.65 (s, 2H, NH<sub>2</sub>). <sup>13</sup>C NMR ((CD<sub>3</sub>)<sub>2</sub>SO, 293 K):  $\delta$  = 164 (s, NH–C(O)–NH<sub>2</sub>), 158 (s, C(O)–O), 150 (s, C9), 130 (s, C8), 127 (s, C5), 124.5 (s, C7), 121 (s, C10), 116 (s, C6), 58 (s, C4).

**S3. Refinement**

The amino-H atoms as well as H-atoms of the water solvate molecule were localized in the difference Fourier map and included in the refinement with fixed positional and isotropic displacement parameters [ $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{N})$  and  $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{O})$ ]. The other hydrogen atoms were placed in calculated positions with C–H = 0.93-0.98 Å and refined in the riding model with fixed isotropic displacement parameters [ $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ ].



**Figure 1**

Molecular structure of **I** with the atom numbering scheme. Displacement ellipsoids are drawn at the 50% probability level. H atoms are presented as a small spheres of arbitrary radius.

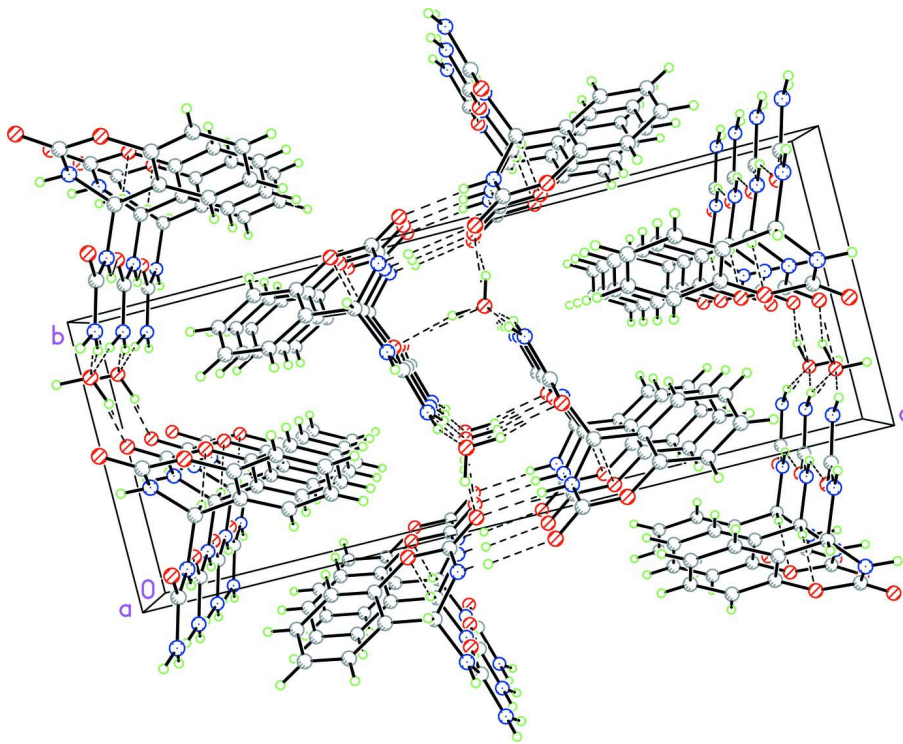


Figure 2

Crystal packing of the molecules of **I** viewed down the *a* axis. Dashed lines show intermolecular hydrogen bonds.

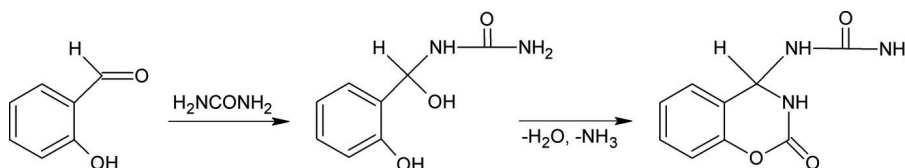


Figure 3

Condensation of salicyl aldehyde with two molecules of urea with reception benzoxazin.

### 1-(2-Oxo-3,4-dihydro-2*H*-1,3-benzoxazin-4-yl)urea monohydrate

#### Crystal data

$C_9H_9N_3O_3 \cdot H_2O$   
 $M_r = 225.21$   
 Monoclinic,  $P2_1/n$   
 Hall symbol: -P 2yn  
 $a = 5.3393$  (10) Å  
 $b = 8.5465$  (16) Å  
 $c = 21.846$  (4) Å  
 $\beta = 95.472$  (4)°  
 $V = 992.3$  (3) Å<sup>3</sup>  
 $Z = 4$

$F(000) = 472$   
 $D_x = 1.508$  Mg m<sup>-3</sup>  
 Melting point: 499 K  
 Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å  
 Cell parameters from 3341 reflections  
 $\theta = 2.6$ – $28.2$ °  
 $\mu = 0.12$  mm<sup>-1</sup>  
 $T = 296$  K  
 Prism, colourless  
 $0.30 \times 0.30 \times 0.30$  mm

*Data collection*

Bruker APEXII CCD diffractometer	10468 measured reflections
Radiation source: fine-focus sealed tube	2448 independent reflections
Graphite monochromator	1942 reflections with $I > 2\sigma(I)$
$\varphi$ and $\omega$ scans	$R_{\text{int}} = 0.058$
Absorption correction: multi-scan (SADABS; Sheldrick, 2003)	$\theta_{\text{max}} = 28.4^\circ$ , $\theta_{\text{min}} = 1.9^\circ$
$T_{\text{min}} = 0.965$ , $T_{\text{max}} = 0.965$	$h = -7 \rightarrow 7$
	$k = -11 \rightarrow 11$
	$l = -29 \rightarrow 28$

*Refinement*

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.043$	H-atom parameters constrained
$wR(F^2) = 0.117$	$w = 1/[\sigma^2(F_o^2) + (0.0598P)^2 + 0.12P]$
$S = 1.00$	where $P = (F_o^2 + 2F_c^2)/3$
2448 reflections	$(\Delta/\sigma)_{\text{max}} < 0.001$
145 parameters	$\Delta\rho_{\text{max}} = 0.28 \text{ e } \text{\AA}^{-3}$
0 restraints	$\Delta\rho_{\text{min}} = -0.29 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	

*Special details*

**Geometry.** All s.u.'s (except the s.u. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell s.u.'s are taken into account individually in the estimation of s.u.'s in distances, angles and torsion angles; correlations between s.u.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell s.u.'s is used for estimating s.u.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.72114 (16)	0.97940 (12)	0.35888 (4)	0.0372 (2)
O2	0.7419 (2)	1.05038 (14)	0.45587 (5)	0.0500 (3)
O3	0.74592 (16)	0.64789 (12)	0.40776 (4)	0.0377 (3)
N1	0.3995 (2)	0.91465 (14)	0.41776 (5)	0.0360 (3)
H1	0.3548	0.9179	0.4565	0.043*
N2	0.32112 (19)	0.63842 (14)	0.40155 (5)	0.0348 (3)
H2A	0.1839	0.5802	0.4065	0.042*
N3	0.5568 (2)	0.42106 (15)	0.43002 (6)	0.0420 (3)
H3A	0.7098	0.3794	0.4466	0.050*
H3B	0.4077	0.3652	0.4378	0.050*
C1	0.6227 (2)	0.98097 (16)	0.41376 (6)	0.0337 (3)
C2	0.2868 (2)	0.79512 (16)	0.37637 (6)	0.0322 (3)
H2	0.1054	0.8154	0.3709	0.039*
C3	0.3850 (2)	0.81370 (15)	0.31454 (6)	0.0316 (3)
C4	0.2696 (3)	0.74090 (18)	0.26224 (7)	0.0418 (3)
H4	0.1274	0.6796	0.2652	0.050*

C5	0.3641 (3)	0.7586 (2)	0.20596 (7)	0.0495 (4)
H5	0.2852	0.7098	0.1713	0.059*
C6	0.5756 (3)	0.8489 (2)	0.20131 (7)	0.0483 (4)
H6	0.6398	0.8600	0.1634	0.058*
C7	0.6926 (3)	0.92281 (19)	0.25240 (7)	0.0418 (3)
H7	0.8347	0.9842	0.2493	0.050*
C8	0.5943 (2)	0.90387 (16)	0.30837 (6)	0.0316 (3)
C9	0.5510 (2)	0.57119 (16)	0.41307 (5)	0.0306 (3)
O4	0.08489 (19)	0.29496 (14)	0.46692 (5)	0.0499 (3)
H4A	0.0070	0.1995	0.4555	0.075*
H4B	0.1116	0.2988	0.5099	0.075*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
O1	0.0276 (4)	0.0451 (6)	0.0392 (5)	-0.0083 (4)	0.0049 (4)	0.0007 (4)
O2	0.0477 (6)	0.0588 (7)	0.0432 (6)	-0.0203 (5)	0.0027 (5)	-0.0081 (5)
O3	0.0229 (4)	0.0452 (6)	0.0455 (5)	-0.0048 (4)	0.0059 (4)	0.0060 (4)
N1	0.0293 (5)	0.0432 (7)	0.0368 (6)	-0.0030 (5)	0.0093 (4)	-0.0067 (5)
N2	0.0218 (5)	0.0415 (6)	0.0419 (6)	-0.0091 (4)	0.0066 (4)	0.0025 (5)
N3	0.0385 (6)	0.0410 (7)	0.0477 (7)	-0.0032 (5)	0.0101 (5)	0.0076 (5)
C1	0.0295 (6)	0.0344 (7)	0.0371 (7)	-0.0026 (5)	0.0024 (5)	0.0001 (5)
C2	0.0177 (5)	0.0416 (7)	0.0375 (7)	-0.0009 (5)	0.0041 (5)	-0.0021 (5)
C3	0.0240 (5)	0.0363 (7)	0.0343 (6)	0.0033 (5)	0.0019 (5)	0.0020 (5)
C4	0.0352 (7)	0.0466 (8)	0.0423 (8)	-0.0025 (6)	-0.0023 (6)	-0.0008 (6)
C5	0.0556 (9)	0.0565 (10)	0.0349 (7)	0.0027 (7)	-0.0032 (6)	-0.0035 (7)
C6	0.0578 (9)	0.0539 (9)	0.0346 (7)	0.0087 (7)	0.0121 (6)	0.0041 (7)
C7	0.0385 (7)	0.0448 (8)	0.0437 (8)	0.0023 (6)	0.0125 (6)	0.0086 (6)
C8	0.0248 (5)	0.0350 (7)	0.0350 (6)	0.0040 (5)	0.0025 (5)	0.0023 (5)
C9	0.0269 (6)	0.0396 (7)	0.0260 (6)	-0.0056 (5)	0.0062 (4)	-0.0013 (5)
O4	0.0416 (6)	0.0568 (7)	0.0506 (6)	-0.0155 (5)	0.0009 (5)	0.0044 (5)

*Geometric parameters (Å, °)*

O1—C1	1.3541 (16)	C2—H2	0.9800
O1—C8	1.3965 (16)	C3—C8	1.3746 (18)
O2—C1	1.2213 (17)	C3—C4	1.3919 (19)
O3—C9	1.2449 (15)	C4—C5	1.381 (2)
N1—C1	1.3300 (16)	C4—H4	0.9300
N1—C2	1.4558 (17)	C5—C6	1.380 (2)
N1—H1	0.9026	C5—H5	0.9300
N2—C9	1.3567 (17)	C6—C7	1.379 (2)
N2—C2	1.4527 (18)	C6—H6	0.9300
N2—H2A	0.9010	C7—C8	1.3853 (19)
N3—C9	1.3350 (19)	C7—H7	0.9300
N3—H3A	0.9321	O4—H4A	0.9390
N3—H3B	0.9572	O4—H4B	0.9361
C2—C3	1.5032 (18)		

C1—O1—C8	120.28 (10)	C4—C3—C2	121.70 (12)
C1—N1—C2	125.32 (11)	C5—C4—C3	120.75 (14)
C1—N1—H1	111.6	C5—C4—H4	119.6
C2—N1—H1	118.3	C3—C4—H4	119.6
C9—N2—C2	122.67 (10)	C6—C5—C4	119.86 (14)
C9—N2—H2A	118.4	C6—C5—H5	120.1
C2—N2—H2A	118.7	C4—C5—H5	120.1
C9—N3—H3A	118.1	C7—C6—C5	120.49 (14)
C9—N3—H3B	122.1	C7—C6—H6	119.8
H3A—N3—H3B	117.0	C5—C6—H6	119.8
O2—C1—N1	124.29 (12)	C6—C7—C8	118.70 (14)
O2—C1—O1	116.96 (12)	C6—C7—H7	120.7
N1—C1—O1	118.66 (11)	C8—C7—H7	120.6
N2—C2—N1	112.46 (11)	C3—C8—C7	122.14 (13)
N2—C2—C3	113.30 (11)	C3—C8—O1	121.28 (11)
N1—C2—C3	108.97 (10)	C7—C8—O1	116.57 (12)
N2—C2—H2	107.3	O3—C9—N3	122.31 (12)
N1—C2—H2	107.3	O3—C9—N2	120.61 (12)
C3—C2—H2	107.3	N3—C9—N2	117.07 (11)
C8—C3—C4	118.06 (12)	H4A—O4—H4B	108.6
C8—C3—C2	120.24 (11)		
C2—N1—C1—O2	-161.10 (14)	C3—C4—C5—C6	0.2 (2)
C2—N1—C1—O1	22.4 (2)	C4—C5—C6—C7	-0.5 (3)
C8—O1—C1—O2	-179.29 (12)	C5—C6—C7—C8	0.3 (2)
C8—O1—C1—N1	-2.56 (18)	C4—C3—C8—C7	-0.5 (2)
C9—N2—C2—N1	-64.31 (15)	C2—C3—C8—C7	179.31 (12)
C9—N2—C2—C3	59.82 (15)	C4—C3—C8—O1	-179.47 (12)
C1—N1—C2—N2	98.81 (15)	C2—C3—C8—O1	0.31 (19)
C1—N1—C2—C3	-27.67 (17)	C6—C7—C8—C3	0.2 (2)
N2—C2—C3—C8	-110.70 (13)	C6—C7—C8—O1	179.23 (13)
N1—C2—C3—C8	15.31 (16)	C1—O1—C8—C3	-8.08 (18)
N2—C2—C3—C4	69.07 (15)	C1—O1—C8—C7	172.86 (12)
N1—C2—C3—C4	-164.92 (12)	C2—N2—C9—O3	6.81 (19)
C8—C3—C4—C5	0.3 (2)	C2—N2—C9—N3	-173.09 (11)
C2—C3—C4—C5	-179.52 (14)		

Hydrogen-bond geometry ( $\text{\AA}$ ,  $^\circ$ )

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N1—H1 $\cdots$ O2 <sup>i</sup>	0.90	2.05	2.946 (2)	174
N2—H2A $\cdots$ O3 <sup>ii</sup>	0.90	2.41	3.088 (2)	132
N3—H3A $\cdots$ O4 <sup>iii</sup>	0.93	2.14	3.054 (2)	169
N3—H3B $\cdots$ O4	0.96	1.99	2.923 (2)	166

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O4—H4A···O2 <sup>iv</sup>	0.94	1.91	2.774 (2)	153
O4—H4B···O3 <sup>v</sup>	0.94	1.94	2.842 (2)	161

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Symmetry codes: (i)  $-x+1, -y+2, -z+1$ ; (ii)  $x-1, y, z$ ; (iii)  $x+1, y, z$ ; (iv)  $x-1, y-1, z$ ; (v)  $-x+1, -y+1, -z+1$ .