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Pirimicarb: 2-dimethylamino-5,6-dimethylpyrimidin-4-yl dimethylcarbamate

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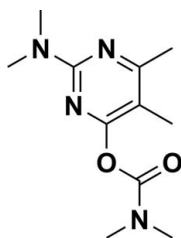
Received 28 June 2010; accepted 7 July 2010

Key indicators: single-crystal X-ray study; $T = 173$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.048; wR factor = 0.146; data-to-parameter ratio = 19.3.

In the title compound, $\text{C}_{11}\text{H}_{18}\text{N}_4\text{O}_2$ (systematic name: 2-dimethylamino-5,6-dimethylpyrimidin-4-yl *N,N*-dimethylcarbamate), the pyrimidine ring and dimethylamino group are almost in the same plane, making a dihedral angle of 1.6 (1)°. The dihedral angle between the mean plane of the pyrimidine ring and that of the dimethylcarbamate group is 83.42 (5)°. In the crystal structure, intermolecular $\text{C}-\text{H}\cdots\text{O}$ hydrogen bonds contribute to the stabilization of the packing.

Related literature

For the toxicity and insecticidal properties of the title compound, see: Pirisi *et al.* (1996). For related structures, see: Dalpozzo *et al.* (2001); Madre *et al.* (2008).



Experimental

Crystal data

$\text{C}_{11}\text{H}_{18}\text{N}_4\text{O}_2$

$M_r = 238.29$

Monoclinic, $P2_1/c$
 $a = 13.5607$ (7) Å
 $b = 7.7868$ (4) Å
 $c = 13.1323$ (7) Å
 $\beta = 114.907$ (3)°
 $V = 1257.72$ (11) Å³

$Z = 4$
 Mo $K\alpha$ radiation
 $\mu = 0.09$ mm⁻¹
 $T = 173$ K
 $0.29 \times 0.25 \times 0.11$ mm

Data collection

Bruker APEXII CCD diffractometer
 Absorption correction: multi-scan (SADABS; Sheldrick, 1996)
 $T_{\min} = 0.975$, $T_{\max} = 0.990$

11979 measured reflections
 3093 independent reflections
 2390 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.031$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.048$
 $wR(F^2) = 0.146$
 $S = 1.05$
 3093 reflections

160 parameters
 H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.27$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.26$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{C5}-\text{H5C}\cdots\text{O2}^i$	0.98	2.60	3.549 (2)	163
$\text{C10}-\text{H10C}\cdots\text{O2}^{ii}$	0.98	2.51	3.431 (2)	157

Symmetry codes: (i) $-x + 1, -y + 1, -z + 1$; (ii) $x, -y + \frac{3}{2}, z + \frac{1}{2}$.

Data collection: APEX2 (Bruker, 2006); cell refinement: SAINT (Bruker, 2006); data reduction: SAINT; program(s) used to solve structure: SHELXTL (Sheldrick, 2008); program(s) used to refine structure: SHELXTL; molecular graphics: SHELXTL and DIAMOND (Brandenburg, 1998); software used to prepare material for publication: SHELXTL.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SJ5033).

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supporting information

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Pirimicarb: 2-dimethylamino-5,6-dimethylpyrimidin-4-yl dimethylcarbamate**Hojin Yang, Tae Ho Kim, Yong Woon Shin, Ki-Min Park and Jineun Kim****S1. Comment**

Pirimicarb (systematic name: 2-dimethylamino-5,6-dimethylpyrimidin-4-yl dimethylcarbamate), is a well known insecticide used to control aphids on vegetable, cereal and orchard crops by inhibiting acetylcholinesterase activity (Pirisi *et al.*, 1996). However its crystal structure has not been reported yet.

In the title compound (Scheme 1, Fig.1), the pyrimidyl ring and dimethylamino group lie in the same plane with a dihedral angle of 1.6 (1)°. This coplanarity may be assisted by the conjugation of π -electrons between pyrimidyl group and nitrogen atom of dimethylamino group. The dihedral angle between the mean plane of the pyrimidyl ring (C1/N1/C2/N2/C3/C4) and the mean plane of the carbamate (O1/O2/C9/N4) is 83.42 (5)° (Fig.1). All bond lengths and bond angles are normal and comparable to those observed in similar structures (Dalpozzo *et al.*, 2001; Madre *et al.*, 2008).

In the crystal structure, weak C—H \cdots O hydrogen bonds are observed [C5—H5C \cdots O2; H5C \cdots O2 = 2.60 Å; C5—H5C \cdots O2 = 163°; C5 \cdots O2 = 3.549 (2) Å; $-x + 1, -y + 1, -z + 1$ and C10—H10C \cdots O2; H10C \cdots O2 = 2.51 Å; C10—H10C \cdots O2 = 157°; C10 \cdots O2 = 3.431 (2) Å; $x, -y + 3/2, z + 1/2$] (Fig. 2).

S2. Experimental

The title compound was purchased from the Dr. Ehrenstorfer GmbH Company. Slow evaporation of a solution in CH₂Cl₂ gave single crystals suitable for X-ray analysis.

S3. Refinement

All H-atoms were positioned geometrically and refined using a riding model with $d(\text{C—H}) = 0.98$ Å, $U_{\text{iso}} = 1.5U_{\text{eq}}(\text{C})$ for the H atoms of the methyl groups.

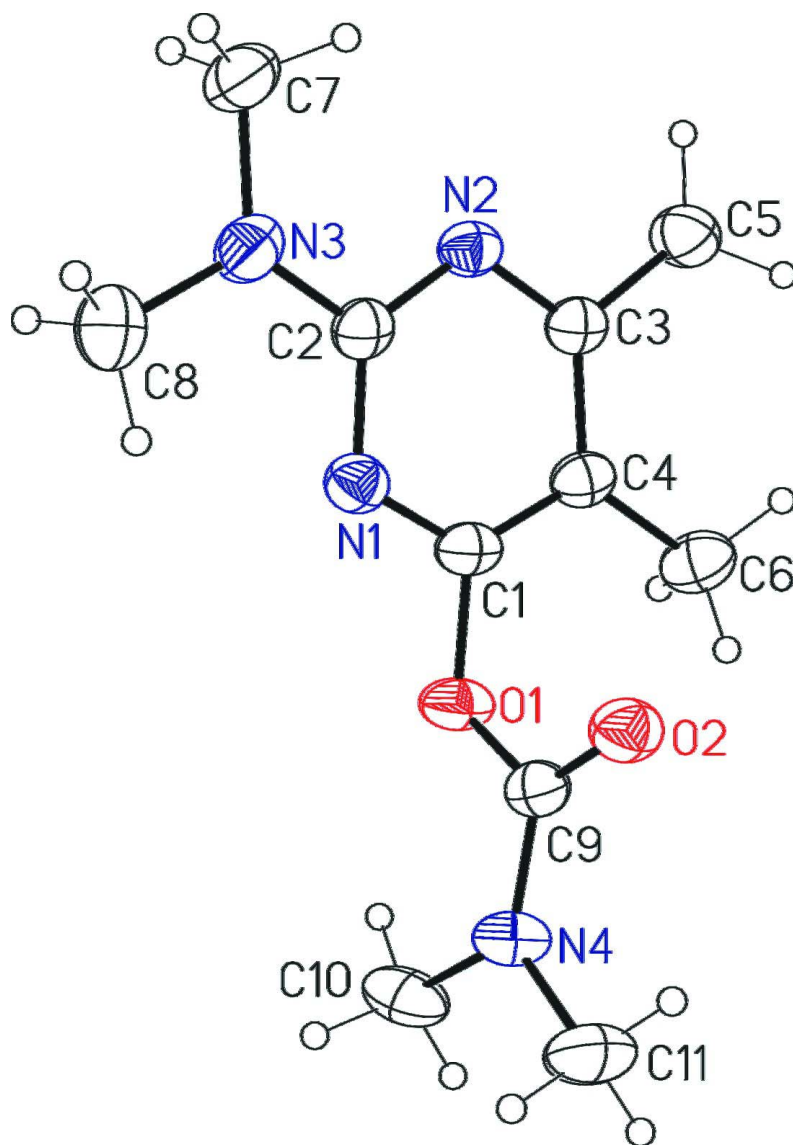


Figure 1

The molecular structure of the title compound, showing displacement ellipsoids drawn at the 50% probability level.

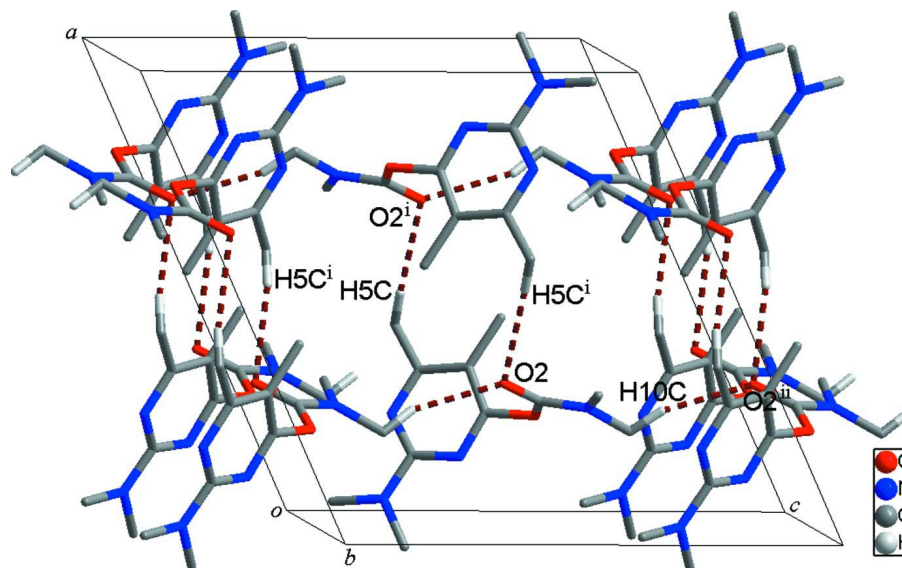


Figure 2

Crystal packing of the title compound with hydrogen bonds shown as dashed lines. H atoms not involved in intermolecular interactions have been omitted for clarity.

2-dimethylamino-5,6-dimethylpyrimidin-4-yl *N,N*-dimethylcarbamate

Crystal data

$C_{11}H_{18}N_4O_2$

$M_r = 238.29$

Monoclinic, $P2_1/c$

Hall symbol: $-P\ 2_1/c$

$a = 13.5607\ (7)\ \text{\AA}$

$b = 7.7868\ (4)\ \text{\AA}$

$c = 13.1323\ (7)\ \text{\AA}$

$\beta = 114.907\ (3)^\circ$

$V = 1257.72\ (11)\ \text{\AA}^3$

$Z = 4$

$F(000) = 512$

$D_x = 1.258\ \text{Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 3964 reflections

$\theta = 3.1\text{--}28.1^\circ$

$\mu = 0.09\ \text{mm}^{-1}$

$T = 173\ \text{K}$

Block, colorless

$0.29 \times 0.25 \times 0.11\ \text{mm}$

Data collection

Bruker APEXII CCD

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

Detector resolution: $10.0\ \text{pixels mm}^{-1}$

φ and ω scans

Absorption correction: multi-scan

(*SADABS*; Sheldrick, 1996)

$T_{\min} = 0.975$, $T_{\max} = 0.990$

11979 measured reflections

3093 independent reflections

2390 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.031$

$\theta_{\max} = 28.3^\circ$, $\theta_{\min} = 1.7^\circ$

$h = -17 \rightarrow 18$

$k = -10 \rightarrow 10$

$l = -15 \rightarrow 17$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.048$

$wR(F^2) = 0.146$

$S = 1.05$

3093 reflections

160 parameters

0 restraints

Primary atom site location: structure-invariant
direct methods

Secondary atom site location: difference Fourier
map

Hydrogen site location: inferred from
neighbouring sites
H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.0757P)^2 + 0.3375P]$$

where $P = (F_o^2 + 2F_c^2)/3$

$$(\Delta/\sigma)_{\max} = 0.001$$

$$\Delta\rho_{\max} = 0.27 \text{ e } \text{\AA}^{-3}$$

$$\Delta\rho_{\min} = -0.26 \text{ e } \text{\AA}^{-3}$$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.24024 (9)	0.53521 (13)	0.53519 (9)	0.0354 (3)
O2	0.32335 (9)	0.74531 (13)	0.48118 (9)	0.0370 (3)
N1	0.14628 (10)	0.43992 (15)	0.35622 (10)	0.0312 (3)
N2	0.22281 (10)	0.23436 (15)	0.27478 (10)	0.0294 (3)
N3	0.05012 (11)	0.34353 (18)	0.17468 (11)	0.0402 (3)
N4	0.27470 (12)	0.78741 (17)	0.62525 (11)	0.0385 (3)
C1	0.23902 (12)	0.43542 (17)	0.44613 (12)	0.0293 (3)
C2	0.14204 (11)	0.33785 (18)	0.27129 (12)	0.0296 (3)
C3	0.31494 (11)	0.23573 (17)	0.36835 (12)	0.0287 (3)
C4	0.32953 (12)	0.33878 (18)	0.46103 (12)	0.0302 (3)
C5	0.40316 (13)	0.1190 (2)	0.36877 (15)	0.0417 (4)
H5A	0.3872	0.0846	0.2916	0.063*
H5B	0.4068	0.0166	0.4137	0.063*
H5C	0.4730	0.1794	0.4013	0.063*
C6	0.43355 (14)	0.3450 (2)	0.56746 (13)	0.0416 (4)
H6A	0.4265	0.2707	0.6243	0.062*
H6B	0.4474	0.4632	0.5955	0.062*
H6C	0.4942	0.3049	0.5516	0.062*
C7	0.03605 (14)	0.2330 (2)	0.08023 (14)	0.0444 (4)
H7A	-0.0065	0.1316	0.0811	0.067*
H7B	0.1074	0.1967	0.0861	0.067*
H7C	-0.0023	0.2961	0.0099	0.067*
C8	-0.04504 (14)	0.4381 (3)	0.16659 (16)	0.0493 (4)
H8A	-0.0234	0.5248	0.2262	0.074*
H8B	-0.0972	0.3587	0.1749	0.074*
H8C	-0.0788	0.4948	0.0932	0.074*
C9	0.28291 (11)	0.69672 (18)	0.54274 (11)	0.0292 (3)
C10	0.22081 (15)	0.7283 (2)	0.69413 (14)	0.0451 (4)
H10A	0.2093	0.6039	0.6851	0.068*
H10B	0.1505	0.7863	0.6705	0.068*

H10C	0.2664	0.7548	0.7731	0.068*
C11	0.31354 (17)	0.9642 (2)	0.64121 (17)	0.0536 (5)
H11A	0.3535	0.9857	0.5953	0.080*
H11B	0.3618	0.9830	0.7205	0.080*
H11C	0.2514	1.0427	0.6185	0.080*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0517 (7)	0.0303 (5)	0.0323 (6)	−0.0064 (4)	0.0258 (5)	−0.0063 (4)
O2	0.0458 (6)	0.0318 (5)	0.0394 (6)	−0.0063 (4)	0.0236 (5)	−0.0044 (4)
N1	0.0350 (6)	0.0273 (6)	0.0343 (7)	−0.0024 (5)	0.0174 (5)	−0.0023 (5)
N2	0.0350 (6)	0.0252 (6)	0.0289 (6)	−0.0022 (5)	0.0145 (5)	−0.0024 (5)
N3	0.0341 (7)	0.0422 (7)	0.0370 (7)	−0.0009 (6)	0.0078 (6)	−0.0073 (6)
N4	0.0487 (8)	0.0343 (7)	0.0350 (7)	0.0003 (6)	0.0200 (6)	−0.0089 (5)
C1	0.0415 (8)	0.0230 (6)	0.0286 (7)	−0.0063 (5)	0.0199 (6)	−0.0022 (5)
C2	0.0327 (7)	0.0258 (6)	0.0315 (7)	−0.0052 (5)	0.0148 (6)	−0.0007 (5)
C3	0.0348 (7)	0.0228 (6)	0.0301 (7)	−0.0024 (5)	0.0151 (6)	0.0010 (5)
C4	0.0371 (7)	0.0259 (7)	0.0275 (7)	−0.0026 (5)	0.0135 (6)	−0.0004 (5)
C5	0.0403 (8)	0.0411 (9)	0.0411 (9)	0.0074 (7)	0.0146 (7)	−0.0054 (7)
C6	0.0461 (9)	0.0405 (8)	0.0305 (8)	0.0011 (7)	0.0088 (7)	−0.0018 (7)
C7	0.0426 (9)	0.0520 (10)	0.0332 (8)	−0.0093 (8)	0.0107 (7)	−0.0084 (7)
C8	0.0358 (9)	0.0564 (11)	0.0496 (10)	0.0051 (7)	0.0120 (8)	0.0006 (8)
C9	0.0311 (7)	0.0275 (7)	0.0259 (7)	0.0024 (5)	0.0090 (6)	−0.0015 (5)
C10	0.0606 (11)	0.0501 (10)	0.0299 (8)	0.0136 (8)	0.0242 (8)	−0.0006 (7)
C11	0.0651 (12)	0.0372 (9)	0.0581 (11)	−0.0025 (8)	0.0257 (10)	−0.0191 (8)

Geometric parameters (Å, °)

O1—C9	1.3704 (18)	C5—H5B	0.9800
O1—C1	1.3985 (16)	C5—H5C	0.9800
O2—C9	1.2128 (18)	C6—H6A	0.9800
N1—C1	1.3138 (19)	C6—H6B	0.9800
N1—C2	1.3510 (18)	C6—H6C	0.9800
N2—C3	1.3333 (19)	C7—H7A	0.9800
N2—C2	1.3449 (19)	C7—H7B	0.9800
N3—C2	1.3540 (19)	C7—H7C	0.9800
N3—C8	1.450 (2)	C8—H8A	0.9800
N3—C7	1.455 (2)	C8—H8B	0.9800
N4—C9	1.3367 (19)	C8—H8C	0.9800
N4—C10	1.456 (2)	C10—H10A	0.9800
N4—C11	1.457 (2)	C10—H10B	0.9800
C1—C4	1.381 (2)	C10—H10C	0.9800
C3—C4	1.401 (2)	C11—H11A	0.9800
C3—C5	1.501 (2)	C11—H11B	0.9800
C4—C6	1.512 (2)	C11—H11C	0.9800
C5—H5A	0.9800		

C9—O1—C1	115.24 (11)	C4—C6—H6C	109.5
C1—N1—C2	114.71 (12)	H6A—C6—H6C	109.5
C3—N2—C2	117.40 (12)	H6B—C6—H6C	109.5
C2—N3—C8	121.66 (14)	N3—C7—H7A	109.5
C2—N3—C7	121.19 (13)	N3—C7—H7B	109.5
C8—N3—C7	116.42 (13)	H7A—C7—H7B	109.5
C9—N4—C10	124.84 (14)	N3—C7—H7C	109.5
C9—N4—C11	117.86 (14)	H7A—C7—H7C	109.5
C10—N4—C11	117.05 (14)	H7B—C7—H7C	109.5
N1—C1—C4	126.79 (13)	N3—C8—H8A	109.5
N1—C1—O1	113.98 (13)	N3—C8—H8B	109.5
C4—C1—O1	119.16 (13)	H8A—C8—H8B	109.5
N2—C2—N1	124.99 (13)	N3—C8—H8C	109.5
N2—C2—N3	117.82 (13)	H8A—C8—H8C	109.5
N1—C2—N3	117.17 (13)	H8B—C8—H8C	109.5
N2—C3—C4	122.69 (13)	O2—C9—N4	126.22 (14)
N2—C3—C5	115.77 (13)	O2—C9—O1	122.26 (12)
C4—C3—C5	121.55 (13)	N4—C9—O1	111.52 (13)
C1—C4—C3	113.39 (13)	N4—C10—H10A	109.5
C1—C4—C6	122.76 (13)	N4—C10—H10B	109.5
C3—C4—C6	123.85 (14)	H10A—C10—H10B	109.5
C3—C5—H5A	109.5	N4—C10—H10C	109.5
C3—C5—H5B	109.5	H10A—C10—H10C	109.5
H5A—C5—H5B	109.5	H10B—C10—H10C	109.5
C3—C5—H5C	109.5	N4—C11—H11A	109.5
H5A—C5—H5C	109.5	N4—C11—H11B	109.5
H5B—C5—H5C	109.5	H11A—C11—H11B	109.5
C4—C6—H6A	109.5	N4—C11—H11C	109.5
C4—C6—H6B	109.5	H11A—C11—H11C	109.5
H6A—C6—H6B	109.5	H11B—C11—H11C	109.5
C2—N1—C1—C4	0.0 (2)	N1—C1—C4—C3	-1.1 (2)
C2—N1—C1—O1	-177.02 (11)	O1—C1—C4—C3	175.74 (11)
C9—O1—C1—N1	-95.73 (15)	N1—C1—C4—C6	178.91 (14)
C9—O1—C1—C4	87.02 (16)	O1—C1—C4—C6	-4.2 (2)
C3—N2—C2—N1	-1.6 (2)	N2—C3—C4—C1	1.0 (2)
C3—N2—C2—N3	176.89 (13)	C5—C3—C4—C1	-178.95 (13)
C1—N1—C2—N2	1.5 (2)	N2—C3—C4—C6	-179.06 (14)
C1—N1—C2—N3	-177.02 (13)	C5—C3—C4—C6	1.0 (2)
C8—N3—C2—N2	173.88 (15)	C10—N4—C9—O2	176.80 (15)
C7—N3—C2—N2	4.0 (2)	C11—N4—C9—O2	2.7 (2)
C8—N3—C2—N1	-7.5 (2)	C10—N4—C9—O1	-3.9 (2)
C7—N3—C2—N1	-177.33 (13)	C11—N4—C9—O1	-178.05 (14)
C2—N2—C3—C4	0.3 (2)	C1—O1—C9—O2	-4.7 (2)
C2—N2—C3—C5	-179.80 (13)	C1—O1—C9—N4	175.97 (12)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
C5—H5C \cdots O2 ⁱ	0.98	2.60	3.549 (2)	163
C10—H10C \cdots O2 ⁱⁱ	0.98	2.51	3.431 (2)	157

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