

Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

## Dibenzoatobis[3-(pyrrol-1-ylmethyl)-pyridine]zinc(II)

Hong Gyu Lee,<sup>a</sup> Jin Hoon Kim,<sup>a</sup> Pan-Gi Kim,<sup>b</sup> Cheal Kim<sup>a\*</sup> and Youngmee Kim<sup>c\*</sup><sup>a</sup>Department of Fine Chemistry and Eco-Product and Materials Education Center, Seoul National University of Technology, Seoul 139-743, Republic of Korea,<sup>b</sup>Department of Forest & Environment Resources, Kyungpook National University, Sangju 742-711, Republic of Korea, and <sup>c</sup>Department of Chemistry and Nano

Science, Ewha Womans University, Seoul 120-750, Republic of Korea

Correspondence e-mail: chealkim@sunt.ac.kr, ymeekim@ewha.ac.kr

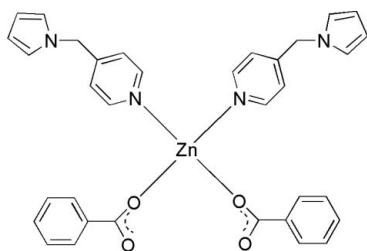
Received 1 July 2010; accepted 16 July 2010

Key indicators: single-crystal X-ray study;  $T = 170$  K; mean  $\sigma(\text{C}-\text{C}) = 0.003$  Å;  $R$  factor = 0.034;  $wR$  factor = 0.071; data-to-parameter ratio = 14.9.

In the title compound,  $[\text{Zn}(\text{C}_7\text{H}_5\text{O}_2)_2(\text{C}_{10}\text{H}_{10}\text{N}_2)_2]$ , the  $\text{Zn}^{\text{II}}$  ion, located on a twofold axis, is coordinated by two N atoms from two 3-(pyrrol-1-ylmethyl)pyridine ligands and two O atoms from two benzoate ligands in a distorted tetrahedral geometry. The pyridine and the pyrrole rings are nearly perpendicular to each other, making a dihedral angle of  $84.83(7)^\circ$ .

## Related literature

For examples of interactions between transition metal ions and biologically active molecules, see: Daniele *et al.* (2008); Parkin (2004); Tshuva & Lippard (2004). For related structures, see: Lee *et al.* (2008); Park *et al.* (2008); Shin *et al.* (2009); Yu *et al.* (2008, 2009, 2010). For a description of the Cambridge Structural Database, see: Allen (2002).



## Experimental

## Crystal data

 $[\text{Zn}(\text{C}_7\text{H}_5\text{O}_2)_2(\text{C}_{10}\text{H}_{10}\text{N}_2)_2]$  $M_r = 623.99$ 

Monoclinic,  $P2_1/c$   
 $a = 14.4347(14)$  Å  
 $b = 9.4399(9)$  Å  
 $c = 11.1959(11)$  Å  
 $\beta = 102.896(2)^\circ$   
 $V = 1487.1(2)$  Å<sup>3</sup>

$Z = 2$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.87$  mm<sup>-1</sup>  
 $T = 170$  K  
 $0.15 \times 0.10 \times 0.03$  mm

## Data collection

Bruker SMART CCD  
 diffractometer  
 8090 measured reflections

2904 independent reflections  
 2204 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.065$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.034$   
 $wR(F^2) = 0.071$   
 $S = 0.91$   
 2904 reflections

195 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.52$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.51$  e Å<sup>-3</sup>

Data collection: *SMART* (Bruker, 1997); cell refinement: *SAINTE* (Bruker, 1997); data reduction: *SAINTE*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-III* (Burnett & Johnson, 1996) and *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *SHELXTL* (Sheldrick, 2008).

Financial support from Korea Ministry Environment "ET-Human resource development Project" and the Cooperative Research Program for Agricultural Science & Technology Development (20070301-036-019-02) is gratefully acknowledged.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: DN2587).

## References

- Allen, F. H. (2002). *Acta Cryst.* **B58**, 380–388.  
 Bruker (1997). *SMART and SAINT*. Bruker AXS Inc., Madison, Wisconsin, USA.  
 Burnett, M. N. & Johnson, C. K. (1996). *ORTEP-III*. Report ORNL-6895. Oak Ridge National Laboratory, Tennessee, USA.  
 Daniele, P. G., Foti, C., Gianguzza, A., Prenesti, E. & Sammartano, S. (2008). *Coord. Chem. Rev.* **252**, 1093–1107.  
 Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.  
 Lee, E. Y., Park, B. K., Kim, C., Kim, S.-J. & Kim, Y. (2008). *Acta Cryst.* **E64**, m286.  
 Park, B. K., Jang, K.-H., Kim, P.-G., Kim, C. & Kim, Y. (2008). *Acta Cryst.* **E64**, m1141.  
 Parkin, G. (2004). *Chem. Rev.* **104**, 699–767.  
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.  
 Shin, D. H., Han, S.-H., Kim, P.-G., Kim, C. & Kim, Y. (2009). *Acta Cryst.* **E65**, m658–m659.  
 Tshuva, E. Y. & Lippard, S. J. (2004). *Chem. Rev.* **104**, 987–1012.  
 Yu, S. M., Koo, K., Kim, P.-G., Kim, C. & Kim, Y. (2010). *Acta Cryst.* **E66**, m61–m62.  
 Yu, S. M., Park, C.-H., Kim, P.-G., Kim, C. & Kim, Y. (2008). *Acta Cryst.* **E64**, m881–m882.  
 Yu, S. M., Shin, D. H., Kim, P.-G., Kim, C. & Kim, Y. (2009). *Acta Cryst.* **E65**, m1045–m1046.

## supporting information

*Acta Cryst.* (2010). E66, m980 [https://doi.org/10.1107/S1600536810028503]

**Dibenzoatobis[3-(pyrrol-1-ylmethyl)pyridine]zinc(II)****Hong Gyu Lee, Jin Hoon Kim, Pan-Gi Kim, Cheal Kim and Youngmee Kim****S1. Comment**

The interaction of transition metal ions with biologically active molecules such as amino acids, proteins, sugars, nucleotides *etc.*, is of great importance in the study of biological systems (Daniele, *et al.*, 2008; Parkin, 2004; Tshuva & Lippard, 2004). As possible models for studying such interaction, the chemistry of transition metal ions with fulvic acids and humic acids has been intensively examined. Our group have reported a variety of structures of copper(II) and zinc(II) benzoates with quinoxaline, 6-methylquinoline, 3-methylquinoline, *trans*-1-(2-pyridyl)-2-(4-pyridyl)ethylene, and di-2-pyridyl ketone (Lee, *et al.*, 2008; Yu, *et al.*, 2008; Park, *et al.*, 2008; Shin, *et al.*, 2009; Yu, *et al.*, 2009; Yu *et al.*, 2010) in order to study the interaction of the transition metal ions with various acids. In this work, we have employed zinc(II) benzoate as a building block and 3-(pyrrol-1-ylmethyl)pyridine as a ligand.

In the title compound, the Zn<sup>II</sup> ion is located on a two fold axis and is coordinated by two nitrogen atoms from two symmetry related 3-(pyrrol-1-ylmethyl)pyridine ligands and two oxygen atoms from two symmetry related benzoate ligands to form a distorted tetrahedral geometry (Fig. 1). Zn—N and Zn—O bond distances are in agreement with reported bond distances in the Cambridge Structural Database (Allen, 2002). The pyridine and the pyrrol rings are nearly perpendicular to each other making a dihedral angle of 84.83 (7)°.

**S2. Experimental**

30.4 mg (0.1 mmol) of Zn(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O and 28.0 mg (0.2 mmol) of C<sub>6</sub>H<sub>5</sub>COONH<sub>4</sub> were dissolved in 4 ml H<sub>2</sub>O and carefully layered by 4 ml acetone solution of 3-(pyrrol-1-ylmethyl)pyridine (31.8 mg, 0.2 mmol). Suitable crystals of the title compound for X-ray analysis were obtained in a few weeks.

**S3. Refinement**

H atoms were placed in calculated positions and treated as riding with C—H distances of 0.95 Å (pyridine and pyrrolidine) and 0.99 Å (methylene) and with U<sub>iso</sub>(H) = 1.2U<sub>eq</sub>(C).

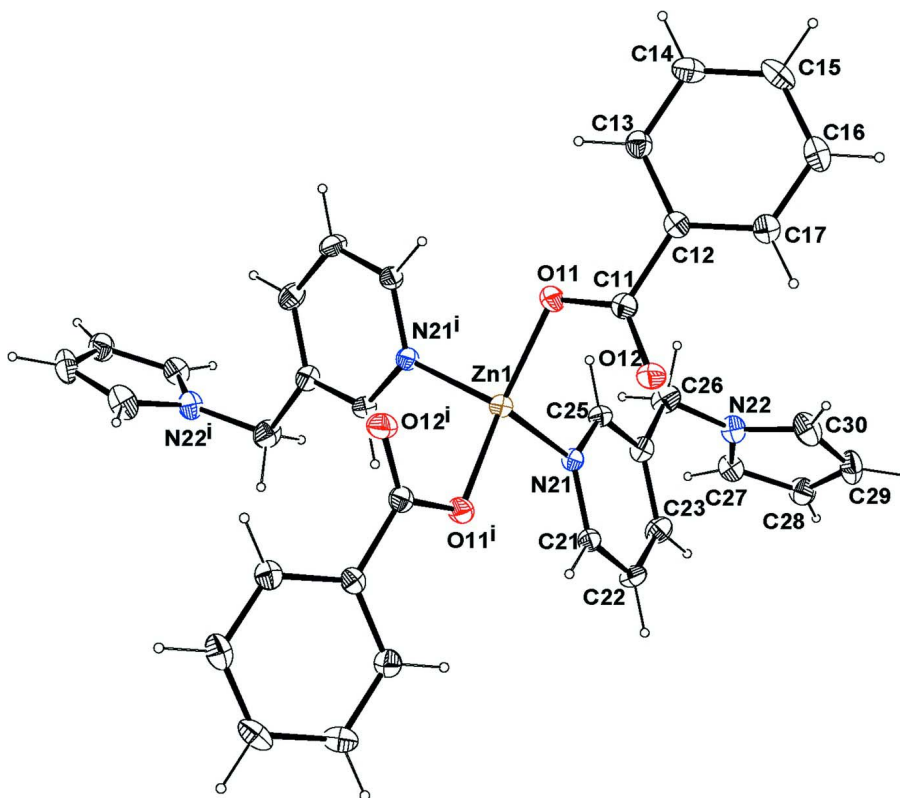


Figure 1

The structure of the title compound showing the atom-labeling scheme. Displacement ellipsoids are shown at the 30% probability level. H atoms are represented as small spheres of arbitrary radii. [Symmetry code: (i) 1-x, y, 3/2-z].

### Dibenzoatobis[3-(pyrrol-1-ylmethyl)pyridine]zinc(II)

#### Crystal data

$[\text{Zn}(\text{C}_7\text{H}_5\text{O}_2)_2(\text{C}_{10}\text{H}_{10}\text{N}_2)_2]$

$M_r = 623.99$

Monoclinic,  $P2/c$

Hall symbol: -P 2yc

$a = 14.4347$  (14) Å

$b = 9.4399$  (9) Å

$c = 11.1959$  (11) Å

$\beta = 102.896$  (2)°

$V = 1487.1$  (2) Å<sup>3</sup>

$Z = 2$

$F(000) = 648$

$D_x = 1.394$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 2361 reflections

$\theta = 2.6\text{--}24.6^\circ$

$\mu = 0.87$  mm<sup>-1</sup>

$T = 170$  K

Plate, colorless

$0.15 \times 0.10 \times 0.03$  mm

#### Data collection

Bruker SMART CCD

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

$\varphi$  and  $\omega$  scans

8090 measured reflections

2904 independent reflections

2204 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.065$

$\theta_{\text{max}} = 26.0^\circ$ ,  $\theta_{\text{min}} = 2.2^\circ$

$h = -9 \rightarrow 17$

$k = -11 \rightarrow 11$

$l = -13 \rightarrow 13$

*Refinement*Refinement on  $F^2$ 

Least-squares matrix: full

 $R[F^2 > 2\sigma(F^2)] = 0.034$  $wR(F^2) = 0.071$  $S = 0.91$ 

2904 reflections

195 parameters

0 restraints

Primary atom site location: structure-invariant  
direct methodsSecondary atom site location: difference Fourier  
mapHydrogen site location: inferred from  
neighbouring sites

H-atom parameters constrained

 $w = 1/[\sigma^2(F_o^2) + (0.0248P)^2]$ where  $P = (F_o^2 + 2F_c^2)/3$  $(\Delta/\sigma)_{\max} < 0.001$  $\Delta\rho_{\max} = 0.52 \text{ e } \text{\AA}^{-3}$  $\Delta\rho_{\min} = -0.51 \text{ e } \text{\AA}^{-3}$ *Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
Zn1	0.5000	0.58373 (4)	0.7500	0.02372 (12)
O11	0.39278 (9)	0.51333 (15)	0.63040 (12)	0.0289 (4)
O12	0.36133 (10)	0.38352 (16)	0.78176 (13)	0.0376 (4)
C11	0.33872 (14)	0.4306 (2)	0.67596 (19)	0.0259 (5)
C12	0.24503 (14)	0.3958 (2)	0.59353 (18)	0.0241 (5)
C13	0.22904 (14)	0.4186 (2)	0.46787 (18)	0.0294 (5)
H13	0.2794	0.4511	0.4330	0.035*
C14	0.14056 (16)	0.3944 (2)	0.3935 (2)	0.0364 (6)
H14	0.1303	0.4094	0.3076	0.044*
C15	0.06679 (16)	0.3482 (2)	0.4441 (2)	0.0395 (6)
H15	0.0056	0.3330	0.3930	0.047*
C16	0.08190 (16)	0.3240 (3)	0.5688 (2)	0.0387 (6)
H16	0.0311	0.2924	0.6034	0.046*
C17	0.17063 (15)	0.3458 (2)	0.6428 (2)	0.0310 (5)
H17	0.1813	0.3266	0.7281	0.037*
N21	0.45103 (11)	0.72476 (17)	0.86239 (14)	0.0238 (4)
N22	0.22689 (12)	1.08801 (19)	0.91482 (16)	0.0316 (4)
C21	0.48598 (14)	0.7216 (2)	0.98335 (18)	0.0258 (5)
H21	0.5324	0.6523	1.0162	0.031*
C22	0.45709 (15)	0.8149 (2)	1.06183 (19)	0.0325 (5)
H22	0.4824	0.8088	1.1476	0.039*
C23	0.39112 (15)	0.9174 (2)	1.01497 (19)	0.0328 (5)
H23	0.3708	0.9829	1.0682	0.039*
C24	0.35458 (14)	0.9243 (2)	0.88950 (19)	0.0283 (5)
C25	0.38588 (14)	0.8249 (2)	0.81762 (18)	0.0258 (5)

H25	0.3602	0.8270	0.7318	0.031*
C26	0.28375 (16)	1.0365 (2)	0.8315 (2)	0.0380 (6)
H26A	0.3184	1.1171	0.8053	0.046*
H26B	0.2410	0.9966	0.7576	0.046*
C27	0.24005 (15)	1.2131 (2)	0.9763 (2)	0.0337 (5)
H27	0.2843	1.2847	0.9675	0.040*
C28	0.17895 (16)	1.2179 (3)	1.0527 (2)	0.0389 (6)
H28	0.1725	1.2932	1.1065	0.047*
C29	0.12722 (16)	1.0912 (3)	1.0374 (2)	0.0443 (6)
H29	0.0790	1.0649	1.0786	0.053*
C30	0.15875 (16)	1.0129 (3)	0.9527 (2)	0.0429 (6)
H30	0.1368	0.9212	0.9248	0.051*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Zn1	0.0232 (2)	0.0247 (2)	0.0228 (2)	0.000	0.00434 (14)	0.000
O11	0.0247 (8)	0.0311 (9)	0.0300 (8)	-0.0069 (7)	0.0045 (7)	-0.0004 (7)
O12	0.0319 (9)	0.0500 (11)	0.0286 (9)	-0.0016 (7)	0.0017 (7)	0.0079 (7)
C11	0.0262 (12)	0.0236 (12)	0.0279 (12)	0.0047 (10)	0.0060 (9)	-0.0044 (10)
C12	0.0244 (11)	0.0186 (11)	0.0291 (12)	0.0026 (9)	0.0053 (9)	-0.0004 (9)
C13	0.0300 (12)	0.0264 (12)	0.0319 (12)	-0.0026 (10)	0.0073 (10)	-0.0006 (10)
C14	0.0398 (14)	0.0333 (14)	0.0307 (13)	-0.0035 (11)	-0.0039 (11)	0.0008 (10)
C15	0.0244 (13)	0.0347 (14)	0.0524 (16)	-0.0013 (11)	-0.0066 (11)	-0.0031 (12)
C16	0.0274 (13)	0.0373 (14)	0.0528 (16)	-0.0030 (11)	0.0121 (12)	-0.0014 (12)
C17	0.0303 (13)	0.0293 (12)	0.0337 (13)	0.0007 (10)	0.0078 (10)	0.0015 (10)
N21	0.0230 (10)	0.0257 (10)	0.0228 (9)	-0.0016 (7)	0.0054 (7)	0.0014 (8)
N22	0.0309 (10)	0.0248 (10)	0.0403 (11)	0.0040 (9)	0.0104 (8)	0.0000 (9)
C21	0.0243 (12)	0.0289 (12)	0.0240 (12)	0.0007 (9)	0.0051 (9)	0.0047 (10)
C22	0.0374 (14)	0.0400 (14)	0.0196 (11)	-0.0015 (11)	0.0052 (10)	0.0004 (10)
C23	0.0416 (14)	0.0297 (12)	0.0290 (12)	0.0012 (11)	0.0121 (10)	-0.0053 (11)
C24	0.0312 (12)	0.0252 (12)	0.0296 (12)	0.0010 (10)	0.0094 (10)	0.0008 (10)
C25	0.0266 (12)	0.0290 (12)	0.0209 (11)	0.0006 (10)	0.0038 (9)	0.0017 (10)
C26	0.0456 (15)	0.0331 (14)	0.0367 (14)	0.0114 (11)	0.0122 (12)	0.0002 (11)
C27	0.0319 (13)	0.0268 (13)	0.0417 (14)	0.0008 (10)	0.0070 (11)	-0.0017 (11)
C28	0.0350 (15)	0.0418 (15)	0.0410 (14)	0.0066 (11)	0.0107 (12)	-0.0063 (12)
C29	0.0298 (13)	0.0471 (16)	0.0604 (17)	0.0038 (12)	0.0197 (12)	0.0083 (14)
C30	0.0353 (15)	0.0275 (14)	0.0649 (18)	-0.0029 (11)	0.0093 (13)	-0.0004 (12)

*Geometric parameters (Å, °)*

Zn1—O11 <sup>i</sup>	1.9248 (13)	N22—C27	1.359 (3)
Zn1—O11	1.9248 (13)	N22—C26	1.457 (3)
Zn1—N21 <sup>i</sup>	2.0621 (16)	C21—C22	1.373 (3)
Zn1—N21	2.0621 (16)	C21—H21	0.9500
O11—C11	1.287 (2)	C22—C23	1.377 (3)
O12—C11	1.239 (2)	C22—H22	0.9500
C11—C12	1.494 (3)	C23—C24	1.387 (3)

C12—C13	1.391 (3)	C23—H23	0.9500
C12—C17	1.395 (3)	C24—C25	1.376 (3)
C13—C14	1.379 (3)	C24—C26	1.514 (3)
C13—H13	0.9500	C25—H25	0.9500
C14—C15	1.383 (3)	C26—H26A	0.9900
C14—H14	0.9500	C26—H26B	0.9900
C15—C16	1.384 (3)	C27—C28	1.359 (3)
C15—H15	0.9500	C27—H27	0.9500
C16—C17	1.377 (3)	C28—C29	1.400 (3)
C16—H16	0.9500	C28—H28	0.9500
C17—H17	0.9500	C29—C30	1.359 (3)
N21—C21	1.336 (2)	C29—H29	0.9500
N21—C25	1.349 (2)	C30—H30	0.9500
N22—C30	1.355 (3)		
O11 <sup>i</sup> —Zn1—O11	139.60 (9)	N21—C21—C22	122.29 (19)
O11 <sup>i</sup> —Zn1—N21 <sup>i</sup>	108.40 (6)	N21—C21—H21	118.9
O11—Zn1—N21 <sup>i</sup>	97.49 (6)	C22—C21—H21	118.9
O11 <sup>i</sup> —Zn1—N21	97.48 (6)	C21—C22—C23	119.3 (2)
O11—Zn1—N21	108.39 (6)	C21—C22—H22	120.3
N21 <sup>i</sup> —Zn1—N21	99.58 (9)	C23—C22—H22	120.3
C11—O11—Zn1	113.47 (13)	C22—C23—C24	119.6 (2)
O12—C11—O11	122.87 (19)	C22—C23—H23	120.2
O12—C11—C12	121.42 (19)	C24—C23—H23	120.2
O11—C11—C12	115.70 (18)	C25—C24—C23	117.37 (19)
C13—C12—C17	118.95 (19)	C25—C24—C26	120.30 (18)
C13—C12—C11	120.95 (18)	C23—C24—C26	122.32 (19)
C17—C12—C11	120.04 (18)	N21—C25—C24	123.58 (18)
C14—C13—C12	120.5 (2)	N21—C25—H25	118.2
C14—C13—H13	119.8	C24—C25—H25	118.2
C12—C13—H13	119.8	N22—C26—C24	112.46 (17)
C13—C14—C15	119.9 (2)	N22—C26—H26A	109.1
C13—C14—H14	120.0	C24—C26—H26A	109.1
C15—C14—H14	120.0	N22—C26—H26B	109.1
C14—C15—C16	120.2 (2)	C24—C26—H26B	109.1
C14—C15—H15	119.9	H26A—C26—H26B	107.8
C16—C15—H15	119.9	N22—C27—C28	108.1 (2)
C17—C16—C15	119.9 (2)	N22—C27—H27	125.9
C17—C16—H16	120.0	C28—C27—H27	125.9
C15—C16—H16	120.0	C27—C28—C29	107.3 (2)
C16—C17—C12	120.5 (2)	C27—C28—H28	126.3
C16—C17—H17	119.8	C29—C28—H28	126.3
C12—C17—H17	119.8	C30—C29—C28	107.3 (2)
C21—N21—C25	117.83 (17)	C30—C29—H29	126.3
C21—N21—Zn1	120.05 (14)	C28—C29—H29	126.3
C25—N21—Zn1	122.09 (13)	N22—C30—C29	108.3 (2)
C30—N22—C27	108.97 (18)	N22—C30—H30	125.9
C30—N22—C26	125.30 (19)	C29—C30—H30	125.9

C27—N22—C26	125.43 (19)		
O11 <sup>i</sup> —Zn1—O11—C11	56.95 (13)	N21 <sup>i</sup> —Zn1—N21—C25	56.41 (14)
N21 <sup>i</sup> —Zn1—O11—C11	-172.84 (13)	C25—N21—C21—C22	0.7 (3)
N21—Zn1—O11—C11	-70.10 (14)	Zn1—N21—C21—C22	178.71 (15)
Zn1—O11—C11—O12	-10.7 (3)	N21—C21—C22—C23	-1.2 (3)
Zn1—O11—C11—C12	168.33 (12)	C21—C22—C23—C24	0.3 (3)
O12—C11—C12—C13	-164.0 (2)	C22—C23—C24—C25	1.0 (3)
O11—C11—C12—C13	17.0 (3)	C22—C23—C24—C26	-178.4 (2)
O12—C11—C12—C17	18.9 (3)	C21—N21—C25—C24	0.8 (3)
O11—C11—C12—C17	-160.13 (18)	Zn1—N21—C25—C24	-177.24 (15)
C17—C12—C13—C14	1.1 (3)	C23—C24—C25—N21	-1.6 (3)
C11—C12—C13—C14	-176.02 (19)	C26—C24—C25—N21	177.83 (19)
C12—C13—C14—C15	0.5 (3)	C30—N22—C26—C24	-70.4 (3)
C13—C14—C15—C16	-1.1 (3)	C27—N22—C26—C24	102.6 (2)
C14—C15—C16—C17	-0.1 (4)	C25—C24—C26—N22	153.59 (19)
C15—C16—C17—C12	1.8 (3)	C23—C24—C26—N22	-27.0 (3)
C13—C12—C17—C16	-2.3 (3)	C30—N22—C27—C28	-0.7 (3)
C11—C12—C17—C16	174.9 (2)	C26—N22—C27—C28	-174.69 (19)
O11 <sup>i</sup> —Zn1—N21—C21	-11.37 (15)	N22—C27—C28—C29	0.2 (3)
O11—Zn1—N21—C21	137.18 (14)	C27—C28—C29—C30	0.3 (3)
N21 <sup>i</sup> —Zn1—N21—C21	-121.55 (16)	C27—N22—C30—C29	0.9 (3)
O11 <sup>i</sup> —Zn1—N21—C25	166.58 (15)	C26—N22—C30—C29	174.90 (19)
O11—Zn1—N21—C25	-44.86 (16)	C28—C29—C30—N22	-0.8 (3)

Symmetry code: (i)  $-x+1, y, -z+3/2$ .