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The trigonal polymorph of strontium tetraborate, β-SrB_4O_7

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Key indicators: single-crystal X-ray study; T = 296 K; mean σ(O–B) = 0.007 Å; R factor = 0.030; wR factor = 0.064; data-to-parameter ratio = 14.0.

The asymmetric unit of the title compound, β-SrB_4O_7, contains five Sr atoms (three located on a threefold rotation axis), twelve B and 21 O atoms. The structure is made up from BO_3 triangles and BO_4 tetrahedra in a 1:1 ratio. Pairs of BO_3 triangles are linked to BO_4 tetrahedra *via* common corners, forming chains. These chains are further linked to adjacent chains through corner-sharing, leading to a three-dimensional framework with channels running parallel to [001]. The Sr^{2+} ions reside in the channels and exhibit strongly distorted polyhedra The density of the β-polymorph is considerably lower than that of α-SrB_4O_7, which is constructed solely from BO_4 tetrahedra.

Related literature

For the orthorhombic α-polymorph, see: Block *et al.* (1964). For the physical properties of this phase, see: Oseledchik *et al.* (1995); Petrov *et al.* (2004); Zaitsev *et al.* (2006); Verwey *et al.* (1992); Machida *et al.* (1979); Pei *et al.* (2000). For other crystalline phases in the system $SrO-B_2O_3$ listed in the ICSD (2009), see: Ross & Angel (1991); Lin *et al.* (1999); Wei *et al.* (2001); Tang *et al.* (2008); Lapshin *et al.* (2007); Kim *et al.* (1996). For glass-phases in this system, see: Imaoka (1959); Polyakova & Litovchik (2008). For a review of B—O bond lengths in BO_3 and BO_4 units, see: Zobetz (1982, 1990).

Experimental

Crystal data

SrB_4O_7
M_r = 242.86
Trigonal, $P3$
a = 17.145 (1) Å
c = 4.2527 (5) Å
V = 1082.61 (16) Å^3
Z = 9
Mo $K\alpha$ radiation
μ = 11.19 mm^{-1}
T = 296 K
0.40 × 0.25 × 0.18 mm

Data collection

Bruker SMART CCD area-detector diffractometer
Absorption correction: multi-scan (*SADABS*; Sheldrick, 2004)
T_{min} = 0.095, T_{max} = 0.242
10350 measured reflections
3709 independent reflections
3202 reflections with $I > 2\sigma(I)$
R_{int} = 0.054

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.030
$wR(F^2)$ = 0.064
S = 0.85
3709 reflections
265 parameters
$\Delta\rho_{max}$ = 1.02 e Å^{-3}
$\Delta\rho_{min}$ = −0.49 e Å^{-3}
Absolute structure: Flack (1983), 1836 Friedel pairs
Flack parameter: −0.030 (7)

Data collection: *SMART* (Bruker, 2001); cell refinement: *SAINT-Plus* (Bruker, 2001); data reduction: *SAINT-Plus*; program(s) used to solve structure: *SHELXTL* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: WM2345).

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S1. Comment

The orthorhombic phase of strontium tetraborate, α-SrB_4O_7 (I), is known for a long time (Block *et al.*, 1964). This compound has attracted attention owing to its interesting physical properties, namely an unprecedented fundamental optical-absorption edge among oxide compounds (~130 nm), high non-linear optical coefficients (Oseledchik *et al.*, 1995; Petrov *et al.*, 2004; Zaitsev *et al.*, 2006), good luminescent characteristics and an ability to stabilize rare-earth elements in divalent state (Verwey *et al.*, 1992; Pei *et al.*, 2000; Machida *et al.*, 1979).

SrB_4O_7 falls in a glass-forming range within the SrO—B_2O_3 system and can simply be obtained as a glass (Imaoka, 1959). The process of glass re-crystallization occurs through complex mechanisms with probabilistic formation of other crystalline phases, specifically of metastable phases. Such a phase was in fact observed and designated as β-SrB_4O_7 (Polyakova & Litovchik, 2008). However, X-ray powder diffraction data of this phase and of two other new compounds described by these authors were not analysed because of impure samples.

The FIZ/NIST Inorganic Crystal Structure Database (release 2009; ICSD, 2009) reveals six phases in the SrO—B_2O_3 system besides (I): strontium diborate, (IIa), SrB_2O_4 (Kim *et al.*, 1996), its high-pressure form (Ross & Angel, 1991), (IIb), distrontium diborate, $Sr_2B_2O_5$, (III), (Lin *et al.*, 1999), tristrontium tetraborate, $Sr_3B_2O_6$, (IV), (Wei *et al.*, 2001), distrontium hexadecaborate, $Sr_2B_{16}O_{26}$, (V), (Tang *et al.*, 2008) and tetrastrontium tetradecaborate, $Sr_4B_{14}O_{25}$, (VI), (Lapshin *et al.*, 2007). Only two of these phases crystallize in non-centrosymmetric space groups, *viz*. (I, $Pmn2_1$ and VI, $Cmc2_1$). The main feature of all these structures are BO_x units (x = 3,4). Isolated (IV) or flat pairs (III) of BO_3 triangles, a framework of BO_4 tetrahedra with shared vertices (I, II) and a framework of triangles and tetrahedra with shared vertices (V, VI) are found in these structures.

In the process of glass re-crystallization of a strontium tetraborate composition at 973–983 K during one day, we obtained β-SrB_4O_7 crystals with dimensions of ~200–400 μm. The crystals were located on the glass surface and were optically homogeneous (i.e. crystals showed homogeneous extinction when observed under a polarizing microscope). The crystals possess strong anisotropy when abrased, and crystals with an elongated ellipsoidal shape were obtained in such a way. As it turned out, the crystals are most firm along the *c*-direction.

The crystal structure of (I) is built up from a three-dimensional framework of connected boron-oxygen tetrahedra. The asymmetric unit of the title structure contains five Sr (three on special positions), twelve B and 21 O atoms. Alternatively, the structural formula of the title compound can thus be written as $Sr_3B_{12}O_{21}$. It consists of BO_3 triangles and BO_4 tetrahedra in an 1:1 ratio (3:1 for structure V and 3:4 for structure VI). They form a three-dimensional framework constructed *via* common vertices. The BO_3 triangles are linked to one another so that two of their vertices and the bridging O atom are located on a straight line (see O1, O3, O5; O6, O7, O10 and O11, O13, O15 in Fig.1). The plane of one triangle in such a pair is tilted relatively to the other one about the line with an angle of ~20°. The remaining two vertices are common with the same tetrahedron (e.g. see O2 and O4 in Fig.1). The BO_4 tetrahedra are connected to one another *via* common vertices and form chains along the *c*-direction (Fig. 2). These chains are connected with pairs of BO_3

triangles, leading to the formation of channels in the structure. The channels are filled with strontium ions (Fig. 3). The coordination polyhedra around the strontium ions are non-regular and defined by six O atoms in the range 2.479 (3)–2.786 (3) Å when a distance < 2.8 Å is considered as relevant.

All vertices in the anionic framework are shared so that every oxygen atom is connected to two boron atoms. The B—O distances fall into the interval 1.323 (6)–1.420 (6)Å (average is 1.367 (6) Å) for BO_3 triangles and into the interval 1.425 (6)–1.538 (6) Å (average is 1.474 (6) Å) for BO_4 tetrahedra. These values compare well with the mean bond lengths calculated for various borate structures (Zobetz, 1982, 1990).

In comparison with α-SrB_4O_7 which is constructued solely from BO_4 tetrahedra, the density of the β-polymorph is considerably lower.

S2. Experimental

Crystals were extracted out of glass by careful dissolving of the latter in a 2% HNO_3 solution. The initial glass has been made from a mixture of $SrCO_3$ (99.8%) and H_3BO_3 (99.98%) in a 1:4 ratio. The mixture was heated up to 353—363 K with addition of a small amount of water and careful mixing until CO_2 gas evolution had stopped. Then the temperature was increased slowly up to 573 K to yield a anhydrous phase. The derived mixture was then placed into a glass-carbon crucible and kept in a molten state at 1323 K during 6 h in a nitrogen atmosphere. The flux was cooled in air down to 773 K and the glass was finally annealed at 723 K during a day to remove strain.

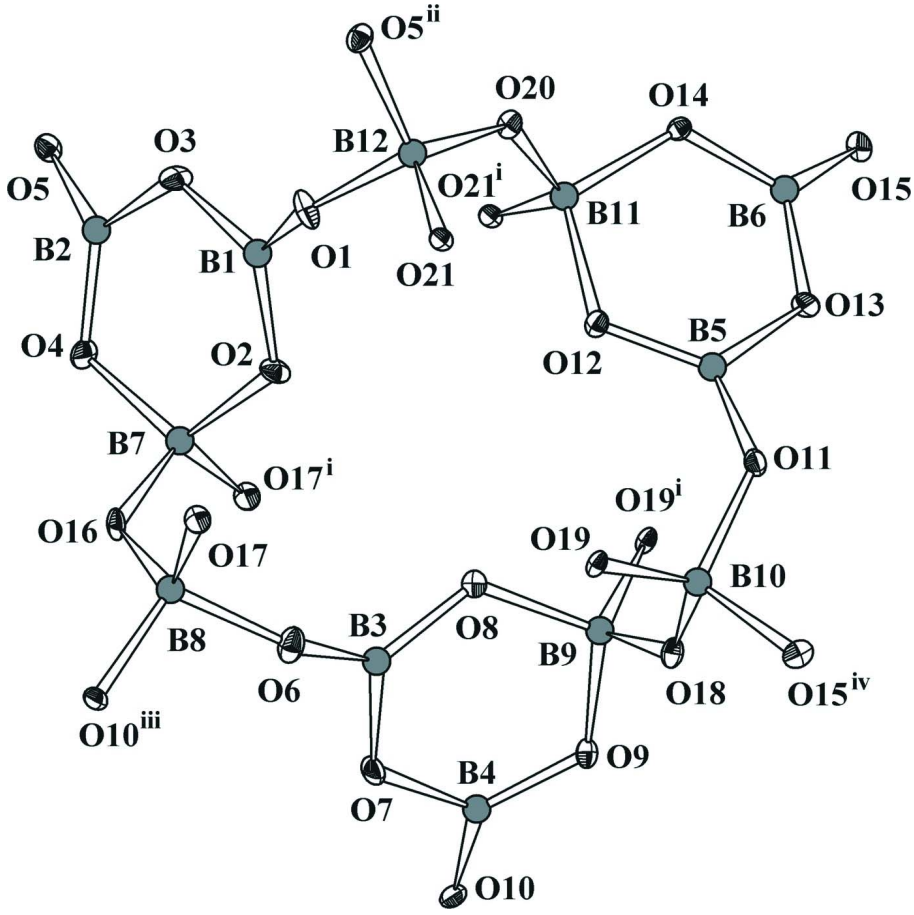

Figure 1

View of the basic structural motif in the title compound with the atom numbering scheme. Displacement ellipsoids are drawn at the 50% probability level. Symmetry codes: (i) x, y, z-1; (ii) 1-y, 1+x-y, 1+z; (iii) -y, x-y, 1+z; (iv) 1-y, x-y, 1+z.

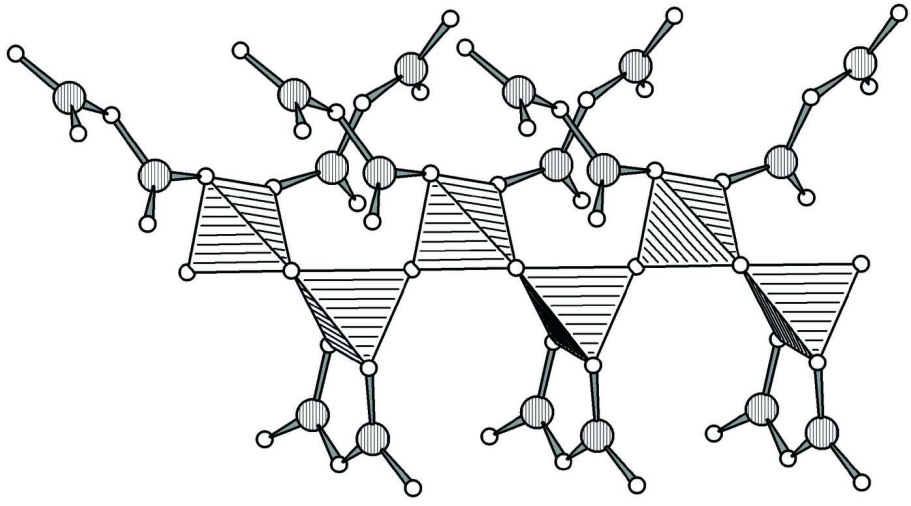

Figure 2

Columns of BO_4 tetrahedra with shared vertices and attached BO_3 triangles extending along the *c*-direction.

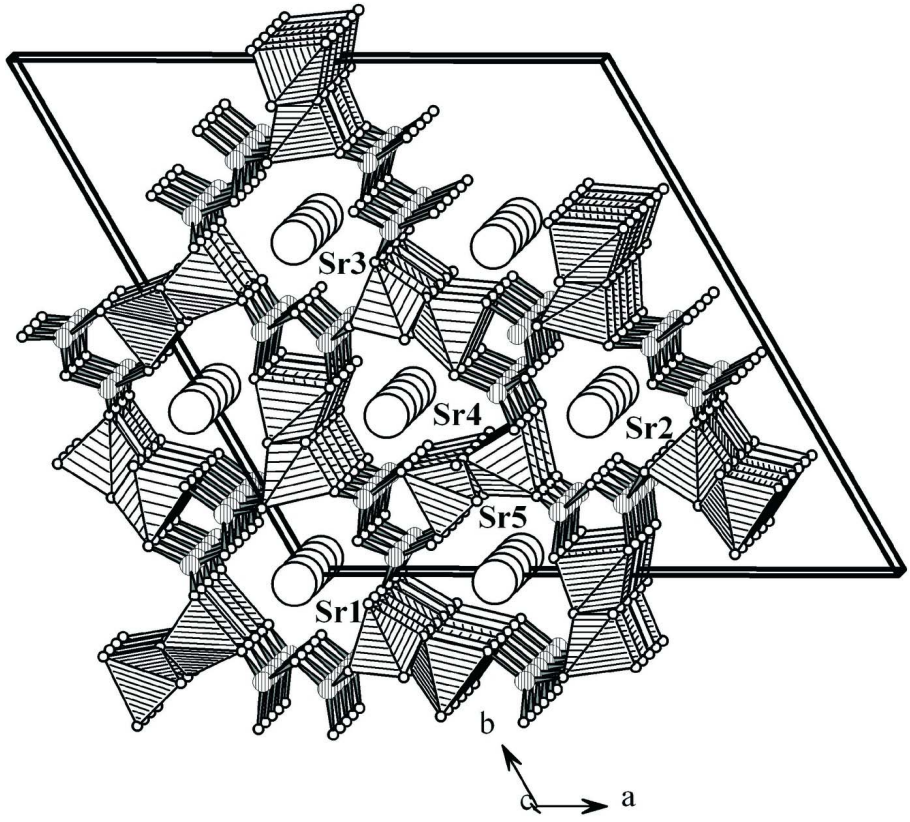

Figure 3

View down [001] of the framework structure of the title compound showing the formation of channels that are filled with Sr^{2+} ions.

strontium tetraborate

Crystal data

SrB_4O_7
M_r = 242.86
Trigonal, $P3$
Hall symbol: P 3
a = 17.145 (1) Å
c = 4.2527 (5) Å
V = 1082.61 (16) Å^3
Z = 9
$F(000)$ = 1026
D_x = 3.353 (1) Mg m^{-3}
Mo $K\alpha$ radiation, λ = 0.71073 Å
Cell parameters from 2820 reflections
θ = 2.4–29.3°
μ = 11.19 mm^{-1}
T = 296 K
Ellipsoidal, colorless
0.40 × 0.25 × 0.18 mm

Data collection

Bruker SMART CCD area-detector diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
φ and ω scans
Absorption correction: multi-scan (*SADABS*; Sheldrick, 2004)
T_{min} = 0.095, T_{max} = 0.242
10350 measured reflections
3709 independent reflections
3202 reflections with $I > 2\sigma(I)$
R_{int} = 0.054
θ_{max} = 28.7°, θ_{min} = 2.4°
h = −22→23
k = −23→23
l = −5→5

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)] = 0.030$
$wR(F^2) = 0.064$
$S = 0.85$
3709 reflections
265 parameters
0 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map

$w = 1/[\sigma^2(F_o^2) + (0.P)^2]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} < 0.001$
$\Delta\rho_{max} = 1.02$ e Å^{-3}
$\Delta\rho_{min} = -0.49$ e Å^{-3}
Extinction correction: *SHELXTL* (Sheldrick, 2008), $Fc^* = kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
Extinction coefficient: 0.0794 (13)
Absolute structure: Flack (1983), 1836 Friedel pairs
Absolute structure parameter: −0.030 (7)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor wR and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) etc. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)

	x	*y*	*z*	U_{iso}*/U_{eq}
Sr1	0.0000	0.0000	0.5000	0.0106 (2)
Sr2	0.6667	0.3333	0.7047 (2)	0.0097 (2)
Sr3	0.3333	0.6667	0.6084 (5)	0.01172 (15)
Sr4	0.32408 (4)	0.33158 (4)	0.7001 (4)	0.00732 (14)
Sr5	−0.00569 (3)	0.33541 (4)	0.5183 (4)	0.00743 (15)
O1	0.3441 (2)	0.5031 (2)	0.5634 (9)	0.0115 (7)
O2	0.2472 (2)	0.3798 (2)	0.2476 (8)	0.0094 (7)
O3	0.2591 (2)	0.5240 (2)	0.1833 (8)	0.0108 (7)
O4	0.1239 (2)	0.3939 (2)	0.0012 (9)	0.0109 (7)
O5	0.1760 (2)	0.5441 (2)	−0.1939 (8)	0.0090 (7)
O6	0.1571 (2)	0.1803 (2)	0.4810 (9)	0.0119 (8)
O7	0.2841 (2)	0.2150 (2)	0.1685 (8)	0.0075 (6)
O8	0.1428 (2)	0.0790 (2)	0.0952 (8)	0.0109 (7)
O9	0.2767 (2)	0.0826 (2)	−0.0917 (8)	0.0083 (7)
O10	0.1278 (2)	−0.0221 (2)	−0.2863 (8)	0.0095 (7)
O11	0.5201 (2)	0.3007 (2)	0.4426 (8)	0.0082 (7)
O12	0.4709 (2)	0.4061 (2)	0.3108 (8)	0.0077 (7)
O13	0.6000 (2)	0.4102 (2)	0.0726 (8)	0.0105 (7)
O14	0.5899 (2)	0.5436 (2)	0.0455 (8)	0.0065 (6)
O15	0.6814 (2)	0.5195 (2)	−0.3106 (8)	0.0082 (7)
O16	0.0899 (2)	0.2724 (2)	0.3487 (8)	0.0084 (7)
O17	0.1632 (2)	0.2814 (2)	0.8498 (8)	0.0085 (7)
O18	0.3864 (2)	0.1604 (2)	0.2995 (8)	0.0082 (7)

O19	0.3945 (2)	0.2362 (2)	0.7992 (8)	0.0062 (6)
O20	0.5007 (2)	0.5578 (2)	0.4336 (8)	0.0073 (7)
O21	0.4320 (2)	0.4828 (2)	0.9349 (8)	0.0067 (7)
B1	0.2862 (4)	0.4681 (4)	0.3246 (14)	0.0079 (10)*
B2	0.1811 (3)	0.4849 (3)	−0.0018 (13)	0.0059 (10)*
B3	0.1948 (3)	0.1620 (3)	0.2422 (13)	0.0075 (10)*
B4	0.1853 (4)	0.0444 (3)	−0.0934 (13)	0.0071 (10)*
B5	0.5250 (3)	0.3704 (3)	0.2794 (12)	0.0052 (9)*
B6	0.6277 (3)	0.4939 (3)	−0.0547 (13)	0.0063 (10)*
B7	0.1537 (4)	0.3283 (3)	0.1104 (13)	0.0048 (10)*
B8	0.1082 (3)	0.2266 (3)	0.5981 (12)	0.0052 (10)*
B9	0.3394 (4)	0.1748 (4)	0.0454 (12)	0.0064 (11)*
B10	0.4409 (3)	0.2160 (3)	0.5607 (12)	0.0063 (10)*
B11	0.4960 (4)	0.4991 (4)	0.1768 (13)	0.0076 (12)*
B12	0.4379 (3)	0.5431 (3)	0.6829 (12)	0.0061 (10)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Sr1	0.0098 (3)	0.0098 (3)	0.0121 (5)	0.00490 (15)	0.000	0.000
Sr2	0.0093 (3)	0.0093 (3)	0.0106 (5)	0.00466 (16)	0.000	0.000
Sr3	0.0100 (2)	0.0100 (2)	0.0151 (4)	0.00502 (10)	0.000	0.000
Sr4	0.0072 (3)	0.0066 (3)	0.0083 (3)	0.0036 (2)	0.0001 (3)	0.0014 (2)
Sr5	0.0072 (3)	0.0076 (3)	0.0082 (3)	0.0043 (2)	0.0007 (2)	0.0015 (2)
O1	0.0055 (16)	0.0191 (19)	0.0082 (17)	0.0048 (15)	−0.0004 (14)	−0.0017 (14)
O2	0.0069 (16)	0.0074 (16)	0.0109 (17)	0.0014 (13)	−0.0039 (13)	0.0001 (13)
O3	0.0099 (16)	0.0065 (16)	0.0141 (18)	0.0027 (14)	−0.0075 (14)	−0.0033 (14)
O4	0.0090 (16)	0.0103 (17)	0.0153 (19)	0.0062 (14)	−0.0051 (14)	−0.0006 (14)
O5	0.0074 (16)	0.0100 (17)	0.0091 (18)	0.0039 (14)	−0.0009 (13)	0.0018 (13)
O6	0.0139 (19)	0.0164 (18)	0.0101 (18)	0.0110 (16)	0.0039 (14)	0.0042 (15)
O7	0.0080 (16)	0.0090 (16)	0.0058 (16)	0.0046 (14)	−0.0001 (13)	−0.0004 (13)
O8	0.0065 (16)	0.0099 (16)	0.0148 (19)	0.0029 (14)	0.0034 (14)	−0.0030 (14)
O9	0.0076 (16)	0.0109 (16)	0.0081 (16)	0.0059 (14)	0.0012 (13)	−0.0006 (13)
O10	0.0087 (17)	0.0060 (16)	0.0148 (19)	0.0044 (14)	−0.0046 (14)	−0.0022 (14)
O11	0.0061 (16)	0.0095 (16)	0.0093 (17)	0.0042 (14)	0.0016 (13)	0.0018 (13)
O12	0.0090 (16)	0.0090 (16)	0.0069 (16)	0.0057 (14)	0.0031 (13)	0.0025 (13)
O13	0.0104 (17)	0.0078 (16)	0.0125 (18)	0.0039 (14)	0.0087 (13)	0.0050 (13)
O14	0.0077 (15)	0.0062 (15)	0.0060 (16)	0.0039 (13)	0.0001 (13)	−0.0009 (12)
O15	0.0086 (18)	0.0114 (18)	0.0048 (17)	0.0053 (15)	0.0040 (14)	0.0046 (14)
O16	0.0069 (16)	0.0130 (17)	0.0075 (16)	0.0066 (14)	0.0036 (13)	0.0026 (13)
O17	0.0082 (16)	0.0092 (16)	0.0077 (17)	0.0041 (14)	−0.0004 (13)	−0.0023 (13)
O18	0.0081 (16)	0.0112 (16)	0.0079 (17)	0.0068 (14)	−0.0016 (13)	0.0003 (13)
O19	0.0089 (15)	0.0083 (16)	0.0043 (16)	0.0065 (13)	0.0021 (12)	0.0017 (12)
O20	0.0099 (16)	0.0099 (16)	0.0038 (15)	0.0061 (14)	0.0024 (13)	0.0000 (12)
O21	0.0057 (15)	0.0066 (15)	0.0076 (17)	0.0029 (13)	−0.0009 (12)	0.0029 (13)

Geometric parameters (Å, º)

Sr1—O10^{i}	2.569 (3)	Sr5—B11^{xii}	3.171 (6)
Sr1—O10^{ii}	2.569 (3)	O1—B1	1.335 (6)
Sr1—O10^{iii}	2.569 (3)	O1—B12	1.486 (6)
Sr1—O8^{iv}	2.734 (3)	O2—B1	1.354 (6)
Sr1—O8^{v}	2.734 (3)	O2—B7	1.509 (6)
Sr1—O8	2.734 (3)	O2—Sr4^{xiv}	2.988 (3)
Sr1—O6	2.913 (4)	O3—B1	1.392 (6)
Sr1—O6^{v}	2.913 (4)	O3—B2	1.401 (6)
Sr1—O6^{iv}	2.913 (4)	O3—Sr3^{xiv}	3.236 (3)
Sr1—B3	3.285 (5)	O4—B2	1.366 (6)
Sr1—B3^{iv}	3.285 (5)	O4—B7	1.521 (6)
Sr1—B3^{v}	3.285 (5)	O4—Sr5^{xiv}	2.816 (3)
Sr2—O11^{vi}	2.542 (3)	O5—B2	1.340 (6)
Sr2—O11^{vii}	2.542 (3)	O5—B12^{xv}	1.489 (6)
Sr2—O11	2.542 (3)	O5—Sr3^{xiv}	2.594 (3)
Sr2—O13^{viii}	2.644 (3)	O6—B3	1.323 (6)
Sr2—O13^{ix}	2.644 (3)	O6—B8	1.499 (6)
Sr2—O13^{ii}	2.644 (3)	O7—B3	1.370 (6)
Sr2—O15^{viii}	3.074 (3)	O7—B9	1.517 (6)
Sr2—O15^{ix}	3.074 (3)	O7—Sr4^{xiv}	2.657 (3)
Sr2—O15^{ii}	3.074 (3)	O8—B3	1.394 (6)
Sr2—B6^{viii}	3.303 (5)	O8—B4	1.399 (6)
Sr2—B6^{ix}	3.303 (5)	O8—Sr1^{xiv}	3.304 (3)
Sr2—B6^{ii}	3.303 (5)	O9—B4	1.364 (6)
Sr3—O5^{x}	2.594 (3)	O9—B9	1.516 (6)
Sr3—O5^{ii}	2.594 (3)	O9—Sr5^{xvi}	2.676 (3)
Sr3—O5^{xi}	2.594 (3)	O10—B4	1.349 (6)
Sr3—O3^{xii}	2.786 (3)	O10—B8^{xvi}	1.487 (5)
Sr3—O3	2.786 (3)	O10—Sr1^{xiv}	2.569 (3)
Sr3—O3^{xiii}	2.786 (3)	O11—B5	1.348 (6)
Sr3—O1^{xii}	2.908 (3)	O11—B10	1.494 (6)
Sr3—O1^{xiii}	2.908 (3)	O12—B5	1.350 (6)
Sr3—O1	2.908 (3)	O12—B11	1.537 (6)
Sr3—O3^{x}	3.236 (3)	O13—B6	1.377 (6)
Sr3—O3^{xi}	3.236 (4)	O13—B5	1.419 (6)
Sr3—O3^{ii}	3.236 (4)	O13—Sr2^{xiv}	2.644 (3)
Sr4—O19	2.507 (3)	O14—B6	1.371 (6)
Sr4—O21	2.519 (3)	O14—B11	1.503 (6)
Sr4—O17	2.526 (3)	O14—Sr5^{xiii}	2.657 (3)
Sr4—O7^{ii}	2.657 (3)	O14—Sr5^{xvii}	2.836 (3)
Sr4—O2	2.685 (3)	O15—B6	1.349 (6)
Sr4—O12	2.737 (3)	O15—B10^{xviii}	1.501 (6)
Sr4—O1	2.845 (4)	O15—Sr5^{xvii}	2.649 (3)
Sr4—O7	2.864 (3)	O15—Sr2^{xiv}	3.074 (3)
Sr4—O6	2.893 (4)	O16—B8	1.445 (6)
Sr4—O2^{ii}	2.988 (3)	O16—B7	1.447 (6)

Sr4—B12	3.145 (5)	O17—B8	1.425 (6)
Sr4—B1	3.158 (5)	O17—B7^{ii}	1.426 (6)
Sr5—O16	2.479 (3)	O18—B9	1.441 (6)
Sr5—O20^{xii}	2.482 (3)	O18—B10	1.458 (6)
Sr5—O18^{v}	2.539 (3)	O18—Sr5^{iv}	2.539 (3)
Sr5—O15^{x}	2.649 (3)	O19—B10	1.434 (6)
Sr5—O14^{xii}	2.657 (3)	O19—B9^{ii}	1.450 (5)
Sr5—O9^{i}	2.676 (3)	O20—B12	1.440 (6)
Sr5—O4^{ii}	2.816 (3)	O20—B11	1.460 (6)
Sr5—O14^{x}	2.836 (3)	O20—Sr5^{xiii}	2.482 (3)
Sr5—O4	2.924 (4)	O21—B11^{ii}	1.425 (6)
Sr5—B10^{v}	3.129 (5)	O21—B12	1.458 (6)
Sr5—B6^{x}	3.149 (5)		

O10^{i}—Sr1—O10^{ii}	108.20 (8)	O7—Sr4—O2^{ii}	145.41 (8)
O10^{i}—Sr1—O10^{iii}	108.20 (8)	O6—Sr4—O2^{ii}	97.10 (9)
O10^{ii}—Sr1—O10^{iii}	108.20 (8)	O16—Sr5—O20^{xii}	116.34 (10)
O10^{i}—Sr1—O8^{iv}	155.68 (11)	O16—Sr5—O18^{v}	104.21 (10)
O10^{ii}—Sr1—O8^{iv}	94.57 (10)	O20^{xii}—Sr5—O18^{v}	119.04 (10)
O10^{iii}—Sr1—O8^{iv}	71.17 (10)	O16—Sr5—O15^{x}	152.34 (10)
O10^{i}—Sr1—O8^{v}	71.17 (10)	O20^{xii}—Sr5—O15^{x}	90.34 (10)
O10^{ii}—Sr1—O8^{v}	155.68 (11)	O18^{v}—Sr5—O15^{x}	52.72 (10)
O10^{iii}—Sr1—O8^{v}	94.57 (10)	O16—Sr5—O14^{xii}	112.32 (11)
O8^{iv}—Sr1—O8^{v}	84.57 (11)	O20^{xii}—Sr5—O14^{xii}	54.17 (10)
O10^{i}—Sr1—O8	94.57 (10)	O18^{v}—Sr5—O14^{xii}	69.52 (10)
O10^{ii}—Sr1—O8	71.17 (10)	O15^{x}—Sr5—O14^{xii}	76.45 (10)
O10^{iii}—Sr1—O8	155.68 (11)	O16—Sr5—O9^{i}	85.40 (10)
O8^{iv}—Sr1—O8	84.57 (11)	O20^{xii}—Sr5—O9^{i}	148.56 (12)
O8^{v}—Sr1—O8	84.57 (11)	O18^{v}—Sr5—O9^{i}	72.45 (10)
O10^{i}—Sr1—O6	49.69 (9)	O15^{x}—Sr5—O9^{i}	73.60 (10)
O10^{ii}—Sr1—O6	76.23 (9)	O14^{xii}—Sr5—O9^{i}	140.89 (9)
O10^{iii}—Sr1—O6	156.14 (11)	O16—Sr5—O4^{ii}	77.96 (10)
O8^{iv}—Sr1—O6	132.54 (10)	O20^{xii}—Sr5—O4^{ii}	83.34 (10)
O8^{v}—Sr1—O6	86.46 (9)	O18^{v}—Sr5—O4^{ii}	151.29 (11)
O8—Sr1—O6	48.18 (9)	O15^{x}—Sr5—O4^{ii}	114.40 (11)
O10^{i}—Sr1—O6^{v}	76.23 (9)	O14^{xii}—Sr5—O4^{ii}	136.93 (9)
O10^{ii}—Sr1—O6^{v}	156.14 (11)	O9^{i}—Sr5—O4^{ii}	79.29 (10)
O10^{iii}—Sr1—O6^{v}	49.69 (9)	O16—Sr5—O14^{x}	142.36 (10)
O8^{iv}—Sr1—O6^{v}	86.46 (9)	O20^{xii}—Sr5—O14^{x}	70.68 (10)
O8^{v}—Sr1—O6^{v}	48.18 (9)	O18^{v}—Sr5—O14^{x}	102.82 (9)
O8—Sr1—O6^{v}	132.54 (10)	O15^{x}—Sr5—O14^{x}	50.85 (9)
O6—Sr1—O6^{v}	119.923 (7)	O14^{xii}—Sr5—O14^{x}	101.42 (10)
O10^{i}—Sr1—O6^{iv}	156.14 (11)	O9^{i}—Sr5—O14^{x}	78.32 (9)
O10^{ii}—Sr1—O6^{iv}	49.69 (9)	O4^{ii}—Sr5—O14^{x}	65.84 (10)
O10^{iii}—Sr1—O6^{iv}	76.23 (9)	O16—Sr5—O4	51.39 (10)
O8^{iv}—Sr1—O6^{iv}	48.18 (9)	O20^{xii}—Sr5—O4	71.17 (10)
O8^{v}—Sr1—O6^{iv}	132.54 (10)	O18^{v}—Sr5—O4	108.16 (10)
O8—Sr1—O6^{iv}	86.46 (9)	O15^{x}—Sr5—O4	142.90 (10)

O6—Sr1—O6^{iv}	119.923 (8)	O14^{xii}—Sr5—O4	66.60 (9)
O6^{v}—Sr1—O6^{iv}	119.923 (8)	O9^{i}—Sr5—O4	136.26 (9)
O11^{vi}—Sr2—O11^{vii}	102.21 (9)	O4^{ii}—Sr5—O4	95.60 (10)
O11^{vi}—Sr2—O11	102.21 (9)	O14^{x}—Sr5—O4	139.09 (9)
O11^{vii}—Sr2—O11	102.21 (9)	B1—O1—B12	149.0 (4)
O11^{vi}—Sr2—O13^{viii}	94.41 (10)	B1—O1—Sr4	90.7 (3)
O11^{vii}—Sr2—O13^{viii}	75.13 (10)	B12—O1—Sr4	87.2 (2)
O11—Sr2—O13^{viii}	163.33 (10)	B1—O1—Sr3	95.2 (3)
O11^{vi}—Sr2—O13^{ix}	75.13 (10)	B12—O1—Sr3	96.2 (3)
O11^{vii}—Sr2—O13^{ix}	163.33 (10)	Sr4—O1—Sr3	161.90 (14)
O11—Sr2—O13^{ix}	94.41 (10)	B1—O2—B7	122.2 (4)
O13^{viii}—Sr2—O13^{ix}	88.58 (11)	B1—O2—Sr4	97.3 (3)
O11^{vi}—Sr2—O13^{ii}	163.33 (10)	B7—O2—Sr4	129.7 (3)
O11^{vii}—Sr2—O13^{ii}	94.41 (10)	B1—O2—Sr4^{xiv}	117.3 (3)
O11—Sr2—O13^{ii}	75.13 (10)	B7—O2—Sr4^{xiv}	91.4 (3)
O13^{viii}—Sr2—O13^{ii}	88.58 (11)	Sr4—O2—Sr4^{xiv}	96.96 (10)
O13^{ix}—Sr2—O13^{ii}	88.58 (11)	B1—O3—B2	118.8 (4)
O11^{vi}—Sr2—O15^{viii}	48.77 (9)	B1—O3—Sr3	99.3 (3)
O11^{vii}—Sr2—O15^{viii}	77.18 (10)	B2—O3—Sr3	133.8 (3)
O11—Sr2—O15^{viii}	148.54 (9)	B1—O3—Sr3^{xiv}	134.9 (3)
O13^{viii}—Sr2—O15^{viii}	47.70 (9)	B2—O3—Sr3^{xiv}	81.8 (3)
O13^{ix}—Sr2—O15^{viii}	89.33 (9)	Sr3—O3—Sr3^{xiv}	89.54 (8)
O13^{ii}—Sr2—O15^{viii}	136.27 (10)	B2—O4—B7	122.2 (4)
O11^{vi}—Sr2—O15^{ix}	77.18 (10)	B2—O4—Sr5^{xiv}	112.8 (3)
O11^{vii}—Sr2—O15^{ix}	148.54 (9)	B7—O4—Sr5^{xiv}	117.1 (3)
O11—Sr2—O15^{ix}	48.77 (9)	B2—O4—Sr5	112.8 (3)
O13^{viii}—Sr2—O15^{ix}	136.27 (10)	B7—O4—Sr5	89.7 (2)
O13^{ix}—Sr2—O15^{ix}	47.70 (9)	Sr5^{xiv}—O4—Sr5	95.60 (10)
O13^{ii}—Sr2—O15^{ix}	89.33 (9)	B2—O5—B12^{xv}	137.7 (4)
O15^{viii}—Sr2—O15^{ix}	119.955 (5)	B2—O5—Sr3^{xiv}	112.1 (3)
O11^{vi}—Sr2—O15^{ii}	148.54 (9)	B12^{xv}—O5—Sr3^{xiv}	110.1 (2)
O11^{vii}—Sr2—O15^{ii}	48.77 (9)	B3—O6—B8	148.2 (4)
O11—Sr2—O15^{ii}	77.18 (10)	B3—O6—Sr4	94.5 (3)
O13^{viii}—Sr2—O15^{ii}	89.33 (9)	B8—O6—Sr4	89.0 (2)
O13^{ix}—Sr2—O15^{ii}	136.27 (10)	B3—O6—Sr1	94.1 (3)
O13^{ii}—Sr2—O15^{ii}	47.70 (9)	B8—O6—Sr1	95.1 (2)
O15^{viii}—Sr2—O15^{ii}	119.955 (6)	Sr4—O6—Sr1	156.60 (14)
O15^{ix}—Sr2—O15^{ii}	119.955 (5)	B3—O7—B9	121.5 (4)
O5^{x}—Sr3—O5^{ii}	110.02 (8)	B3—O7—Sr4^{xiv}	116.9 (3)
O5^{x}—Sr3—O5^{xi}	110.02 (8)	B9—O7—Sr4^{xiv}	95.6 (2)
O5^{ii}—Sr3—O5^{xi}	110.02 (8)	B3—O7—Sr4	94.6 (3)
O5^{x}—Sr3—O3^{xii}	70.47 (10)	B9—O7—Sr4	127.1 (3)
O5^{ii}—Sr3—O3^{xii}	94.31 (9)	Sr4^{xiv}—O7—Sr4	100.68 (10)
O5^{xi}—Sr3—O3^{xii}	152.90 (11)	B3—O8—B4	119.2 (4)
O5^{x}—Sr3—O3	152.90 (11)	B3—O8—Sr1	100.4 (3)
O5^{ii}—Sr3—O3	70.47 (10)	B4—O8—Sr1	132.7 (3)
O5^{xi}—Sr3—O3	94.31 (9)	B3—O8—Sr1^{xiv}	136.6 (3)
O3^{xii}—Sr3—O3	82.43 (11)	B4—O8—Sr1^{xiv}	79.8 (3)

O5^{x}—Sr3—O3^{xiii}	94.31 (9)	Sr1—O8—Sr1^{xiv}	89.02 (8)
O5^{ii}—Sr3—O3^{xiii}	152.90 (11)	B4—O9—B9	123.2 (4)
O5^{xi}—Sr3—O3^{xiii}	70.47 (10)	B4—O9—Sr5^{xvi}	114.8 (3)
O3^{xii}—Sr3—O3^{xiii}	82.43 (11)	B9—O9—Sr5^{xvi}	120.0 (2)
O3—Sr3—O3^{xiii}	82.43 (11)	B4—O10—B8^{xvi}	135.4 (4)
O5^{x}—Sr3—O1^{xii}	77.41 (9)	B4—O10—Sr1^{xiv}	114.0 (3)
O5^{ii}—Sr3—O1^{xii}	49.47 (9)	B8^{xvi}—O10—Sr1^{xiv}	110.6 (3)
O5^{xi}—Sr3—O1^{xii}	158.70 (11)	B5—O11—B10	131.2 (4)
O3^{xii}—Sr3—O1^{xii}	48.06 (9)	B5—O11—Sr2	114.5 (3)
O3—Sr3—O1^{xii}	84.14 (10)	B10—O11—Sr2	112.6 (3)
O3^{xiii}—Sr3—O1^{xii}	129.98 (11)	B5—O12—B11	122.3 (4)
O5^{x}—Sr3—O1^{xiii}	49.47 (9)	B5—O12—Sr4	121.9 (3)
O5^{ii}—Sr3—O1^{xiii}	158.70 (10)	B11—O12—Sr4	115.1 (3)
O5^{xi}—Sr3—O1^{xiii}	77.41 (9)	B6—O13—B5	120.1 (4)
O3^{xii}—Sr3—O1^{xiii}	84.14 (10)	B6—O13—Sr2^{xiv}	106.1 (3)
O3—Sr3—O1^{xiii}	129.98 (11)	B5—O13—Sr2^{xiv}	129.8 (3)
O3^{xiii}—Sr3—O1^{xiii}	48.06 (9)	B6—O14—B11	121.3 (4)
O1^{xii}—Sr3—O1^{xiii}	119.572 (18)	B6—O14—Sr5^{xiii}	119.7 (3)
O5^{x}—Sr3—O1	158.70 (11)	B11—O14—Sr5^{xiii}	95.3 (3)
O5^{ii}—Sr3—O1	77.41 (9)	B6—O14—Sr5^{xvii}	89.9 (3)
O5^{xi}—Sr3—O1	49.47 (9)	B11—O14—Sr5^{xvii}	130.0 (3)
O3^{xii}—Sr3—O1	129.98 (11)	Sr5^{xiii}—O14—Sr5^{xvii}	101.42 (10)
O3—Sr3—O1	48.06 (9)	B6—O15—B10^{xviii}	147.5 (4)
O3^{xiii}—Sr3—O1	84.14 (9)	B6—O15—Sr5^{xvii}	98.7 (3)
O1^{xii}—Sr3—O1	119.572 (19)	B10^{xviii}—O15—Sr5^{xvii}	93.8 (2)
O1^{xiii}—Sr3—O1	119.572 (19)	B6—O15—Sr2^{xiv}	87.5 (3)
O5^{x}—Sr3—O3^{x}	44.45 (9)	B10^{xviii}—O15—Sr2^{xiv}	89.5 (2)
O5^{ii}—Sr3—O3^{x}	68.82 (9)	Sr5^{xvii}—O15—Sr2^{xiv}	162.45 (13)
O5^{xi}—Sr3—O3^{x}	109.89 (10)	B8—O16—B7	125.3 (4)
O3^{xii}—Sr3—O3^{x}	89.54 (8)	B8—O16—Sr5	113.2 (3)
O3—Sr3—O3^{x}	137.69 (4)	B7—O16—Sr5	111.1 (3)
O3^{xiii}—Sr3—O3^{x}	137.69 (4)	B8—O17—B7^{ii}	137.0 (4)
O1^{xii}—Sr3—O3^{x}	60.73 (9)	B8—O17—Sr4	106.5 (3)
O1^{xiii}—Sr3—O3^{x}	89.91 (9)	B7^{ii}—O17—Sr4	114.7 (3)
O1—Sr3—O3^{x}	129.76 (10)	B9—O18—B10	133.6 (4)
O5^{x}—Sr3—O3^{xi}	68.82 (9)	B9—O18—Sr5^{iv}	123.3 (3)
O5^{ii}—Sr3—O3^{xi}	109.89 (10)	B10—O18—Sr5^{iv}	99.5 (3)
O5^{xi}—Sr3—O3^{xi}	44.45 (9)	B10—O19—B9^{ii}	125.1 (4)
O3^{xii}—Sr3—O3^{xi}	137.69 (4)	B10—O19—Sr4	123.0 (3)
O3—Sr3—O3^{xi}	137.69 (4)	B9^{ii}—O19—Sr4	104.1 (3)
O3^{xiii}—Sr3—O3^{xi}	89.54 (8)	B12—O20—B11	130.6 (4)
O1^{xii}—Sr3—O3^{xi}	129.76 (10)	B12—O20—Sr5^{xiii}	120.1 (3)
O1^{xiii}—Sr3—O3^{xi}	60.73 (9)	B11—O20—Sr5^{xiii}	104.0 (3)
O1—Sr3—O3^{xi}	89.91 (9)	B11^{ii}—O21—B12	128.2 (4)
O3^{x}—Sr3—O3^{xi}	69.12 (9)	O1—B1—O2	121.8 (5)
O5^{x}—Sr3—O3^{ii}	109.89 (10)	O1—B1—O3	116.7 (4)
O5^{ii}—Sr3—O3^{ii}	44.45 (9)	O2—B1—O3	121.1 (4)
O5^{xi}—Sr3—O3^{ii}	68.82 (9)	O5—B2—O4	126.2 (4)

O3^{xii}—Sr3—O3^{ii}	137.69 (4)	O5—B2—O3	112.9 (4)
O3—Sr3—O3^{ii}	89.54 (8)	O4—B2—O3	120.6 (4)
O3^{xiii}—Sr3—O3^{ii}	137.69 (4)	O6—B3—O7	122.4 (4)
O1^{xii}—Sr3—O3^{ii}	89.91 (9)	O6—B3—O8	116.6 (4)
O1^{xiii}—Sr3—O3^{ii}	129.76 (10)	O7—B3—O8	120.5 (4)
O1—Sr3—O3^{ii}	60.73 (9)	O10—B4—O9	125.9 (4)
O3^{x}—Sr3—O3^{ii}	69.12 (9)	O10—B4—O8	113.1 (4)
O3^{xi}—Sr3—O3^{ii}	69.12 (9)	O9—B4—O8	120.7 (4)
O19—Sr4—O21	105.01 (10)	O11—B5—O12	126.8 (4)
O19—Sr4—O17	122.38 (10)	O11—B5—O13	112.6 (4)
O21—Sr4—O17	111.46 (10)	O12—B5—O13	120.4 (4)
O19—Sr4—O7^{ii}	53.90 (9)	O15—B6—O14	120.4 (4)
O21—Sr4—O7^{ii}	104.90 (10)	O15—B6—O13	118.6 (4)
O17—Sr4—O7^{ii}	74.20 (10)	O14—B6—O13	120.1 (4)
O19—Sr4—O2	142.35 (11)	O17^{xiv}—B7—O16	115.6 (4)
O21—Sr4—O2	100.85 (10)	O17^{xiv}—B7—O2	104.0 (4)
O17—Sr4—O2	70.44 (10)	O16—B7—O2	110.3 (4)
O7^{ii}—Sr4—O2	141.90 (9)	O17^{xiv}—B7—O4	110.8 (4)
O19—Sr4—O12	77.05 (9)	O16—B7—O4	106.4 (4)
O21—Sr4—O12	74.24 (10)	O2—B7—O4	109.7 (4)
O17—Sr4—O12	154.69 (11)	O17—B8—O16	116.4 (4)
O7^{ii}—Sr4—O12	129.48 (9)	O17—B8—O10^{i}	111.8 (4)
O2—Sr4—O12	84.31 (10)	O16—B8—O10^{i}	109.5 (4)
O19—Sr4—O1	149.31 (10)	O17—B8—O6	103.8 (4)
O21—Sr4—O1	50.97 (9)	O16—B8—O6	112.0 (4)
O17—Sr4—O1	87.17 (10)	O10^{i}—B8—O6	102.2 (3)
O7^{ii}—Sr4—O1	141.53 (10)	O18—B9—O19^{xiv}	116.7 (4)
O2—Sr4—O1	50.18 (10)	O18—B9—O9	105.9 (4)
O12—Sr4—O1	77.65 (9)	O19^{xiv}—B9—O9	110.2 (4)
O19—Sr4—O7	72.72 (9)	O18—B9—O7	110.4 (4)
O21—Sr4—O7	145.82 (10)	O19^{xiv}—B9—O7	104.4 (3)
O17—Sr4—O7	96.98 (10)	O9—B9—O7	109.2 (4)
O7^{ii}—Sr4—O7	100.68 (10)	O19—B10—O18	116.4 (4)
O2—Sr4—O7	70.49 (10)	O19—B10—O11	110.3 (4)
O12—Sr4—O7	72.05 (9)	O18—B10—O11	110.7 (4)
O1—Sr4—O7	115.02 (10)	O19—B10—O15^{ix}	111.6 (4)
O19—Sr4—O6	93.03 (10)	O18—B10—O15^{ix}	102.4 (3)
O21—Sr4—O6	159.89 (10)	O11—B10—O15^{ix}	104.5 (3)
O17—Sr4—O6	49.69 (10)	O21^{xiv}—B11—O20	116.5 (4)
O7^{ii}—Sr4—O6	78.44 (10)	O21^{xiv}—B11—O14	110.8 (4)
O2—Sr4—O6	68.12 (10)	O20—B11—O14	104.6 (4)
O12—Sr4—O6	119.38 (10)	O21^{xiv}—B11—O12	106.0 (4)
O1—Sr4—O6	114.52 (9)	O20—B11—O12	109.2 (4)
O7—Sr4—O6	48.38 (9)	O14—B11—O12	109.7 (4)
O19—Sr4—O2^{ii}	118.18 (10)	O20—B12—O21	116.5 (4)
O21—Sr4—O2^{ii}	66.73 (10)	O20—B12—O1	111.7 (4)
O17—Sr4—O2^{ii}	48.74 (9)	O21—B12—O1	104.4 (4)
O7^{ii}—Sr4—O2^{ii}	68.92 (9)	O20—B12—O5^{xi}	109.0 (4)

O2—Sr4—O2^{ii}	96.96 (10)	O21—B12—O5^{xi}	111.8 (4)
O12—Sr4—O2^{ii}	140.51 (9)	O1—B12—O5^{xi}	102.5 (3)
O1—Sr4—O2^{ii}	73.46 (9)		

Symmetry codes: (i) $-y, x-y, z+1$; (ii) $x, y, z+1$; (iii) $-x+y, -x, z+1$; (iv) $-x+y, -x, z$; (v) $-y, x-y, z$; (vi) $-y+1, x-y, z$; (vii) $-x+y+1, -x+1, z$; (viii) $-x+y+1, -x+1, z+1$; (ix) $-y+1, x-y, z+1$; (x) $-x+y, -x+1, z+1$; (xi) $-y+1, x-y+1, z+1$; (xii) $-x+y, -x+1, z$; (xiii) $-y+1, x-y+1, z$; (xiv) $x, y, z-1$; (xv) $-x+y, -x+1, z-1$; (xvi) $-x+y, -x, z-1$; (xvii) $-y+1, x-y+1, z-1$; (xviii) $-x+y+1, -x+1, z-1$.