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(3*S*,*NR*)-3-Hydroxymethyl-2-methyl-2-(3-methylbut-2-en-1-yl)-1,2,3,4-tetrahydroisoquinolinium bromide–1,1'-bi-2-naphthol (1/1)

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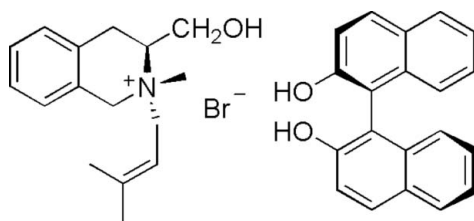
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Key indicators: single-crystal X-ray study; $T = 153$ K; mean $\sigma(\text{C}-\text{C}) = 0.008$ Å; R factor = 0.045; wR factor = 0.180; data-to-parameter ratio = 16.4.

In the title compound, $\text{C}_{16}\text{H}_{24}\text{NO}^+\cdot\text{Br}^-\cdot\text{C}_{20}\text{H}_{14}\text{O}_2$, the *N*-heterocyclic six-membered ring assumes a half-chair conformation. The two naphthalene ring systems are nearly perpendicular to one another, making a dihedral angle of 89.5 (2)°. Intermolecular $\text{O}-\text{H}\cdots\text{Br}$ hydrogen bonding helps to stabilize the crystal structure.

Related literature

For the optical properties of binaphthalen-2-ol, see: Tayama & Tanaka (2007). For related structures with nearly perpendicular naphthyl rings, see: Fukushima *et al.* (1999); Mori *et al.* (1993). For the synthesis, see: Schultz *et al.* (1998); Tayama & Tanaka (2007).



Experimental

Crystal data

$\text{C}_{16}\text{H}_{24}\text{NO}^+\cdot\text{C}_{20}\text{H}_{14}\text{O}_2\cdot\text{Br}^-$
 $\text{C}_{20}\text{H}_{14}\text{O}_2$
 $M_r = 612.58$

Monoclinic, $P2_1$
 $a = 11.3972$ (5) Å
 $b = 10.1329$ (4) Å

$c = 13.8151$ (5) Å
 $\beta = 107.976$ (1)°
 $V = 1517.58$ (11) Å³
 $Z = 2$

Mo $K\alpha$ radiation
 $\mu = 1.39$ mm⁻¹
 $T = 153$ K
 $0.42 \times 0.31 \times 0.18$ mm

Data collection

Rigaku R-Axis RAPID
diffractometer
Absorption correction: multi-scan
(*ABSCOR*; Higashi, 1995)
 $T_{\text{min}} = 0.593$, $T_{\text{max}} = 0.788$

15042 measured reflections
6319 independent reflections
4162 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.039$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.045$
 $wR(F^2) = 0.180$
 $S = 1.00$
6319 reflections
386 parameters
1 restraint

H atoms treated by a mixture of
independent and constrained
refinement
 $\Delta\rho_{\text{max}} = 0.68$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.68$ e Å⁻³
Absolute structure: Flack (1983),
2637 Friedel pairs
Flack parameter: -0.002 (13)

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{O1}-\text{H1O}\cdots\text{Br1}^i$	0.74 (7)	2.56 (7)	3.294 (5)	170 (6)
$\text{O2}-\text{H2O}\cdots\text{Br1}$	0.83 (6)	2.50 (6)	3.291 (4)	161 (6)
$\text{O3}-\text{H3O}\cdots\text{Br1}^{ii}$	0.85 (10)	2.50 (10)	3.344 (4)	174 (9)

Symmetry codes: (i) $x - 1, y, z$; (ii) $-x + 1, y - \frac{1}{2}, -z + 1$.

Data collection: *PROCESS-AUTO* (Rigaku, 1998); cell refinement: *PROCESS-AUTO*; data reduction: *CrystalStructure* (Rigaku/MS, 2002); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: XU2720).

References

- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
Flack, H. D. (1983). *Acta Cryst.* **A39**, 876–881.
Fukushima, S., Hosomi, H., Ohba, S. & Kawashima, M. (1999). *Acta Cryst.* **C55**, 120–123.
Higashi, T. (1995). *ABSCOR*. Rigaku Corporation, Tokyo, Japan.
Mori, K., Masuda, Y. & Kashino, S. (1993). *Acta Cryst.* **C49**, 1224–1227.
Rigaku (1998). *PROCESS-AUTO*. Rigaku Corporation, Tokyo, Japan.
Rigaku/MS (2002). *CrystalStructure*. Rigaku/MS, The Woodlands, Texas, USA.
Schultz, A. G., Guzi, T. J., Larsson, E., Rahm, R., Thakkar, K. & Bidlack, J. M. (1998). *J. Org. Chem.* **63**, 7795–7804.
Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
Tayama, E. & Tanaka, H. (2007). *Tetrahedron Lett.* **48**, 4183–4185.

supporting information

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(3*S*,*NR*)-3-Hydroxymethyl-2-methyl-2-(3-methylbut-2-en-1-yl)-1,2,3,4-tetrahydroisoquinolinium bromide–1,1'-bi-2-naphthol (1/1)

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S1. Comment

The molecular is formed with (3*S*,*NR*)-1,2,3,4-tetrahydro-3-(hydroxymethyl)-2-methyl-2-(3-methylbut-2-en-1-yl)isoquinolinium bromide and (*S*)-1,1'-binaphthalenol. The molecular structure is shown in Fig. 1.

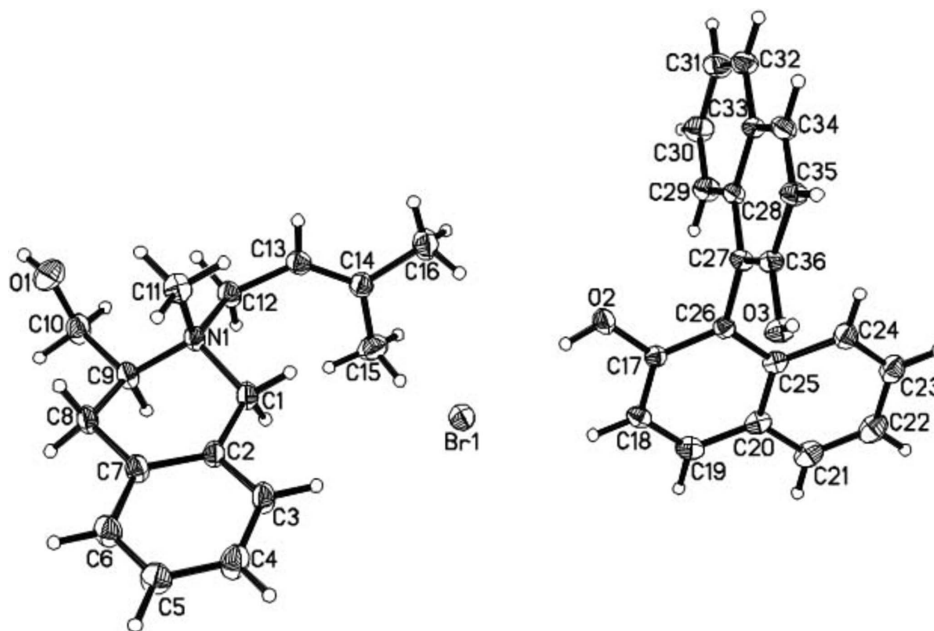
In the molecular, the N-Chiral molecule contain two six-membered rings A (atom C2–C7) and B (C1,C2,C7–C9,N1), and the ring A is benzene ring, the ring B exists in half-chair conformation with N1, C9 at the flap. The C1–C8; C8,C9,N1 and C1,N1,C9 respectively form least square plane D, E, F, and the dihedral angle between the plane D and E is $-47.4 (7)^\circ$ [C7–C8–C9–N1], between the plane E and F is $63.9 (6)^\circ$ [C1–N1–C9–C8], between the plane D and F is $-51.6 (6)^\circ$ [C9–N1–C1–C2]. The N-Chiral QASs molecule have two chiral center. The optical resolution can be efficiently carried out by an isomerization/crystallization method utilizing the axial chiral diol (*S*)-1,1'-binaphthyl-2,2'-diol (Tayama & Tanaka, 2007). The two naphthyl rings of (*S*)-1,1'-binaphthyl-2,2'-diol are approximately perpendicular to each other (Fig. 1) as observed in crystals (Mori *et al.*, 1993; Fukushima *et al.*, 1999). Intermolecular O—H...Br hydrogen bonding present in the crystal structure (Table 1).

S2. Experimental

(3*S*)-(2-Methyl-1,2,3,4-tetrahydroisoquinolin-3-yl)methanol (Schultz *et al.*, 1998) (1 mmol) and 3-methylbut-2-en-1-yl bromide (2 mmol) were dissolved in absolute acetonitrile (5 ml), refluxed for 42 h. After being cooled to room temperature, the excess 3-methylbut-2-en-1-yl bromide and acetonitrile were removed under reduced pressure. The residue was purified by flash chromatography eluted with ethyl acetate/methanol (8:1) to afford the diastereomeric mixtures of the salt. The two diastereomers of the salt was separated by Tayama and Tanaka's Method (Tayama & Tanaka, 2007).

S3. Refinement

Hydroxy H atoms were located in a difference Fourier map and were refined isotropically. Other H atoms were placed in calculated positions with C—H = 0.95–1.00 Å, and refined in riding mode with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

**Figure 1**

View of the molecule of (1) showing the atom-numbering scheme. Displacement ellipsoids are drawn at the 30% probability level.

(3*S*,*NR*)-3-Hydroxymethyl-2-methyl-2-(3-methylbut-2-en-1-yl)- 1,2,3,4-tetrahydroisoquinolinium bromide–1,1'-bi-2-naphthol (1/1)

Crystal data

$C_{16}H_{24}NO^+ \cdot C_{20}H_{14}O_2 \cdot Br^- \cdot C_{20}H_{14}O_2$
 $M_r = 612.58$
 Monoclinic, $P2_1$
 Hall symbol: P 2yb
 $a = 11.3972$ (5) Å
 $b = 10.1329$ (4) Å
 $c = 13.8151$ (5) Å
 $\beta = 107.976$ (1)°
 $V = 1517.58$ (11) Å³
 $Z = 2$

$F(000) = 640$
 $D_x = 1.341$ Mg m⁻³
 Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
 Cell parameters from 10834 reflections
 $\theta = 3.1$ – 27.5 °
 $\mu = 1.39$ mm⁻¹
 $T = 153$ K
 Block, colorless
 $0.42 \times 0.31 \times 0.18$ mm

Data collection

Rigaku R-AXIS RAPID
 diffractometer
 Radiation source: Rotating Anode
 Graphite monochromator
 ω scans
 Absorption correction: multi-scan
 (ABSCOR; Higashi, 1995)
 $T_{\min} = 0.593$, $T_{\max} = 0.788$

15042 measured reflections
 6319 independent reflections
 4162 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.039$
 $\theta_{\max} = 27.5$ °, $\theta_{\min} = 3.1$ °
 $h = -14 \rightarrow 14$
 $k = -12 \rightarrow 13$
 $l = -17 \rightarrow 17$

*Refinement*Refinement on F^2

Least-squares matrix: full

 $R[F^2 > 2\sigma(F^2)] = 0.045$ $wR(F^2) = 0.180$ $S = 1.00$

6319 reflections

386 parameters

1 restraint

Primary atom site location: structure-invariant
direct methodsSecondary atom site location: difference Fourier
mapHydrogen site location: inferred from
neighbouring sitesH atoms treated by a mixture of independent
and constrained refinement $w = 1/[\sigma^2(F_o^2) + (0.1P)^2 + 1.096P]$ where $P = (F_o^2 + 2F_c^2)/3$ $(\Delta/\sigma)_{\max} < 0.001$ $\Delta\rho_{\max} = 0.68 \text{ e } \text{Å}^{-3}$ $\Delta\rho_{\min} = -0.68 \text{ e } \text{Å}^{-3}$ Extinction correction: *SHELXL97* (Sheldrick,
2008), $F_c^* = kF_c[1 + 0.001 \times F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$

Extinction coefficient: 0.031 (3)

Absolute structure: Flack (1983), 2637 Friedel
pairsAbsolute structure parameter: -0.002 (13)*Special details***Experimental.** IR (KBr): 3421, 2976, 2928, 1638, 1454, 1082, 766 cm^{-1} ; ^{13}C -NMR (150 MHz, CDCl_3): 149.2; 130.0; 128.9; 128.7; 127.5; 127.3; 126.1; 110.8; 66.8; 61.8; 59.4; 44.1; 27.6; 26.6; 19.6. HR-ESI-MS: 246.1850 ($[\text{M}-\text{Br}]^+$, $\text{C}_{16}\text{H}_{24}\text{NO}^+$; calc. 246.1852).**Geometry.** All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.**Refinement.** Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)*

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
Br1	0.24874 (5)	0.38085 (9)	0.48816 (4)	0.0613 (2)
O1	-0.4886 (4)	0.3748 (8)	0.4353 (4)	0.0836 (17)
O2	0.4637 (4)	0.2940 (4)	0.6986 (3)	0.0518 (10)
O3	0.7346 (4)	0.1223 (4)	0.6721 (3)	0.0472 (9)
N1	-0.2234 (4)	0.4165 (4)	0.4580 (3)	0.0380 (10)
C1	-0.1158 (5)	0.4553 (7)	0.4230 (4)	0.0459 (13)
H1A	-0.0485	0.3905	0.4496	0.055*
H1B	-0.0850	0.5425	0.4522	0.055*
C2	-0.1468 (5)	0.4623 (6)	0.3087 (4)	0.0424 (12)
C3	-0.0493 (5)	0.4536 (7)	0.2679 (4)	0.0571 (17)
H3	0.0319	0.4344	0.3104	0.068*
C4	-0.0728 (6)	0.4734 (8)	0.1637 (5)	0.0650 (18)
H4	-0.0072	0.4692	0.1350	0.078*
C5	-0.1902 (6)	0.4989 (8)	0.1029 (4)	0.0628 (18)
H5	-0.2053	0.5144	0.0322	0.075*
C6	-0.2862 (6)	0.5022 (7)	0.1429 (4)	0.0539 (15)
H6	-0.3676	0.5175	0.0994	0.065*
C7	-0.2658 (5)	0.4833 (6)	0.2466 (4)	0.0438 (12)
C8	-0.3700 (5)	0.4857 (6)	0.2910 (4)	0.0449 (12)

H8A	-0.4153	0.4011	0.2751	0.054*
H8B	-0.4278	0.5570	0.2577	0.054*
C9	-0.3304 (5)	0.5070 (6)	0.4055 (4)	0.0437 (12)
H9	-0.2974	0.5990	0.4173	0.052*
C10	-0.4372 (6)	0.5020 (8)	0.4468 (5)	0.0599 (17)
H10A	-0.4092	0.5261	0.5198	0.072*
H10B	-0.5007	0.5665	0.4103	0.072*
C11	-0.2518 (6)	0.2732 (6)	0.4345 (5)	0.0513 (14)
H11A	-0.2676	0.2580	0.3615	0.062*
H11B	-0.3249	0.2485	0.4535	0.062*
H11C	-0.1814	0.2195	0.4732	0.062*
C12	-0.1876 (5)	0.4348 (7)	0.5745 (4)	0.0473 (13)
H12A	-0.1747	0.5300	0.5903	0.057*
H12B	-0.2573	0.4053	0.5976	0.057*
C13	-0.0749 (5)	0.3619 (7)	0.6328 (3)	0.0468 (12)
H13	-0.0843	0.2699	0.6414	0.056*
C14	0.0378 (5)	0.4109 (6)	0.6745 (4)	0.0476 (14)
C15	0.0735 (7)	0.5505 (7)	0.6669 (5)	0.0665 (18)
H15A	0.1320	0.5553	0.6278	0.080*
H15B	0.1122	0.5860	0.7353	0.080*
H15C	-0.0001	0.6024	0.6326	0.080*
C16	0.1391 (6)	0.3246 (8)	0.7354 (5)	0.0677 (19)
H16A	0.1079	0.2346	0.7362	0.081*
H16B	0.1700	0.3583	0.8052	0.081*
H16C	0.2062	0.3240	0.7049	0.081*
C17	0.5660 (4)	0.3709 (6)	0.7316 (3)	0.0363 (9)
C18	0.5603 (5)	0.5051 (5)	0.7061 (4)	0.0399 (11)
H18	0.4852	0.5432	0.6653	0.048*
C19	0.6647 (6)	0.5809 (5)	0.7408 (4)	0.0467 (13)
H19	0.6610	0.6715	0.7223	0.056*
C20	0.7767 (5)	0.5286 (5)	0.8031 (4)	0.0412 (12)
C21	0.8858 (6)	0.6058 (6)	0.8398 (5)	0.0541 (14)
H21	0.8838	0.6967	0.8223	0.065*
C22	0.9932 (6)	0.5514 (7)	0.8997 (5)	0.0615 (17)
H22	1.0651	0.6043	0.9240	0.074*
C23	0.9971 (5)	0.4173 (6)	0.9254 (4)	0.0531 (16)
H23	1.0720	0.3799	0.9671	0.064*
C24	0.8950 (5)	0.3401 (6)	0.8912 (4)	0.0463 (13)
H24	0.8998	0.2493	0.9093	0.056*
C25	0.7809 (4)	0.3932 (5)	0.8287 (3)	0.0356 (9)
C26	0.6748 (4)	0.3142 (5)	0.7930 (3)	0.0327 (10)
C27	0.6773 (4)	0.1713 (5)	0.8183 (3)	0.0341 (10)
C28	0.6523 (4)	0.1276 (5)	0.9080 (3)	0.0339 (10)
C29	0.6179 (5)	0.2139 (6)	0.9743 (4)	0.0460 (13)
H29	0.6121	0.3057	0.9597	0.055*
C30	0.5926 (6)	0.1699 (7)	1.0584 (4)	0.0548 (15)
H30	0.5694	0.2309	1.1015	0.066*
C31	0.6007 (6)	0.0339 (7)	1.0825 (4)	0.0555 (15)

H31	0.5831	0.0035	1.1416	0.067*
C32	0.6337 (5)	-0.0525 (6)	1.0211 (4)	0.0478 (13)
H32	0.6395	-0.1438	1.0376	0.057*
C33	0.6598 (5)	-0.0093 (5)	0.9326 (4)	0.0369 (10)
C34	0.6936 (5)	-0.0977 (5)	0.8673 (4)	0.0413 (11)
H34	0.6988	-0.1891	0.8831	0.050*
C35	0.7191 (5)	-0.0567 (5)	0.7825 (4)	0.0404 (11)
H35	0.7426	-0.1185	0.7402	0.048*
C36	0.7100 (5)	0.0801 (5)	0.7579 (4)	0.0364 (10)
H10	-0.542 (6)	0.378 (9)	0.455 (5)	0.061 (18)*
H20	0.405 (6)	0.329 (6)	0.655 (5)	0.053 (18)*
H30	0.741 (9)	0.057 (10)	0.635 (7)	0.12 (4)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Br1	0.0427 (3)	0.0882 (5)	0.0533 (3)	0.0066 (3)	0.0154 (2)	0.0232 (4)
O1	0.051 (3)	0.115 (5)	0.102 (4)	0.008 (4)	0.049 (3)	0.028 (4)
O2	0.036 (2)	0.049 (2)	0.059 (2)	-0.0024 (17)	-0.0014 (18)	0.014 (2)
O3	0.059 (3)	0.044 (2)	0.045 (2)	-0.0069 (18)	0.0260 (18)	-0.0015 (18)
N1	0.031 (2)	0.048 (3)	0.0339 (18)	0.0060 (17)	0.0091 (15)	0.0000 (18)
C1	0.032 (3)	0.070 (4)	0.035 (2)	0.004 (2)	0.009 (2)	0.002 (2)
C2	0.030 (3)	0.064 (4)	0.032 (2)	-0.001 (2)	0.0084 (19)	0.001 (2)
C3	0.033 (3)	0.099 (5)	0.041 (3)	0.010 (3)	0.012 (2)	0.006 (3)
C4	0.047 (4)	0.104 (6)	0.050 (3)	-0.001 (4)	0.023 (3)	-0.011 (4)
C5	0.060 (4)	0.099 (5)	0.031 (3)	-0.005 (4)	0.017 (3)	-0.007 (3)
C6	0.051 (4)	0.074 (4)	0.033 (3)	-0.001 (3)	0.008 (2)	-0.008 (3)
C7	0.038 (3)	0.054 (3)	0.036 (2)	0.003 (2)	0.008 (2)	-0.002 (2)
C8	0.031 (3)	0.063 (3)	0.039 (3)	0.009 (2)	0.007 (2)	0.002 (2)
C9	0.037 (3)	0.053 (3)	0.038 (3)	0.010 (2)	0.007 (2)	-0.001 (2)
C10	0.039 (3)	0.095 (5)	0.049 (3)	0.012 (3)	0.018 (2)	0.003 (3)
C11	0.053 (4)	0.046 (3)	0.052 (3)	0.003 (3)	0.013 (3)	-0.007 (3)
C12	0.042 (3)	0.068 (4)	0.031 (2)	0.005 (2)	0.010 (2)	0.001 (2)
C13	0.047 (3)	0.057 (3)	0.036 (2)	0.003 (3)	0.0123 (19)	0.006 (3)
C14	0.042 (3)	0.065 (4)	0.036 (2)	0.000 (2)	0.011 (2)	0.003 (3)
C15	0.069 (5)	0.065 (4)	0.064 (4)	-0.023 (3)	0.020 (3)	0.002 (3)
C16	0.047 (3)	0.095 (5)	0.055 (3)	0.002 (3)	0.007 (3)	0.009 (3)
C17	0.036 (2)	0.040 (3)	0.0346 (19)	-0.002 (2)	0.0119 (16)	0.003 (3)
C18	0.048 (3)	0.034 (3)	0.038 (2)	0.006 (2)	0.013 (2)	0.005 (2)
C19	0.063 (4)	0.034 (3)	0.046 (3)	0.000 (2)	0.021 (3)	-0.002 (2)
C20	0.053 (3)	0.035 (3)	0.040 (3)	-0.007 (2)	0.020 (2)	-0.007 (2)
C21	0.060 (4)	0.051 (3)	0.055 (3)	-0.017 (3)	0.023 (3)	-0.014 (3)
C22	0.054 (4)	0.069 (4)	0.062 (4)	-0.023 (3)	0.019 (3)	-0.018 (3)
C23	0.042 (3)	0.074 (5)	0.040 (2)	-0.008 (3)	0.008 (2)	-0.007 (3)
C24	0.043 (3)	0.057 (3)	0.038 (2)	-0.002 (2)	0.010 (2)	-0.002 (2)
C25	0.044 (2)	0.034 (2)	0.0299 (18)	-0.007 (2)	0.0132 (16)	-0.007 (2)
C26	0.034 (2)	0.035 (2)	0.028 (2)	-0.0021 (18)	0.0088 (17)	-0.0005 (18)
C27	0.039 (3)	0.033 (2)	0.030 (2)	0.0022 (19)	0.0096 (18)	0.0035 (19)

C28	0.036 (2)	0.032 (2)	0.033 (2)	0.0035 (19)	0.0102 (18)	0.004 (2)
C29	0.058 (3)	0.044 (3)	0.040 (3)	0.003 (2)	0.021 (2)	-0.001 (2)
C30	0.071 (4)	0.061 (4)	0.038 (3)	0.000 (3)	0.026 (3)	-0.002 (3)
C31	0.065 (4)	0.063 (4)	0.041 (3)	0.001 (3)	0.019 (3)	0.009 (3)
C32	0.053 (3)	0.051 (3)	0.038 (2)	-0.002 (3)	0.012 (2)	0.011 (2)
C33	0.034 (2)	0.041 (3)	0.035 (2)	0.000 (2)	0.0092 (19)	0.010 (2)
C34	0.043 (3)	0.034 (3)	0.043 (2)	0.003 (2)	0.0088 (19)	0.006 (2)
C35	0.050 (3)	0.030 (2)	0.043 (3)	0.001 (2)	0.018 (2)	0.001 (2)
C36	0.041 (3)	0.038 (3)	0.032 (2)	-0.002 (2)	0.0137 (19)	0.003 (2)

Geometric parameters (Å, °)

O1—C10	1.404 (10)	C15—H15A	0.9800
O1—H1O	0.74 (7)	C15—H15B	0.9800
O2—C17	1.359 (6)	C15—H15C	0.9800
O2—H2O	0.83 (6)	C16—H16A	0.9800
O3—C36	1.367 (6)	C16—H16B	0.9800
O3—H3O	0.85 (10)	C16—H16C	0.9800
N1—C11	1.501 (7)	C17—C26	1.392 (6)
N1—C1	1.503 (7)	C17—C18	1.402 (8)
N1—C9	1.519 (6)	C18—C19	1.372 (8)
N1—C12	1.544 (6)	C18—H18	0.9500
C1—C2	1.510 (7)	C19—C20	1.405 (8)
C1—H1A	0.9900	C19—H19	0.9500
C1—H1B	0.9900	C20—C25	1.414 (8)
C2—C7	1.379 (7)	C20—C21	1.423 (8)
C2—C3	1.396 (8)	C21—C22	1.364 (10)
C3—C4	1.394 (8)	C21—H21	0.9500
C3—H3	0.9500	C22—C23	1.402 (9)
C4—C5	1.367 (9)	C22—H22	0.9500
C4—H4	0.9500	C23—C24	1.360 (8)
C5—C6	1.371 (8)	C23—H23	0.9500
C5—H5	0.9500	C24—C25	1.426 (7)
C6—C7	1.392 (7)	C24—H24	0.9500
C6—H6	0.9500	C25—C26	1.407 (7)
C7—C8	1.496 (8)	C26—C27	1.488 (7)
C8—C9	1.520 (7)	C27—C36	1.372 (7)
C8—H8A	0.9900	C27—C28	1.425 (6)
C8—H8B	0.9900	C28—C29	1.407 (7)
C9—C10	1.497 (8)	C28—C33	1.425 (7)
C9—H9	1.0000	C29—C30	1.356 (8)
C10—H10A	0.9900	C29—H29	0.9500
C10—H10B	0.9900	C30—C31	1.414 (9)
C11—H11A	0.9800	C30—H30	0.9500
C11—H11B	0.9800	C31—C32	1.352 (9)
C11—H11C	0.9800	C31—H31	0.9500
C12—C13	1.486 (7)	C32—C33	1.413 (7)
C12—H12A	0.9900	C32—H32	0.9500

C12—H12B	0.9900	C33—C34	1.407 (7)
C13—C14	1.331 (8)	C34—C35	1.355 (7)
C13—H13	0.9500	C34—H34	0.9500
C14—C16	1.485 (8)	C35—C36	1.424 (7)
C14—C15	1.485 (9)	C35—H35	0.9500
C10—O1—H1O	106 (7)	H15A—C15—H15B	109.5
C17—O2—H2O	114 (4)	C14—C15—H15C	109.5
C36—O3—H3O	111 (7)	H15A—C15—H15C	109.5
C11—N1—C1	109.0 (4)	H15B—C15—H15C	109.5
C11—N1—C9	113.1 (4)	C14—C16—H16A	109.5
C1—N1—C9	107.7 (4)	C14—C16—H16B	109.5
C11—N1—C12	108.1 (4)	H16A—C16—H16B	109.5
C1—N1—C12	108.9 (4)	C14—C16—H16C	109.5
C9—N1—C12	109.9 (4)	H16A—C16—H16C	109.5
N1—C1—C2	113.5 (4)	H16B—C16—H16C	109.5
N1—C1—H1A	108.9	O2—C17—C26	118.6 (5)
C2—C1—H1A	108.9	O2—C17—C18	120.3 (4)
N1—C1—H1B	108.9	C26—C17—C18	121.1 (5)
C2—C1—H1B	108.9	C19—C18—C17	119.2 (5)
H1A—C1—H1B	107.7	C19—C18—H18	120.4
C7—C2—C3	120.8 (5)	C17—C18—H18	120.4
C7—C2—C1	121.7 (5)	C18—C19—C20	121.9 (5)
C3—C2—C1	117.5 (5)	C18—C19—H19	119.0
C4—C3—C2	118.9 (5)	C20—C19—H19	119.0
C4—C3—H3	120.5	C19—C20—C25	118.2 (5)
C2—C3—H3	120.5	C19—C20—C21	122.7 (5)
C5—C4—C3	120.2 (6)	C25—C20—C21	119.1 (5)
C5—C4—H4	119.9	C22—C21—C20	121.0 (6)
C3—C4—H4	119.9	C22—C21—H21	119.5
C4—C5—C6	120.5 (5)	C20—C21—H21	119.5
C4—C5—H5	119.7	C21—C22—C23	119.9 (6)
C6—C5—H5	119.7	C21—C22—H22	120.1
C5—C6—C7	120.7 (6)	C23—C22—H22	120.1
C5—C6—H6	119.6	C24—C23—C22	120.9 (6)
C7—C6—H6	119.6	C24—C23—H23	119.6
C2—C7—C6	118.8 (5)	C22—C23—H23	119.6
C2—C7—C8	120.0 (5)	C23—C24—C25	121.0 (5)
C6—C7—C8	121.2 (5)	C23—C24—H24	119.5
C7—C8—C9	114.2 (4)	C25—C24—H24	119.5
C7—C8—H8A	108.7	C26—C25—C20	120.4 (4)
C9—C8—H8A	108.7	C26—C25—C24	121.5 (5)
C7—C8—H8B	108.7	C20—C25—C24	118.2 (5)
C9—C8—H8B	108.7	C17—C26—C25	119.1 (4)
H8A—C8—H8B	107.6	C17—C26—C27	119.6 (4)
C10—C9—N1	115.1 (5)	C25—C26—C27	121.3 (4)
C10—C9—C8	112.2 (5)	C36—C27—C28	119.0 (4)
N1—C9—C8	110.4 (4)	C36—C27—C26	120.1 (4)

C10—C9—H9	106.2	C28—C27—C26	120.8 (4)
N1—C9—H9	106.2	C29—C28—C33	117.3 (4)
C8—C9—H9	106.2	C29—C28—C27	123.0 (4)
O1—C10—C9	110.3 (5)	C33—C28—C27	119.8 (4)
O1—C10—H10A	109.6	C30—C29—C28	122.0 (5)
C9—C10—H10A	109.6	C30—C29—H29	119.0
O1—C10—H10B	109.6	C28—C29—H29	119.0
C9—C10—H10B	109.6	C29—C30—C31	120.4 (6)
H10A—C10—H10B	108.1	C29—C30—H30	119.8
N1—C11—H11A	109.5	C31—C30—H30	119.8
N1—C11—H11B	109.5	C32—C31—C30	119.6 (5)
H11A—C11—H11B	109.5	C32—C31—H31	120.2
N1—C11—H11C	109.5	C30—C31—H31	120.2
H11A—C11—H11C	109.5	C31—C32—C33	121.1 (6)
H11B—C11—H11C	109.5	C31—C32—H32	119.5
C13—C12—N1	114.3 (4)	C33—C32—H32	119.5
C13—C12—H12A	108.7	C34—C33—C32	122.0 (5)
N1—C12—H12A	108.7	C34—C33—C28	118.3 (4)
C13—C12—H12B	108.7	C32—C33—C28	119.7 (5)
N1—C12—H12B	108.7	C35—C34—C33	122.3 (5)
H12A—C12—H12B	107.6	C35—C34—H34	118.9
C14—C13—C12	127.4 (6)	C33—C34—H34	118.9
C14—C13—H13	116.3	C34—C35—C36	119.1 (5)
C12—C13—H13	116.3	C34—C35—H35	120.5
C13—C14—C16	120.2 (6)	C36—C35—H35	120.5
C13—C14—C15	125.0 (6)	O3—C36—C27	118.9 (4)
C16—C14—C15	114.8 (6)	O3—C36—C35	119.5 (4)
C14—C15—H15A	109.5	C27—C36—C35	121.6 (4)
C14—C15—H15B	109.5		
C11—N1—C1—C2	71.5 (6)	C22—C23—C24—C25	-0.3 (8)
C9—N1—C1—C2	-51.6 (6)	C19—C20—C25—C26	-0.4 (7)
C12—N1—C1—C2	-170.7 (5)	C21—C20—C25—C26	-179.6 (4)
N1—C1—C2—C7	24.1 (8)	C19—C20—C25—C24	179.5 (4)
N1—C1—C2—C3	-159.1 (5)	C21—C20—C25—C24	0.3 (7)
C7—C2—C3—C4	3.4 (10)	C23—C24—C25—C26	180.0 (5)
C1—C2—C3—C4	-173.4 (7)	C23—C24—C25—C20	0.1 (7)
C2—C3—C4—C5	-1.0 (12)	O2—C17—C26—C25	-179.1 (4)
C3—C4—C5—C6	-1.5 (12)	C18—C17—C26—C25	-0.8 (7)
C4—C5—C6—C7	1.8 (11)	O2—C17—C26—C27	1.7 (6)
C3—C2—C7—C6	-3.2 (10)	C18—C17—C26—C27	-179.9 (4)
C1—C2—C7—C6	173.5 (6)	C20—C25—C26—C17	0.3 (6)
C3—C2—C7—C8	177.0 (6)	C24—C25—C26—C17	-179.5 (4)
C1—C2—C7—C8	-6.3 (9)	C20—C25—C26—C27	179.5 (4)
C5—C6—C7—C2	0.6 (10)	C24—C25—C26—C27	-0.4 (7)
C5—C6—C7—C8	-179.6 (6)	C17—C26—C27—C36	91.4 (6)
C2—C7—C8—C9	18.2 (8)	C25—C26—C27—C36	-87.7 (6)
C6—C7—C8—C9	-161.6 (6)	C17—C26—C27—C28	-92.0 (6)

C11—N1—C9—C10	71.5 (6)	C25—C26—C27—C28	88.9 (6)
C1—N1—C9—C10	-167.9 (5)	C36—C27—C28—C29	-179.9 (5)
C12—N1—C9—C10	-49.4 (6)	C26—C27—C28—C29	3.5 (7)
C11—N1—C9—C8	-56.7 (6)	C36—C27—C28—C33	-1.0 (7)
C1—N1—C9—C8	63.9 (6)	C26—C27—C28—C33	-177.6 (4)
C12—N1—C9—C8	-177.6 (5)	C33—C28—C29—C30	0.0 (8)
C7—C8—C9—C10	-177.2 (5)	C27—C28—C29—C30	178.9 (5)
C7—C8—C9—N1	-47.4 (7)	C28—C29—C30—C31	0.2 (9)
N1—C9—C10—O1	-61.7 (6)	C29—C30—C31—C32	-0.1 (10)
C8—C9—C10—O1	65.6 (7)	C30—C31—C32—C33	-0.3 (10)
C11—N1—C12—C13	62.5 (6)	C31—C32—C33—C34	-179.6 (5)
C1—N1—C12—C13	-55.8 (6)	C31—C32—C33—C28	0.5 (8)
C9—N1—C12—C13	-173.5 (5)	C29—C28—C33—C34	179.8 (5)
N1—C12—C13—C14	102.3 (6)	C27—C28—C33—C34	0.8 (7)
C12—C13—C14—C16	176.7 (5)	C29—C28—C33—C32	-0.4 (7)
C12—C13—C14—C15	-2.3 (9)	C27—C28—C33—C32	-179.3 (5)
O2—C17—C18—C19	179.6 (5)	C32—C33—C34—C35	-179.8 (5)
C26—C17—C18—C19	1.2 (7)	C28—C33—C34—C35	0.1 (7)
C17—C18—C19—C20	-1.3 (8)	C33—C34—C35—C36	-0.7 (8)
C18—C19—C20—C25	0.9 (8)	C28—C27—C36—O3	-179.6 (4)
C18—C19—C20—C21	-180.0 (5)	C26—C27—C36—O3	-3.0 (7)
C19—C20—C21—C22	-179.7 (6)	C28—C27—C36—C35	0.3 (8)
C25—C20—C21—C22	-0.6 (8)	C26—C27—C36—C35	177.0 (5)
C20—C21—C22—C23	0.4 (9)	C34—C35—C36—O3	-179.5 (5)
C21—C22—C23—C24	0.1 (9)	C34—C35—C36—C27	0.6 (8)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
O1—H1O...Br1 ⁱ	0.74 (7)	2.56 (7)	3.294 (5)	170 (6)
O2—H2O...Br1	0.83 (6)	2.50 (6)	3.291 (4)	161 (6)
O3—H3O...Br1 ⁱⁱ	0.85 (10)	2.50 (10)	3.344 (4)	174 (9)

Symmetry codes: (i) $x-1, y, z$; (ii) $-x+1, y-1/2, -z+1$.