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Poly[[(μ-2,2′-bipyrimidine-$\kappa^4N^1,N^{1'}$:$N^3,N^{3'}$)(μ-sulfato-κ^2O:O')zinc(II)] monohydrate]

Aaron Oxendine,[a] Jennifer Kelley,[a]* LeRoy Peterson Jr,[a] Mark D. Smith[b] and Hans-Conrad zur Loye[b]

[a]Chemistry Department, Francis Marion University, Florence, South Carolina 29502, USA, and [b]Department of Chemistry and Biochemistry, University of South Carolina, Columbia, South Carolina 29208, USA
Correspondence e-mail: jkelley@fmarion.edu

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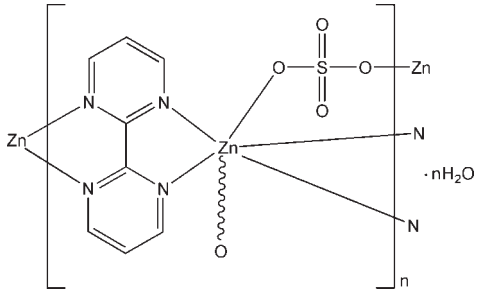
Key indicators: single-crystal X-ray study; T = 294 K; mean σ(C–C) = 0.003 Å; R factor = 0.023; wR factor = 0.063; data-to-parameter ratio = 12.5.

In the title compound, $\{[Zn(SO_4)(C_8H_6N_4)]\cdot H_2O\}_n$, the Zn^{II} atom is in a distorted octahedral environment. The Zn^{II} atoms are bridged by both 2,2′-bipyrimidine and sulfate ligands, thus forming a three-dimensional polymeric metal–organic solid that contains uncoordinated water molecules in the interstitial space. O—H···O hydrogen bonding consolidates the crystal structure.

Related literature

For general background to metal-organic polymers with 2,2′-bipyrimidine ligands, see: De Munno *et al.* (1995); Kawata *et al.* (1998); Marshall *et al.* (2000); Wang *et al.* (2007). For a related structure, see: De Munno & Julve (1994).



Experimental

Crystal data

$[Zn(SO_4)(C_8H_6N_4)]\cdot H_2O$
M_r = 337.61
Monoclinic, $P2_1/c$
a = 8.9935 (3) Å
b = 13.9783 (5) Å
c = 9.8459 (4) Å
β = 117.007 (1)°
V = 1102.79 (7) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 2.44 mm^{-1}
T = 294 K
0.20 × 0.15 × 0.08 mm

Data collection

Bruker SMART APEX CCD diffractometer
Absorption correction: multi-scan (*SADABS*; Bruker, 2001)
T_{min} = 0.874, T_{max} = 1.000
12167 measured reflections
2254 independent reflections
2087 reflections with $I > 2\sigma(I)$
R_{int} = 0.030

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.023
$wR(F^2)$ = 0.063
S = 1.05
2254 reflections
180 parameters
1 restraint
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.40 e Å^{-3}
$\Delta\rho_{min}$ = −0.52 e Å^{-3}

Table 1
Selected bond lengths (Å).

Zn1—N1	2.2646 (16)	Zn1—N4^{ii}	2.2852 (16)
Zn1—N2^{i}	2.1228 (15)	Zn1—O1	2.0302 (14)
Zn1—N3	2.1403 (17)	Zn1—O2^{iii}	2.0371 (14)

Symmetry codes: (i) $-x+2, -y+1, -z+1$; (ii) $-x+1, -y+1, -z$; (iii) $x, -y+\frac{3}{2}, z-\frac{1}{2}$.

Table 2
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O5—H5A···O3	0.78 (2)	2.07 (2)	2.838 (3)	166 (3)
O5—H5B···O2^{iv}	0.78 (2)	2.11 (2)	2.883 (2)	173 (3)

Symmetry code: (iv) $x, -y+\frac{3}{2}, z+\frac{1}{2}$.

Data collection: *SMART* (Bruker, 2007); cell refinement: *SAINT* (Bruker, 2007); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *DIAMOND* (Brandenburg, 1999); software used to prepare material for publication: *SHELXTL* (Sheldrick, 2008).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HY2264).

References

Brandenburg, K. (1999). *DIAMOND*. Crystal Impact GbR, Bonn, Germany.
Bruker (2001). *SADABS*. Bruker AXS Inc., Madison, Wisconsin, USA.
Bruker (2007). *SMART* and *SAINT*. Bruker AXS Inc., Madison, Wisconsin, USA.
De Munno, G. & Julve, M. (1994). *Acta Cryst.* C**50**, 1034–1037.
De Munno, G., Ruiz, R., Lloret, F., Faus, J., Sessoli, R. & Julve, M. (1995). *Inorg. Chem.* **34**, 408–411.
Kawata, S., Kitagawa, S., Enomoto, M., Kumagai, H. & Katada, M. (1998). *Inorg. Chim. Acta*, **283**, 80–90.
Marshall, S. R., Incarvito, C. D., Manson, J. L., Rheingold, A. L. & Miller, J. S. (2000). *Inorg. Chem.* **39**, 1969–1973.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Wang, C.-C., Kuo, C.-T., Yang, J.-C., Lee, G.-H., Shih, W.-J. & Sheu, H.-S. (2007). *Cryst. Growth Des.* **7**, 1476–1482.

supporting information

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Poly[[(μ-2,2′-bipyrimidine-$\kappa^4N^1,N^{1'}:N^3,N^{3'}$)($\mu$-sulfato-$\kappa^2O:O'$)zinc(II)] monohydrate]

Aaron Oxendine, Jennifer Kelley, LeRoy Peterson, Mark D. Smith and Hans-Conrad zur Loye

S1. Comment

Metal-organic polymers utilizing the 2,2′-bipyrimidine (bpm) ligand are being studied due to the ability of bpm to produce interesting and potentially useful materials (Kawata *et al.*, 1998; Marshall *et al.*, 2000; Wang *et al.*, 2007). Such features are often associated with the ability of this ligand to link metal centers through the bis-bidentate coordination mode (De Munno *et al.*, 1995). Herein we report the crystal structure of the title compound, (I), that is a three-dimensional metal-organic framework where bpm binds Zn^{II} atoms in this fashion.

The crystal structure of (I), which incidentally is isostructural with $[Cu(bpm)(SO_4)].H_2O$ (Kawata *et al.*, 1998), is a three-dimensional polymeric solid with an asymmetric unit consisting of one Zn^{II} atom, two half-bpm ligands, a sulfate ligand, and one lattice water. The Zn^{II} atom resides in a distorted octahedral environment composed of four N donors from a pair of equivalent bpm ligands, and two O atoms from two equivalent sulfate anions (Fig. 1). All of the Zn—N and Zn—O bond distances are in a normal range (Table 1).

The bpm ligand bridges Zn^{II} atoms in a bis-bidentate fashion, producing undulating chains running along the [101] direction. Further, the sulfate ligand serves to bridge neighboring chains, thus forming a three-dimensional microporous solid. The pores are occupied by lattice waters that are hydrogen bonded to uncoordinated O2 and O3 atoms of nearby sulfate anions (Table 2 and Fig. 2).

It is also interesting to note that the crystal structure of (I) differs from that of $[Zn_2(\mu\text{-bpm})(H_2O)_8](SO_4)_2.2H_2O$ (II) (De Munno & Julve, 1994), which contains the same chemical components as (I), but was synthesized under different synthetic conditions (see below).

S2. Experimental

All starting chemicals were purchased from commercial sources and used as received. An aqueous solution of zinc sulfate heptahydrate (0.10 mmol, 10 ml) was slowly added to 10 ml of an ethanolic solution composed of bpm (0.050 mmol) and 4,4′-bipyridine (bpy) (0.050 mmol). Colorless, plate-like crystals formed within several weeks after slow evaporation of all the solvent under ambient conditions. Although bpy was not incorporated into the crystal structure of (I), it was required for synthesis of the crystalline product, as no such crystals were formed without it with all other conditions being the same. The synthesis in water alone using only zinc sulfate heptahydrate and bpm was reported to produce (II), as previously mentioned.

S3. Refinement

H atoms bonded to C atoms were placed in geometrically idealized positions and refined as riding atoms, with C—H = 0.93 Å and U_{iso}(H) = 1.2U_{eq}(C). H atoms of water molecule were located from a difference Fourier map and refined isotropically, with their O—H distances restrained to 0.84 (2) Å.

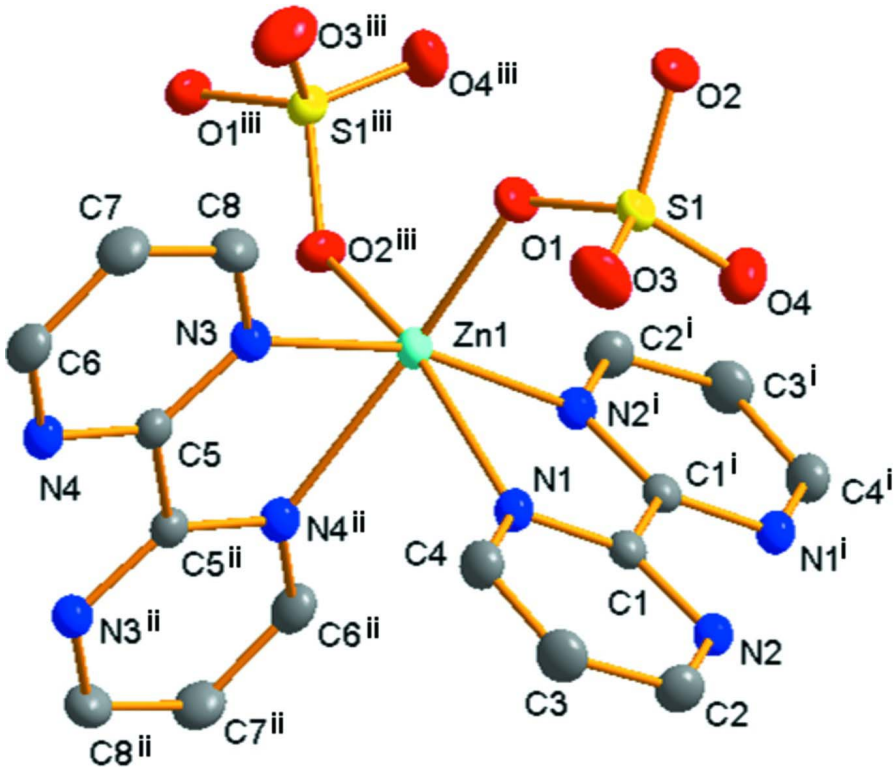

Figure 1

Coordination environment of Zn^{II} atom in (I). Displacement ellipsoids are shown at the 50% probability level. H atoms and water molecule have been omitted for clarity. [Symmetry codes: (i) 2-x, 1-y, 1-z; (ii) 1-x, 1-y, -z; (iii) x, 3/2-y, -1/2+z.]

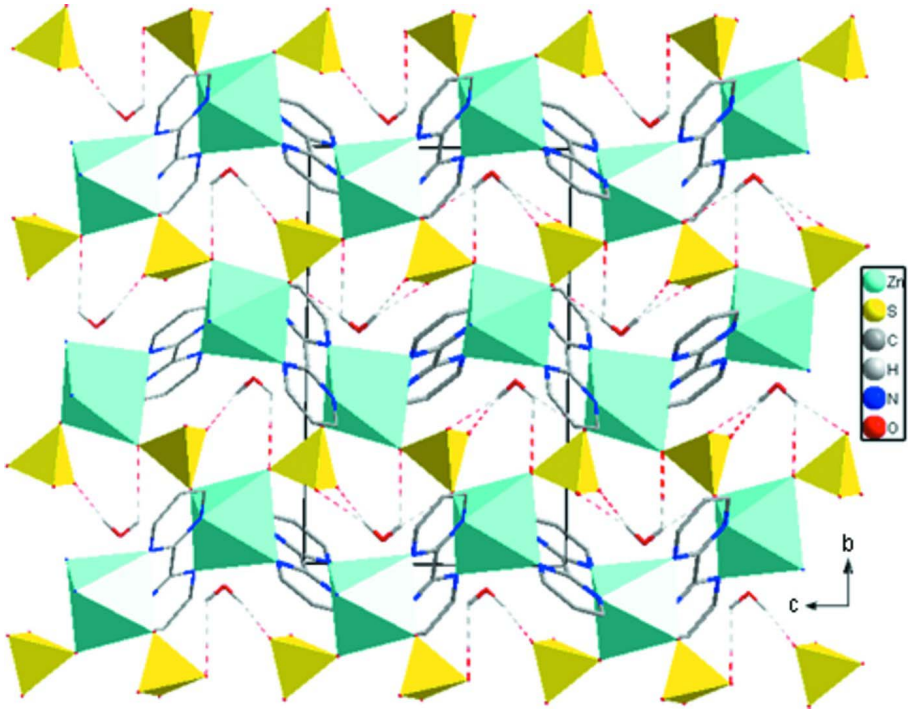

Figure 2

Polyhedral and wireframe representation of the crystal packing in (I). All H atoms except those of water have been omitted for clarity. Hydrogen bonds are represented by dashed lines.

Poly[[(μ-2,2′-bipyrimidine-$\kappa^4N^1,N^{1'}:N^3,N^{3'}$)(μ-sulfato-$\kappa^2O:O'$)zinc(II)] monohydrate]

Crystal data

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$b = 13.9783\ (5)$ Å
$c = 9.8459\ (4)$ Å
$\beta = 117.007\ (1)°$
$V = 1102.79\ (7)$ Å^3
$Z = 4$
$F(000) = 680$
$D_x = 2.033$ Mg m^{-3}
Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Cell parameters from 6452 reflections
$\theta = 2.5–26.4°$
$\mu = 2.44$ mm^{-1}
$T = 294$ K
Prism, colorless
0.20 × 0.15 × 0.08 mm

Data collection

Bruker SMART APEX CCD diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
φ and ω scans
Absorption correction: multi-scan (*SADABS*; Bruker, 2001)
$T_{min} = 0.874$, $T_{max} = 1.000$
12167 measured reflections
2254 independent reflections
2087 reflections with $I > 2\sigma(I)$
$R_{int} = 0.030$
$\theta_{max} = 26.4°$, $\theta_{min} = 2.5°$
$h = -11\rightarrow11$
$k = -17\rightarrow17$
$l = -12\rightarrow12$

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)] = 0.023$
$wR(F^2) = 0.063$
$S = 1.05$
2254 reflections
180 parameters
1 restraint
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: mixed
H atoms treated by a mixture of independent and constrained refinement
$w = 1/[\sigma^2(F_o^2) + (0.0366P)^2 + 0.505P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} = 0.001$
$\Delta\rho_{max} = 0.40$ e Å^{-3}
$\Delta\rho_{min} = -0.52$ e Å^{-3}

Special details

Refinement. Water molecule O—H bonds restrained to be approximately equal.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)

	x	*y*	*z*	U_{iso}*/U_{eq}
Zn1	0.77696 (3)	0.609668 (15)	0.23367 (2)	0.01838 (9)
S1	0.81536 (6)	0.72465 (3)	0.52650 (5)	0.01914 (12)
C1	0.9291 (2)	0.47924 (13)	0.5096 (2)	0.0177 (4)
C2	0.8397 (3)	0.38425 (14)	0.6454 (2)	0.0243 (4)
H2	0.8615	0.3449	0.7285	0.029*
C3	0.6764 (3)	0.40060 (15)	0.5389 (3)	0.0268 (5)
H3	0.5878	0.3722	0.5478	0.032*
C4	0.6501 (2)	0.46036 (15)	0.4197 (2)	0.0243 (4)
H4	0.5412	0.4735	0.3478	0.029*
C5	0.4422 (2)	0.53564 (13)	0.0071 (2)	0.0180 (4)
C6	0.1759 (2)	0.58714 (15)	−0.0698 (2)	0.0247 (4)
H6	0.0615	0.5802	−0.1296	0.030*
C7	0.2350 (3)	0.66300 (16)	0.0306 (2)	0.0274 (4)
H7	0.1626	0.7078	0.0381	0.033*
C8	0.4051 (3)	0.66992 (15)	0.1193 (2)	0.0261 (4)
H8	0.4480	0.7202	0.1882	0.031*
N1	0.77690 (19)	0.50046 (12)	0.40355 (18)	0.0195 (3)
N2	0.96733 (19)	0.42423 (12)	0.63088 (18)	0.0185 (3)
N3	0.5106 (2)	0.60557 (11)	0.10833 (19)	0.0210 (4)
N4	0.2798 (2)	0.52321 (12)	−0.08307 (18)	0.0203 (3)
O1	0.78154 (19)	0.72359 (10)	0.36342 (16)	0.0271 (3)
O2	0.81664 (19)	0.82843 (10)	0.56562 (16)	0.0263 (3)
O3	0.6819 (2)	0.67544 (12)	0.54284 (19)	0.0374 (4)
O4	0.97797 (19)	0.68377 (12)	0.62188 (17)	0.0337 (4)
O5	0.7510 (2)	0.55805 (14)	0.7998 (2)	0.0384 (4)
H5A	0.747 (3)	0.5945 (17)	0.738 (3)	0.034 (8)*
H5B	0.769 (4)	0.5931 (19)	0.867 (3)	0.046 (9)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Zn1	0.01726 (14)	0.01966 (14)	0.01514 (13)	−0.00064 (8)	0.00468 (10)	0.00031 (8)
S1	0.0218 (2)	0.0193 (2)	0.0160 (2)	−0.00226 (18)	0.00831 (19)	−0.00235 (18)
C1	0.0188 (9)	0.0163 (9)	0.0171 (9)	0.0000 (7)	0.0073 (7)	−0.0026 (7)
C2	0.0248 (10)	0.0255 (11)	0.0250 (10)	−0.0019 (8)	0.0135 (9)	0.0039 (8)
C3	0.0202 (10)	0.0308 (12)	0.0317 (11)	−0.0050 (8)	0.0138 (9)	−0.0002 (9)
C4	0.0169 (9)	0.0268 (11)	0.0256 (10)	0.0000 (8)	0.0066 (8)	−0.0002 (8)
C5	0.0192 (9)	0.0183 (9)	0.0155 (9)	0.0007 (7)	0.0071 (7)	0.0018 (7)
C6	0.0184 (9)	0.0281 (11)	0.0252 (10)	0.0034 (8)	0.0079 (8)	0.0051 (9)
C7	0.0266 (10)	0.0267 (11)	0.0306 (11)	0.0079 (8)	0.0145 (9)	0.0026 (9)
C8	0.0304 (11)	0.0213 (10)	0.0254 (10)	0.0026 (8)	0.0116 (9)	−0.0033 (8)
N1	0.0160 (7)	0.0211 (8)	0.0185 (8)	0.0004 (6)	0.0053 (6)	0.0003 (7)
N2	0.0180 (8)	0.0199 (8)	0.0171 (8)	−0.0005 (6)	0.0075 (6)	−0.0003 (6)
N3	0.0203 (8)	0.0204 (9)	0.0195 (8)	0.0003 (6)	0.0066 (7)	−0.0006 (6)
N4	0.0180 (7)	0.0218 (8)	0.0189 (8)	−0.0007 (6)	0.0064 (6)	0.0013 (6)
O1	0.0408 (9)	0.0226 (8)	0.0170 (7)	0.0002 (6)	0.0124 (6)	−0.0035 (6)
O2	0.0393 (8)	0.0204 (7)	0.0207 (7)	−0.0009 (6)	0.0150 (6)	−0.0040 (6)
O3	0.0378 (9)	0.0417 (10)	0.0404 (9)	−0.0170 (7)	0.0245 (8)	−0.0086 (8)
O4	0.0314 (8)	0.0363 (9)	0.0263 (8)	0.0089 (7)	0.0067 (7)	−0.0002 (7)
O5	0.0487 (11)	0.0327 (10)	0.0320 (9)	−0.0069 (8)	0.0167 (8)	−0.0042 (8)

Geometric parameters (Å, º)

Zn1—N1	2.2646 (16)	C3—C4	1.371 (3)
Zn1—N2^{i}	2.1228 (15)	C3—H3	0.9300
Zn1—N3	2.1403 (17)	C4—N1	1.343 (3)
Zn1—N4^{ii}	2.2852 (16)	C4—H4	0.9300
Zn1—O1	2.0302 (14)	C5—N3	1.331 (2)
Zn1—O2^{iii}	2.0371 (14)	C5—N4	1.332 (2)
S1—O4	1.4491 (15)	C5—C5^{ii}	1.492 (4)
S1—O3	1.4547 (15)	C6—N4	1.341 (3)
S1—O1	1.4924 (14)	C6—C7	1.381 (3)
S1—O2	1.4996 (15)	C6—H6	0.9300
C1—N1	1.324 (2)	C7—C8	1.378 (3)
C1—N2	1.327 (2)	C7—H7	0.9300
C1—C1^{i}	1.489 (4)	C8—N3	1.346 (3)
C2—N2	1.341 (3)	C8—H8	0.9300
C2—C3	1.382 (3)	O5—H5A	0.78 (2)
C2—H2	0.9300	O5—H5B	0.78 (2)
O1—Zn1—O2^{iii}	102.53 (6)	N1—C4—C3	121.99 (18)
O1—Zn1—N2^{i}	94.20 (6)	N1—C4—H4	119.0
O2^{iii}—Zn1—N2^{i}	93.77 (6)	C3—C4—H4	119.0
O1—Zn1—N3	94.66 (6)	N3—C5—N4	126.10 (17)
O2^{iii}—Zn1—N3	96.09 (6)	N3—C5—C5^{ii}	117.1 (2)
N2^{i}—Zn1—N3	165.00 (6)	N4—C5—C5^{ii}	116.8 (2)

O1—Zn1—N1	94.05 (6)	N4—C6—C7	121.50 (18)
O2^{iii}—Zn1—N1	160.92 (6)	N4—C6—H6	119.2
N2^{i}—Zn1—N1	75.46 (6)	C7—C6—H6	119.2
N3—Zn1—N1	91.85 (6)	C8—C7—C6	117.60 (18)
O1—Zn1—N4^{ii}	168.69 (6)	C8—C7—H7	121.2
O2^{iii}—Zn1—N4^{ii}	83.58 (6)	C6—C7—H7	121.2
N2^{i}—Zn1—N4^{ii}	94.89 (6)	N3—C8—C7	121.46 (19)
N3—Zn1—N4^{ii}	75.07 (6)	N3—C8—H8	119.3
N1—Zn1—N4^{ii}	81.74 (6)	C7—C8—H8	119.3
O4—S1—O3	112.49 (10)	C1—N1—C4	116.24 (17)
O4—S1—O1	110.24 (9)	C1—N1—Zn1	112.75 (12)
O3—S1—O1	109.82 (9)	C4—N1—Zn1	130.81 (13)
O4—S1—O2	109.10 (9)	C1—N2—C2	116.91 (16)
O3—S1—O2	109.85 (9)	C1—N2—Zn1^{i}	117.30 (12)
O1—S1—O2	105.07 (8)	C2—N2—Zn1^{i}	125.12 (14)
N1—C1—N2	126.29 (17)	C5—N3—C8	116.61 (17)
N1—C1—C1^{i}	116.8 (2)	C5—N3—Zn1	117.85 (13)
N2—C1—C1^{i}	117.0 (2)	C8—N3—Zn1	125.50 (14)
N2—C2—C3	121.07 (19)	C5—N4—C6	116.72 (17)
N2—C2—H2	119.5	C5—N4—Zn1^{ii}	113.14 (12)
C3—C2—H2	119.5	C6—N4—Zn1^{ii}	130.12 (14)
C4—C3—C2	117.47 (18)	S1—O1—Zn1	128.37 (9)
C4—C3—H3	121.3	S1—O2—Zn1^{iv}	129.47 (9)
C2—C3—H3	121.3	H5A—O5—H5B	100 (3)

N2—C2—C3—C4	−1.1 (3)	C7—C8—N3—C5	0.4 (3)
C2—C3—C4—N1	1.4 (3)	C7—C8—N3—Zn1	178.11 (15)
N4—C6—C7—C8	−1.2 (3)	O1—Zn1—N3—C5	−176.93 (14)
C6—C7—C8—N3	0.4 (3)	O2^{iii}—Zn1—N3—C5	79.91 (14)
N2—C1—N1—C4	−1.5 (3)	N2^{i}—Zn1—N3—C5	−50.9 (3)
C1^{i}—C1—N1—C4	178.1 (2)	N1—Zn1—N3—C5	−82.71 (14)
N2—C1—N1—Zn1	173.85 (15)	N4^{ii}—Zn1—N3—C5	−1.76 (13)
C1^{i}—C1—N1—Zn1	−6.5 (3)	O1—Zn1—N3—C8	5.40 (17)
C3—C4—N1—C1	−0.2 (3)	O2^{iii}—Zn1—N3—C8	−97.75 (17)
C3—C4—N1—Zn1	−174.55 (15)	N2^{i}—Zn1—N3—C8	131.4 (2)
O1—Zn1—N1—C1	−85.87 (13)	N1—Zn1—N3—C8	99.63 (17)
O2^{iii}—Zn1—N1—C1	64.6 (2)	N4^{ii}—Zn1—N3—C8	−179.43 (18)
N2^{i}—Zn1—N1—C1	7.42 (12)	N3—C5—N4—C6	−0.3 (3)
N3—Zn1—N1—C1	179.32 (13)	C5^{ii}—C5—N4—C6	179.8 (2)
N4^{ii}—Zn1—N1—C1	104.70 (13)	N3—C5—N4—Zn1^{ii}	−178.55 (15)
O1—Zn1—N1—C4	88.67 (18)	C5^{ii}—C5—N4—Zn1^{ii}	1.5 (2)
O2^{iii}—Zn1—N1—C4	−120.8 (2)	C7—C6—N4—C5	1.1 (3)
N2^{i}—Zn1—N1—C4	−178.04 (18)	C7—C6—N4—Zn1^{ii}	179.05 (15)
N3—Zn1—N1—C4	−6.14 (18)	O4—S1—O1—Zn1	57.96 (14)
N4^{ii}—Zn1—N1—C4	−80.76 (18)	O3—S1—O1—Zn1	−66.53 (14)
N1—C1—N2—C2	1.8 (3)	O2—S1—O1—Zn1	175.38 (10)
C1^{i}—C1—N2—C2	−177.8 (2)	O2^{iii}—Zn1—O1—S1	−158.01 (11)
N1—C1—N2—Zn1^{i}	172.93 (15)	N2^{i}—Zn1—O1—S1	−63.20 (12)

C1^{i}—C1—N2—Zn1^{i}	−6.7 (3)	N3—Zn1—O1—S1	104.68 (12)
C3—C2—N2—C1	−0.4 (3)	N1—Zn1—O1—S1	12.49 (12)
C3—C2—N2—Zn1^{i}	−170.74 (15)	N4^{ii}—Zn1—O1—S1	80.2 (3)
N4—C5—N3—C8	−0.5 (3)	O4—S1—O2—Zn1^{iv}	−84.54 (13)
C5^{ii}—C5—N3—C8	179.5 (2)	O3—S1—O2—Zn1^{iv}	39.20 (15)
N4—C5—N3—Zn1	−178.36 (14)	O1—S1—O2—Zn1^{iv}	157.27 (10)
C5^{ii}—C5—N3—Zn1	1.6 (3)		

Symmetry codes: (i) −*x*+2, −*y*+1, −*z*+1; (ii) −*x*+1, −*y*+1, −*z*; (iii) *x*, −*y*+3/2, *z*−1/2; (iv) *x*, −*y*+3/2, *z*+1/2.

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O5—H5*A*···O3	0.78 (2)	2.07 (2)	2.838 (3)	166 (3)
O5—H5*B*···O2^{iv}	0.78 (2)	2.11 (2)	2.883 (2)	173 (3)

Symmetry code: (iv) *x*, −*y*+3/2, *z*+1/2.