

2-Hydroxy-5-nitrobenzaldehyde

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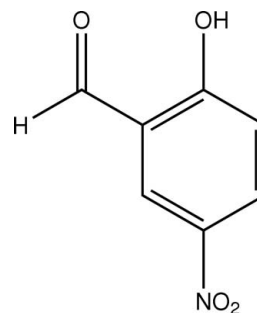
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Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(C-C) = 0.003$ Å; R factor = 0.050; wR factor = 0.119; data-to-parameter ratio = 12.5.

The title compound, $C_7H_5NO_4$, is essentially planar, with a maximum deviation from the mean plane of 0.0116 (11) Å for the hydroxy O atom. The molecular and crystal structure are stabilized by intra- and intermolecular interactions. An intramolecular $O-H\cdots O$ hydrogen bond generates a six-membered ring, producing an $S(6)$ ring motif. The $C-H\cdots O$ interactions result in the formation of $C(5)$ chains and $R_2^2(8)$ rings forming an approximately planar network parallel to $(10\bar{1})$. These planes are interconnected through $\pi-\pi$ interactions [centroid-centroid distance 3.582 (2) Å].

Related literature

Nitroaromatics are widely used as intermediates in explosives, dyestuffs, pesticides and organic synthesis, see: Yan *et al.* (2006). They occur in industrial wastes and as direct pollutants in the environment and are relatively soluble in water and detectable in rivers, ponds and soil, see: Yan *et al.* (2006); Soojhawon *et al.* (2005). Aromatic compounds with multiple nitro substituents are known to be resistant to electrophilic attack by oxygenases, see: Halas *et al.* (1983). For comparison bond lengths and angles in related structures, see: Rizal *et al.* (2008); Garden *et al.* (2004). For hydrogen-bond motifs, see: Bernstein *et al.* (1995).



Experimental

Crystal data

$C_7H_5NO_4$
 $M_r = 167.12$
 Monoclinic, $P2_1/n$
 $a = 7.2580$ (17) Å
 $b = 8.3960$ (13) Å
 $c = 11.704$ (3) Å
 $\beta = 95.165$ (18)°

$V = 710.3$ (3) Å³
 $Z = 4$
 Mo $K\alpha$ radiation
 $\mu = 0.13$ mm⁻¹
 $T = 296$ K
 $0.54 \times 0.28 \times 0.15$ mm

Data collection

Stoe IPDS II diffractometer
 Absorption correction: integration
 ($X-RED32$; Stoe & Cie, 2002)
 $T_{min} = 0.979$, $T_{max} = 0.992$

4345 measured reflections
 1396 independent reflections
 944 reflections with $I > 2\sigma(I)$
 $R_{int} = 0.062$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.050$
 $wR(F^2) = 0.119$
 $S = 1.06$
 1396 reflections
 112 parameters

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{max} = 0.16$ e Å⁻³
 $\Delta\rho_{min} = -0.15$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$C2-H2\cdots O4^i$	0.93	2.50	3.427 (3)	175
$C6-H6\cdots O3^{ii}$	0.93	2.53	3.433 (3)	163
$C7-H7\cdots O2^{iii}$	0.93	2.50	3.176 (3)	130
$O3-H3A\cdots O4$	0.93 (3)	1.73 (3)	2.613 (3)	157 (3)

Symmetry codes: (i) $x, y - 1, z$; (ii) $x + \frac{1}{2}, -y + \frac{1}{2}, z + \frac{1}{2}$; (iii) $x, y + 1, z$.

Data collection: $X-AREA$ (Stoe & Cie, 2002); cell refinement: $X-AREA$; data reduction: $X-RED32$ (Stoe & Cie, 2002); program(s) used to solve structure: $SHELXS97$ (Sheldrick, 2008); program(s) used to refine structure: $SHELXL97$ (Sheldrick, 2008); molecular graphics: $ORTEP-3$ for Windows (Farrugia, 1997) and $PLATON$ (Spek, 2009); software used to prepare material for publication: $WinGX$ (Farrugia, 1999).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: DN2510).

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supporting information

Acta Cryst. (2009). E65, o3056–o3057 [doi:10.1107/S1600536809046807]

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S1. Comment

Nitroaromatics are widely used either as materials or as intermediates in explosives, dyestuffs, pesticides and organic synthesis (Yan *et al.*, 2006). Nitroaromatics occur as industrial wastes and direct pollutants in the environment, and are relatively soluble in water and detectable in rivers, ponds and soil (Yan *et al.*, 2006; Soojhawon *et al.*, 2005). Moreover, aromatic compounds with multiple nitro substituents are known to be resistant to electrophilic attack by oxygenases (Halas *et al.*, 1983).

In the title compound (I, Fig. 1), the molecule is essentially planar with a maximum deviation from the mean plane of 0.0116 (11) Å for atom O3. The bond lengths and angles in (I) have normal values, and are comparable with those in the related structures (Rizal *et al.*, 2008; Garden *et al.*, 2004). The dihedral angle between the aromatic ring and the nitro group is 3.83 (3)°.

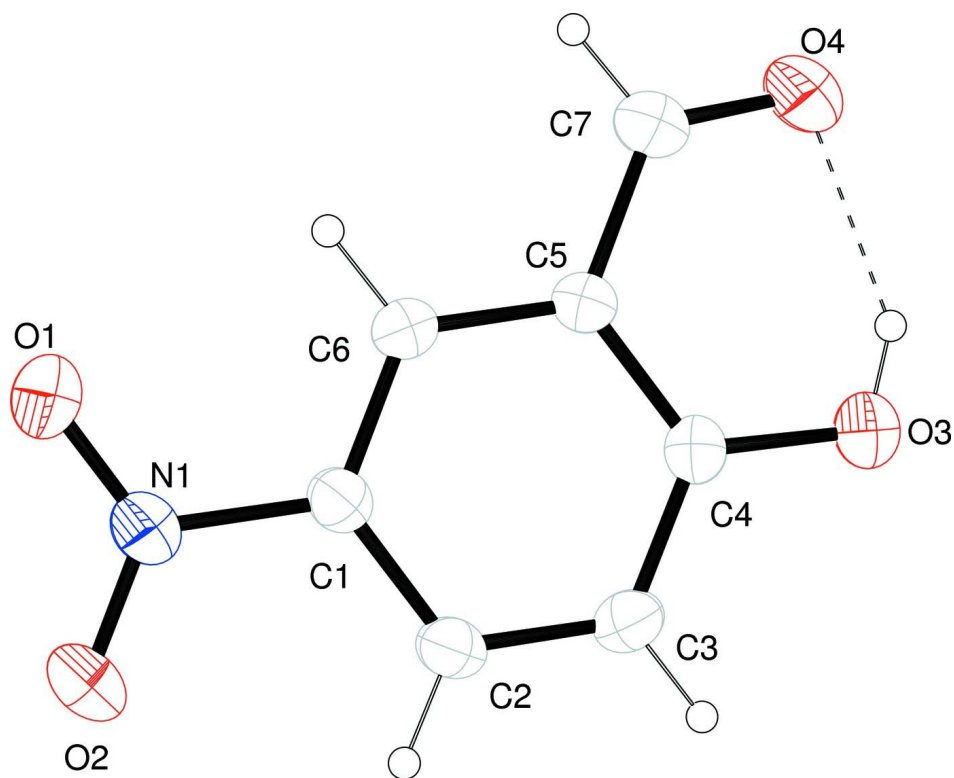
An intramolecular O3-H33...O4 interaction (Table 1, and Fig. 1) generates an S(6) ring motif (Bernstein *et al.*, 1995). In the crystal structure, the molecules are linked by intermolecular C2-H2...O4, C6-H6...O3 and C7-H7...O7 interactions into a three-dimensional framework. The C-H...O interactions result in the formation of C(5) chain but also R₂²(8) ring forming an approximately planar network parallel to the (1 0 -1) plane (Fig. 2). These planes are interconnected through π - π interaction which occurs between Cg1 (the centroid of the C1-C6 ring) and its symmetry equivalent at (-x,-y,-z), with a centroid-to-centroid distance of 3.582 (2) Å, a plane-to-plane separation of 3.367 (1) Å and a slippage of 1.22 Å.

S2. Experimental

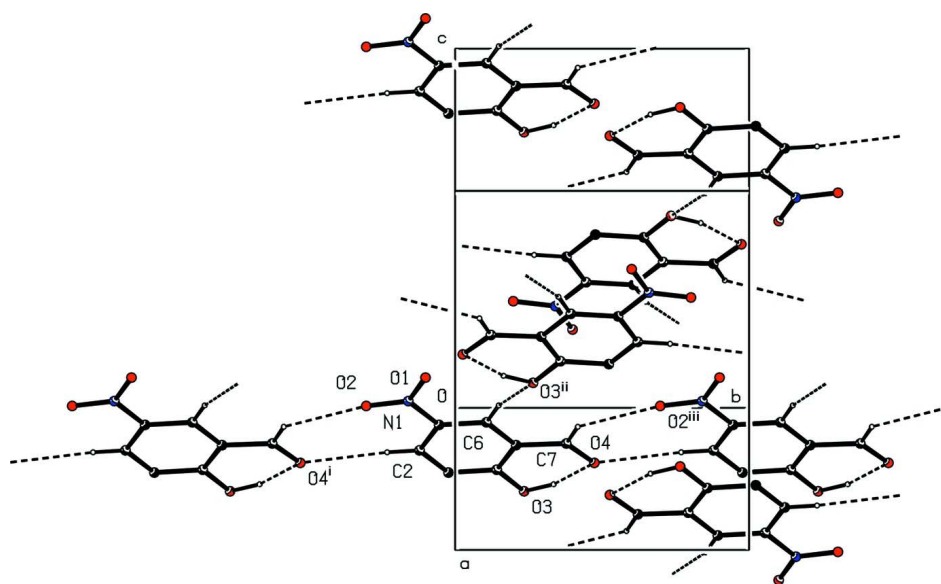
The commercially available compound (Acros Organics) was recrystallized from ethanol.

S3. Refinement

C-bound H atoms were positioned geometrically and refined using a riding model, with C—H = 0.93 Å and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$. The position of the H3A atom was obtained from a difference map of the electron density in the unit-cell and its coordinates were refined freely with $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{O})$.

**Figure 1**

The molecular structure of the title compound, showing the atom-numbering scheme and 50% probability displacement ellipsoids. H atoms are represented as small spheres of arbitrary radii.

**Figure 2**

Partial packing view showing the formation of C-H...O hydrogen bonds represented as dashed lines. H atoms not involved in hydrogen bonds have been omitted for clarity. [Symmetry codes: (i) $x, y-1, z$; (ii) $x+1/2, -y+1/2, z+1/2$; (iii) $x, y+1, z$]

2-Hydroxy-5-nitrobenzaldehyde

Crystal data

C₇H₅NO₄ $M_r = 167.12$ Monoclinic, $P2_1/n$

Hall symbol: -P 2yn

 $a = 7.2580 (17) \text{ \AA}$ $b = 8.3960 (13) \text{ \AA}$ $c = 11.704 (3) \text{ \AA}$ $\beta = 95.165 (18)^\circ$ $V = 710.3 (3) \text{ \AA}^3$ $Z = 4$ $F(000) = 344$ $D_x = 1.563 \text{ Mg m}^{-3}$ Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 8368 reflections

 $\theta = 1.8\text{--}27.3^\circ$ $\mu = 0.13 \text{ mm}^{-1}$ $T = 296 \text{ K}$

Prism., red

 $0.54 \times 0.28 \times 0.15 \text{ mm}$

Data collection

Stoe IPDS II

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

Detector resolution: 6.67 pixels mm^{-1}

rotation method scans

Absorption correction: integration

 $(X\text{-RED32}; \text{Stoe \& Cie, 2002})$ $T_{\min} = 0.979, T_{\max} = 0.992$

4345 measured reflections

1396 independent reflections

944 reflections with $I > 2\sigma(I)$ $R_{\text{int}} = 0.062$ $\theta_{\max} = 26.0^\circ, \theta_{\min} = 3.0^\circ$ $h = -8 \rightarrow 8$ $k = -10 \rightarrow 10$ $l = -14 \rightarrow 14$

Refinement

Refinement on F^2

Least-squares matrix: full

 $R[F^2 > 2\sigma(F^2)] = 0.050$ $wR(F^2) = 0.119$ $S = 1.06$

1396 reflections

112 parameters

0 restraints

Primary atom site location: structure-invariant

direct methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H atoms treated by a mixture of independent and constrained refinement

 $w = 1/[\sigma^2(F_o^2) + (0.05P)^2 + 0.0581P]$ where $P = (F_o^2 + 2F_c^2)/3$ $(\Delta/\sigma)_{\max} < 0.001$ $\Delta\rho_{\max} = 0.16 \text{ e \AA}^{-3}$ $\Delta\rho_{\min} = -0.15 \text{ e \AA}^{-3}$

Special details

Experimental. 168 frames, detector distance = 120 mm**Geometry.** All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.**Refinement.** Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.4340 (3)	-0.1006 (2)	0.25586 (14)	0.0788 (6)
O2	0.3658 (3)	-0.3022 (2)	0.15003 (16)	0.0819 (6)

O3	0.1019 (3)	0.2346 (2)	-0.19209 (14)	0.0702 (6)
H3A	0.130 (4)	0.336 (4)	-0.162 (3)	0.105*
O4	0.2251 (3)	0.4736 (2)	-0.06415 (16)	0.0895 (7)
N1	0.3709 (3)	-0.1591 (2)	0.16564 (16)	0.0524 (5)
C1	0.3019 (3)	-0.0545 (2)	0.07198 (17)	0.0425 (5)
C2	0.2232 (3)	-0.1209 (2)	-0.03011 (17)	0.0459 (5)
H2	0.2154	-0.2309	-0.0385	0.055*
C3	0.1574 (3)	-0.0228 (3)	-0.11799 (18)	0.0494 (5)
H3	0.1037	-0.0659	-0.1863	0.059*
C4	0.1711 (3)	0.1417 (2)	-0.10464 (18)	0.0471 (5)
C5	0.2551 (3)	0.2073 (2)	-0.00295 (17)	0.0437 (5)
C6	0.3190 (3)	0.1067 (2)	0.08585 (17)	0.0431 (5)
H6	0.3730	0.1485	0.1544	0.052*
C7	0.2771 (3)	0.3784 (3)	0.0104 (2)	0.0631 (7)
H7	0.3338	0.4172	0.0792	0.076*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.1123 (15)	0.0673 (12)	0.0522 (10)	-0.0024 (11)	-0.0171 (10)	0.0069 (9)
O2	0.1202 (16)	0.0391 (10)	0.0829 (13)	0.0026 (9)	-0.0101 (11)	0.0101 (8)
O3	0.0914 (13)	0.0556 (11)	0.0590 (11)	0.0048 (9)	-0.0195 (9)	0.0061 (8)
O4	0.1338 (17)	0.0423 (10)	0.0871 (14)	0.0022 (11)	-0.0199 (12)	0.0094 (9)
N1	0.0593 (11)	0.0433 (12)	0.0547 (11)	-0.0005 (9)	0.0047 (9)	0.0071 (9)
C1	0.0418 (11)	0.0382 (12)	0.0475 (12)	-0.0002 (9)	0.0041 (9)	0.0023 (9)
C2	0.0486 (12)	0.0355 (10)	0.0535 (12)	-0.0007 (9)	0.0046 (9)	-0.0025 (9)
C3	0.0526 (12)	0.0492 (13)	0.0453 (12)	-0.0032 (10)	-0.0013 (9)	-0.0086 (10)
C4	0.0463 (11)	0.0458 (13)	0.0480 (12)	0.0029 (9)	-0.0018 (9)	0.0028 (10)
C5	0.0468 (11)	0.0366 (11)	0.0470 (12)	-0.0014 (9)	0.0007 (9)	-0.0008 (9)
C6	0.0457 (11)	0.0418 (12)	0.0413 (11)	-0.0031 (9)	0.0009 (9)	-0.0049 (9)
C7	0.0805 (17)	0.0424 (13)	0.0643 (15)	-0.0020 (12)	-0.0045 (12)	-0.0008 (12)

Geometric parameters (Å, °)

O1—N1	1.216 (2)	C2—H2	0.9300
O2—N1	1.215 (2)	C3—C4	1.392 (3)
O3—C4	1.348 (2)	C3—H3	0.9300
O3—H3A	0.93 (3)	C4—C5	1.401 (3)
O4—C7	1.218 (3)	C5—C6	1.386 (3)
N1—C1	1.458 (3)	C5—C7	1.452 (3)
C1—C6	1.367 (3)	C6—H6	0.9300
C1—C2	1.394 (3)	C7—H7	0.9300
C2—C3	1.370 (3)		
C4—O3—H3A	100.5 (19)	O3—C4—C3	118.12 (19)
O2—N1—O1	122.27 (19)	O3—C4—C5	121.43 (19)
O2—N1—C1	118.59 (19)	C3—C4—C5	120.45 (19)
O1—N1—C1	119.13 (18)	C6—C5—C4	119.19 (18)

C6—C1—C2	121.56 (19)	C6—C5—C7	119.77 (19)
C6—C1—N1	119.05 (18)	C4—C5—C7	121.04 (19)
C2—C1—N1	119.38 (18)	C1—C6—C5	119.59 (19)
C3—C2—C1	119.5 (2)	C1—C6—H6	120.2
C3—C2—H2	120.3	C5—C6—H6	120.2
C1—C2—H2	120.3	O4—C7—C5	123.3 (2)
C2—C3—C4	119.71 (19)	O4—C7—H7	118.4
C2—C3—H3	120.1	C5—C7—H7	118.4
C4—C3—H3	120.1		

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
C2—H2...O4 ⁱ	0.93	2.50	3.427 (3)	175
C6—H6...O3 ⁱⁱ	0.93	2.53	3.433 (3)	163
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