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Bis[μ -pentane-2,4-dionato(1-)]bis[aqua-[1,1,1,5,5,5-hexafluoropentane-2,4-dionato(1-)]cobalt(II)]

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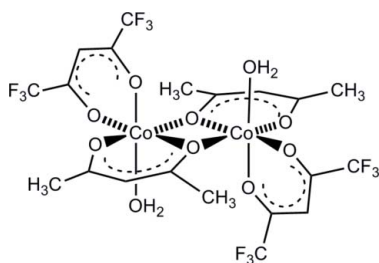
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Key indicators: single-crystal X-ray study; $T = 100$ K; mean $\sigma(\text{C}-\text{C}) = 0.007$ Å; R factor = 0.058; wR factor = 0.119; data-to-parameter ratio = 16.3.

The title complex, $[\text{Co}_2(\text{C}_5\text{HF}_6\text{O}_2)_2(\text{C}_5\text{H}_7\text{O}_2)_2(\text{H}_2\text{O})_2]$, is centrosymmetric with a crystallographic inversion center in the middle of the molecule. The octahedrally coordinated Co^{II} atoms are bridged by two chelating acetylacetonate (acac) ligands and two more electron-poor 1,1,1,5,5,5-hexafluoropentane-2,4-dionato (hfac) ligands are bonded terminally in a solely chelating manner. The coordinated water molecules form intermolecular $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds with electron-rich acac O atoms of neighboring molecules, leading to strings of molecules along the a axis.

Related literature

For mass spectrometry of β -diketonates, see: Reichert & Westmore (1969); Westmore (1976); Lerach & Leskiw (2008). For applications of β -diketonate complexes, see: Condorelli *et al.* (2007); Silvennoinen *et al.* (2007); Fahlmen (2006). For related structures, see: Hunter *et al.* (2009a,b); Lerach *et al.* (2007); Cotton & Elder (1966); McCann *et al.* (2001).



Experimental

Crystal data

 $[\text{Co}_2(\text{C}_5\text{HF}_6\text{O}_2)_2(\text{C}_5\text{H}_7\text{O}_2)_2(\text{H}_2\text{O})_2]$
 $M_r = 766.22$ Triclinic, $P\bar{1}$ $a = 7.563$ (3) Å $b = 9.541$ (4) Å $c = 9.716$ (4) Å $\alpha = 94.865$ (6)° $\beta = 92.792$ (6)°
 $\gamma = 93.622$ (6)°
 $V = 696.2$ (5) Å³
 $Z = 1$
 Mo $K\alpha$ radiation

 $\mu = 1.32$ mm⁻¹
 $T = 100$ K
 $0.20 \times 0.16 \times 0.08$ mm

Data collection

 Bruker APEXII CCD
 diffractometer
 Absorption correction: multi-scan
 (SADABS; Bruker, 2008)
 $T_{\text{min}} = 0.717$, $T_{\text{max}} = 0.900$

 6687 measured reflections
 3378 independent reflections
 1937 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.073$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.058$ $wR(F^2) = 0.119$ $S = 0.97$

3378 reflections

207 parameters

2 restraints

 H atoms treated by a mixture of
 independent and constrained
 refinement
 $\Delta\rho_{\text{max}} = 0.63$ e Å⁻³ $\Delta\rho_{\text{min}} = -0.73$ e Å⁻³
Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O5}-\text{H5A}\cdots\text{O2}^{\text{i}}$	0.83 (4)	2.25 (3)	2.973 (4)	147 (5)
$\text{O5}-\text{H5B}\cdots\text{O3}^{\text{ii}}$	0.84 (2)	1.87 (2)	2.703 (4)	169 (5)

Symmetry codes: (i) $-x, -y, -z + 1$; (ii) $-x + 1, -y, -z + 1$.

Data collection: APEX2 (Bruker, 2008); cell refinement: SAINT (Bruker, 2008); data reduction: SAINT; program(s) used to solve structure: SHELXTL (Sheldrick, 2008); program(s) used to refine structure: SHELXTL; molecular graphics: SHELXTL; software used to prepare material for publication: SHELXTL.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: LX2118).

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supporting information

Acta Cryst. (2009). E65, m1476 [https://doi.org/10.1107/S1600536809044389]

Bis[μ -pentane-2,4-dionato(1-)]bis{aqua[1,1,1,5,5,5-hexafluoropentane-2,4-dionato(1-)]cobalt(II)}

Gerald O. Hunter, Matthias Zeller and Brian D. Leskiw

S1. Comment

Our interest in β -diketonate complexes, specifically fluorinated acetylacetonate (acac) derivatives, stems from their volatility and most notably their ability to undergo both partial and complete ligand exchange reactions. Recent applications of β -diketonate complexes can be seen in the areas of catalysis and microelectronics (Silvennoinen *et al.* 2007), and the deposition of metallic or ceramic thin films (Condorelli *et al.* 2007). Furthermore, β -diketonates are ideally suited as precursors for vapor deposition processes (Fahlmen 2006). Research in the area of ligand exchange reactions is being conducted with the goal of observing gas-phase reactions and ligand exchange *via* mass spectrometry (Lerach & Leskiw 2008).

In previous structure reports on complexes with the 1,1,1-trifluoro-5,5-dimethyl-2,4-hexanedione (tftm) ligand, we reported three monometallic complexes with zinc, nickel and cobalt as the metal (Hunter *et al.* 2009a and 2009b, Lerach *et al.* 2007). Using a mixture of two acetylacetonate ligands, parent pentane-2,4-dionate (acac) and the hexafluoro derivative 1,1,1,5,5,5-hexafluoropentane-2,4-dionate (hfac), the title compound was obtained as a dimeric biscobalt complex after purification by sublimation and recrystallization (Figure 1). The complex is centrosymmetric with the center of the complex being located on a crystallographic inversion center, and each of the metal centers exhibits an only slightly distorted octahedral coordination environment of six oxygen atoms as expected for Co(II) complexes. The ligand environment of each metal center is composed of a chelating hfac ligand, one coordinated water molecule and two chelating bridging acac ligands. The connection between the two metal ions is facilitated by the two μ -oxygen atoms from these two acac ligands. The coordination modes of the two types of β -diketonate ligands are thus quite distinct based on the electron densities available at the oxygen atoms of the ligands. The less electron rich hfac ligand is a chelating terminal ligand, the more electron rich acac ligands are chelating and bridging between the metal centers. One of the oxygen atoms of the hfac ligand is *trans* to the coordinated water molecule, the other *trans* to the bridging acac μ -O4 atom.

The general motif for this dimeric structure is not unknown. For cobalt, dimeric complexes similar to the title compound were for example reported with only acetyl acetonate as the ligand rather than two different acac derivatives. Structures are known for bis(aqua-(μ^2 -pentane-2,4-dionato-*O,O'*)-(pentane-2,4-dionato)-cobalt(II)) itself (Cotton & Elder, 1966) (but the quality of the structure is very low) and as a co-crystal with tetra-aqua-(acetylacetonato)-cobalt(II) perchlorate (McCann *et al.*, 2001). In both structures the dimeric complexes exhibit the same centrosymmetric structural motif with the same coordination arrangement of acac and water ligands as in the title compound. The metal-oxygen bonding distances in the title compound and the well resolved structure are the same within 0.04 Å.

The coordinated water molecules are involved in hydrogen bonding interactions (Table 1). An intramolecular hydrogen bond stabilizes the dimeric structure in both the title structure and the acac parent complex. In the title compound these

hydrogen bonds are oriented towards the neighboring oxygen atom $O2^i$ of the hfac ligand (symmetry operator (i): $-x, -y, -z + 1$). The other H atom of the water molecule makes a strong intermolecular H bond to $O3^{ii}$ in a neighboring molecule (symmetry operator (ii): $-x + 1, -y, -z + 1$). The intermolecular hydrogen bonds are arranged in inversion symmetric pairs that connect molecules along the a -axis leading to strongly hydrogen bonded strings of molecules along that axis (Figure 2). Individual interactions between these strings of molecules, on the other hand, are weak and are mostly based on shape recognition of the acac and hfac ligands (Figure 3).

It should be stated that this oxygen atom O3 is probably the most electron rich in the dimer (being an acac O atom and not bridging) and the O—H \cdots O hydrogen bond formed is thus the strongest one possible in this system. It could therefore be assumed that the packing of the molecules is at least partially based on the ability to form this strong hydrogen bond (rather than a weaker one towards one of the less electron rich O atoms). This is however at least partially ruled out by the fact that the acac-only complex (Cotton & Elder, 1966) adopts the same hydrogen bonding motif with infinite hydrogen bonded chains where the hydrogen bonding acceptor is the monodentate oxygen atom of the bridging acac ligand. Other influences than only the electron donor ability of hydrogen atom acceptor thus must play an important role as well, which might be found among the ability to form pairwise hydrogen bonds, shape recognition between the molecules, or preassembly of hydrogen bonded chains in solution before crystallization.

S2. Experimental

The synthesis of the title compound involved equal molar concentrations (1.0 mmol) of both cobalt acetylacetonate and hexafluoroacetylacetonate, dissolved in concentrated methanol, and being reacted under a steady reflux for forty-eight hours. The solvent was subsequently removed *in vacuo*, and the desired product was purified *via* sublimation under vacuum. The desired product was re-crystallized overnight by vapor diffusion of hexanes into a solution of diethyl ether.

S3. Refinement

The water H atoms were located in a difference density Fourier map. The O—H distances were restrained to 0.84 (2) Å. All other H atoms were placed in calculated positions with C—H distances of 0.98 (methyl) and 0.95 Å (CH). The methyl and hydroxyl H's were refined with an isotropic displacement parameter U_{iso} of 1.5 times U_{eq} of the adjacent carbon or oxygen atom, and the C—H hydrogen atom with $U_{iso} = 1.2 U_{eq}(C)$. Methyl hydrogen atoms were allowed to rotate to best fit the experimental electron density.

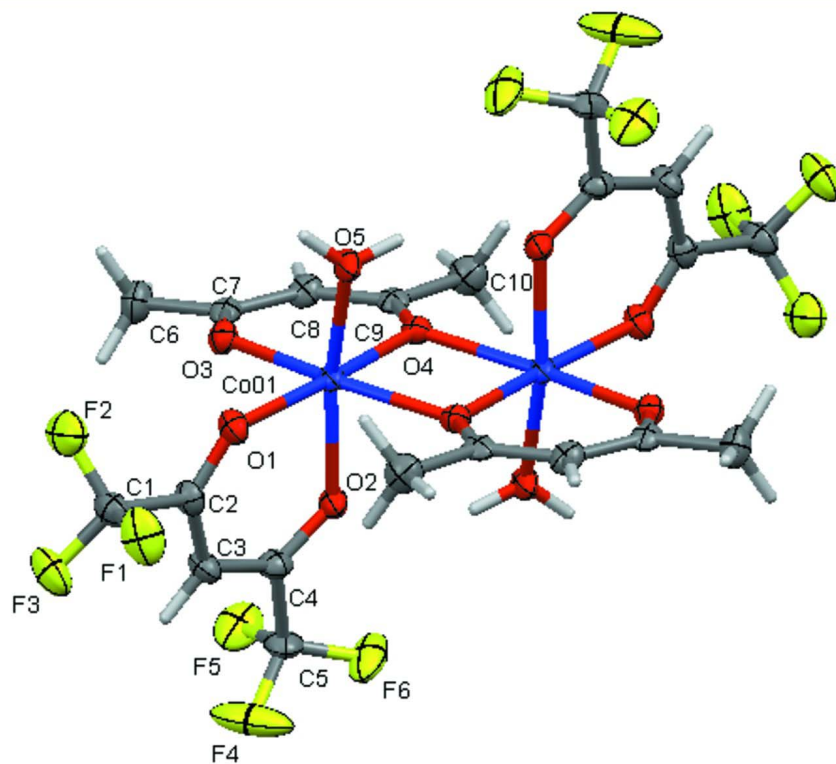
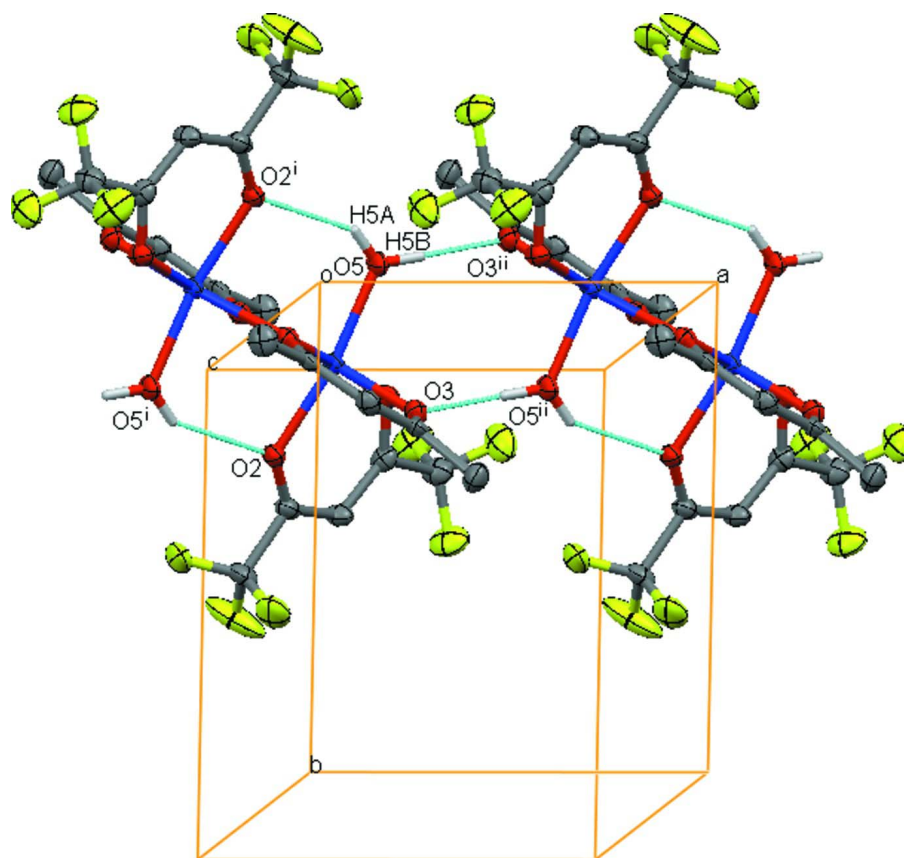


Figure 1

ORTEP representation of the title compound (50% probability displacement ellipsoids)

**Figure 2**

View of a section of one of the hydrogen bonded chains along the *a* axis. Hydrogen bonds are symbolized by blue dashed lines. Symmetry codes: (i) $-x, -y, -z + 1$; (ii) $-x + 1, -y, -z + 1$.

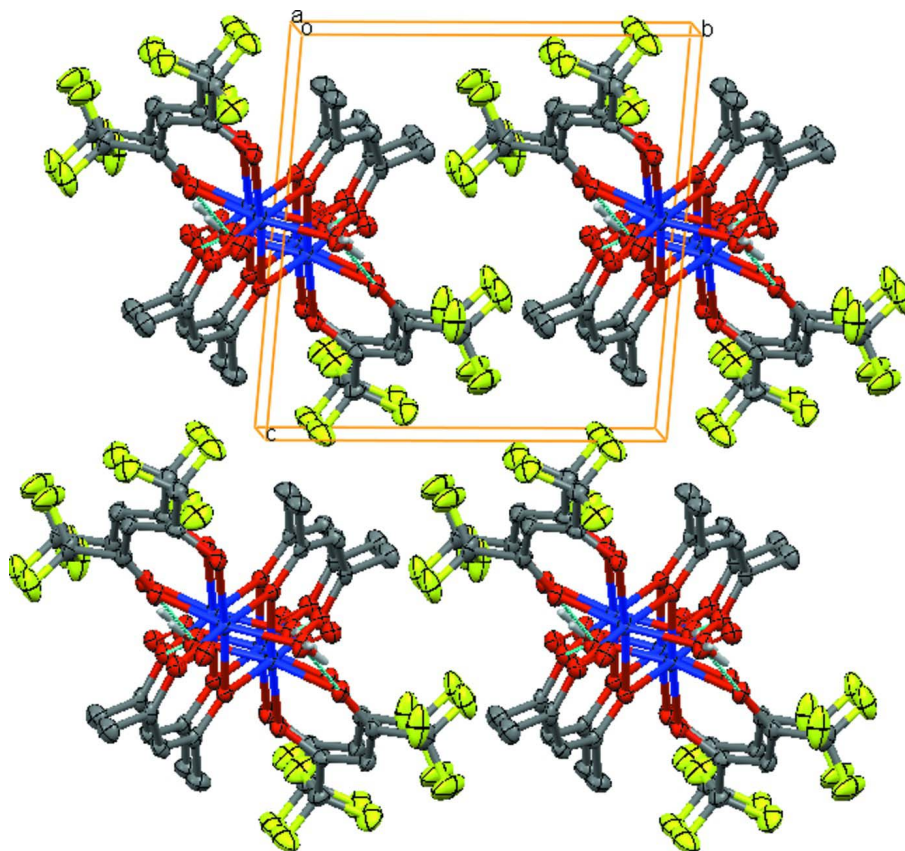


Figure 3

Packing view of the title structure, view down the *a* axis. Hydrogen bonds are symbolized by blue dashed lines.

Bis[μ -pentane-2,4-dionato(1-)]bis[aqua[1,1,1,5,5,5-hexafluoropentane-2,4-dionato(1-)]cobalt(II)]

Crystal data

[Co₂(C₅HF₆O₂)₂(C₅H₇O₂)₂(H₂O)₂]

$M_r = 766.22$

Triclinic, $P\bar{1}$

Hall symbol: -P 1

$a = 7.563$ (3) Å

$b = 9.541$ (4) Å

$c = 9.716$ (4) Å

$\alpha = 94.865$ (6)°

$\beta = 92.792$ (6)°

$\gamma = 93.622$ (6)°

$V = 696.2$ (5) Å³

$Z = 1$

$F(000) = 382$

$D_x = 1.828$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 1148 reflections

$\theta = 2.7$ – 30.6 °

$\mu = 1.32$ mm⁻¹

$T = 100$ K

Block, red

$0.20 \times 0.16 \times 0.08$ mm

Data collection

Bruker APEXII CCD
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

ω scans

Absorption correction: multi-scan

(*SADABS*; Bruker, 2008)

$T_{\min} = 0.717$, $T_{\max} = 0.900$

6687 measured reflections

3378 independent reflections

1937 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.073$

$\theta_{\max} = 28.3$ °, $\theta_{\min} = 2.1$ °

$h = -9 \rightarrow 10$

$k = -12 \rightarrow 12$

$l = -12 \rightarrow 12$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.058$
 $wR(F^2) = 0.119$
 $S = 0.97$
 3378 reflections
 207 parameters
 2 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: difference Fourier map
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.0425P)^2]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} < 0.001$
 $\Delta\rho_{\max} = 0.63 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.73 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.5543 (6)	0.2269 (5)	0.8934 (5)	0.0302 (11)
C2	0.3920 (6)	0.2188 (5)	0.7936 (5)	0.0242 (10)
C3	0.2836 (6)	0.3329 (5)	0.7993 (5)	0.0259 (11)
H3	0.3165	0.4122	0.8634	0.031*
C4	0.1302 (6)	0.3346 (5)	0.7153 (5)	0.0247 (10)
C5	0.0245 (6)	0.4664 (5)	0.7323 (5)	0.0315 (12)
C6	0.4853 (6)	0.3448 (5)	0.2998 (5)	0.0323 (12)
H6A	0.5072	0.4195	0.3754	0.048*
H6B	0.4443	0.3858	0.2157	0.048*
H6C	0.5953	0.2986	0.2830	0.048*
C7	0.3458 (5)	0.2382 (4)	0.3383 (5)	0.0228 (10)
C8	0.1962 (5)	0.2021 (5)	0.2477 (5)	0.0236 (10)
H8	0.1960	0.2434	0.1621	0.028*
C9	0.0496 (5)	0.1142 (5)	0.2678 (5)	0.0212 (10)
C10	-0.0963 (6)	0.0916 (5)	0.1565 (5)	0.0325 (12)
H10A	-0.1091	-0.0083	0.1220	0.049*
H10B	-0.0674	0.1487	0.0804	0.049*
H10C	-0.2079	0.1196	0.1943	0.049*
Co01	0.19236 (7)	0.06799 (6)	0.54981 (6)	0.01780 (18)
F1	0.5216 (4)	0.1512 (3)	0.9985 (3)	0.0505 (9)
F2	0.6953 (4)	0.1786 (3)	0.8342 (3)	0.0498 (9)
F3	0.6031 (4)	0.3593 (3)	0.9485 (3)	0.0465 (8)
F4	0.0423 (5)	0.5320 (4)	0.8555 (3)	0.0806 (13)
F5	0.0781 (4)	0.5585 (3)	0.6446 (3)	0.0502 (9)
F6	-0.1465 (4)	0.4379 (3)	0.7000 (4)	0.0579 (10)

O1	0.3727 (4)	0.1059 (3)	0.7171 (3)	0.0241 (7)
O2	0.0665 (4)	0.2433 (3)	0.6216 (3)	0.0229 (7)
O3	0.3733 (4)	0.1857 (3)	0.4531 (3)	0.0243 (7)
O4	0.0264 (3)	0.0478 (3)	0.3757 (3)	0.0198 (7)
O5	0.2844 (4)	-0.1274 (3)	0.5039 (4)	0.0271 (8)
H5A	0.214 (5)	-0.188 (4)	0.462 (5)	0.041*
H5B	0.393 (3)	-0.143 (5)	0.506 (5)	0.041*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.033 (3)	0.029 (3)	0.027 (3)	0.003 (2)	-0.006 (2)	-0.001 (2)
C2	0.025 (2)	0.026 (3)	0.021 (3)	-0.001 (2)	-0.004 (2)	0.004 (2)
C3	0.030 (2)	0.019 (3)	0.027 (3)	0.002 (2)	-0.004 (2)	-0.003 (2)
C4	0.028 (2)	0.018 (2)	0.028 (3)	0.005 (2)	0.005 (2)	0.002 (2)
C5	0.033 (3)	0.022 (3)	0.040 (3)	0.010 (2)	0.005 (2)	-0.006 (2)
C6	0.025 (2)	0.033 (3)	0.039 (3)	-0.003 (2)	0.001 (2)	0.009 (2)
C7	0.021 (2)	0.014 (2)	0.034 (3)	0.0067 (18)	0.005 (2)	0.002 (2)
C8	0.023 (2)	0.024 (3)	0.026 (3)	0.0076 (19)	0.001 (2)	0.008 (2)
C9	0.021 (2)	0.017 (2)	0.026 (3)	0.0105 (18)	-0.0004 (19)	-0.001 (2)
C10	0.033 (3)	0.039 (3)	0.027 (3)	0.005 (2)	-0.002 (2)	0.007 (2)
Co01	0.0150 (3)	0.0169 (3)	0.0214 (3)	0.0046 (2)	-0.0020 (2)	0.0003 (2)
F1	0.056 (2)	0.054 (2)	0.0414 (19)	-0.0057 (16)	-0.0206 (15)	0.0201 (17)
F2	0.0334 (16)	0.063 (2)	0.050 (2)	0.0093 (15)	-0.0129 (15)	-0.0097 (17)
F3	0.0498 (18)	0.0393 (19)	0.0461 (19)	-0.0049 (14)	-0.0192 (15)	-0.0035 (15)
F4	0.133 (3)	0.071 (3)	0.040 (2)	0.070 (2)	-0.013 (2)	-0.017 (2)
F5	0.0483 (19)	0.0323 (18)	0.075 (2)	0.0155 (14)	0.0118 (17)	0.0211 (17)
F6	0.0275 (16)	0.0332 (19)	0.115 (3)	0.0105 (13)	0.0148 (17)	0.0020 (19)
O1	0.0226 (16)	0.0200 (18)	0.0289 (18)	0.0034 (13)	-0.0072 (14)	0.0019 (15)
O2	0.0187 (15)	0.0201 (17)	0.0294 (18)	0.0041 (13)	-0.0019 (13)	-0.0006 (15)
O3	0.0184 (16)	0.0259 (18)	0.0290 (19)	0.0031 (13)	-0.0012 (14)	0.0040 (15)
O4	0.0199 (15)	0.0203 (17)	0.0194 (17)	0.0046 (13)	-0.0004 (13)	0.0016 (14)
O5	0.0153 (16)	0.0252 (19)	0.039 (2)	0.0047 (14)	-0.0057 (15)	-0.0057 (16)

Geometric parameters (Å, °)

C1—F1	1.324 (5)	C7—C8	1.406 (6)
C1—F2	1.325 (5)	C8—C9	1.380 (6)
C1—F3	1.352 (5)	C8—H8	0.9500
C1—C2	1.521 (6)	C9—O4	1.284 (5)
C2—O1	1.252 (5)	C9—C10	1.499 (6)
C2—C3	1.402 (6)	C10—H10A	0.9800
C3—C4	1.388 (6)	C10—H10B	0.9800
C3—H3	0.9500	C10—H10C	0.9800
C4—O2	1.261 (5)	Co01—O3	2.032 (3)
C4—C5	1.534 (6)	Co01—O4	2.045 (3)
C5—F4	1.300 (5)	Co01—O5	2.052 (3)
C5—F6	1.321 (5)	Co01—O1	2.063 (3)

C5—F5	1.334 (6)	Co01—O2	2.064 (3)
C6—C7	1.501 (6)	Co01—O4 ⁱ	2.122 (3)
C6—H6A	0.9800	O4—Co01 ⁱ	2.122 (3)
C6—H6B	0.9800	O5—H5A	0.83 (4)
C6—H6C	0.9800	O5—H5B	0.84 (2)
C7—O3	1.274 (5)		
F1—C1—F2	107.6 (4)	O4—C9—C8	125.2 (4)
F1—C1—F3	106.5 (4)	O4—C9—C10	116.0 (4)
F2—C1—F3	106.4 (4)	C8—C9—C10	118.9 (4)
F1—C1—C2	110.1 (4)	C9—C10—H10A	109.5
F2—C1—C2	112.5 (4)	C9—C10—H10B	109.5
F3—C1—C2	113.3 (4)	H10A—C10—H10B	109.5
O1—C2—C3	128.6 (4)	C9—C10—H10C	109.5
O1—C2—C1	113.3 (4)	H10A—C10—H10C	109.5
C3—C2—C1	118.1 (4)	H10B—C10—H10C	109.5
C4—C3—C2	122.6 (4)	O3—Co01—O4	90.52 (12)
C4—C3—H3	118.7	O3—Co01—O5	99.03 (13)
C2—C3—H3	118.7	O4—Co01—O5	91.97 (12)
O2—C4—C3	128.8 (4)	O3—Co01—O1	83.91 (12)
O2—C4—C5	114.1 (4)	O4—Co01—O1	174.10 (12)
C3—C4—C5	117.1 (4)	O5—Co01—O1	90.82 (12)
F4—C5—F6	108.5 (4)	O3—Co01—O2	92.44 (12)
F4—C5—F5	106.8 (4)	O4—Co01—O2	89.58 (11)
F6—C5—F5	105.1 (4)	O5—Co01—O2	168.41 (12)
F4—C5—C4	113.9 (4)	O1—Co01—O2	88.73 (12)
F6—C5—C4	112.2 (4)	O3—Co01—O4 ⁱ	170.59 (12)
F5—C5—C4	109.8 (4)	O4—Co01—O4 ⁱ	80.47 (12)
C7—C6—H6A	109.5	O5—Co01—O4 ⁱ	84.06 (12)
C7—C6—H6B	109.5	O1—Co01—O4 ⁱ	105.00 (11)
H6A—C6—H6B	109.5	O2—Co01—O4 ⁱ	84.87 (11)
C7—C6—H6C	109.5	C2—O1—Co01	124.6 (3)
H6A—C6—H6C	109.5	C4—O2—Co01	124.7 (3)
H6B—C6—H6C	109.5	C7—O3—Co01	125.7 (3)
O3—C7—C8	124.3 (4)	C9—O4—Co01	125.1 (3)
O3—C7—C6	116.5 (4)	C9—O4—Co01 ⁱ	134.2 (3)
C8—C7—C6	119.2 (4)	Co01—O4—Co01 ⁱ	99.53 (12)
C9—C8—C7	128.0 (4)	Co01—O5—H5A	117 (4)
C9—C8—H8	116.0	Co01—O5—H5B	123 (3)
C7—C8—H8	116.0	H5A—O5—H5B	118 (5)
F1—C1—C2—O1	-77.7 (5)	C3—C4—O2—Co01	-7.2 (7)
F2—C1—C2—O1	42.3 (6)	C5—C4—O2—Co01	170.3 (3)
F3—C1—C2—O1	163.1 (4)	O3—Co01—O2—C4	-71.2 (3)
F1—C1—C2—C3	101.1 (5)	O4—Co01—O2—C4	-161.7 (3)
F2—C1—C2—C3	-138.9 (4)	O5—Co01—O2—C4	100.5 (7)
F3—C1—C2—C3	-18.1 (6)	O1—Co01—O2—C4	12.7 (3)
O1—C2—C3—C4	0.2 (7)	O4 ⁱ —Co01—O2—C4	117.8 (3)

C1—C2—C3—C4	-178.4 (4)	C8—C7—O3—Co01	-12.1 (6)
C2—C3—C4—O2	-2.4 (8)	C6—C7—O3—Co01	168.2 (3)
C2—C3—C4—C5	-179.9 (4)	O4—Co01—O3—C7	12.0 (3)
O2—C4—C5—F4	154.2 (4)	O5—Co01—O3—C7	104.1 (3)
C3—C4—C5—F4	-27.9 (6)	O1—Co01—O3—C7	-166.0 (3)
O2—C4—C5—F6	30.5 (6)	O2—Co01—O3—C7	-77.6 (3)
C3—C4—C5—F6	-151.7 (4)	C8—C9—O4—Co01	4.9 (6)
O2—C4—C5—F5	-86.0 (5)	C10—C9—O4—Co01	-175.4 (3)
C3—C4—C5—F5	91.8 (5)	C8—C9—O4—Co01 ⁱ	169.5 (3)
O3—C7—C8—C9	4.8 (7)	C10—C9—O4—Co01 ⁱ	-10.8 (6)
C6—C7—C8—C9	-175.6 (4)	O3—Co01—O4—C9	-8.5 (3)
C7—C8—C9—O4	-0.8 (7)	O5—Co01—O4—C9	-107.5 (3)
C7—C8—C9—C10	179.5 (4)	O2—Co01—O4—C9	84.0 (3)
C3—C2—O1—Co01	11.1 (6)	O4 ⁱ —Co01—O4—C9	168.8 (4)
C1—C2—O1—Co01	-170.2 (3)	O3—Co01—O4—Co01 ⁱ	-177.30 (13)
O3—Co01—O1—C2	78.1 (3)	O5—Co01—O4—Co01 ⁱ	83.64 (13)
O5—Co01—O1—C2	177.1 (3)	O2—Co01—O4—Co01 ⁱ	-84.87 (12)
O2—Co01—O1—C2	-14.5 (3)	O4 ⁱ —Co01—O4—Co01 ⁱ	0.0
O4 ⁱ —Co01—O1—C2	-98.8 (3)		

Symmetry code: (i) $-x, -y, -z+1$.

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
O5—H5A \cdots O2 ⁱ	0.83 (4)	2.25 (3)	2.973 (4)	147 (5)
O5—H5B \cdots O3 ⁱⁱ	0.84 (2)	1.87 (2)	2.703 (4)	169 (5)

Symmetry codes: (i) $-x, -y, -z+1$; (ii) $-x+1, -y, -z+1$.