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## Structure Reports

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## 2-Nitrobenzyl 2-chloroacetate

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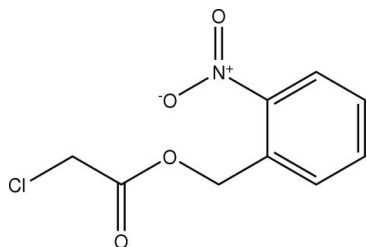
Received 6 August 2009; accepted 10 August 2009

 Key indicators: single-crystal X-ray study;  $T = 294$  K; mean  $\sigma(\text{C}-\text{C}) = 0.007$  Å;  $R$  factor = 0.067;  $wR$  factor = 0.195; data-to-parameter ratio = 13.8.

In the molecule of the title compound,  $\text{C}_9\text{H}_8\text{ClNO}_4$ , an intramolecular  $\text{C}-\text{H}\cdots\text{O}$  interaction results in the formation of a near-planar (r.m.s. deviation 0.002 Å) five-membered ring, which is oriented at a dihedral angle of 4.07 (4)° with respect to the adjacent aromatic ring. In the crystal structure, intermolecular  $\text{C}-\text{H}\cdots\text{O}$  interactions link the molecules into a two-dimensional network.

## Related literature

For a related structure, see: Pyun *et al.* (2001). For bond-length data, see: Allen *et al.* (1987).



## Experimental

## Crystal data

$\text{C}_9\text{H}_8\text{ClNO}_4$   
 $M_r = 229.61$   
 Monoclinic,  $P2_1/c$   
 $a = 8.0270$  (16) Å

$b = 6.7530$  (14) Å  
 $c = 19.266$  (4) Å  
 $\beta = 92.52$  (3)°  
 $V = 1043.3$  (4) Å<sup>3</sup>

$Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.36$  mm<sup>-1</sup>

$T = 294$  K  
 $0.30 \times 0.20 \times 0.10$  mm

## Data collection

Enraf-Nonius CAD-4 diffractometer  
 Absorption correction:  $\psi$  scan (North *et al.*, 1968)  
 $T_{\min} = 0.900$ ,  $T_{\max} = 0.965$   
 2036 measured reflections

1893 independent reflections  
 891 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.025$   
 3 standard reflections  
 frequency: 120 min  
 intensity decay: 1%

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.067$   
 $wR(F^2) = 0.195$   
 $S = 1.00$   
 1893 reflections

137 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.29$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.23$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{C1}-\text{H1A}\cdots\text{O3}^{\text{i}}$	0.97	2.43	3.372 (6)	166
$\text{C7}-\text{H7A}\cdots\text{O1}^{\text{ii}}$	0.93	2.58	3.374 (6)	143

Symmetry codes: (i)  $x - 1, y, z$ ; (ii)  $x, -y + \frac{1}{2}, z + \frac{1}{2}$ .

Data collection: *CAD-4 Software* (Enraf-Nonius, 1989); cell refinement: *CAD-4 Software*; data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997) and *PLATON* (Spek, 2009); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HK2752).

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## supporting information

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## 2-Nitrobenzyl 2-chloroacetate

Kai Zhu, Hui Liu, Yan-Hua Wang, Ping-Fang Han and Ping Wei

### S1. Comment

Some derivatives of *p*-nitrobenzyl alcohol are important chemical materials. We report herein the crystal structure of the title compound.

In the molecule of the title compound, (Fig. 1), the bond lengths (Allen *et al.*, 1987) and angles are within normal ranges. Ring A (C4–C9) is, of course, planar. Intramolecular C–H $\cdots$ O interaction (Table 1) results in the formation of a planar five-membered ring B (O2/C3/C4/C9/H9A), which is oriented with respect to the adjacent ring A at a dihedral angle of A/B = 4.07 (4) $^\circ$ .

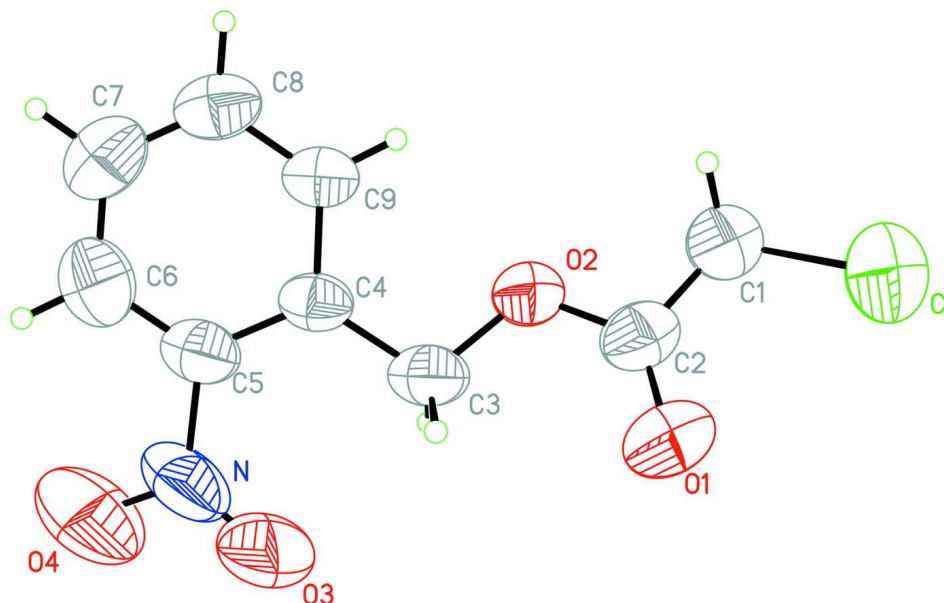
In the crystal structure, intermolecular C–H $\cdots$ O interactions (Table 1) link the molecules into a two dimensional network (Fig. 2), in which they may be effective in the stabilization of the structure.

### S2. Experimental

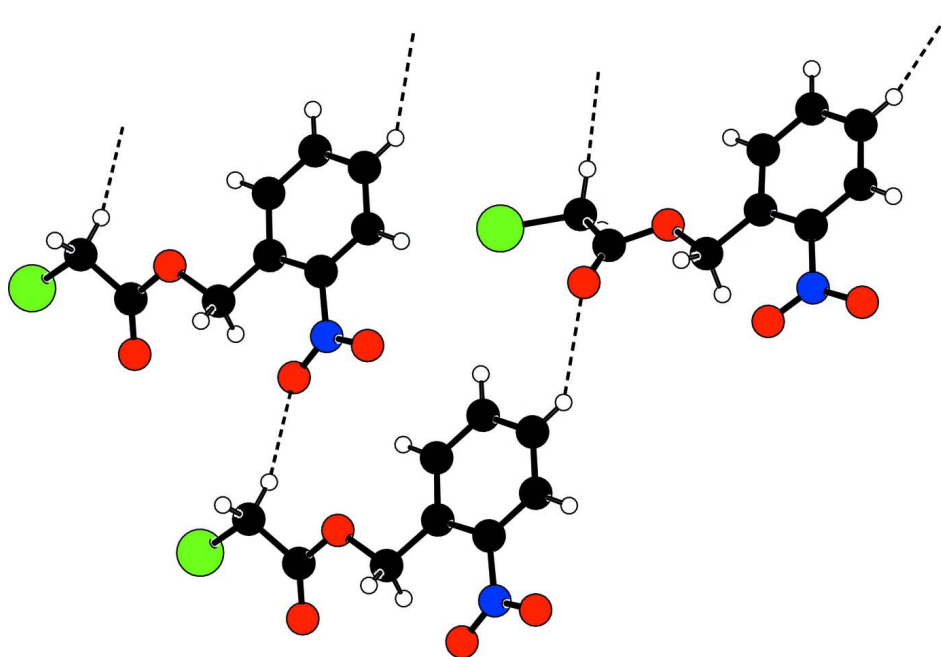
For the preparation of the title compound, chloroacetyl chloride (1.1 g) and 2-nitrobenzyl alcohol (1.53 g) were added into the mixture of pyridine (15 ml) and dichloromethane (30 ml) at 273–278 K. The gross products were extracted with *n*-hexane, washed with water and dried under vacuum, and then recrystallized from dichloromethane. Finally the title compound was obtained (yield; 0.61 g) (Pyun *et al.*, 2001). Crystals suitable for X-ray analysis were obtained by slow evaporation of a methanol solution.

### S3. Refinement

H atoms were positioned geometrically, with C–H = 0.93 and 0.97 Å for aromatic and methylene H, respectively, and constrained to ride on their parent atoms, with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ .

**Figure 1**

The molecular structure of the title molecule, with the atom-numbering scheme. Displacement ellipsoids are drawn at 30% probability level.

**Figure 2**

A partial packing diagram of the title compound. Hydrogen bonds are shown as dashed lines.

### 2-Nitrobenzyl 2-chloroacetate

#### *Crystal data*

$C_9H_8ClNO_4$   
 $M_r = 229.61$

Monoclinic,  $P2_1/c$   
Hall symbol: -P 2ybc

$a = 8.0270$  (16) Å  
 $b = 6.7530$  (14) Å  
 $c = 19.266$  (4) Å  
 $\beta = 92.52$  (3)°  
 $V = 1043.3$  (4) Å<sup>3</sup>  
 $Z = 4$   
 $F(000) = 472$   
 $D_x = 1.462$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å  
 Cell parameters from 25 reflections  
 $\theta = 9\text{--}13^\circ$   
 $\mu = 0.36$  mm<sup>-1</sup>  
 $T = 294$  K  
 Block, yellow  
 $0.30 \times 0.20 \times 0.10$  mm

*Data collection*

Enraf–Nonius CAD-4  
 diffractometer  
 Radiation source: fine-focus sealed tube  
 Graphite monochromator  
 $\omega/2\theta$  scans  
 Absorption correction:  $\psi$  scan  
 (North *et al.*, 1968)  
 $T_{\min} = 0.900$ ,  $T_{\max} = 0.965$   
 2036 measured reflections

1893 independent reflections  
 891 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.025$   
 $\theta_{\max} = 25.3^\circ$ ,  $\theta_{\min} = 2.1^\circ$   
 $h = 0 \rightarrow 9$   
 $k = 0 \rightarrow 8$   
 $l = -23 \rightarrow 23$   
 3 standard reflections every 120 min  
 intensity decay: 1%

*Refinement*

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.067$   
 $wR(F^2) = 0.195$   
 $S = 1.00$   
 1893 reflections  
 137 parameters  
 0 restraints  
 Primary atom site location: structure-invariant  
 direct methods  
 Secondary atom site location: difference Fourier  
 map

Hydrogen site location: inferred from  
 neighbouring sites  
 H-atom parameters constrained  
 $w = 1/[\sigma^2(F_o^2) + (0.07P)^2 + 0.84P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} < 0.001$   
 $\Delta\rho_{\max} = 0.29$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.23$  e Å<sup>-3</sup>  
 Extinction correction: *SHELXL97* (Sheldrick,  
 2008),  $F_c^* = kF_c[1 + 0.001x F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$   
 Extinction coefficient: 0.031 (4)

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å<sup>2</sup>)*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cl	-0.3781 (2)	0.3294 (4)	0.19943 (9)	0.1458 (10)
O1	-0.0286 (4)	0.3868 (6)	0.25207 (18)	0.0985 (12)
O2	-0.0679 (3)	0.2382 (5)	0.35365 (15)	0.0743 (9)
O3	0.4281 (5)	0.3113 (9)	0.3977 (2)	0.151 (2)
O4	0.5546 (5)	0.2492 (8)	0.4935 (3)	0.161 (2)
N	0.4264 (5)	0.2744 (8)	0.4576 (3)	0.1028 (16)

C1	-0.2908 (6)	0.2341 (10)	0.2755 (2)	0.1002 (18)
H1A	-0.3577	0.2745	0.3137	0.120*
H1B	-0.2939	0.0906	0.2730	0.120*
C2	-0.1141 (6)	0.2988 (7)	0.2904 (2)	0.0722 (13)
C3	0.1048 (5)	0.2696 (7)	0.3759 (2)	0.0722 (13)
H3A	0.1424	0.3976	0.3600	0.087*
H3B	0.1746	0.1680	0.3566	0.087*
C4	0.1165 (5)	0.2610 (6)	0.4535 (2)	0.0559 (10)
C5	0.2673 (5)	0.2629 (7)	0.4927 (2)	0.0706 (13)
C6	0.2753 (7)	0.2548 (8)	0.5642 (3)	0.0889 (15)
H6A	0.3780	0.2567	0.5885	0.107*
C7	0.1300 (8)	0.2440 (7)	0.5994 (3)	0.0874 (15)
H7A	0.1332	0.2379	0.6477	0.105*
C8	-0.0187 (6)	0.2423 (7)	0.5624 (2)	0.0735 (13)
H8A	-0.1173	0.2364	0.5859	0.088*
C9	-0.0259 (5)	0.2491 (6)	0.4913 (2)	0.0599 (11)
H9A	-0.1294	0.2456	0.4677	0.072*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Cl	0.1014 (12)	0.224 (2)	0.1105 (13)	0.0086 (13)	-0.0150 (9)	0.0395 (14)
O1	0.101 (3)	0.116 (3)	0.080 (2)	-0.014 (2)	0.0205 (19)	0.024 (2)
O2	0.0632 (18)	0.096 (2)	0.0641 (19)	-0.0135 (17)	0.0092 (14)	0.0080 (17)
O3	0.074 (2)	0.267 (7)	0.115 (3)	-0.035 (3)	0.041 (2)	-0.036 (4)
O4	0.055 (2)	0.222 (6)	0.205 (5)	0.002 (3)	0.001 (3)	0.003 (4)
N	0.049 (3)	0.122 (4)	0.138 (4)	-0.010 (3)	0.013 (3)	-0.032 (4)
C1	0.077 (3)	0.148 (5)	0.076 (3)	-0.006 (4)	0.004 (2)	0.008 (3)
C2	0.080 (3)	0.075 (3)	0.064 (3)	0.004 (3)	0.018 (2)	0.000 (3)
C3	0.062 (3)	0.080 (3)	0.076 (3)	-0.008 (2)	0.021 (2)	-0.005 (3)
C4	0.051 (2)	0.050 (2)	0.068 (2)	0.000 (2)	0.0161 (19)	-0.005 (2)
C5	0.057 (2)	0.071 (3)	0.086 (3)	-0.005 (2)	0.016 (2)	-0.007 (3)
C6	0.081 (3)	0.091 (4)	0.093 (4)	0.003 (3)	-0.015 (3)	0.004 (3)
C7	0.114 (4)	0.080 (4)	0.069 (3)	-0.001 (4)	0.015 (3)	0.001 (3)
C8	0.078 (3)	0.064 (3)	0.080 (3)	-0.003 (3)	0.028 (3)	0.001 (3)
C9	0.056 (2)	0.056 (3)	0.070 (3)	-0.001 (2)	0.0164 (19)	-0.001 (2)

*Geometric parameters (Å, °)*

Cl—C1	1.721 (5)	C3—H3B	0.9700
O1—C2	1.189 (5)	C4—C9	1.385 (5)
O2—C2	1.324 (5)	C4—C5	1.398 (6)
O2—C3	1.448 (5)	C5—C6	1.378 (6)
N—O3	1.181 (6)	C6—C7	1.377 (7)
N—O4	1.227 (6)	C6—H6A	0.9300
N—C5	1.472 (6)	C7—C8	1.363 (7)
C1—C2	1.499 (7)	C7—H7A	0.9300
C1—H1A	0.9700	C8—C9	1.369 (6)

C1—H1B	0.9700	C8—H8A	0.9300
C3—C4	1.494 (6)	C9—H9A	0.9300
C3—H3A	0.9700		
C2—O2—C3	117.0 (3)	C9—C4—C5	115.6 (4)
O3—N—O4	122.3 (5)	C9—C4—C3	120.8 (4)
O3—N—C5	120.5 (5)	C5—C4—C3	123.7 (4)
O4—N—C5	117.2 (6)	C6—C5—C4	122.7 (4)
C2—C1—C1	113.6 (4)	C6—C5—N	117.2 (5)
C2—C1—H1A	108.8	C4—C5—N	120.0 (4)
C1—C1—H1A	108.8	C7—C6—C5	119.5 (5)
C2—C1—H1B	108.8	C7—C6—H6A	120.3
C1—C1—H1B	108.8	C5—C6—H6A	120.3
H1A—C1—H1B	107.7	C8—C7—C6	119.0 (5)
O1—C2—O2	125.4 (5)	C8—C7—H7A	120.5
O1—C2—C1	126.5 (5)	C6—C7—H7A	120.5
O2—C2—C1	108.1 (4)	C7—C8—C9	121.4 (4)
O2—C3—C4	107.9 (3)	C7—C8—H8A	119.3
O2—C3—H3A	110.1	C9—C8—H8A	119.3
C4—C3—H3A	110.1	C8—C9—C4	121.9 (4)
O2—C3—H3B	110.1	C8—C9—H9A	119.0
C4—C3—H3B	110.1	C4—C9—H9A	119.0
H3A—C3—H3B	108.4		
C3—O2—C2—O1	-4.6 (7)	O3—N—C5—C6	-168.6 (6)
C3—O2—C2—C1	174.4 (4)	O4—N—C5—C6	9.6 (8)
C1—C1—C2—O1	-9.0 (8)	O3—N—C5—C4	11.2 (8)
C1—C1—C2—O2	172.1 (3)	O4—N—C5—C4	-170.6 (5)
C2—O2—C3—C4	159.9 (4)	C4—C5—C6—C7	0.2 (8)
O2—C3—C4—C9	-7.0 (6)	N—C5—C6—C7	180.0 (5)
O2—C3—C4—C5	172.5 (4)	C5—C6—C7—C8	-0.3 (8)
C9—C4—C5—C6	-0.4 (7)	C6—C7—C8—C9	0.7 (8)
C3—C4—C5—C6	-180.0 (5)	C7—C8—C9—C4	-1.0 (7)
C9—C4—C5—N	179.8 (4)	C5—C4—C9—C8	0.8 (6)
C3—C4—C5—N	0.2 (7)	C3—C4—C9—C8	-179.6 (4)

Hydrogen-bond geometry ( $\text{\AA}$ ,  $^\circ$ )

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
C1—H1A...O3 <sup>i</sup>	0.97	2.43	3.372 (6)	166
C7—H7A...O1 <sup>ii</sup>	0.93	2.58	3.374 (6)	143
C9—H9A...O2	0.93	2.27	2.660 (5)	104

Symmetry codes: (i)  $x-1, y, z$ ; (ii)  $x, -y+1/2, z+1/2$ .