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## Structure Reports

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4-(2*H*-Tetrazol-5-yl)pyridinium hydrogen sulfate

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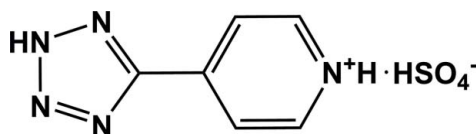
Received 22 June 2009; accepted 18 August 2009

Key indicators: single-crystal X-ray study;  $T = 298$  K; mean  $\sigma(\text{C}-\text{C}) = 0.003$  Å;  
R factor = 0.050; wR factor = 0.134; data-to-parameter ratio = 14.9.

In the cation of the title compound,  $\text{C}_6\text{H}_6\text{N}_5^+\cdot\text{HSO}_4^-$ , the pyridine and tetrazole rings are close to being co-planar [dihedral angle =  $3.98(7)^\circ$ ]. In the crystal, the ions are linked by  $\text{O}-\text{H}\cdots\text{O}$ ,  $\text{N}-\text{H}\cdots\text{O}$  and  $\text{N}-\text{H}\cdots(\text{O},\text{O})$  hydrogen bonds, resulting in chains.

## Related literature

Tetrazoles are excellent ligands for the construction of metal-organic frameworks because of their various coordination modes, see: Fu *et al.* (2008); Wang *et al.* (2005). For the applications of metal-organic coordination compounds, see: Fu *et al.* (2007); Huang *et al.* (1999); Liu *et al.* (1999); Xie *et al.* (2003); Zhang *et al.* (2001); Zhang *et al.* (2000).



## Experimental

## Crystal data

$\text{C}_6\text{H}_6\text{N}_5^+\cdot\text{HSO}_4^-$   
 $M_r = 245.23$   
Triclinic,  $P\bar{1}$   
 $a = 6.6515(13)$  Å  
 $b = 7.5507(15)$  Å  
 $c = 10.072(2)$  Å  
 $\alpha = 77.72(3)^\circ$   
 $\beta = 76.88(3)^\circ$

$\gamma = 79.71(3)^\circ$   
 $V = 476.84(16)$  Å<sup>3</sup>  
 $Z = 2$   
Mo  $K\alpha$  radiation  
 $\mu = 0.35$  mm<sup>-1</sup>  
 $T = 298$  K  
 $0.30 \times 0.25 \times 0.20$  mm

## Data collection

Rigaku Mercury2 diffractometer  
Absorption correction: multi-scan  
(*CrystalClear*; Rigaku, 2005)  
 $T_{\min} = 0.910$ ,  $T_{\max} = 1.000$   
(expected range = 0.848–0.932)

4917 measured reflections  
2175 independent reflections  
1961 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.055$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.050$   
 $wR(F^2) = 0.134$   
 $S = 1.07$   
2175 reflections

146 parameters  
H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.34$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.41$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O1}-\text{H1B}\cdots\text{O2}^{\text{i}}$	0.82	1.90	2.694 (2)	163
$\text{N1}-\text{H1A}\cdots\text{O4}^{\text{ii}}$	0.86	1.91	2.736 (2)	159
$\text{N4}-\text{H4A}\cdots\text{O2}^{\text{iii}}$	0.86	2.57	3.033 (3)	115
$\text{N4}-\text{H4A}\cdots\text{O3}$	0.86	1.97	2.741 (2)	150

Symmetry codes: (i)  $-x, -y + 2, -z + 1$ ; (ii)  $x, y - 1, z - 1$ ; (iii)  $-x, -y + 1, -z + 1$ .

Data collection: *CrystalClear* (Rigaku, 2005); cell refinement: *CrystalClear*; data reduction: *CrystalClear*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

This work was supported by a start-up grant from Southeast University to Professor Ren-Gen Xiong.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BX2220).

## References

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## supporting information

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## 4-(2*H*-Tetrazol-5-yl)pyridinium hydrogen sulfate

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### S1. Comment

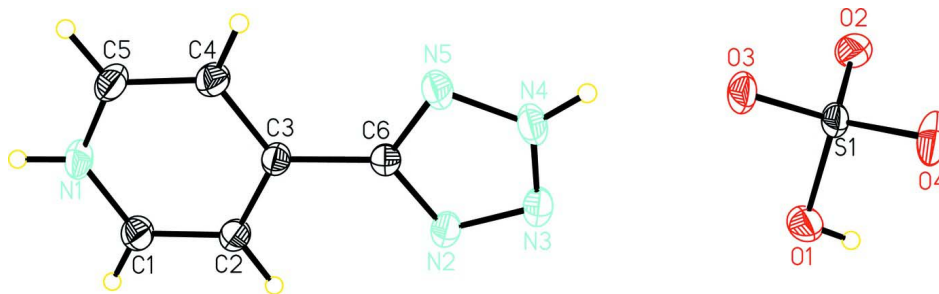
The construction of metal-organic coordination compounds has attracted much attention owing to the potential functions, such as permittivity, fluorescence, magnetism and optical properties. (Fu *et al.*, 2007; Huang *et al.*, 1999; Liu *et al.*, 1999; Xie *et al.*, 2003; Zhang *et al.*, 2001; Zhang *et al.*, 2000) Tetrazole compounds are a class of excellent ligands for the construction of novel metal-organic frameworks, because of its various coordination modes. (Wang, *et al.* 2005; Fu *et al.*, 2008). We report here the crystal and molecular structure of the title compound, 4-(2*H*-tetrazol-5-yl)pyridinium bi-sulfate), (I), Fig.1, The pyridine N atoms are protonated. The pyridine and tetrazole rings are nearly coplanar and only twisted from each other by a dihedral angle of 3.98 (7) °. The geometric parameters of the tetrazole rings are comparable to those in related molecules (Wang, *et al.* 2005; Fu *et al.*, 2008). The crystal packing is stabilized by coulombic forces, one O—H···O and three N—H···O hydrogen bonds (Table 1 and Fig.2).

### S2. Experimental

Isonicotinonitrile (30 mmol), NaN<sub>3</sub> (45 mmol), NH<sub>4</sub>Cl (33 mmol) and DMF (50 ml) were added in a flask under nitrogen atmosphere and the mixture stirred at 110°C for 20 h. The resulting solution was then poured into ice-water (100 ml), and a white solid was obtained after adding HCl (6 M) still pH=6. The precipitate was filtered and washed with distilled water. Colourless block-shaped crystals suitable for X-ray analysis were obtained from the crude product by slow evaporation of an ethanol/H<sub>2</sub>SO<sub>4</sub> (50:1 v/v) solution.

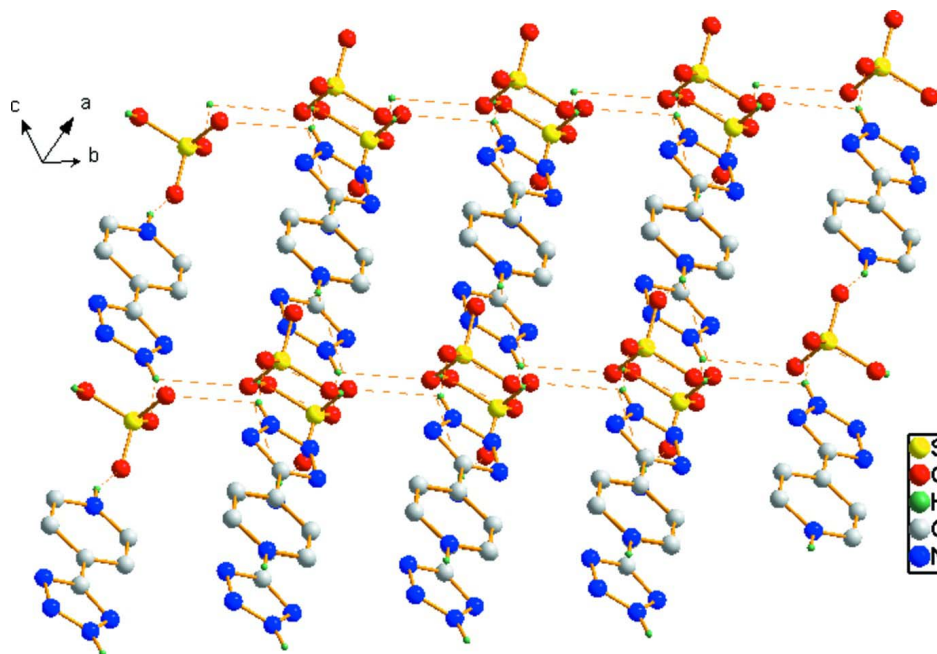
### S3. Refinement

All H atoms attached to C, O and N atoms were fixed geometrically and treated as riding with C—H = 0.93 Å, O—H = 0.82 Å and N—H = 0.86 Å with  $U_{iso}(H) = 1.2U_{eq}(C \text{ or } N)$  and  $U_{iso}(H) = 1.5U_{eq}(O)$ .



**Figure 1**

A view of the title compound with the atomic numbering scheme. Displacement ellipsoids were drawn at the 30% probability level.

**Figure 2**

The crystal packing of the title compound showing the two dimensional hydrogen bondings network (dashed line). Hydrogen atoms not involved in hydrogen bonding have been omitted for clarity.

#### 4-(2*H*-Tetrazol-5-yl)pyridinium hydrogen sulfate

##### Crystal data

$C_6H_6N_5^+ \cdot HSO_4^-$

$M_r = 245.23$

Triclinic,  $P\bar{1}$

Hall symbol: -P 1

$a = 6.6515$  (13) Å

$b = 7.5507$  (15) Å

$c = 10.072$  (2) Å

$\alpha = 77.72$  (3)°

$\beta = 76.88$  (3)°

$\gamma = 79.71$  (3)°

$V = 476.84$  (16) Å<sup>3</sup>

$Z = 2$

$F(000) = 252$

$D_x = 1.708$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 1961 reflections

$\theta = 3.2$ – $27.5$ °

$\mu = 0.35$  mm<sup>-1</sup>

$T = 298$  K

Block, colorless

$0.30 \times 0.25 \times 0.20$  mm

##### Data collection

Rigaku Mercury2  
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

Detector resolution: 13.6612 pixels mm<sup>-1</sup>

CCD profile fitting scans

Absorption correction: multi-scan

(*CrystalClear*; Rigaku, 2005)

$T_{\min} = 0.910$ ,  $T_{\max} = 1.000$

4917 measured reflections

2175 independent reflections

1961 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.055$

$\theta_{\max} = 27.5$ °,  $\theta_{\min} = 3.2$ °

$h = -8 \rightarrow 8$

$k = -9 \rightarrow 9$

$l = -13 \rightarrow 13$

*Refinement*Refinement on  $F^2$ 

Least-squares matrix: full

 $R[F^2 > 2\sigma(F^2)] = 0.050$  $wR(F^2) = 0.134$  $S = 1.07$ 

2175 reflections

146 parameters

0 restraints

Primary atom site location: structure-invariant  
direct methodsSecondary atom site location: difference Fourier  
mapHydrogen site location: inferred from  
neighbouring sites

H-atom parameters constrained

 $w = 1/[\sigma^2(F_o^2) + (0.0597P)^2 + 0.1736P]$ where  $P = (F_o^2 + 2F_c^2)/3$  $(\Delta/\sigma)_{\max} < 0.001$  $\Delta\rho_{\max} = 0.34 \text{ e } \text{\AA}^{-3}$  $\Delta\rho_{\min} = -0.41 \text{ e } \text{\AA}^{-3}$ Extinction correction: *SHELXL97* (Sheldrick,  
2008),  $F_c^* = kF_c[1 + 0.001x F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$ 

Extinction coefficient: 0.137 (14)

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.17610 (8)	0.76548 (6)	0.48163 (5)	0.0345 (2)
O2	-0.0454 (3)	0.7896 (2)	0.54236 (17)	0.0469 (4)
O1	0.2068 (3)	0.9101 (2)	0.34498 (15)	0.0469 (4)
H1B	0.1631	1.0128	0.3627	0.070*
C6	0.2581 (3)	0.3789 (3)	0.0215 (2)	0.0321 (4)
N2	0.2528 (3)	0.5615 (2)	-0.00724 (18)	0.0394 (4)
O3	0.2479 (3)	0.5928 (2)	0.43634 (16)	0.0445 (4)
N1	0.2800 (3)	0.0732 (2)	-0.28273 (19)	0.0397 (4)
H1A	0.2824	0.0112	-0.3456	0.048*
O4	0.3022 (3)	0.8031 (2)	0.56937 (19)	0.0537 (5)
N3	0.2423 (3)	0.6128 (3)	0.1106 (2)	0.0438 (5)
N4	0.2414 (3)	0.4632 (3)	0.20321 (19)	0.0447 (5)
H4A	0.2349	0.4630	0.2895	0.054*
N5	0.2511 (3)	0.3125 (3)	0.15493 (18)	0.0444 (5)
C2	0.2855 (3)	0.3527 (3)	-0.2218 (2)	0.0363 (5)
H2	0.2935	0.4774	-0.2472	0.044*
C3	0.2684 (3)	0.2687 (3)	-0.0841 (2)	0.0305 (4)
C4	0.2597 (3)	0.0814 (3)	-0.0490 (2)	0.0374 (5)
H4	0.2498	0.0217	0.0426	0.045*
C5	0.2661 (4)	-0.0136 (3)	-0.1518 (2)	0.0415 (5)
H5	0.2606	-0.1389	-0.1303	0.050*
C1	0.2904 (4)	0.2516 (3)	-0.3202 (2)	0.0395 (5)
H1	0.3010	0.3072	-0.4128	0.047*

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
S1	0.0519 (4)	0.0259 (3)	0.0276 (3)	-0.0071 (2)	-0.0060 (2)	-0.00971 (19)
O2	0.0546 (10)	0.0378 (8)	0.0457 (9)	-0.0136 (7)	0.0045 (7)	-0.0116 (7)
O1	0.0696 (11)	0.0312 (8)	0.0346 (8)	-0.0076 (7)	0.0002 (7)	-0.0046 (6)
C6	0.0385 (10)	0.0295 (9)	0.0286 (9)	-0.0064 (8)	-0.0048 (7)	-0.0064 (7)
N2	0.0534 (11)	0.0322 (9)	0.0353 (9)	-0.0087 (8)	-0.0078 (8)	-0.0105 (7)
O3	0.0711 (11)	0.0288 (8)	0.0365 (8)	-0.0053 (7)	-0.0095 (7)	-0.0146 (6)
N1	0.0495 (10)	0.0379 (10)	0.0384 (10)	-0.0063 (8)	-0.0112 (8)	-0.0181 (8)
O4	0.0735 (12)	0.0439 (9)	0.0561 (10)	-0.0003 (8)	-0.0281 (9)	-0.0259 (8)
N3	0.0573 (11)	0.0383 (10)	0.0399 (10)	-0.0082 (8)	-0.0081 (8)	-0.0160 (8)
N4	0.0634 (13)	0.0434 (10)	0.0290 (9)	-0.0042 (9)	-0.0081 (8)	-0.0141 (7)
N5	0.0654 (13)	0.0396 (10)	0.0284 (9)	-0.0074 (9)	-0.0058 (8)	-0.0099 (7)
C2	0.0487 (12)	0.0312 (10)	0.0309 (10)	-0.0111 (9)	-0.0078 (8)	-0.0052 (8)
C3	0.0349 (9)	0.0283 (9)	0.0304 (9)	-0.0050 (7)	-0.0076 (7)	-0.0082 (7)
C4	0.0495 (12)	0.0300 (10)	0.0335 (10)	-0.0082 (9)	-0.0098 (8)	-0.0032 (8)
C5	0.0549 (13)	0.0261 (10)	0.0462 (12)	-0.0066 (9)	-0.0130 (10)	-0.0080 (8)
C1	0.0524 (12)	0.0380 (11)	0.0305 (10)	-0.0098 (9)	-0.0101 (9)	-0.0062 (8)

Geometric parameters ( $\text{\AA}$ ,  $^\circ$ )

S1—O3	1.4378 (15)	N3—N4	1.302 (3)
S1—O4	1.4469 (17)	N4—N5	1.315 (3)
S1—O2	1.4564 (17)	N4—H4A	0.8600
S1—O1	1.5623 (16)	C2—C1	1.365 (3)
O1—H1B	0.8200	C2—C3	1.385 (3)
C6—N5	1.324 (3)	C2—H2	0.9300
C6—N2	1.343 (3)	C3—C4	1.392 (3)
C6—C3	1.466 (3)	C4—C5	1.369 (3)
N2—N3	1.309 (2)	C4—H4	0.9300
N1—C1	1.329 (3)	C5—H5	0.9300
N1—C5	1.332 (3)	C1—H1	0.9300
N1—H1A	0.8600		
O3—S1—O4	112.84 (10)	N5—N4—H4A	122.5
O3—S1—O2	113.35 (10)	N4—N5—C6	100.98 (18)
O4—S1—O2	112.29 (11)	C1—C2—C3	119.79 (19)
O3—S1—O1	104.09 (9)	C1—C2—H2	120.1
O4—S1—O1	106.97 (11)	C3—C2—H2	120.1
O2—S1—O1	106.53 (10)	C2—C3—C4	119.02 (18)
S1—O1—H1B	109.5	C2—C3—C6	119.51 (18)
N5—C6—N2	112.09 (18)	C4—C3—C6	121.47 (18)
N5—C6—C3	124.79 (18)	C5—C4—C3	118.83 (19)
N2—C6—C3	123.12 (18)	C5—C4—H4	120.6
N3—N2—C6	106.28 (18)	C3—C4—H4	120.6
C1—N1—C5	122.79 (18)	N1—C5—C4	120.08 (19)
C1—N1—H1A	118.6	N1—C5—H5	120.0

C5—N1—H1A	118.6	C4—C5—H5	120.0
N4—N3—N2	105.63 (17)	N1—C1—C2	119.5 (2)
N3—N4—N5	115.02 (17)	N1—C1—H1	120.3
N3—N4—H4A	122.5	C2—C1—H1	120.3
C6—N2—N3—N4	-0.1 (2)	C3—C4—C5—N1	-0.1 (3)
N2—N3—N4—N5	0.1 (3)	N4—N5—C6—N2	0.0 (2)
N3—N4—N5—C6	-0.1 (3)	N4—N5—C6—C3	-179.4 (2)
C5—N1—C1—C2	-0.4 (3)	N3—N2—C6—N5	0.0 (2)
N1—C1—C2—C3	-0.4 (3)	N3—N2—C6—C3	179.51 (19)
C1—C2—C3—C4	0.9 (3)	C2—C3—C6—N5	-176.8 (2)
C1—C2—C3—C6	-178.5 (2)	C4—C3—C6—N5	3.8 (3)
C2—C3—C4—C5	-0.7 (3)	C2—C3—C6—N2	3.7 (3)
C6—C3—C4—C5	178.7 (2)	C4—C3—C6—N2	-175.68 (19)
C1—N1—C5—C4	0.7 (3)		

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
O1—H1B...O2 <sup>i</sup>	0.82	1.90	2.694 (2)	163
N1—H1A...O4 <sup>ii</sup>	0.86	1.91	2.736 (2)	159
N4—H4A...O2 <sup>iii</sup>	0.86	2.57	3.033 (3)	115
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