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## 2,3-Diaminopyridinium 4-hydroxybenzoate

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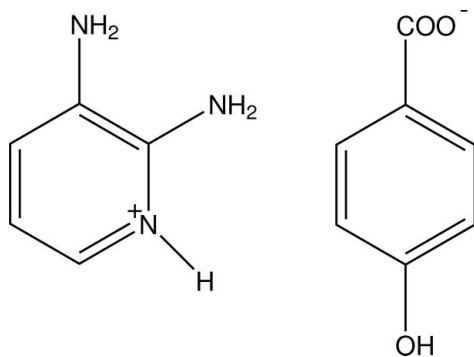
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Key indicators: single-crystal X-ray study;  $T = 100$  K; mean  $\sigma(\text{C}-\text{C}) = 0.001$  Å;  $R$  factor = 0.049;  $wR$  factor = 0.138; data-to-parameter ratio = 24.6.

In the title compound,  $\text{C}_5\text{H}_8\text{N}_3^+\cdot\text{C}_7\text{H}_5\text{O}_3^-$ , the pyridine N atom is protonated. In the 4-hydroxybenzoate anion, the carboxylate group is twisted slightly out of the benzene ring plane by an angle of  $3.77$  (5)°. The protonated N atom and one of the two amino groups are hydrogen-bonded to the 4-hydroxybenzoate anion through a pair of  $\text{N}-\text{H}\cdots\text{O}$  hydrogen bonds, forming an  $R_2^2(8)$  ring motif. The crystal structure is further stabilized by  $\text{O}-\text{H}\cdots\text{O}$  and  $\text{C}-\text{H}\cdots\text{O}$  hydrogen bonds and  $\pi-\pi$  interactions involving the pyridinium rings [centroid-centroid distance of  $3.6277$  (5) Å], leading to the formation of a three-dimensional network.

## Related literature

For general background to substituted pyridines, see: Pozharski *et al.* (1997); Katritzky *et al.* (1996). For details of hydrogen bonding, see: Jeffrey & Saenger (1991); Jeffrey (1997); Scheiner (1997). For hydrogen-bond motifs, see: Bernstein *et al.* (1995). For bond-length data, see: Allen *et al.* (1987). For the stability of the temperature controller used in the data collection, see: Cosier & Glazer (1986).



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## Experimental

## Crystal data

$\text{C}_5\text{H}_8\text{N}_3^+\cdot\text{C}_7\text{H}_5\text{O}_3^-$   
 $M_r = 247.25$   
Monoclinic,  $P2_1/c$   
 $a = 10.2915$  (2) Å  
 $b = 11.4946$  (2) Å  
 $c = 11.0921$  (2) Å  
 $\beta = 112.644$  (1)°

$V = 1211.01$  (4) Å<sup>3</sup>  
 $Z = 4$   
Mo  $K\alpha$  radiation  
 $\mu = 0.10$  mm<sup>-1</sup>  
 $T = 100$  K  
 $0.51 \times 0.39 \times 0.14$  mm

## Data collection

Bruker SMART APEXII CCD  
area-detector diffractometer  
Absorption correction: multi-scan  
(*SADABS*; Bruker, 2005)  
 $T_{\min} = 0.950$ ,  $T_{\max} = 0.986$

25821 measured reflections  
5296 independent reflections  
4257 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.034$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.049$   
 $wR(F^2) = 0.138$   
 $S = 1.04$   
5296 reflections  
215 parameters

H atoms treated by a mixture of independent and constrained refinement  
 $\Delta\rho_{\text{max}} = 0.65$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.19$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N1}-\text{H1N1}\cdots\text{O2}$	0.89 (2)	1.903 (15)	2.7874 (9)	173 (1)
$\text{N2}-\text{H2N2}\cdots\text{O3}$	0.89 (1)	1.898 (14)	2.7843 (9)	176 (1)
$\text{N2}-\text{H1N2}\cdots\text{O2}^i$	0.87 (1)	2.014 (15)	2.8689 (9)	168 (1)
$\text{N3}-\text{H1N3}\cdots\text{O2}^j$	0.91 (2)	2.071 (16)	2.9790 (10)	174 (1)
$\text{N3}-\text{H2N3}\cdots\text{O3}^{ii}$	0.89 (2)	2.057 (15)	2.9285 (10)	166 (1)
$\text{O1}-\text{H1O1}\cdots\text{O3}^{iii}$	0.90 (2)	1.775 (19)	2.6595 (8)	168 (2)
$\text{C2}-\text{H2}\cdots\text{O3}^{iii}$	1.00 (1)	2.500 (14)	3.2104 (10)	128 (1)

Symmetry codes: (i)  $x, -y + \frac{3}{2}, z - \frac{1}{2}$ ; (ii)  $-x + 2, y + \frac{1}{2}, -z + \frac{1}{2}$ ; (iii)  $-x + 1, y - \frac{1}{2}, -z + \frac{1}{2}$ .

Data collection: *APEX2* (Bruker, 2005); cell refinement: *SAINT* (Bruker, 2005); data reduction: *SAINT*; program(s) used to solve structure: *SHELXTL* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL* and *PLATON* (Spek, 2009).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CI2813).

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## supporting information

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## 2,3-Diaminopyridinium 4-hydroxybenzoate

Hoong-Kun Fun and Kasthuri Balasubramani

### S1. Comment

Pyridine and its derivatives play an important role in heterocyclic chemistry (Pozharski *et al.*, 1997; Katritzky *et al.*, 1996). Pyridine and its substituted derivatives are often involved in hydrogen-bond interactions (Jeffrey & Saenger, 1991; Jeffrey, 1997; Scheiner, 1997). Since our aim is to study some interesting hydrogen-bonding interactions, the crystal structure of the title compound is presented here.

The asymmetric unit of the title compound (Fig 1), contains one 2,3-diaminopyridinium cation and one 4-hydroxybenzoate anion. The bond lengths (Allen *et al.*, 1987) and angles are normal. The 2,3-diaminopyridinium cation is planar to within  $\pm 0.015$  (1) Å. The protonation of N1 atom resulted in a slight increase in the C8—N1—C12 angle [123.47 (7)°]. In the anion, the carboxylate group is twisted slightly away from the attached ring; the dihedral angle between C1-C6 and O2/O3/C7/C6 planes is 3.77 (5)°.

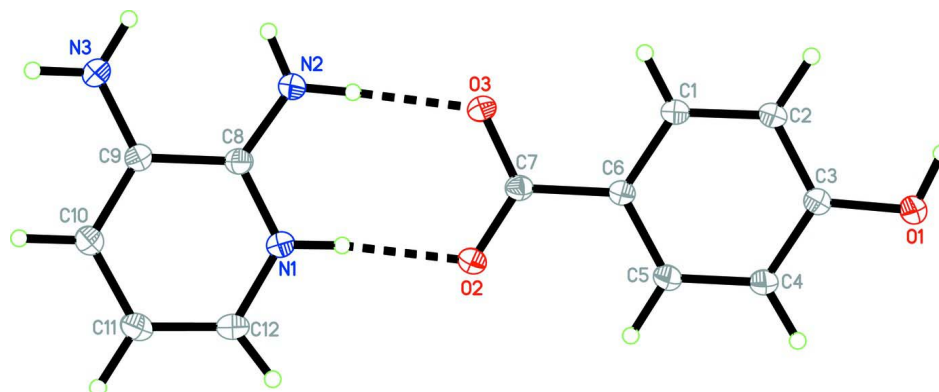
In the crystal packing (Fig. 2), the protonated N1 atom and the 2-amino group (N2) is hydrogen-bonded to the carboxylate oxygen atoms (O2 and O3) *via* a pair of N—H $\cdots$ O hydrogen bonds forming a ring motif  $R_2^2(8)$  (Bernstein *et al.*, 1995). The amino groups (N2 and N3) are involved in N—H $\cdots$ O hydrogen bonding interactions to form a  $R_1^2(7)$  ring motif. The hydroxyl group hydrogen atom is also hydrogen-bonded to the carboxylate oxygen atom through O—H $\cdots$ O hydrogen bonds. Moreover O—H $\cdots$ O and C—H $\cdots$ O hydrogen bonds form a  $R_1^2(6)$  ring motif (Table 1 and Fig 2). The crystal structure is further stabilized by  $\pi$ - $\pi$  stacking interactions between the pyridinium rings (C8—C12/N1) at (x, y, z) and (2-x, 2-y, 1-z), with a centroid-to-centroid distance of 3.6277 (5) Å.

### S2. Experimental

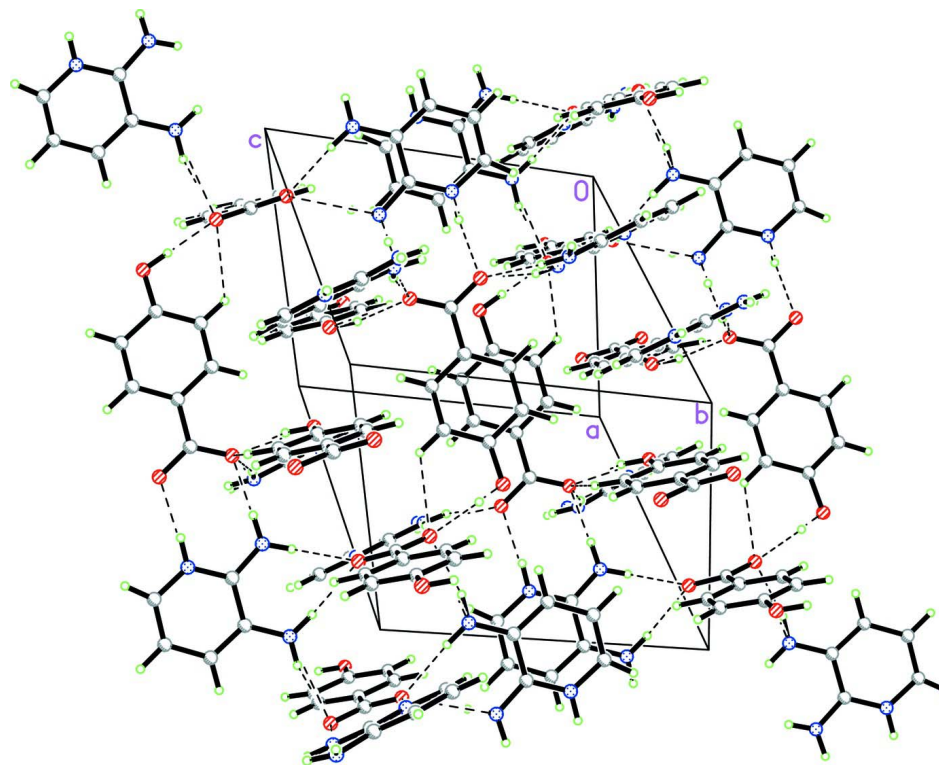
Hot methanol solutions (20 ml) of 2,3-diaminopyridine (27 mg, Aldrich) and 4-hydroxybenzoic acid (35 mg, Merck) were mixed and warmed over a heating magnetic stirrer for 5 minutes. The resulting solution was allowed to cool slowly at room temperature. Crystals of the title compound appeared from the mother liquor after a few days.

### S3. Refinement

All the H atoms were located in a difference Fourier map and allowed to refine freely [N—H = 0.86–0.91 Å, C—H = 0.95–1.01 (15) Å and O—H = 0.89 (18) Å]

**Figure 1**

The molecular structure of the title compound, showing 50% probability displacement ellipsoids and the atom numbering scheme. Dashed lines indicate hydrogen bonds.

**Figure 2**

Crystal packing of the title compound, showing a part of the three-dimensional network. Hydrogen bonds are shown as dashed lines.

**(I)***Crystal data*
 $C_5H_8N_3^+ \cdot C_7H_5O_3^-$ 
 $M_r = 247.25$ 

 Monoclinic,  $P2_1/c$ 

 Hall symbol:  $-P\ 2_1/c$ 
 $a = 10.2915(2)\ \text{\AA}$ 
 $b = 11.4946(2)\ \text{\AA}$ 
 $c = 11.0921(2)\ \text{\AA}$ 
 $\beta = 112.644(1)^\circ$

$V = 1211.01 (4) \text{ \AA}^3$   
 $Z = 4$   
 $F(000) = 520$   
 $D_x = 1.356 \text{ Mg m}^{-3}$   
 Mo  $K\alpha$  radiation,  $\lambda = 0.71073 \text{ \AA}$   
 Cell parameters from 9979 reflections

$\theta = 2.7\text{--}37.8^\circ$   
 $\mu = 0.10 \text{ mm}^{-1}$   
 $T = 100 \text{ K}$   
 Plate, brown  
 $0.51 \times 0.39 \times 0.14 \text{ mm}$

*Data collection*

Bruker SMART APEXII CCD area-detector  
 diffractometer  
 Radiation source: fine-focus sealed tube  
 Graphite monochromator  
 $\varphi$  and  $\omega$  scans  
 Absorption correction: multi-scan  
 (SADABS; Bruker, 2005)  
 $T_{\min} = 0.950$ ,  $T_{\max} = 0.986$

25821 measured reflections  
 5296 independent reflections  
 4257 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.034$   
 $\theta_{\max} = 35.0^\circ$ ,  $\theta_{\min} = 2.1^\circ$   
 $h = -16 \rightarrow 16$   
 $k = -16 \rightarrow 18$   
 $l = -17 \rightarrow 17$

*Refinement*

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.049$   
 $wR(F^2) = 0.138$   
 $S = 1.04$   
 5296 reflections  
 215 parameters  
 0 restraints  
 Primary atom site location: structure-invariant  
 direct methods

Secondary atom site location: difference Fourier  
 map  
 Hydrogen site location: inferred from  
 neighbouring sites  
 H atoms treated by a mixture of independent  
 and constrained refinement  
 $w = 1/[\sigma^2(F_o^2) + (0.0817P)^2 + 0.1673P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} = 0.001$   
 $\Delta\rho_{\max} = 0.65 \text{ e \AA}^{-3}$   
 $\Delta\rho_{\min} = -0.19 \text{ e \AA}^{-3}$

*Special details*

**Experimental.** The crystal was placed in the cold stream of an Oxford Cyrosystems Cobra open-flow nitrogen cryostat (Cosier & Glazer, 1986) operating at 100.0 (1) K.

**Geometry.** All s.u.'s (except the s.u. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell s.u.'s are taken into account individually in the estimation of s.u.'s in distances, angles and torsion angles; correlations between s.u.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell s.u.'s is used for estimating s.u.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted R-factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional R-factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and R-factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.45467 (8)	0.16900 (6)	0.43336 (6)	0.02563 (15)
O2	0.84786 (7)	0.61461 (5)	0.52034 (5)	0.02074 (13)
O3	0.73136 (6)	0.62156 (5)	0.30541 (5)	0.01784 (12)
N1	1.01845 (8)	0.79427 (6)	0.49674 (6)	0.01857 (14)
N2	0.87767 (8)	0.81703 (7)	0.27844 (7)	0.02195 (15)
N3	1.05767 (9)	0.99458 (8)	0.25977 (7)	0.02666 (17)
C1	0.57827 (9)	0.42070 (7)	0.30431 (7)	0.01757 (14)
C2	0.50323 (9)	0.32084 (7)	0.30618 (7)	0.01879 (15)

C3	0.52248 (9)	0.26810 (7)	0.42538 (7)	0.01899 (15)
C4	0.61617 (11)	0.31656 (8)	0.54174 (8)	0.02518 (18)
C5	0.69083 (10)	0.41624 (8)	0.53881 (7)	0.02324 (17)
C6	0.67313 (8)	0.46970 (7)	0.41995 (7)	0.01604 (14)
C7	0.75521 (8)	0.57533 (6)	0.41592 (7)	0.01540 (14)
C8	0.99181 (9)	0.84838 (7)	0.38181 (7)	0.01678 (14)
C9	1.08674 (9)	0.93716 (7)	0.37590 (7)	0.01803 (15)
C10	1.20049 (9)	0.96332 (8)	0.48910 (8)	0.02092 (16)
C11	1.22331 (9)	0.90438 (8)	0.60700 (8)	0.02314 (17)
C12	1.13117 (10)	0.82008 (8)	0.60867 (8)	0.02224 (17)
H1	0.5647 (15)	0.4565 (11)	0.2226 (13)	0.031 (3)*
H2	0.4376 (14)	0.2854 (12)	0.2229 (13)	0.029 (3)*
H4	0.6280 (18)	0.2759 (15)	0.6232 (17)	0.051 (4)*
H5	0.7615 (17)	0.4479 (13)	0.6225 (15)	0.039 (4)*
H10	1.2669 (14)	1.0227 (12)	0.4848 (12)	0.028 (3)*
H11	1.3091 (16)	0.9246 (13)	0.6883 (14)	0.039 (4)*
H12	1.1341 (17)	0.7745 (14)	0.6809 (15)	0.043 (4)*
H1N1	0.9588 (15)	0.7391 (12)	0.4987 (13)	0.031 (3)*
H1N2	0.8601 (17)	0.8454 (12)	0.2011 (15)	0.033 (3)*
H2N2	0.8275 (15)	0.7565 (12)	0.2854 (13)	0.027 (3)*
H1N3	0.9924 (17)	0.9660 (13)	0.1840 (15)	0.040 (4)*
H2N3	1.1284 (16)	1.0365 (12)	0.2548 (14)	0.034 (3)*
H1O1	0.3995 (19)	0.1445 (15)	0.3531 (18)	0.051 (5)*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
O1	0.0299 (3)	0.0248 (3)	0.0172 (3)	-0.0123 (2)	0.0034 (2)	0.0013 (2)
O2	0.0238 (3)	0.0217 (3)	0.0141 (2)	-0.0059 (2)	0.0044 (2)	-0.00153 (19)
O3	0.0188 (3)	0.0178 (3)	0.0149 (2)	0.0002 (2)	0.00434 (19)	0.00308 (19)
N1	0.0200 (3)	0.0194 (3)	0.0148 (3)	-0.0020 (2)	0.0051 (2)	0.0018 (2)
N2	0.0232 (3)	0.0234 (3)	0.0154 (3)	-0.0072 (3)	0.0031 (2)	0.0021 (2)
N3	0.0284 (4)	0.0324 (4)	0.0171 (3)	-0.0118 (3)	0.0065 (3)	0.0037 (3)
C1	0.0190 (3)	0.0187 (3)	0.0131 (3)	-0.0014 (3)	0.0041 (2)	0.0005 (2)
C2	0.0198 (4)	0.0205 (3)	0.0131 (3)	-0.0037 (3)	0.0031 (2)	-0.0007 (2)
C3	0.0202 (4)	0.0199 (3)	0.0150 (3)	-0.0046 (3)	0.0047 (3)	-0.0001 (2)
C4	0.0316 (5)	0.0273 (4)	0.0130 (3)	-0.0114 (3)	0.0046 (3)	0.0008 (3)
C5	0.0286 (4)	0.0251 (4)	0.0131 (3)	-0.0099 (3)	0.0049 (3)	-0.0015 (3)
C6	0.0173 (3)	0.0167 (3)	0.0131 (3)	-0.0016 (2)	0.0048 (2)	-0.0007 (2)
C7	0.0167 (3)	0.0152 (3)	0.0140 (3)	0.0013 (2)	0.0055 (2)	-0.0002 (2)
C8	0.0177 (3)	0.0174 (3)	0.0143 (3)	-0.0003 (3)	0.0052 (2)	0.0001 (2)
C9	0.0183 (3)	0.0204 (3)	0.0155 (3)	-0.0023 (3)	0.0066 (2)	0.0000 (2)
C10	0.0192 (4)	0.0249 (4)	0.0178 (3)	-0.0044 (3)	0.0062 (3)	-0.0010 (3)
C11	0.0200 (4)	0.0294 (4)	0.0166 (3)	-0.0029 (3)	0.0032 (3)	-0.0007 (3)
C12	0.0230 (4)	0.0262 (4)	0.0145 (3)	-0.0014 (3)	0.0038 (3)	0.0023 (3)

*Geometric parameters (Å, °)*

O1—C3	1.3565 (10)	C2—C3	1.3976 (11)
O1—H1O1	0.897 (18)	C2—H2	0.997 (13)
O2—C7	1.2666 (9)	C3—C4	1.3956 (11)
O3—C7	1.2702 (9)	C4—C5	1.3869 (12)
N1—C8	1.3484 (10)	C4—H4	0.982 (17)
N1—C12	1.3661 (11)	C5—C6	1.4014 (11)
N1—H1N1	0.888 (15)	C5—H5	1.001 (15)
N2—C8	1.3369 (10)	C6—C7	1.4896 (11)
N2—H1N2	0.869 (15)	C8—C9	1.4316 (11)
N2—H2N2	0.887 (14)	C9—C10	1.3805 (11)
N3—C9	1.3739 (10)	C10—C11	1.4100 (12)
N3—H1N3	0.911 (16)	C10—H10	0.981 (14)
N3—H2N3	0.892 (15)	C11—C12	1.3608 (13)
C1—C2	1.3881 (11)	C11—H11	1.017 (15)
C1—C6	1.3969 (10)	C12—H12	0.948 (15)
C1—H1	0.955 (13)		
C3—O1—H1O1	110.1 (11)	C4—C5—H5	119.3 (8)
C8—N1—C12	123.47 (7)	C6—C5—H5	119.9 (9)
C8—N1—H1N1	117.5 (9)	C1—C6—C5	118.61 (7)
C12—N1—H1N1	119.1 (9)	C1—C6—C7	120.33 (6)
C8—N2—H1N2	121.7 (10)	C5—C6—C7	121.05 (7)
C8—N2—H2N2	119.1 (9)	O2—C7—O3	122.08 (7)
H1N2—N2—H2N2	118.2 (13)	O2—C7—C6	119.94 (6)
C9—N3—H1N3	120.6 (10)	O3—C7—C6	117.96 (6)
C9—N3—H2N3	115.4 (9)	N2—C8—N1	118.57 (7)
H1N3—N3—H2N3	117.8 (13)	N2—C8—C9	122.86 (7)
C2—C1—C6	121.05 (7)	N1—C8—C9	118.57 (7)
C2—C1—H1	119.4 (8)	N3—C9—C10	123.40 (8)
C6—C1—H1	119.5 (8)	N3—C9—C8	118.64 (7)
C1—C2—C3	119.72 (7)	C10—C9—C8	117.92 (7)
C1—C2—H2	120.4 (8)	C9—C10—C11	121.27 (8)
C3—C2—H2	119.9 (8)	C9—C10—H10	118.1 (8)
O1—C3—C4	117.67 (7)	C11—C10—H10	120.6 (8)
O1—C3—C2	122.47 (7)	C12—C11—C10	119.12 (8)
C4—C3—C2	119.85 (7)	C12—C11—H11	121.5 (8)
C5—C4—C3	119.99 (7)	C10—C11—H11	119.4 (8)
C5—C4—H4	122.8 (10)	C11—C12—N1	119.65 (7)
C3—C4—H4	117.2 (10)	C11—C12—H12	127.5 (10)
C4—C5—C6	120.78 (7)	N1—C12—H12	112.8 (10)
C6—C1—C2—C3	-0.01 (13)	C5—C6—C7—O3	-178.02 (8)
C1—C2—C3—O1	-178.47 (8)	C12—N1—C8—N2	-179.40 (8)
C1—C2—C3—C4	0.46 (14)	C12—N1—C8—C9	0.41 (12)
O1—C3—C4—C5	178.40 (9)	N2—C8—C9—N3	1.55 (13)
C2—C3—C4—C5	-0.57 (15)	N1—C8—C9—N3	-178.25 (8)

C3—C4—C5—C6	0.24 (15)	N2—C8—C9—C10	179.31 (8)
C2—C1—C6—C5	-0.32 (13)	N1—C8—C9—C10	-0.49 (12)
C2—C1—C6—C7	178.59 (7)	N3—C9—C10—C11	177.91 (9)
C4—C5—C6—C1	0.21 (14)	C8—C9—C10—C11	0.26 (13)
C4—C5—C6—C7	-178.70 (9)	C9—C10—C11—C12	0.06 (14)
C1—C6—C7—O2	-175.21 (8)	C10—C11—C12—N1	-0.16 (14)
C5—C6—C7—O2	3.67 (12)	C8—N1—C12—C11	-0.08 (13)
C1—C6—C7—O3	3.09 (11)		

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1N1...O2	0.89 (2)	1.903 (15)	2.7874 (9)	173 (1)
N2—H2N2...O3	0.89 (1)	1.898 (14)	2.7843 (9)	176 (1)
N2—H1N2...O2 <sup>i</sup>	0.87 (1)	2.014 (15)	2.8689 (9)	168 (1)
N3—H1N3...O2 <sup>i</sup>	0.91 (2)	2.071 (16)	2.9790 (10)	174 (1)
N3—H2N3...O3 <sup>ii</sup>	0.89 (2)	2.057 (15)	2.9285 (10)	166 (1)
O1—H1O1...O3 <sup>iii</sup>	0.90 (2)	1.775 (19)	2.6595 (8)	168 (2)
C2—H2...O3 <sup>iii</sup>	1.00 (1)	2.500 (14)	3.2104 (10)	128 (1)

Symmetry codes: (i)  $x, -y+3/2, z-1/2$ ; (ii)  $-x+2, y+1/2, -z+1/2$ ; (iii)  $-x+1, y-1/2, -z+1/2$ .