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## Structure Reports

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## Ethylenediammonium dichloride

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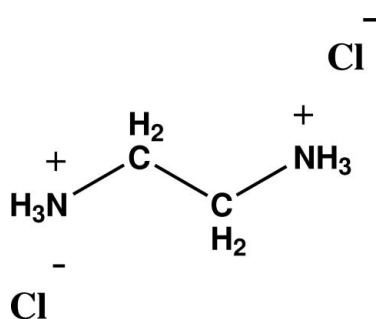
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Key indicators: single-crystal X-ray study;  $T = 100$  K; mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å;  $R$  factor = 0.028;  $wR$  factor = 0.070; data-to-parameter ratio = 11.8.

The title ionic compound,  $\text{C}_2\text{H}_{10}\text{N}_2^{2+} \cdot 2\text{Cl}^-$ , crystallizes with a center of symmetry within the cation. Each of the positively charged ammonium ends of the molecule is trigonally hydrogen bonded to three different chloride counter-ions, while each of the chloride ions is trigonally hydrogen bonded to three different ethylenediammonium cations. The hydrogen-bonding network leads to stabilization of the structure.

## Related literature

For the applications of ethylenediamine, see: Kotti *et al.* (2006); Warner (1912).



## Experimental

## Crystal data

$\text{C}_2\text{H}_{10}\text{N}_2^{2+} \cdot 2\text{Cl}^-$   
 $M_r = 133.02$   
 Monoclinic,  $P2_1/c$   
 $a = 4.3807$  (3) Å  
 $b = 6.8569$  (4) Å  
 $c = 9.9464$  (5) Å  
 $\beta = 91.192$  (2)°

$V = 298.71$  (3) Å<sup>3</sup>  
 $Z = 2$   
 Cu  $K\alpha$  radiation  
 $\mu = 8.71$  mm<sup>-1</sup>  
 $T = 100$  K  
 $0.45 \times 0.30 \times 0.29$  mm

## Data collection

Bruker SMART CCD APEXII  
 area-detector diffractometer  
 Absorption correction: multi-scan  
 (SADABS; Sheldrick, 2001)  
 $T_{\min} = 0.085$ ,  $T_{\max} = 0.090$

1654 measured reflections  
 521 independent reflections  
 520 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.022$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.028$   
 $wR(F^2) = 0.070$   
 $S = 1.15$   
 521 reflections

44 parameters  
 Only H-atom coordinates refined  
 $\Delta\rho_{\text{max}} = 0.41$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.31$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H \cdots A$	$D-H$	$H \cdots A$	$D \cdots A$	$D-H \cdots A$
$\text{N1}-\text{H1E} \cdots \text{Cl1}^i$	0.89 (2)	2.27 (2)	3.1514 (15)	175 (2)
$\text{N1}-\text{H1D} \cdots \text{Cl1}^i$	0.80 (3)	2.39 (3)	3.1770 (15)	170 (2)
$\text{N1}-\text{H1C} \cdots \text{Cl1}^ii$	0.91 (2)	2.29 (2)	3.1922 (15)	171 (2)

Symmetry codes: (i)  $-x + 1, y + \frac{1}{2}, -z + \frac{3}{2}$ ; (ii)  $x, -y + \frac{1}{2}, z + \frac{1}{2}$ .

Data collection: APEX2 (Bruker, 2006); cell refinement: APEX2; data reduction: SAINT (Bruker, 2005); program(s) used to solve structure: SHELXTL (Sheldrick, 2008); program(s) used to refine structure: SHELXTL; molecular graphics: SHELXTL; software used to prepare material for publication: SHELXTL.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: LH2823).

## References

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## supporting information

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## Ethylenediammonium dichloride

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### S1. Comment

Ethylenediamine has been used for approximately one century in the preparation of many metal coordination complexes, such as tris(ethylenediamine)cobalt(III) chloride (Warner, 1912). This is an important precursor to many polymers, chelating agents and pharmaceuticals, including drug design (Kotti, *et al.*, 2006). Since it is a widely used building block in the synthesis of many materials, its structure is of interest. However, it exists as a liquid at room temperature, with a melting point of 282K. We report herein the crystal structure of the dichloride salt of ethylenediamine.

Ethylenediammonium dichloride (I) crystallizes with a center of symmetry in the ethylene moiety. Fig. 1 shows the dication with one chloride counterion hydrogen bonded at each terminal nitrogen atom; however, there are three such chloride ions surrounding each N atom. The angles around both N1 and C1 are essentially tetrahedral, with the N1—C1—C1[1 - x, 1/2 + y, 2.5 - z] angle = 109.68 (18)°, and the angles around N1 range from 102 (2) to 115 (2)°.

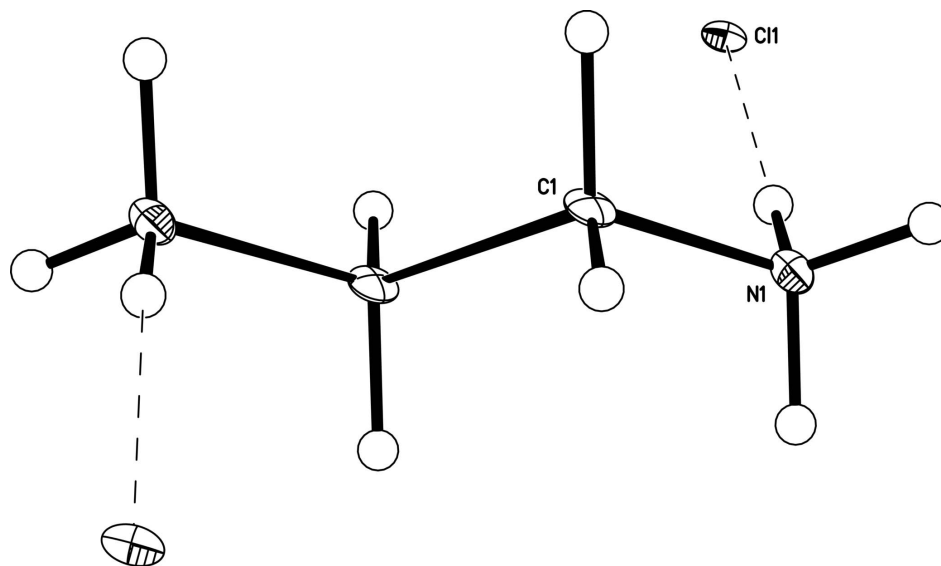
Fig. 2 illustrates the packing of (I). Each of the chloride counterions is trigonally H bonded to three different ethylenediammonium cations with N...Cl bond distances of 3.1516 (15), 3.1931 (16) & 3.1749 (16) Å and angles N—H...Cl of 175 (2), 170 (2) & 173 (2)° (see Table 1). Protonation occurs at both ammonium sites in the molecule (the 2nd is centrosymmetrically related); as a result, each nitrogen is also trigonally H bonded to three symmetry-related chlorides. This H bonding fixes both the chloride anions and the organic dication very rigidly in the lattice. Therefore, through symmetry, there exist six N—H...Cl bonds for each molecule, which leads to a great degree of stabilization in the structure.

### S2. Experimental

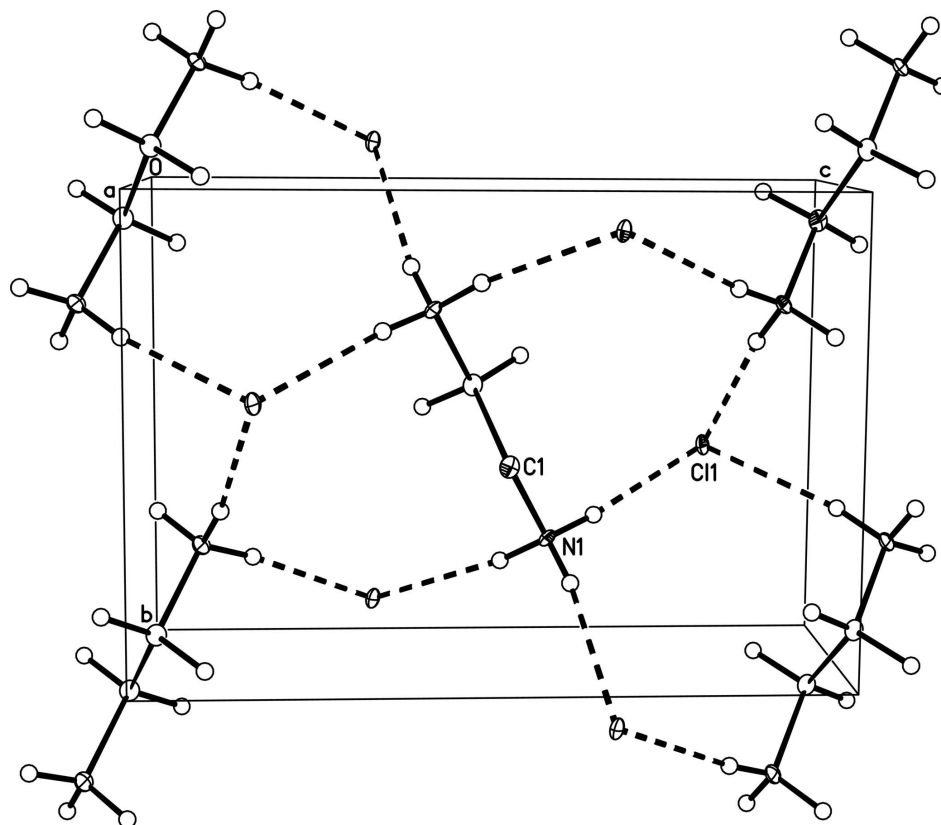
Compound (I) was prepared by mixing 2.5 ml of ethylenediamine with 62 ml of water. Then 7.5 ml of 12 M HCl were added and this mixture was stirred in an ice bath at 273K until a white precipitate formed. The white precipitate was filtered and washed 3 times with methyl alcohol. The product was dissolved in water and then 12 M HCl was added until precipitation just began; a small quantity of water was then added to redissolve the precipitate. This mixture was allowed to evaporate slowly and large colorless needles of (I) formed, which were used directly for X-ray analysis.

### S3. Refinement

All H atoms for (I) were found in electron density difference maps. The ammonium and methylene Hs' fractional coordinates were allowed to refine, but their isotropic thermal parameters were set at  $U_{\text{iso}}(\text{H}) = 1.5U_{\text{eq}}(\text{N})$  and  $1.2U_{\text{eq}}(\text{C})$ .

**Figure 1**

The molecular structure of (I), with its numbering; one-half of the molecule is generated through a center of symmetry at  $1/2, 1/2, 1/2$  in the chosen unit cell. The H bonds to the chlorides are shown as dashed lines. Displacement ellipsoids are drawn at the 80% probability level.



**Figure 2**

A partial packing diagram for (I), with extracellular molecules, illustrating the trigonal hydrogen bonding of the chloride ion to three different ethylenediammonium cations. Also the three H atoms of the ammonium are bound to three different chloride counterions. All of the cations lie on centers of symmetry, at  $1/2, 1/2, 1/2$  and  $1/2, 0, 0$ . Displacement ellipsoids are drawn at the 80% probability level.

### Ethylenediammonium dichloride

#### Crystal data

$C_2H_{10}N_2^{2+} \cdot 2Cl^-$   
 $M_r = 133.02$   
 Monoclinic,  $P2_1/c$   
 Hall symbol:  $-P\ 2_1/c$   
 $a = 4.3807(3)\ \text{\AA}$   
 $b = 6.8569(4)\ \text{\AA}$   
 $c = 9.9464(5)\ \text{\AA}$   
 $\beta = 91.192(2)^\circ$   
 $V = 298.71(3)\ \text{\AA}^3$   
 $Z = 2$

$F(000) = 140$   
 $D_x = 1.479\ \text{Mg m}^{-3}$   
 Cu  $K\alpha$  radiation,  $\lambda = 1.54178\ \text{\AA}$   
 Cell parameters from 1684 reflections  
 $\theta = 4.5\text{--}67.4^\circ$   
 $\mu = 8.71\ \text{mm}^{-1}$   
 $T = 100\ \text{K}$   
 Parallelepiped, colourless  
 $0.45 \times 0.30 \times 0.29\ \text{mm}$

#### Data collection

Bruker SMART CCD APEXII area-detector  
 diffractometer  
 Radiation source: fine-focus sealed tube  
 Graphite monochromator  
 $\varphi$  and  $\omega$  scans

Absorption correction: multi-scan  
 (*SADABS*; Sheldrick, 2001)  
 $T_{\min} = 0.085$ ,  $T_{\max} = 0.090$   
 1654 measured reflections  
 521 independent reflections

520 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.022$   
 $\theta_{\text{max}} = 67.8^\circ$ ,  $\theta_{\text{min}} = 7.9^\circ$

$h = -5 \rightarrow 4$   
 $k = -8 \rightarrow 8$   
 $l = -11 \rightarrow 11$

*Refinement*

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.028$   
 $wR(F^2) = 0.070$   
 $S = 1.15$   
 521 reflections  
 44 parameters  
 0 restraints  
 Primary atom site location: structure-invariant  
 direct methods  
 Secondary atom site location: difference Fourier  
 map

Hydrogen site location: inferred from  
 neighbouring sites  
 Only H-atom coordinates refined  
 $w = 1/[\sigma^2(F_o^2) + (0.0415P)^2 + 0.1865P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\text{max}} < 0.001$   
 $\Delta\rho_{\text{max}} = 0.41 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\text{min}} = -0.31 \text{ e } \text{\AA}^{-3}$   
 Extinction correction: *SHELXTL* (Sheldrick,  
 2008),  $F_c^* = kFc[1 + 0.001x \text{Fc}^2 \lambda^3 / \sin(2\theta)]^{-1/4}$   
 Extinction coefficient: 0.066 (5)

*Special details*

**Experimental.** crystal mounted on cryoloop using Paratone-N

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cl1	0.91302 (8)	0.07997 (5)	0.67014 (3)	0.0043 (3)
C1	0.3856 (4)	0.0762 (2)	0.97532 (18)	0.0037 (4)
H1A	0.254 (5)	0.019 (3)	0.903 (2)	0.004*
H1B	0.269 (5)	0.131 (3)	1.047 (2)	0.004*
N1	0.5508 (3)	0.2442 (2)	0.91610 (14)	0.0044 (4)
H1C	0.660 (5)	0.305 (3)	0.983 (2)	0.007*
H1D	0.444 (5)	0.328 (4)	0.885 (2)	0.007*
H1E	0.654 (5)	0.206 (3)	0.845 (2)	0.007*

*Atomic displacement parameters ( $\text{\AA}^2$ )*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Cl1	0.0057 (3)	0.0048 (3)	0.0023 (3)	-0.00036 (11)	-0.00138 (18)	-0.00050 (11)
C1	0.0017 (8)	0.0051 (9)	0.0041 (8)	-0.0003 (6)	-0.0013 (7)	0.0001 (5)
N1	0.0067 (7)	0.0032 (7)	0.0031 (7)	0.0013 (6)	-0.0015 (6)	0.0009 (5)

*Geometric parameters (Å, °)*

C1—N1	1.488 (2)	N1—H1C	0.91 (2)
C1—C1 <sup>i</sup>	1.522 (3)	N1—H1D	0.80 (3)
C1—H1A	0.99 (2)	N1—H1E	0.89 (2)
C1—H1B	0.96 (2)		
N1—C1—C1 <sup>i</sup>	109.68 (18)	C1—N1—H1C	108.8 (13)
N1—C1—H1A	107.3 (12)	C1—N1—H1D	114.8 (16)
C1 <sup>i</sup> —C1—H1A	109.4 (13)	H1C—N1—H1D	104 (2)
N1—C1—H1B	105.1 (13)	C1—N1—H1E	110.4 (15)
C1 <sup>i</sup> —C1—H1B	112.8 (13)	H1C—N1—H1E	116 (2)
H1A—C1—H1B	112.3 (17)	H1D—N1—H1E	102 (2)

Symmetry code: (i)  $-x+1, -y, -z+2$ .*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1E...C11	0.89 (2)	2.27 (2)	3.1514 (15)	175 (2)
N1—H1D...C11 <sup>ii</sup>	0.80 (3)	2.39 (3)	3.1770 (15)	170 (2)
N1—H1C...C11 <sup>iii</sup>	0.91 (2)	2.29 (2)	3.1922 (15)	171 (2)

Symmetry codes: (ii)  $-x+1, y+1/2, -z+3/2$ ; (iii)  $x, -y+1/2, z+1/2$ .