

Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

## 4-Nitrobenzyl 2-chloroacetate

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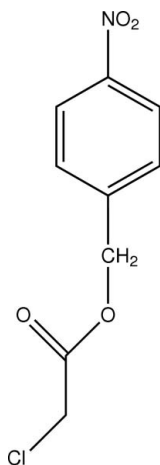
Received 20 May 2009; accepted 21 May 2009

 Key indicators: single-crystal X-ray study;  $T = 294$  K; mean  $\sigma(\text{C}-\text{C}) = 0.006$  Å;  $R$  factor = 0.068;  $wR$  factor = 0.191; data-to-parameter ratio = 13.3.

In the molecule of the title compound,  $\text{C}_9\text{H}_8\text{ClNO}_4$ , the nearly planar acetate moiety [maximum deviation = 0.015 (3) Å for an O atom] is oriented with respect to the plane of the aromatic ring at a dihedral angle of 73.03 (3)°. In the crystal structure, intermolecular  $\text{C}-\text{H}\cdots\text{O}$  interactions link molecules into a network.  $\pi-\pi$  contacts between benzene rings [centroid-centroid distance = 4.000 (1) Å] may further stabilize the structure.

## Related literature

For a related structure, see: Pyun *et al.* (2001). For bond-length data, see: Allen *et al.* (1987).



## Experimental

## Crystal data

$\text{C}_9\text{H}_8\text{ClNO}_4$	$V = 1993.5$ (8) Å <sup>3</sup>
$M_r = 229.61$	$Z = 8$
Monoclinic, $C2/c$	Mo $K\alpha$ radiation
$a = 13.636$ (3) Å	$\mu = 0.38$ mm <sup>-1</sup>
$b = 8.1570$ (16) Å	$T = 294$ K
$c = 18.878$ (4) Å	$0.30 \times 0.20 \times 0.10$ mm
$\beta = 108.30$ (3)°	

## Data collection

Enraf-Nonius CAD-4 diffractometer	1814 independent reflections
Absorption correction: $\psi$ scan (North <i>et al.</i> , 1968)	1132 reflections with $I > 2\sigma(I)$
$T_{\min} = 0.896$ , $T_{\max} = 0.963$	$R_{\text{int}} = 0.055$
1892 measured reflections	3 standard reflections
	frequency: 120 min
	intensity decay: 1%

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.068$	136 parameters
$wR(F^2) = 0.191$	H-atom parameters constrained
$S = 1.00$	$\Delta\rho_{\max} = 0.31$ e Å <sup>-3</sup>
1814 reflections	$\Delta\rho_{\min} = -0.28$ e Å <sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{C1}-\text{H1B}\cdots\text{O1}^{\text{i}}$	0.97	2.35	3.275 (5)	160
$\text{C3}-\text{H3A}\cdots\text{O1}^{\text{ii}}$	0.97	2.58	3.456 (5)	151

 Symmetry codes: (i)  $-x + \frac{3}{2}, y + \frac{1}{2}, -z + \frac{3}{2}$ ; (ii)  $-x + 1, y, -z + \frac{3}{2}$ .

Data collection: *CAD-4 Software* (Enraf-Nonius, 1989); cell refinement: *CAD-4 Software*; data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *SHELXL97* and *PLATON* (Spek, 2009).

The authors thank the Center for Testing and Analysis, Nanjing University, for support.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HK2695).

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## supporting information

*Acta Cryst.* (2009). E65, o1423 [doi:10.1107/S1600536809019278]

## 4-Nitrobenzyl 2-chloroacetate

Kai Zhu, Yan-Hua Wang, Hui Liu, Ping-Fang Han and Ping Wei

### S1. Comment

Some derivatives of *p*-nitrobenzyl alcohol are important chemical materials. We report herein the crystal structure of the title compound.

In the molecule of the title compound (Fig 1), the bond lengths (Allen *et al.*, 1987) and angles are within normal ranges. Ring A (C4-C9) is, of course, planar. Atoms N, C3, O3 and O4 are 0.005 (3), 0.027 (3), -0.146 (3) and 0.144 (4) Å away from the ring plane, respectively. On the other hand, (O1/O2/C1-C3) moiety is nearly planar with a maximum deviation of 0.015 (3) Å for atom O2 and it is oriented with respect to ring A at a dihedral angle of 73.03 (3)°.

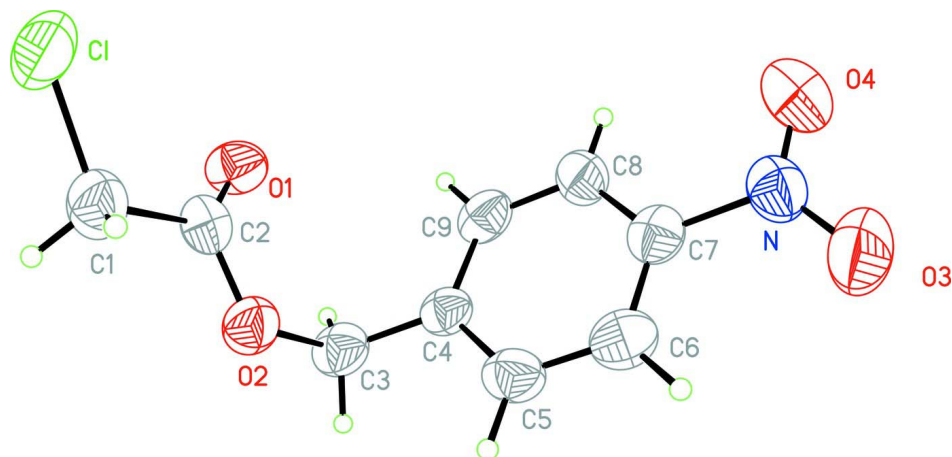
In the crystal structure, intermolecular C-H...O interactions (Table 1) link the molecules into a network (Fig. 2), in which they may be effective in the stabilization of the structure. The  $\pi$ - $\pi$  contact between the benzene rings, Cg1—Cg1<sup>i</sup> [symmetry code: (i) 1 - x, -y, -z, where Cg1 is centroid of the ring A (C4-C9)] may further stabilize the structure, with centroid-centroid distance of 4.000 (1) Å.

### S2. Experimental

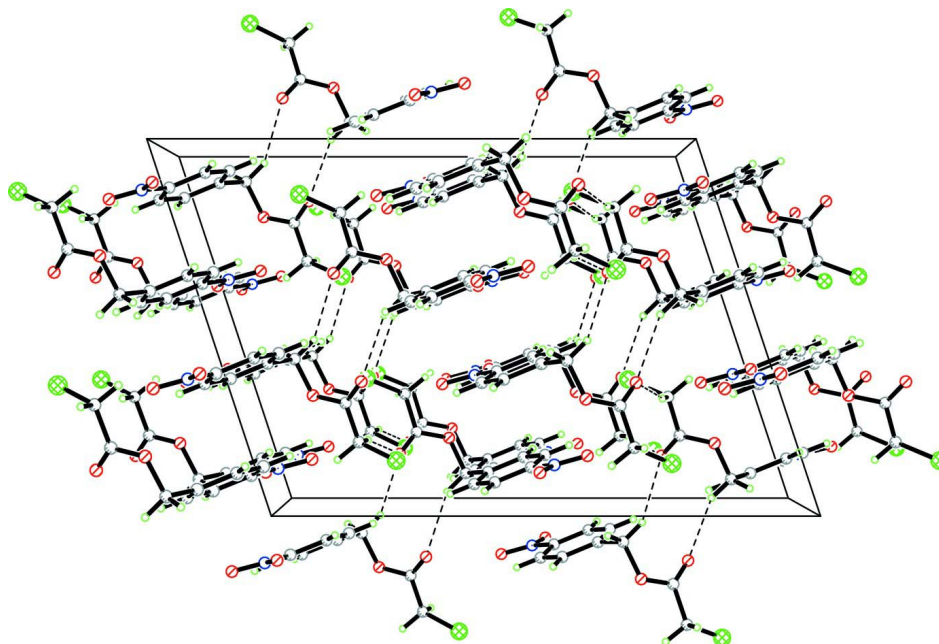
For the preparation of the title compound, chloroacetyl chloride (1.1 g) and *p*-nitrobenzyl alcohol (1.53 g) were added into the mixture of pyridine (15 ml) and dichloromethane (30 ml) at 273–278 K. The gross products were extracted with *n*-hexane, washed with water, and dried under vacuum, and then recrystallized in dichloromethane (yield; 0.916 g) (Pyun *et al.*, 2001). Crystals suitable for X-ray analysis were obtained by slow evaporation of a methanol solution.

### S3. Refinement

H atoms were positioned geometrically, with C-H = 0.93 and 0.97 Å for aromatic and methylene H, respectively, and constrained to ride on their parent atoms, with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ .

**Figure 1**

The molecular structure of the title molecule, with the atom-numbering scheme. Displacement ellipsoids are drawn at the 30% probability level.

**Figure 2**

A partial packing diagram of the title compound. Hydrogen bonds are shown as dashed lines.

#### 4-Nitrobenzyl 2-chloroacetate

##### *Crystal data*

$C_9H_8ClNO_4$

$M_r = 229.61$

Monoclinic,  $C2/c$

Hall symbol:  $-C 2yc$

$a = 13.636 (3) \text{ \AA}$

$b = 8.1570 (16) \text{ \AA}$

$c = 18.878 (4) \text{ \AA}$

$\beta = 108.30 (3)^\circ$

$V = 1993.5 (8) \text{ \AA}^3$

$Z = 8$

$F(000) = 944$

$D_x = 1.530 \text{ Mg m}^{-3}$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 25 reflections

$\theta = 9\text{--}13^\circ$

$\mu = 0.38 \text{ mm}^{-1}$

$T = 294$  K  $0.30 \times 0.20 \times 0.10$  mm  
 Block, colorless

*Data collection*

Enraf–Nonius CAD-4 diffractometer	1814 independent reflections 1132 reflections with $I > 2\sigma(I)$
Radiation source: fine-focus sealed tube	$R_{\text{int}} = 0.055$
Graphite monochromator	$\theta_{\text{max}} = 25.3^\circ$ , $\theta_{\text{min}} = 2.3^\circ$
$\omega/2\theta$ scans	$h = -16 \rightarrow 0$
Absorption correction: $\psi$ scan (North <i>et al.</i> , 1968)	$k = -9 \rightarrow 0$
$T_{\text{min}} = 0.896$ , $T_{\text{max}} = 0.963$	$l = -21 \rightarrow 22$
1892 measured reflections	3 standard reflections every 120 min intensity decay: 1%

*Refinement*

Refinement on $F^2$	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.068$	H-atom parameters constrained
$wR(F^2) = 0.191$	$w = 1/[\sigma^2(F_o^2) + (0.1P)^2 + 2P]$
$S = 1.00$	where $P = (F_o^2 + 2F_c^2)/3$
1814 reflections	$(\Delta/\sigma)_{\text{max}} < 0.001$
136 parameters	$\Delta\rho_{\text{max}} = 0.31 \text{ e } \text{\AA}^{-3}$
0 restraints	$\Delta\rho_{\text{min}} = -0.28 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	

*Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
Cl	0.85639 (10)	0.10201 (17)	0.74171 (7)	0.0745 (5)
O1	0.6449 (2)	0.0812 (4)	0.75755 (16)	0.0561 (8)
O2	0.6848 (2)	0.2403 (4)	0.86053 (15)	0.0528 (8)
O3	0.6555 (4)	-0.3688 (5)	1.1134 (2)	0.1110 (17)
O4	0.6159 (3)	-0.5146 (4)	1.0154 (2)	0.0854 (11)
N	0.6314 (3)	-0.3823 (5)	1.0467 (2)	0.0567 (10)
C1	0.8107 (3)	0.2211 (5)	0.8028 (3)	0.0578 (11)
H1A	0.8582	0.2114	0.8531	0.069*
H1B	0.8085	0.3355	0.7883	0.069*
C2	0.7037 (3)	0.1679 (5)	0.8022 (2)	0.0444 (9)
C3	0.5855 (3)	0.2023 (5)	0.8709 (2)	0.0531 (11)
H3A	0.5334	0.1868	0.8228	0.064*
H3B	0.5641	0.2927	0.8961	0.064*

C4	0.5956 (3)	0.0483 (5)	0.9169 (2)	0.0411 (9)
C5	0.6365 (3)	0.0540 (5)	0.9945 (2)	0.0518 (10)
H5A	0.6568	0.1545	1.0177	0.062*
C6	0.6476 (3)	-0.0835 (5)	1.0369 (2)	0.0544 (11)
H6A	0.6741	-0.0768	1.0886	0.065*
C7	0.6192 (3)	-0.2331 (5)	1.0024 (2)	0.0465 (10)
C8	0.5789 (3)	-0.2437 (5)	0.9254 (2)	0.0505 (10)
H8A	0.5596	-0.3446	0.9023	0.061*
C9	0.5679 (3)	-0.1045 (5)	0.8843 (2)	0.0507 (10)
H9A	0.5411	-0.1116	0.8326	0.061*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Cl	0.0709 (9)	0.0894 (9)	0.0776 (9)	-0.0099 (7)	0.0438 (7)	-0.0052 (7)
O1	0.0460 (17)	0.0621 (18)	0.0581 (17)	-0.0091 (14)	0.0132 (14)	-0.0194 (15)
O2	0.0569 (18)	0.0581 (17)	0.0450 (15)	-0.0098 (14)	0.0182 (13)	-0.0066 (13)
O3	0.195 (5)	0.079 (3)	0.056 (2)	-0.008 (3)	0.036 (3)	0.008 (2)
O4	0.104 (3)	0.056 (2)	0.097 (3)	0.005 (2)	0.033 (2)	0.009 (2)
N	0.055 (2)	0.052 (2)	0.066 (3)	0.0018 (18)	0.0231 (19)	0.0086 (19)
C1	0.054 (3)	0.061 (3)	0.060 (3)	-0.012 (2)	0.020 (2)	-0.004 (2)
C2	0.040 (2)	0.043 (2)	0.046 (2)	-0.0012 (19)	0.0078 (18)	0.0066 (19)
C3	0.045 (2)	0.059 (3)	0.059 (3)	0.0026 (19)	0.023 (2)	-0.004 (2)
C4	0.036 (2)	0.047 (2)	0.044 (2)	-0.0013 (17)	0.0165 (16)	-0.0054 (17)
C5	0.053 (3)	0.048 (2)	0.055 (3)	-0.007 (2)	0.018 (2)	-0.012 (2)
C6	0.052 (3)	0.061 (3)	0.046 (2)	-0.003 (2)	0.0091 (19)	-0.010 (2)
C7	0.035 (2)	0.056 (2)	0.051 (2)	0.0016 (19)	0.0182 (18)	0.002 (2)
C8	0.056 (3)	0.044 (2)	0.052 (2)	-0.011 (2)	0.019 (2)	-0.004 (2)
C9	0.051 (2)	0.063 (3)	0.039 (2)	-0.008 (2)	0.0163 (18)	-0.009 (2)

*Geometric parameters (Å, °)*

Cl—C1	1.765 (4)	C3—H3B	0.9700
O1—C2	1.195 (5)	C4—C9	1.389 (5)
O2—C2	1.344 (5)	C4—C5	1.394 (5)
O2—C3	1.461 (5)	C5—C6	1.359 (6)
N—O3	1.203 (5)	C5—H5A	0.9300
N—O4	1.217 (5)	C6—C7	1.381 (6)
N—C7	1.457 (5)	C6—H6A	0.9300
C1—C2	1.519 (6)	C7—C8	1.385 (6)
C1—H1A	0.9700	C8—C9	1.357 (6)
C1—H1B	0.9700	C8—H8A	0.9300
C3—C4	1.509 (6)	C9—H9A	0.9300
C3—H3A	0.9700		
C2—O2—C3	116.2 (3)	C9—C4—C5	117.4 (4)
O3—N—O4	122.7 (4)	C9—C4—C3	121.9 (3)
O3—N—C7	118.0 (4)	C5—C4—C3	120.7 (4)

O4—N—C7	119.4 (4)	C6—C5—C4	121.7 (4)
C2—C1—C1	111.9 (3)	C6—C5—H5A	119.2
C2—C1—H1A	109.2	C4—C5—H5A	119.2
C1—C1—H1A	109.2	C5—C6—C7	119.3 (4)
C2—C1—H1B	109.2	C5—C6—H6A	120.4
C1—C1—H1B	109.2	C7—C6—H6A	120.4
H1A—C1—H1B	107.9	C6—C7—C8	120.7 (4)
O1—C2—O2	125.3 (4)	C6—C7—N	120.2 (4)
O1—C2—C1	127.2 (4)	C8—C7—N	119.1 (4)
O2—C2—C1	107.4 (3)	C9—C8—C7	118.9 (4)
O2—C3—C4	109.5 (3)	C9—C8—H8A	120.5
O2—C3—H3A	109.8	C7—C8—H8A	120.5
C4—C3—H3A	109.8	C8—C9—C4	122.1 (4)
O2—C3—H3B	109.8	C8—C9—H9A	119.0
C4—C3—H3B	109.8	C4—C9—H9A	119.0
H3A—C3—H3B	108.2		
C3—O2—C2—O1	2.8 (6)	C5—C6—C7—N	179.3 (4)
C3—O2—C2—C1	-178.9 (3)	O3—N—C7—C6	8.3 (6)
C1—C1—C2—O1	-14.4 (6)	O4—N—C7—C6	-172.6 (4)
C1—C1—C2—O2	167.3 (3)	O3—N—C7—C8	-171.9 (5)
C2—O2—C3—C4	86.8 (4)	O4—N—C7—C8	7.3 (6)
O2—C3—C4—C9	-96.3 (4)	C6—C7—C8—C9	0.1 (6)
O2—C3—C4—C5	81.6 (5)	N—C7—C8—C9	-179.7 (4)
C9—C4—C5—C6	-1.1 (6)	C7—C8—C9—C4	-0.2 (6)
C3—C4—C5—C6	-179.1 (4)	C5—C4—C9—C8	0.7 (6)
C4—C5—C6—C7	1.0 (7)	C3—C4—C9—C8	178.6 (4)
C5—C6—C7—C8	-0.5 (6)		

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
C1—H1B...O1 <sup>i</sup>	0.97	2.35	3.275 (5)	160
C3—H3A...O1 <sup>ii</sup>	0.97	2.58	3.456 (5)	151

Symmetry codes: (i)  $-x+3/2, y+1/2, -z+3/2$ ; (ii)  $-x+1, y, -z+3/2$ .