Acta Crystallographica Section E
Structure Reports Online
ISSN 1600-5368

Dichlorido{[2-(diphenylphosphino)-phenyliminomethyl]ferrocene-κ^2N,P}-platinum(II) dichloromethane hemisolvate

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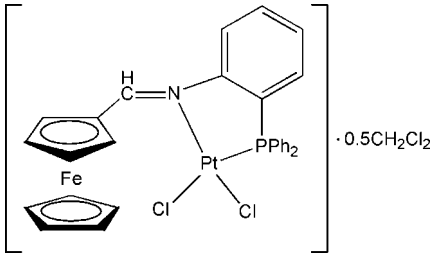
Received 26 November 2008; accepted 20 December 2008

Key indicators: single-crystal X-ray study; T = 293 K; mean σ(C–C) = 0.008 Å; disorder in solvent or counterion; R factor = 0.043; wR factor = 0.116; data-to-parameter ratio = 17.1.

In the title compound, $[FePt(C_5H_5)(C_{24}H_{19}NP)Cl_2]\cdot 0.5CH_2Cl_2$, the Pt^{II} atom adopts a distorted square-planar geometry defined by one P atom and one N atom from the bidentate [2-(diphenylphosphino)phenyliminomethyl]ferrocene ligand and two Cl atoms. Two disordered dichloromethane solvent molecules are each 0.25-occupied on a twofold rotation axis.

Related literature

For general background, see: Cullen & Woolins (1981); Farrell *et al.* (2002); Gul *et al.* (2002); Wu *et al.* (2001). For the ligand synthesis, see: Gong *et al.* (2006); Zhang *et al.* (2006).



Experimental

Crystal data

$[FePt(C_5H_5)(C_{24}H_{19}NP)Cl_2]\cdot 0.5CH_2Cl_2$
M_r = 781.75
Orthorhombic, $P2_12_12$
a = 19.5045 (11) Å
b = 12.0182 (7) Å
c = 12.9800 (8) Å
V = 3042.6 (3) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 5.40 mm^{-1}
T = 293 (2) K
0.3 × 0.1 × 0.1 mm

Data collection

Siemens SMART CCD diffractometer
Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)
T_{min} = 0.532, T_{max} = 0.588
19964 measured reflections
5902 independent reflections
4954 reflections with $I > 2\sigma(I)$
R_{int} = 0.058

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.043
$wR(F^2)$ = 0.116
S = 1.04
5902 reflections
345 parameters
24 restraints
H-atom parameters constrained
$\Delta\rho_{max}$ = 1.18 e Å^{-3}
$\Delta\rho_{min}$ = −0.93 e Å^{-3}
Absolute structure: Flack (1983), 856 Friedel pairs
Flack parameter: 0.002 (11)

Table 1
Selected bond lengths (Å).

Pt—N	2.031 (3)	Pt—Cl1	2.2996 (13)
Pt—P	2.2089 (10)	Pt—Cl2	2.3673 (11)

Data collection: *SMART* (Siemens, 1996); cell refinement: *SAINT* (Siemens, 1996); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

The authors thank Jiangxi Provincial Department of Education (2007168) and the Natural Science Foundation of Jiangxi Province (0620071) for financial support for this work. In addition, the Materials Chemistry Department of Nanchang Hangkong University is acknowledged.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HY2171).

References

Cullen, W. R. & Woolins, J. D. (1981). *Coord. Chem. Rev.* **39**, 1–47.
Farrell, A., Goddard, R. & Guiry, P. J. (2002). *J. Org. Chem.* **67**, 4209–4217.
Flack, H. D. (1983). *Acta Cryst.* A**39**, 876–881.
Gong, J. F., Wang, D. W., Zhang, Y. H., Zhu, Y. & Wu, Y. J. (2006). *Inorg. Chim. Acta*, **359**, 2115–2119.
Gul, N., Nelson, J. H., Willis, A. C. & Ross, A. D. (2002). *Organometallics*, **21**, 2041–2045.
Sheldrick, G. M. (1996). *SADABS*. University of Göttingen, Germany.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Siemens (1996). *SMART* and *SAINT*. Siemens Analytical X-ray Instruments Inc., Madison, Wisconsin, USA.
Wu, Y., Hou, J., Cui, X. & Yuan, R. (2001). *J. Organomet. Chem.* **637–639**, 793–798.
Zhang, X.-L., Kuang, R.-Y., Cao, Q.-Y., Li, Y.-X. & Luo, L.-T. (2006). *Acta Cryst.* E**62**, m2491–m2492.

supporting information

Acta Cryst. (2009). E**65**, m138 [doi:10.1107/S1600536808043493]

Dichlorido{[2-(diphenylphosphino)phenyliminomethyl]ferrocene-κ^2N,P}platinum(II) dichloromethane hemisolvate

Yu Xie, Ya Wei, Jin-Gang Hu, Jin-Sheng Zhao and Feng-Xing Jiang

S1. Comment

Ferrocene derivatives containing heteroatoms with good donor abilities have attracted great interest during the last few years in organometallic chemistry, since the coordination of a metal to these heteroatoms produces polymetallic molecules (Cullen & Woolins, 1981). Some examples showing the utility of these polymetallic compounds in homogeneous catalysis (Farrell *et al.*, 2002; Wu *et al.*, 2001) or for the design of new materials with outstanding properties (Gul *et al.*, 2002) have been described. Here we report a new bimetallic platinum(II) complex with a ferrocenyliminophosphine ligand.

In the title compound, the Pt^{II} atom adopts a distorted square planar geometry, defined by one P atom and one N atom from the bidentate ferrocenyliminophosphine ligand (*L*) and two Cl atoms (Fig. 1), with N—Pt—Cl1 = 176.29 (9)° and P—Pt—Cl2 = 170.19 (4)°. The difference between the Pt—Cl bond lengths [2.2996 (13) and 2.3673 (11) Å] (Table 1) reflects the stronger *trans* influence of the tertiary phosphine compared with the imino group. The ligand *L* adopts a five-membered chelating ring. The benzene ring and the cyclopentadienyl ring are *trans*, with respect to the C11—N double bond.

S2. Experimental

The ligand *L* was prepared by literature method (Gong *et al.*, 2006; Zhang *et al.*, 2006). The title compound was prepared by reacting equal molar K_2PtCl_4 and *L* in CH_2Cl_2 (yield 83%). Brown needle crystals suitable for X-ray analysis were obtained by vapor diffusion of diethyl ether into a solution of the title compound in CH_2Cl_2. Analysis calculated for $C_{29.5}H_{25}Cl_3FeNPPt$: C 45.32, H 3.22, N 1.79%; found: C 45.61, H 3.33, N 1.70%.

S3. Refinement

H atoms were positioned geomertrically and treated as riding, with C—H = 0.98 (cyclopentadienyl), 0.93 (phenyl) and 0.96 (CH_2) Å and with $U_{iso}(H) = 1.2U_{eq}(C)$. The highest residual electron density was found 0.67 Å from atom Cl4 and the deepest hole 0.79 Å from atom Pt. The disordered dichloromethane solvent molecules are each 0.25-occupied on a twofold rotation axis.

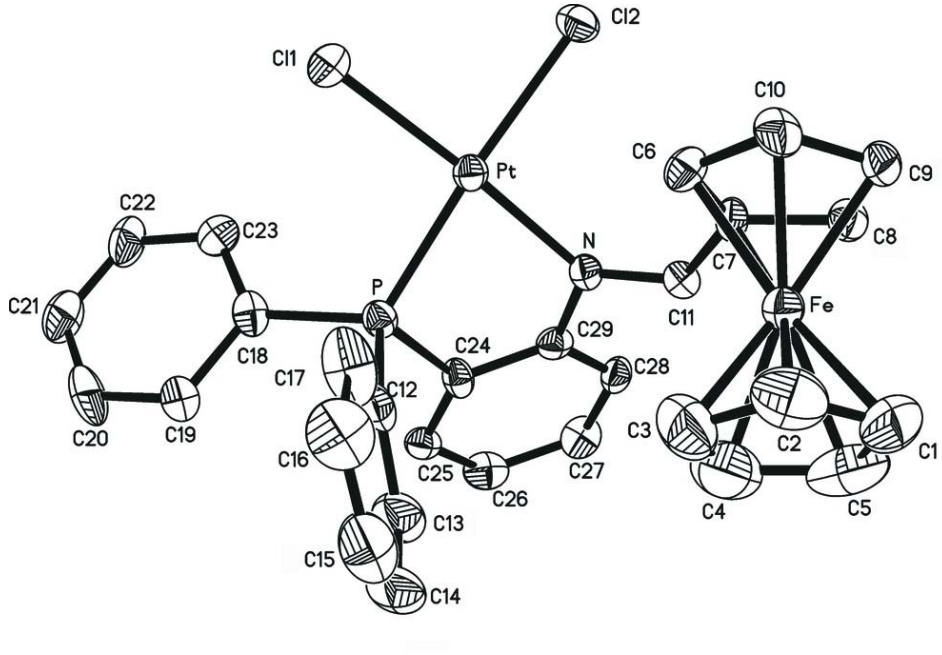


Figure 1

The molecular structure of the title compound. Displacement ellipsoids are drawn at the 30% probability level. H atoms and disordered dichloromethane solvent molecules have been omitted for clarity.

Dichlorido{[2-(diphenylphosphino)phenyliminomethyl]ferrocene- κ^2N,P}platinum(II) dichloromethane hemisolvate

Crystal data

$[FePt(C_5H_5)(C_{24}H_{19}NP)Cl_2]\cdot 0.5CH_2Cl_2$
$M_r = 781.75$
Orthorhombic, $P2_12_12$
Hall symbol: P 2 2ab
$a = 19.5045\ (11)$ Å
$b = 12.0182\ (7)$ Å
$c = 12.9800\ (8)$ Å
$V = 3042.6\ (3)$ Å^3
$Z = 4$
$F(000) = 1516$
$D_x = 1.707$ Mg m^{-3}
Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Cell parameters from 892 reflections
$\theta = 3.2–25.8°$
$\mu = 5.40$ mm^{-1}
$T = 293$ K
Needle, brown–yellow
0.3 × 0.1 × 0.1 mm

Data collection

Siemens SMART CCD diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
ω scans
Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)
$T_{min} = 0.532$, $T_{max} = 0.588$
19964 measured reflections
5902 independent reflections
4954 reflections with $I > 2\sigma(I)$
$R_{int} = 0.058$
$\theta_{max} = 26.0°$, $\theta_{min} = 1.9°$
$h = -23\rightarrow24$
$k = -14\rightarrow14$
$l = -16\rightarrow15$

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)] = 0.043$
$wR(F^2) = 0.116$
$S = 1.04$
5902 reflections
345 parameters
24 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H-atom parameters constrained
$w = 1/[\sigma^2(F_o^2) + (0.05P)^2]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} = 0.024$
$\Delta\rho_{max} = 1.18$ e Å^{-3}
$\Delta\rho_{min} = -0.93$ e Å^{-3}
Absolute structure: Flack (1983), 856 Friedel pairs
Absolute structure parameter: 0.002 (11)

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)

	x	*y*	*z*	U_{iso}*/U_{eq}	Occ. (<1)
Pt	0.270357 (7)	0.911967 (14)	0.143789 (12)	0.03558 (4)	
Fe	0.34183 (3)	0.52770 (6)	0.19921 (5)	0.04447 (19)	
Cl1	0.32009 (6)	1.06905 (11)	0.21198 (10)	0.0568 (4)	
Cl2	0.35362 (5)	0.90898 (12)	0.01112 (9)	0.0538 (3)	
Cl3	0.4874 (2)	0.8826 (5)	0.3249 (4)	0.167 (2)	0.50
Cl4	0.06774 (11)	0.54821 (19)	0.23620 (19)	0.0506 (7)	0.50
P	0.17978 (5)	0.91268 (11)	0.24594 (8)	0.0378 (3)	
N	0.22282 (15)	0.7723 (3)	0.0921 (2)	0.0337 (9)	
C1	0.3477 (4)	0.3728 (6)	0.2664 (5)	0.097 (3)	
H1A	0.3771	0.3110	0.2447	0.117*	
C2	0.3629 (4)	0.4509 (6)	0.3371 (4)	0.096 (3)	
H2A	0.4066	0.4563	0.3742	0.115*	
C3	0.3087 (4)	0.5257 (7)	0.3468 (5)	0.102 (2)	
H3A	0.3067	0.5896	0.3936	0.123*	
C4	0.2585 (4)	0.4921 (8)	0.2825 (6)	0.112 (2)	
H4A	0.2137	0.5282	0.2749	0.135*	
C5	0.2797 (4)	0.3974 (6)	0.2289 (6)	0.110 (3)	
H5A	0.2530	0.3543	0.1789	0.132*	
C6	0.3711 (2)	0.6848 (4)	0.1581 (3)	0.0451 (13)	
H6A	0.3666	0.7517	0.2007	0.054*	
C7	0.3221 (2)	0.6444 (4)	0.0886 (3)	0.0402 (12)	
C8	0.3478 (2)	0.5451 (4)	0.0462 (3)	0.0442 (13)	
H8A	0.3246	0.4984	−0.0050	0.053*	
C9	0.4145 (2)	0.5275 (4)	0.0860 (4)	0.0482 (13)	
H9A	0.4452	0.4658	0.0686	0.058*	
C10	0.4278 (2)	0.6129 (4)	0.1561 (4)	0.0539 (15)	
H10A	0.4696	0.6210	0.1973	0.065*	
C11	0.25027 (18)	0.6757 (4)	0.0722 (3)	0.0360 (11)	
H11A	0.2215	0.6215	0.0449	0.043*	
C12	0.1890 (2)	0.8272 (4)	0.3601 (4)	0.0516 (13)	
C13	0.1388 (3)	0.7539 (5)	0.3915 (4)	0.0752 (19)	
H13A	0.0985	0.7462	0.3539	0.090*	

C14	0.1495 (4)	0.6915 (6)	0.4806 (4)	0.090 (2)	
H14A	0.1174	0.6392	0.5022	0.108*	
C15	0.2068 (3)	0.7088 (6)	0.5338 (5)	0.092 (2)	
H15A	0.2139	0.6660	0.5925	0.111*	
C16	0.2540 (3)	0.7823 (6)	0.5090 (5)	0.0902 (19)	
H16A	0.2923	0.7946	0.5502	0.108*	
C17	0.2437 (3)	0.8398 (6)	0.4191 (4)	0.086 (2)	
H17A	0.2770	0.8904	0.3985	0.104*	
C18	0.1426 (2)	1.0435 (4)	0.2854 (3)	0.0443 (13)	
C19	0.1014 (3)	1.0460 (5)	0.3785 (4)	0.0586 (16)	
H19A	0.0966	0.9831	0.4195	0.070*	
C20	0.0699 (3)	1.1442 (5)	0.4040 (5)	0.0793 (19)	
H20A	0.0416	1.1462	0.4617	0.095*	
C21	0.0784 (3)	1.2379 (4)	0.3481 (5)	0.0659 (17)	
H21A	0.0579	1.3039	0.3693	0.079*	
C22	0.1165 (3)	1.2358 (5)	0.2618 (5)	0.0641 (17)	
H22A	0.1198	1.2999	0.2219	0.077*	
C23	0.1507 (2)	1.1413 (4)	0.2307 (4)	0.0505 (14)	
H23A	0.1792	1.1434	0.1732	0.061*	
C24	0.1200 (2)	0.8442 (4)	0.1607 (3)	0.0378 (11)	
C25	0.0481 (2)	0.8598 (4)	0.1615 (3)	0.0404 (12)	
H25A	0.0273	0.9038	0.2115	0.049*	
C26	0.0092 (2)	0.8074 (4)	0.0854 (4)	0.0536 (15)	
H26A	−0.0381	0.8164	0.0841	0.064*	
C27	0.0420 (2)	0.7399 (4)	0.0095 (3)	0.0506 (14)	
H27A	0.0159	0.7044	−0.0406	0.061*	
C28	0.1120 (2)	0.7269 (4)	0.0100 (3)	0.0410 (12)	
H28A	0.1336	0.6825	−0.0389	0.049*	
C29	0.14973 (19)	0.7811 (3)	0.0845 (3)	0.0372 (11)	
C30	0.5000	1.0000	0.2754 (18)	0.104 (8)	0.50
H30A	0.5386	0.9885	0.2308	0.125*	0.50
C31	0.0000	0.5000	0.1697 (9)	0.051 (4)	0.50
H31A	−0.0159	0.5590	0.1260	0.061*	0.50

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Pt	0.03605 (6)	0.03424 (7)	0.03646 (7)	0.00216 (7)	−0.00032 (7)	0.00005 (8)
Fe	0.0432 (3)	0.0405 (4)	0.0497 (4)	0.0021 (3)	0.0018 (3)	−0.0005 (3)
Cl1	0.0584 (6)	0.0514 (7)	0.0608 (7)	−0.0109 (6)	0.0017 (5)	−0.0103 (6)
Cl2	0.0508 (5)	0.0506 (6)	0.0598 (6)	0.0014 (6)	0.0200 (5)	0.0055 (6)
Cl3	0.131 (3)	0.146 (5)	0.223 (5)	−0.030 (3)	−0.087 (3)	0.041 (4)
Cl4	0.0373 (10)	0.0439 (13)	0.0706 (14)	−0.0113 (9)	−0.0071 (10)	−0.0009 (11)
P	0.0425 (5)	0.0380 (6)	0.0329 (5)	0.0046 (6)	0.0017 (4)	0.0009 (5)
N	0.0341 (15)	0.0340 (17)	0.0331 (15)	0.0043 (15)	0.0016 (14)	−0.0008 (14)
C1	0.132 (5)	0.060 (4)	0.100 (5)	−0.006 (4)	0.029 (4)	0.016 (3)
C2	0.119 (5)	0.115 (6)	0.054 (3)	−0.001 (4)	0.009 (3)	0.029 (3)
C3	0.124 (4)	0.104 (4)	0.078 (3)	0.008 (3)	0.032 (3)	−0.005 (3)

C4	0.097 (4)	0.117 (4)	0.123 (4)	0.006 (3)	0.030 (3)	0.021 (4)
C5	0.123 (5)	0.113 (5)	0.095 (5)	−0.060 (5)	−0.009 (4)	0.023 (4)
C6	0.046 (2)	0.038 (2)	0.051 (3)	−0.004 (2)	−0.010 (2)	−0.004 (2)
C7	0.051 (2)	0.031 (2)	0.039 (2)	0.009 (2)	0.001 (2)	−0.0032 (19)
C8	0.038 (2)	0.045 (3)	0.049 (2)	0.004 (2)	−0.0016 (19)	−0.003 (2)
C9	0.036 (2)	0.042 (3)	0.066 (3)	0.005 (2)	0.002 (2)	−0.006 (2)
C10	0.0300 (18)	0.065 (3)	0.067 (3)	0.004 (2)	−0.006 (2)	−0.011 (3)
C11	0.0337 (19)	0.044 (2)	0.0302 (19)	0.0036 (17)	−0.0006 (16)	0.0032 (19)
C12	0.063 (2)	0.047 (3)	0.044 (2)	0.020 (2)	0.016 (2)	−0.005 (2)
C13	0.095 (4)	0.068 (4)	0.063 (3)	0.015 (3)	0.018 (3)	0.018 (3)
C14	0.129 (5)	0.075 (4)	0.067 (3)	0.004 (4)	0.045 (3)	0.023 (3)
C15	0.121 (4)	0.087 (4)	0.069 (3)	0.035 (3)	0.006 (3)	0.011 (3)
C16	0.093 (3)	0.095 (4)	0.083 (3)	0.028 (3)	−0.014 (3)	0.018 (3)
C17	0.108 (4)	0.114 (5)	0.037 (2)	0.045 (4)	−0.034 (3)	−0.007 (3)
C18	0.040 (2)	0.049 (3)	0.044 (2)	0.013 (2)	−0.0078 (19)	−0.007 (2)
C19	0.065 (3)	0.058 (3)	0.053 (3)	−0.002 (3)	0.011 (2)	−0.011 (2)
C20	0.070 (3)	0.093 (4)	0.075 (3)	0.038 (3)	0.015 (3)	−0.029 (3)
C21	0.071 (3)	0.037 (3)	0.090 (4)	0.016 (2)	−0.010 (3)	−0.006 (3)
C22	0.064 (3)	0.042 (3)	0.085 (4)	0.010 (3)	−0.007 (3)	−0.017 (3)
C23	0.052 (2)	0.048 (3)	0.051 (3)	−0.003 (2)	−0.007 (2)	0.002 (2)
C24	0.048 (2)	0.039 (2)	0.026 (2)	0.0119 (19)	0.0011 (18)	0.0000 (18)
C25	0.0375 (19)	0.042 (2)	0.042 (2)	−0.0011 (19)	0.0028 (18)	0.003 (2)
C26	0.039 (2)	0.056 (3)	0.065 (3)	0.005 (2)	−0.004 (2)	0.017 (3)
C27	0.047 (2)	0.051 (3)	0.054 (3)	0.003 (2)	−0.018 (2)	0.001 (2)
C28	0.040 (2)	0.028 (2)	0.055 (3)	0.0042 (19)	0.003 (2)	−0.0047 (19)
C29	0.0346 (18)	0.029 (2)	0.048 (2)	0.0037 (18)	0.0009 (18)	0.0088 (19)
C30	0.017 (6)	0.121 (17)	0.174 (19)	0.001 (8)	0.000	0.000
C31	0.093 (9)	0.020 (6)	0.040 (7)	−0.019 (7)	0.000	0.000

Geometric parameters (Å, º)

Pt—N	2.031 (3)	C9—H9A	0.9800
Pt—P	2.2089 (10)	C10—H10A	0.9800
Pt—Cl1	2.2996 (13)	C11—H11A	0.9300
Pt—Cl2	2.3673 (11)	C12—C17	1.323 (7)
Fe—C1	2.059 (7)	C12—C13	1.379 (8)
Fe—C2	2.055 (6)	C13—C14	1.394 (8)
Fe—C3	2.022 (7)	C13—H13A	0.9300
Fe—C4	1.998 (8)	C14—C15	1.330 (9)
Fe—C5	2.017 (7)	C14—H14A	0.9300
Fe—C6	2.043 (5)	C15—C16	1.317 (10)
Fe—C7	2.044 (4)	C15—H15A	0.9300
Fe—C8	2.000 (5)	C16—C17	1.371 (9)
Fe—C9	2.041 (5)	C16—H16A	0.9300
Fe—C10	2.042 (5)	C17—H17A	0.9300
Cl3—C30	1.570 (11)	C18—C23	1.382 (7)
Cl4—C31	1.681 (6)	C18—C19	1.451 (6)
P—C18	1.806 (5)	C19—C20	1.371 (8)

P—C24	1.806 (4)	C19—H19A	0.9300
P—C12	1.812 (5)	C20—C21	1.350 (8)
N—C11	1.304 (5)	C20—H20A	0.9300
N—C29	1.433 (5)	C21—C22	1.345 (8)
C1—C2	1.346 (10)	C21—H21A	0.9300
C1—C5	1.443 (10)	C22—C23	1.378 (7)
C1—H1A	0.9800	C22—H22A	0.9300
C2—C3	1.393 (10)	C23—H23A	0.9300
C2—H2A	0.9800	C24—C29	1.375 (6)
C3—C4	1.348 (11)	C24—C25	1.413 (6)
C3—H3A	0.9800	C25—C26	1.397 (6)
C4—C5	1.396 (11)	C25—H25A	0.9300
C4—H4A	0.9800	C26—C27	1.427 (7)
C5—H5A	0.9800	C26—H26A	0.9300
C6—C7	1.401 (6)	C27—C28	1.375 (6)
C6—C10	1.404 (6)	C27—H27A	0.9300
C6—H6A	0.9800	C28—C29	1.378 (6)
C7—C8	1.407 (6)	C28—H28A	0.9300
C7—C11	1.466 (6)	C30—Cl3^{i}	1.570 (11)
C8—C9	1.415 (6)	C30—H30A	0.9600
C8—H8A	0.9800	C31—Cl4^{ii}	1.681 (6)
C9—C10	1.396 (7)	C31—H31A	0.9600

N—Pt—P	80.59 (9)	C10—C6—Fe	69.9 (3)
N—Pt—Cl1	176.29 (9)	C7—C6—H6A	125.9
P—Pt—Cl1	95.92 (4)	C10—C6—H6A	125.9
N—Pt—Cl2	93.46 (9)	Fe—C6—H6A	125.9
P—Pt—Cl2	170.19 (4)	C6—C7—C8	107.6 (4)
Cl1—Pt—Cl2	90.17 (4)	C6—C7—C11	131.0 (4)
C4—Fe—C8	127.4 (3)	C8—C7—C11	120.1 (4)
C4—Fe—C5	40.7 (3)	C6—C7—Fe	69.9 (3)
C8—Fe—C5	107.8 (3)	C8—C7—Fe	68.0 (3)
C4—Fe—C3	39.2 (3)	C11—C7—Fe	117.3 (3)
C8—Fe—C3	163.9 (2)	C7—C8—C9	108.2 (4)
C5—Fe—C3	67.5 (3)	C7—C8—Fe	71.3 (3)
C4—Fe—C9	162.5 (3)	C9—C8—Fe	71.1 (3)
C8—Fe—C9	40.96 (18)	C7—C8—H8A	125.9
C5—Fe—C9	123.6 (3)	C9—C8—H8A	125.9
C3—Fe—C9	154.7 (2)	Fe—C8—H8A	125.9
C4—Fe—C10	157.4 (3)	C10—C9—C8	107.4 (4)
C8—Fe—C10	68.14 (19)	C10—C9—Fe	70.1 (3)
C5—Fe—C10	159.2 (3)	C8—C9—Fe	68.0 (3)
C3—Fe—C10	121.9 (3)	C10—C9—H9A	126.3
C9—Fe—C10	39.98 (19)	C8—C9—H9A	126.3
C4—Fe—C6	124.5 (3)	Fe—C9—H9A	126.3
C8—Fe—C6	68.16 (18)	C9—C10—C6	108.6 (4)
C5—Fe—C6	159.2 (2)	C9—C10—Fe	70.0 (3)
C3—Fe—C6	110.3 (3)	C6—C10—Fe	69.9 (3)

C9—Fe—C6	67.63 (19)	C9—C10—H10A	125.7
C10—Fe—C6	40.19 (18)	C6—C10—H10A	125.7
C4—Fe—C7	111.9 (3)	Fe—C10—H10A	125.7
C8—Fe—C7	40.71 (18)	N—C11—C7	126.3 (4)
C5—Fe—C7	123.6 (2)	N—C11—H11A	116.9
C3—Fe—C7	127.8 (3)	C7—C11—H11A	116.9
C9—Fe—C7	68.04 (18)	C17—C12—C13	118.3 (5)
C10—Fe—C7	67.55 (18)	C17—C12—P	119.3 (4)
C6—Fe—C7	40.09 (17)	C13—C12—P	122.3 (4)
C4—Fe—C2	66.2 (3)	C12—C13—C14	118.9 (6)
C8—Fe—C2	154.1 (2)	C12—C13—H13A	120.6
C5—Fe—C2	66.8 (3)	C14—C13—H13A	120.6
C3—Fe—C2	40.0 (3)	C15—C14—C13	118.2 (6)
C9—Fe—C2	119.2 (2)	C15—C14—H14A	120.9
C10—Fe—C2	107.4 (3)	C13—C14—H14A	120.9
C6—Fe—C2	125.9 (2)	C16—C15—C14	124.5 (7)
C7—Fe—C2	163.2 (2)	C16—C15—H15A	117.8
C4—Fe—C1	67.8 (3)	C14—C15—H15A	117.8
C8—Fe—C1	120.8 (2)	C15—C16—C17	116.3 (6)
C5—Fe—C1	41.5 (3)	C15—C16—H16A	121.8
C3—Fe—C1	66.8 (3)	C17—C16—H16A	121.8
C9—Fe—C1	105.4 (2)	C12—C17—C16	123.7 (6)
C10—Fe—C1	121.6 (3)	C12—C17—H17A	118.2
C6—Fe—C1	158.5 (2)	C16—C17—H17A	118.2
C7—Fe—C1	158.2 (2)	C23—C18—C19	118.3 (4)
C2—Fe—C1	38.2 (3)	C23—C18—P	123.3 (3)
C18—P—C24	108.1 (2)	C19—C18—P	118.4 (4)
C18—P—C12	107.6 (2)	C20—C19—C18	117.9 (5)
C24—P—C12	107.9 (2)	C20—C19—H19A	121.1
C18—P—Pt	119.65 (15)	C18—C19—H19A	121.1
C24—P—Pt	98.46 (13)	C21—C20—C19	122.2 (5)
C12—P—Pt	114.18 (15)	C21—C20—H20A	118.9
C11—N—C29	117.4 (3)	C19—C20—H20A	118.9
C11—N—Pt	127.8 (3)	C22—C21—C20	120.0 (5)
C29—N—Pt	114.6 (2)	C22—C21—H21A	120.0
C2—C1—C5	106.8 (7)	C20—C21—H21A	120.0
C2—C1—Fe	70.7 (4)	C21—C22—C23	121.8 (5)
C5—C1—Fe	67.7 (4)	C21—C22—H22A	119.1
C2—C1—H1A	126.6	C23—C22—H22A	119.1
C5—C1—H1A	126.6	C22—C23—C18	119.6 (5)
Fe—C1—H1A	126.6	C22—C23—H23A	120.2
C1—C2—C3	110.1 (7)	C18—C23—H23A	120.2
C1—C2—Fe	71.1 (4)	C29—C24—C25	119.8 (4)
C3—C2—Fe	68.7 (4)	C29—C24—P	114.8 (3)
C1—C2—H2A	124.9	C25—C24—P	125.1 (3)
C3—C2—H2A	124.9	C26—C25—C24	118.3 (4)
Fe—C2—H2A	124.9	C26—C25—H25A	120.9
C4—C3—C2	107.6 (7)	C24—C25—H25A	120.9

C4—C3—Fe	69.5 (4)	C25—C26—C27	120.1 (4)
C2—C3—Fe	71.3 (4)	C25—C26—H26A	120.0
C4—C3—H3A	126.2	C27—C26—H26A	120.0
C2—C3—H3A	126.2	C28—C27—C26	120.4 (4)
Fe—C3—H3A	126.2	C28—C27—H27A	119.8
C3—C4—C5	109.7 (7)	C26—C27—H27A	119.8
C3—C4—Fe	71.3 (5)	C27—C28—C29	118.7 (4)
C5—C4—Fe	70.4 (4)	C27—C28—H28A	120.7
C3—C4—H4A	125.1	C29—C28—H28A	120.7
C5—C4—H4A	125.1	C24—C29—C28	122.7 (4)
Fe—C4—H4A	125.1	C24—C29—N	114.3 (4)
C4—C5—C1	105.7 (7)	C28—C29—N	123.0 (4)
C4—C5—Fe	68.9 (5)	Cl3^{i}—C30—Cl3	131.7 (16)
C1—C5—Fe	70.8 (4)	Cl3^{i}—C30—H30A	104.6
C4—C5—H5A	127.1	Cl3—C30—H30A	104.0
C1—C5—H5A	127.1	Cl4—C31—Cl4^{ii}	118.2 (7)
Fe—C5—H5A	127.1	Cl4—C31—H31A	107.6
C7—C6—C10	108.2 (4)	Cl4^{ii}—C31—H31A	107.7
C7—C6—Fe	70.0 (3)		

Symmetry codes: (i) $-x+1, -y+2, z$; (ii) $-x, -y+1, z$.