

Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

# Diaqua(5-methylpyrazine-2-carboxylato- $\kappa^2N^1,O$ )iron(II)

 Guang Fan,<sup>a</sup> Jia-Juan Sun,<sup>a</sup> Jun-Cai Zhang,<sup>a</sup> Zhan-Ying Ma<sup>a</sup> and Sheng-Li Gao<sup>b\*</sup>
<sup>a</sup>Department of Chemistry, Xianyang Normal University, Xianyang 712000, Shaanxi, People's Republic of China, and <sup>b</sup>Department of Chemistry, Northwest University, Xi'an 710069, Shaanxi, People's Republic of China

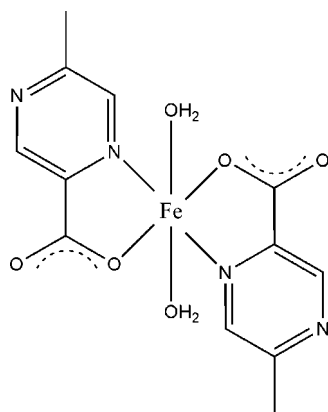
Correspondence e-mail: fanguang2004@163.com

Received 13 December 2008; accepted 13 December 2008

 Key indicators: single-crystal X-ray study;  $T = 298$  K; mean  $\sigma(C-C) = 0.005$  Å;  $R$  factor = 0.041;  $wR$  factor = 0.095; data-to-parameter ratio = 11.8.

In the neutral title complex,  $[Fe(C_6H_5N_2O_2)_2(H_2O)_2]$ , the coordination geometry around the  $Fe^{II}$  atom, which lies on an inversion centre, is distorted octahedral comprising two N atoms and two O atoms from two 5-methylpyrazine-2-carboxylate ligands, and two water molecules. The crystal structure is stabilized by a network of  $O-H\cdots O$  hydrogen bonds, resulting in a two-dimensional supramolecular structure.

## Related literature

 For background to this study, see: Fan *et al.* (2007).


## Experimental

## Crystal data

 $[Fe(C_6H_5N_2O_2)_2(H_2O)_2]$   
 $M_r = 366.12$   
 Triclinic,  $P\bar{1}$   
 $a = 5.068$  (1) Å  
 $b = 6.401$  (1) Å  
 $c = 12.381$  (1) Å  
 $\alpha = 103.851$  (2)°  
 $\beta = 91.790$ (10)°

 $\gamma = 108.340$  (2)°  
 $V = 368.22$  (10) Å<sup>3</sup>  
 $Z = 1$   
 Mo  $K\alpha$  radiation  
 $\mu = 1.06$  mm<sup>-1</sup>  
 $T = 298$  (2) K  
 $0.18 \times 0.09 \times 0.05$  mm

## Data collection

 Bruker SMART CCD area-detector diffractometer  
 Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)  
 $T_{min} = 0.832$ ,  $T_{max} = 0.949$ 

 1916 measured reflections  
 1260 independent reflections  
 1061 reflections with  $I > 2\sigma(I)$   
 $R_{int} = 0.025$ 

## Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.041$   
 $wR(F^2) = 0.095$   
 $S = 1.00$   
 1260 reflections

 107 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{max} = 0.33$  e Å<sup>-3</sup>  
 $\Delta\rho_{min} = -0.28$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$O3-H3A\cdots O1^i$	0.85	1.93	2.720 (3)	155
$O3-H3B\cdots O2^{ii}$	0.85	1.86	2.673 (3)	159

 Symmetry codes: (i)  $x + 1, y, z$ ; (ii)  $-x + 1, -y + 2, -z + 1$ .

Data collection: *SMART* (Bruker, 2002); cell refinement: *SAINT* (Bruker, 2002); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

We gratefully acknowledge the Special Research Fund of Xianyang Normal University for Talent Introduction (08XSYK305) and the Financial Support Fund from the Education Department of Shaanxi Province (No. 07JK424).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NG2527).

## References

- Bruker (2002). *SMART* and *SAINT*. Bruker AXS Inc., Madison, Wisconsin, USA.  
 Fan, G., Chen, S.-P. & Gao, S.-L. (2007). *Acta Cryst.* **E63**, m772–m773.  
 Sheldrick, G. M. (1996). *SADABS*. University of Göttingen, Germany.  
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.

## supporting information

*Acta Cryst.* (2009). E65, m107 [doi:10.1107/S1600536808042529]

## Diaqua(5-methylpyrazine-2-carboxylato- $\kappa^2N^1,O$ )iron(II)

Guang Fan, Jia-Juan Sun, Jun-Cai Zhang, Zhan-Ying Ma and Sheng-Li Gao

### S1. Comment

The background to this study is set out in the preceding paper (Fan *et al.*, 2007). Here we report the crystal structure of a mononuclear Fe<sup>II</sup> (Fig. 1).

The asymmetric unit consists of a Fe<sup>II</sup> atom, which lies on an inversion centre, one 2mpac ligand and two water molecules. A ring nitrogen atom and an oxygen atom of the carboxylate group from 2mpac ligand with Fe1—O1 = 2.103 (2) Å and Fe1—N1 = 2.167 (3) Å are involved in coordination to the Fe<sup>II</sup> atom; these form a square. The coordination of the two water molecules with Fe1—O3 = 2.114 (2) Å occupied the axial sites results in the formation of a distorted octahedral geometry.

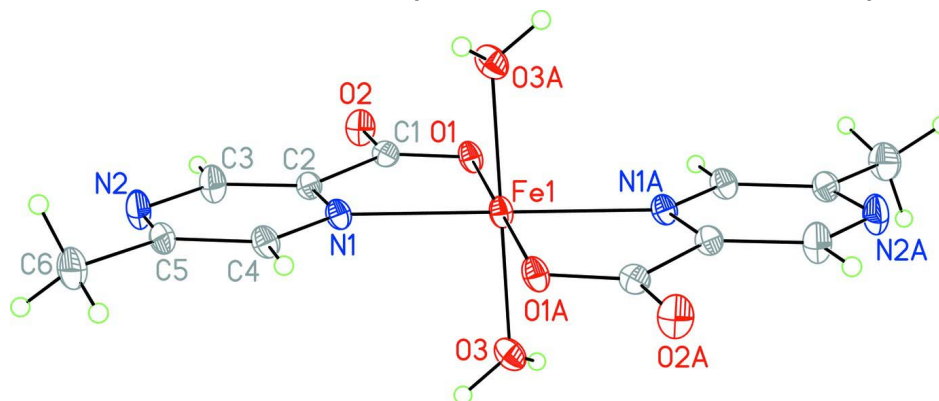
In the crystal structure, hydrogen bonding interactions are observed between the hydrogen atoms of the coordinated water molecules and the oxygen atoms of the carboxyl groups of a neighbouring unit, affording a two-dimensional supramolecular structure (Figure 2).

### S2. Experimental

The title compound was obtained from the mixture of ferrous ammonium sulfate hexahydrate(0.10 g, 0.25 mmol), 5-methylpyrazine-2-carboxylic acid (0.70 g, 0.5 mmol) and distilled water (20 ml), which was placed at room temperature for two weeks and red single crystals were obtained finally.

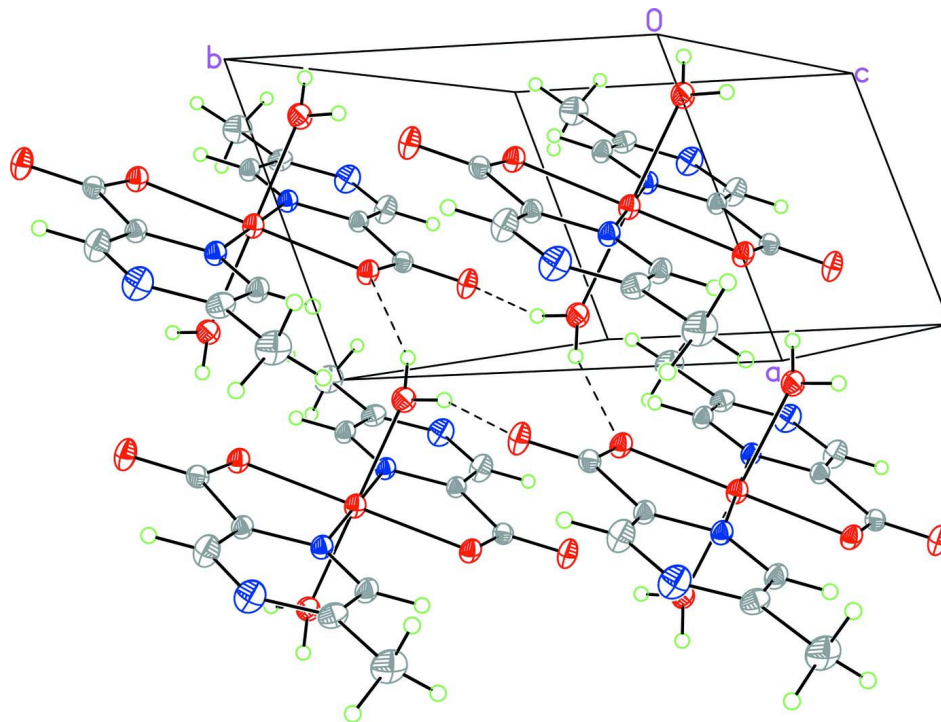
### S3. Refinement

All H atoms attached to C atoms from the organic ligands were generated in idealized positions and constrained to ride on their parent atoms, with  $d(C-H) = 0.93$  Å,  $U_{iso} = 1.2U_{eq}(C)$  for aromatic and 0.96 Å,  $U_{iso} = 1.5U_{eq}(C)$  for CH<sub>3</sub> atoms.



**Figure 1**

A view of the molecular structure of title complex with the atom-labelling scheme. Displacement ellipsoids are drawn at the 30% probability level. Symmetry code for A: 1-X, 1-Y, 1-Z.

**Figure 2**

Two-dimensional supramolecular structure of the title complex.

### Diaqua(5-methylpyrazine-2-carboxylato- $k^2N^1,O$ )iron(II)

#### Crystal data

$[\text{Fe}(\text{C}_6\text{H}_5\text{N}_2\text{O}_2)_2(\text{H}_2\text{O})_2]$

$M_r = 366.12$

Triclinic,  $P\bar{1}$

$a = 5.068$  (1) Å

$b = 6.401$  (1) Å

$c = 12.3810$  (12) Å

$\alpha = 103.851$  (2)°

$\beta = 91.079$  (1)°

$\gamma = 108.340$  (2)°

$V = 368.22$  (10) Å<sup>3</sup>

$Z = 1$

$F(000) = 188$

$D_x = 1.651$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 715 reflections

$\theta = 3.4\text{--}26.8^\circ$

$\mu = 1.06$  mm<sup>-1</sup>

$T = 298$  K

Block, red

$0.18 \times 0.09 \times 0.05$  mm

#### Data collection

Bruker SMART CCD area-detector  
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

$\varphi$  and  $\omega$  scans

Absorption correction: multi-scan

(*SADABS*; Sheldrick, 1996)

$T_{\min} = 0.832$ ,  $T_{\max} = 0.949$

1916 measured reflections

1260 independent reflections

1061 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.025$

$\theta_{\max} = 25.0^\circ$ ,  $\theta_{\min} = 1.7^\circ$

$h = -5 \rightarrow 6$

$k = -7 \rightarrow 7$

$l = -14 \rightarrow 12$

Refinement

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.041$   
 $wR(F^2) = 0.095$   
 $S = 1.00$   
 1260 reflections  
 107 parameters  
 0 restraints  
 Primary atom site location: structure-invariant  
 direct methods

Secondary atom site location: difference Fourier  
 map  
 Hydrogen site location: inferred from  
 neighbouring sites  
 H-atom parameters constrained  
 $w = 1/[\sigma^2(F_o^2) + (0.0456P)^2]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} < 0.001$   
 $\Delta\rho_{\max} = 0.33 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\min} = -0.28 \text{ e } \text{\AA}^{-3}$

Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) etc. and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
Fe1	0.5000	0.5000	0.5000	0.0326 (3)
N1	0.5923 (6)	0.4742 (4)	0.3281 (2)	0.0323 (7)
N2	0.6791 (7)	0.4921 (6)	0.1096 (2)	0.0508 (8)
O1	0.3162 (5)	0.7203 (4)	0.45178 (18)	0.0359 (6)
O2	0.2550 (5)	0.8729 (4)	0.3133 (2)	0.0472 (7)
O3	0.8721 (5)	0.7772 (4)	0.56033 (19)	0.0423 (6)
H3A	1.0121	0.7975	0.5226	0.051*
H3B	0.8701	0.9077	0.5961	0.051*
C1	0.3463 (7)	0.7489 (5)	0.3543 (3)	0.0330 (8)
C2	0.5049 (7)	0.6138 (5)	0.2826 (3)	0.0313 (8)
C3	0.5494 (8)	0.6191 (7)	0.1744 (3)	0.0482 (10)
H3	0.4856	0.7167	0.1445	0.058*
C4	0.7211 (7)	0.3461 (6)	0.2642 (3)	0.0360 (8)
H4	0.7855	0.2487	0.2940	0.043*
C5	0.7624 (7)	0.3530 (6)	0.1548 (3)	0.0394 (8)
C6	0.9003 (9)	0.2011 (7)	0.0829 (3)	0.0582 (11)
H6A	0.7600	0.0722	0.0347	0.087*
H6B	1.0059	0.1495	0.1297	0.087*
H6C	1.0232	0.2847	0.0384	0.087*

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Fe1	0.0405 (5)	0.0352 (4)	0.0282 (4)	0.0197 (3)	0.0122 (3)	0.0092 (3)
N1	0.0366 (16)	0.0321 (16)	0.0329 (15)	0.0170 (13)	0.0109 (12)	0.0089 (13)

N2	0.064 (2)	0.068 (2)	0.0329 (17)	0.0347 (19)	0.0192 (15)	0.0171 (17)
O1	0.0419 (14)	0.0395 (14)	0.0356 (14)	0.0238 (11)	0.0170 (10)	0.0118 (11)
O2	0.0649 (18)	0.0445 (16)	0.0469 (15)	0.0355 (14)	0.0114 (12)	0.0151 (13)
O3	0.0427 (15)	0.0346 (14)	0.0527 (16)	0.0186 (12)	0.0178 (11)	0.0082 (12)
C1	0.0330 (19)	0.0275 (18)	0.037 (2)	0.0099 (15)	0.0054 (14)	0.0044 (15)
C2	0.0337 (19)	0.0296 (18)	0.0322 (19)	0.0137 (15)	0.0061 (14)	0.0066 (15)
C3	0.063 (3)	0.059 (3)	0.038 (2)	0.035 (2)	0.0160 (18)	0.020 (2)
C4	0.039 (2)	0.037 (2)	0.039 (2)	0.0206 (16)	0.0122 (16)	0.0104 (16)
C5	0.038 (2)	0.047 (2)	0.0333 (19)	0.0198 (17)	0.0102 (15)	0.0036 (17)
C6	0.065 (3)	0.063 (3)	0.047 (2)	0.030 (2)	0.019 (2)	-0.001 (2)

*Geometric parameters (Å, °)*

Fe1—O1 <sup>i</sup>	2.103 (2)	O3—H3A	0.8500
Fe1—O1	2.103 (2)	O3—H3B	0.8499
Fe1—O3 <sup>i</sup>	2.114 (2)	C1—C2	1.510 (5)
Fe1—O3	2.114 (2)	C2—C3	1.369 (5)
Fe1—N1 <sup>i</sup>	2.167 (3)	C3—H3	0.9300
Fe1—N1	2.167 (3)	C4—C5	1.384 (5)
N1—C4	1.331 (4)	C4—H4	0.9300
N1—C2	1.339 (4)	C5—C6	1.505 (5)
N2—C5	1.323 (4)	C6—H6A	0.9600
N2—C3	1.335 (5)	C6—H6B	0.9600
O1—C1	1.268 (4)	C6—H6C	0.9600
O2—C1	1.232 (4)		
O1 <sup>i</sup> —Fe1—O1	180.00 (11)	H3A—O3—H3B	107.6
O1 <sup>i</sup> —Fe1—O3 <sup>i</sup>	89.75 (9)	O2—C1—O1	126.0 (3)
O1—Fe1—O3 <sup>i</sup>	90.25 (9)	O2—C1—C2	117.9 (3)
O1 <sup>i</sup> —Fe1—O3	90.25 (9)	O1—C1—C2	116.0 (3)
O1—Fe1—O3	89.75 (9)	N1—C2—C3	119.7 (3)
O3 <sup>i</sup> —Fe1—O3	180.00 (10)	N1—C2—C1	116.5 (3)
O1 <sup>i</sup> —Fe1—N1 <sup>i</sup>	77.21 (9)	C3—C2—C1	123.7 (3)
O1—Fe1—N1 <sup>i</sup>	102.79 (9)	N2—C3—C2	123.6 (3)
O3 <sup>i</sup> —Fe1—N1 <sup>i</sup>	92.02 (9)	N2—C3—H3	118.2
O3—Fe1—N1 <sup>i</sup>	87.98 (9)	C2—C3—H3	118.2
O1 <sup>i</sup> —Fe1—N1	102.79 (9)	N1—C4—C5	122.1 (3)
O1—Fe1—N1	77.21 (9)	N1—C4—H4	118.9
O3 <sup>i</sup> —Fe1—N1	87.98 (9)	C5—C4—H4	118.9
O3—Fe1—N1	92.02 (9)	N2—C5—C4	121.0 (3)
N1 <sup>i</sup> —Fe1—N1	180.000 (1)	N2—C5—C6	117.8 (3)
C4—N1—C2	117.2 (3)	C4—C5—C6	121.2 (3)
C4—N1—Fe1	130.3 (2)	C5—C6—H6A	109.5
C2—N1—Fe1	112.5 (2)	C5—C6—H6B	109.5
C5—N2—C3	116.3 (3)	H6A—C6—H6B	109.5
C1—O1—Fe1	117.6 (2)	C5—C6—H6C	109.5
Fe1—O3—H3A	121.1	H6A—C6—H6C	109.5
Fe1—O3—H3B	121.9	H6B—C6—H6C	109.5

O1 <sup>i</sup> —Fe1—N1—C4	2.7 (3)	C4—N1—C2—C3	0.2 (5)
O1—Fe1—N1—C4	-177.3 (3)	Fe1—N1—C2—C3	179.5 (3)
O3 <sup>i</sup> —Fe1—N1—C4	-86.6 (3)	C4—N1—C2—C1	177.0 (3)
O3—Fe1—N1—C4	93.4 (3)	Fe1—N1—C2—C1	-3.7 (3)
N1 <sup>i</sup> —Fe1—N1—C4	131 (100)	O2—C1—C2—N1	-177.5 (3)
O1 <sup>i</sup> —Fe1—N1—C2	-176.5 (2)	O1—C1—C2—N1	1.4 (4)
O1—Fe1—N1—C2	3.5 (2)	O2—C1—C2—C3	-0.9 (5)
O3 <sup>i</sup> —Fe1—N1—C2	94.2 (2)	O1—C1—C2—C3	178.0 (3)
O3—Fe1—N1—C2	-85.8 (2)	C5—N2—C3—C2	1.1 (6)
N1 <sup>i</sup> —Fe1—N1—C2	-48 (100)	N1—C2—C3—N2	-0.5 (6)
O1 <sup>i</sup> —Fe1—O1—C1	29 (100)	C1—C2—C3—N2	-177.0 (3)
O3 <sup>i</sup> —Fe1—O1—C1	-90.8 (2)	C2—N1—C4—C5	-0.6 (5)
O3—Fe1—O1—C1	89.2 (2)	Fe1—N1—C4—C5	-179.8 (2)
N1 <sup>i</sup> —Fe1—O1—C1	177.1 (2)	C3—N2—C5—C4	-1.5 (5)
N1—Fe1—O1—C1	-2.9 (2)	C3—N2—C5—C6	177.9 (3)
Fe1—O1—C1—O2	-179.3 (3)	N1—C4—C5—N2	1.3 (5)
Fe1—O1—C1—C2	1.9 (4)	N1—C4—C5—C6	-178.0 (3)

Symmetry code: (i)  $-x+1, -y+1, -z+1$ .

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
O3—H3A $\cdots$ O1 <sup>ii</sup>	0.85	1.93	2.720 (3)	155
O3—H3B $\cdots$ O2 <sup>iii</sup>	0.85	1.86	2.673 (3)	159

Symmetry codes: (ii)  $x+1, y, z$ ; (iii)  $-x+1, -y+2, -z+1$ .