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4-[(4-Amino-3-pyridyl)iminomethyl]-benzonitrile

Hoong-Kun Fun,^{a*} Hadi Kargar^{b‡} and Reza Kia^a

^aX-ray Crystallography Unit, School of Physics, Universiti Sains Malaysia, 11800 USM, Penang, Malaysia, and ^bDepartment of Chemistry, School of Science, Payame Noor University (PNU), Ardakan, Yazd, Iran

Correspondence e-mail: hkfun@usm.my

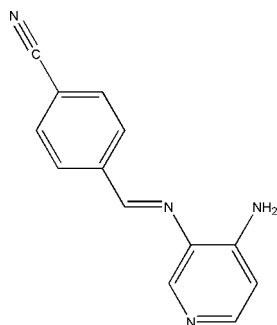
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Key indicators: single-crystal X-ray study; $T = 100$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; R factor = 0.074; wR factor = 0.219; data-to-parameter ratio = 11.8.

The asymmetric unit of the potential mono-Schiff base ligand title compound, $\text{C}_{13}\text{H}_{10}\text{N}_4$, contains two crystallographically independent molecules, *A* and *B*. In molecule *A*, the two rings are twisted from each other by 13.90 (18)°. By contrast, the dihedral angle between the two rings in molecule *B* is 0.67 (19)°. In the crystal structure, molecules are linked through intermolecular $\text{N}-\text{H}\cdots\text{N}$ interactions via $R_4^4(32)$ motifs, forming two-dimensional arrays. The short distances between the centroids of the six-membered rings indicate the existence of $\pi-\pi$ interactions [centroid-centroid distances = 3.6880 (17)- 3.7466 (15) Å].

Related literature

For details of hydrogen-bond motifs, see: Bernstein *et al.* (1995). For related structures, see: Li *et al.* (2005); Bomfim *et al.* (2005); Glidewell *et al.* (2005, 2006); Sun *et al.* (2004); Fun *et al.* (2008).



Experimental

Crystal data

$\text{C}_{13}\text{H}_{10}\text{N}_4$
 $M_r = 222.25$

Monoclinic, $P2_1/c$
 $a = 13.5560$ (8) Å

‡ Additional correspondence author, e-mail: hkargar@pnu.ac.ir.

$b = 12.3000$ (7) Å
 $c = 15.7514$ (8) Å
 $\beta = 124.651$ (2)°
 $V = 2160.5$ (2) Å³
 $Z = 8$

Mo $K\alpha$ radiation
 $\mu = 0.09$ mm⁻¹
 $T = 100.0$ (1) K
 $0.45 \times 0.09 \times 0.07$ mm

Data collection

Bruker SMART APEXII CCD
area-detector diffractometer
Absorption correction: multi-scan
(SADABS; Bruker, 2005)
 $T_{\min} = 0.962$, $T_{\max} = 0.994$

18251 measured reflections
3810 independent reflections
2372 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.060$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.074$
 $wR(F^2) = 0.219$
 $S = 1.03$
3810 reflections
323 parameters

H atoms treated by a mixture of
independent and constrained
refinement
 $\Delta\rho_{\text{max}} = 0.63$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.29$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{N3A}-\text{H2NA}\cdots\text{N4A}^i$	0.92 (4)	2.26 (4)	3.155 (4)	165 (3)
$\text{N3A}-\text{H1NA}\cdots\text{N1B}^{ii}$	0.89 (4)	2.33 (5)	3.080 (4)	143 (4)
$\text{N3B}-\text{H2NB}\cdots\text{N1A}^{iii}$	0.89 (4)	2.42 (4)	3.112 (4)	136 (4)
$\text{N3B}-\text{H1NB}\cdots\text{N4B}^i$	0.90 (4)	2.36 (4)	3.220 (4)	159 (2)

Symmetry codes: (i) $x, y - 1, z$; (ii) $-x + 2, -y, -z + 1$; (iii) $-x + 1, -y, -z + 1$.

Data collection: APEX2 (Bruker, 2005); cell refinement: APEX2; data reduction: SAINT (Bruker, 2005); program(s) used to solve structure: SIR2004 (Burla *et al.*, 2003); program(s) used to refine structure: SHELXTL (Sheldrick, 2008); molecular graphics: SHELXTL; software used to prepare material for publication: SHELXTL and PLATON (Spek, 2003).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: TK2327).

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supporting information

Acta Cryst. (2008). E64, o2342 [doi:10.1107/S1600536808037070]

4-[(4-Amino-3-pyridyl)iminomethyl]benzonitrile

Hoong-Kun Fun, Hadi Kargar and Reza Kia

S1. Comment

Schiff bases are one of most prevalent mixed-donor ligands in the field of coordination chemistry and play an important role in the development of coordination chemistry (Fun *et al.*, 2008). Structures of Schiff bases derived from substituted benzaldehydes have been reported previously (Li *et al.*, 2005; Bomfim *et al.*, 2005; Glidewell *et al.*, 2005, 2006; Sun *et al.*, 2004; Fun *et al.*, 2008).

Each imino (C=N) functional group is co-planar with the adjacent benzene ring in (I), Fig. 1. Two independent molecules, A and B, comprise the crystallographic asymmetric unit. In molecule A, the two phenyl rings are twisted from each other by 13.90 (18)°. The dihedral angle between the two phenyl rings in molecule B is 0.67 (19)° which indicates the molecule is planar. In the crystal structure, molecules are linked together through intermolecular N—H...N interactions via $R_4^4(32)$ motifs to form 2-D arrays parallel to the *ab*-plane, Fig. 2 & Table 1. The short distances between the centroids of the six-membered rings prove the existence of π - π interactions [Cg1...Cg4ⁱ = 3.7466 (15) Å, (i) *x*, 1/2 - *y*, -1/2 + *z*; Cg2...Cg3ⁱ = 3.6894 (14) Å; Cg3...Cg4ⁱⁱ, (ii) 1 - *x*, -*y*, 1 - *z*; Cg1, Cg2, Cg3, and Cg4 are the centroids of the N1A/C1A–C5A, C7A–C12A, N1B/C1B–C5B, and C7B–C12B rings, respectively.

S2. Experimental

The synthetic method has been described earlier (Fun *et al.*, 2008). Single crystals suitable for X-ray diffraction were obtained by evaporation of an ethanol solution of (I) held at room temperature.

S3. Refinement

The hydrogen atoms of the amino groups were located from the difference Fourier map and refined freely. The remaining hydrogen atoms were positioned geometrically and refined using a riding model with C—H = 0.93 Å, and with $U_{\text{iso}}(\text{H}) = 1.2 U_{\text{eq}}(\text{H})$.

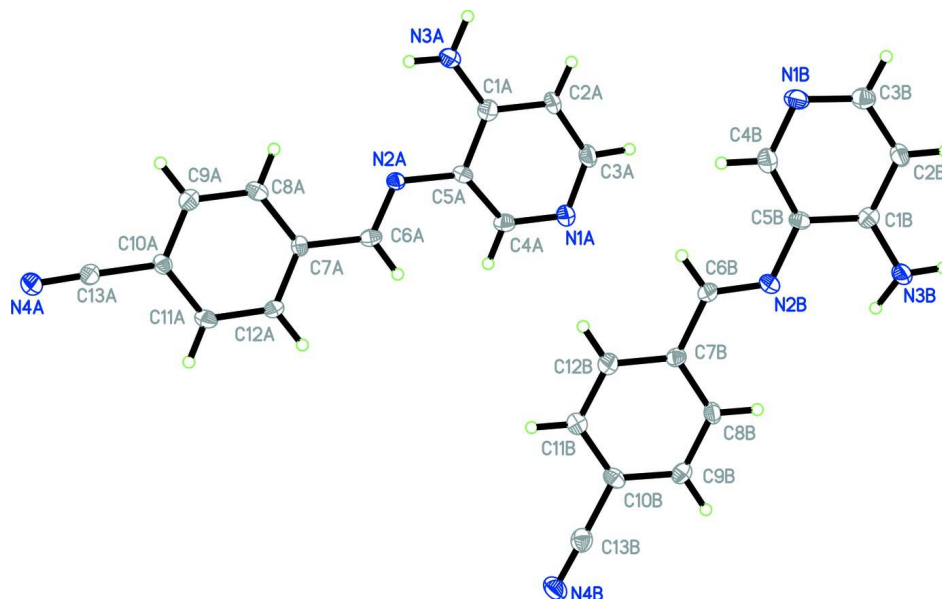


Figure 1

The molecular structure of (I) with atom labels and 50% displacement ellipsoids for non-H atoms.

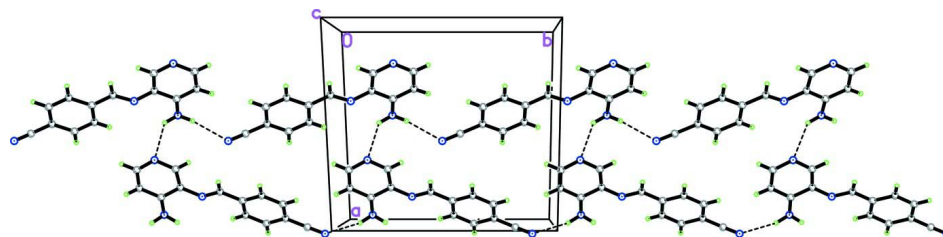


Figure 2

The crystal packing of (I), viewed down the *c*-axis showing a part of the 2-D array and $R_2^2(32)$ ring motifs. Intermolecular hydrogen bonds are shown as dashed lines.

4-[(4-Amino-3-pyridyl)iminomethyl]benzonitrile

Crystal data

$C_{13}H_{10}N_4$

$M_r = 222.25$

Monoclinic, $P2_1/c$

Hall symbol: -P 2ybc

$a = 13.5560$ (8) Å

$b = 12.3000$ (7) Å

$c = 15.7514$ (8) Å

$\beta = 124.651$ (2)°

$V = 2160.5$ (2) Å³

$Z = 8$

$F(000) = 928$

$D_x = 1.367$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 1699 reflections

$\theta = 2.3$ – 30.1 °

$\mu = 0.09$ mm⁻¹

$T = 100$ K

Needle, yellow

$0.45 \times 0.09 \times 0.07$ mm

Data collection

Bruker SMART APEXII CCD area-detector
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

φ and ω scans

Absorption correction: multi-scan

(*SADABS*; Bruker, 2005)

$T_{\min} = 0.962$, $T_{\max} = 0.994$

18251 measured reflections
 3810 independent reflections
 2372 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.060$

$\theta_{\text{max}} = 25.0^\circ$, $\theta_{\text{min}} = 1.8^\circ$
 $h = -16 \rightarrow 16$
 $k = -14 \rightarrow 13$
 $l = -18 \rightarrow 18$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.074$
 $wR(F^2) = 0.219$
 $S = 1.03$
 3810 reflections
 323 parameters
 0 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: inferred from
 neighbouring sites
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.1298P)^2 + 0.1507P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\text{max}} < 0.001$
 $\Delta\rho_{\text{max}} = 0.63 \text{ e } \text{Å}^{-3}$
 $\Delta\rho_{\text{min}} = -0.29 \text{ e } \text{Å}^{-3}$

Special details

Experimental. The low-temperature data was collected with the Oxford Cyrosystem Cobra low-temperature attachment.

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
N1A	0.6639 (2)	0.12555 (19)	0.38263 (17)	0.0217 (6)
N2A	0.85668 (19)	0.33048 (19)	0.36498 (16)	0.0148 (6)
N3A	0.9556 (2)	0.1387 (2)	0.3572 (2)	0.0277 (7)
N4A	1.0578 (2)	0.9061 (2)	0.36363 (17)	0.0227 (6)
C1A	0.8618 (2)	0.1349 (2)	0.36579 (19)	0.0185 (7)
C2A	0.8126 (2)	0.0361 (2)	0.36880 (19)	0.0200 (7)
H2AA	0.8447	-0.0292	0.3654	0.024*
C3A	0.7163 (2)	0.0359 (2)	0.3769 (2)	0.0210 (7)
H3AA	0.6856	-0.0312	0.3784	0.025*
C4A	0.7119 (2)	0.2198 (2)	0.37969 (19)	0.0184 (7)
H4AA	0.6773	0.2833	0.3834	0.022*
C5A	0.8084 (2)	0.2316 (2)	0.37159 (19)	0.0156 (6)
C6A	0.8305 (2)	0.4207 (2)	0.38827 (19)	0.0173 (7)
H6AA	0.7799	0.4207	0.4104	0.021*
C7A	0.8790 (2)	0.5238 (2)	0.38050 (18)	0.0140 (6)
C8A	0.9575 (2)	0.5272 (2)	0.34913 (18)	0.0189 (7)
H8AA	0.9781	0.4632	0.3313	0.023*
C9A	1.0037 (2)	0.6245 (2)	0.34472 (19)	0.0203 (7)
H9AA	1.0551	0.6263	0.3234	0.024*

C10A	0.9739 (2)	0.7212 (2)	0.37224 (19)	0.0156 (6)
C11A	0.8961 (2)	0.7180 (2)	0.40366 (19)	0.0185 (7)
H11A	0.8755	0.7819	0.4217	0.022*
C12A	0.8499 (2)	0.6198 (2)	0.40791 (19)	0.0168 (7)
H12A	0.7987	0.6179	0.4294	0.020*
C13A	1.0218 (2)	0.8236 (2)	0.36773 (19)	0.0173 (6)
C3B	0.7831 (2)	-0.3515 (2)	0.62159 (19)	0.0202 (7)
H3BA	0.8150	-0.4183	0.6211	0.024*
N2B	0.63747 (19)	-0.05867 (19)	0.62738 (16)	0.0183 (6)
N3B	0.5661 (2)	-0.2465 (2)	0.66987 (19)	0.0236 (6)
N4B	0.4350 (2)	0.5212 (2)	0.62229 (18)	0.0269 (6)
C1B	0.6482 (2)	-0.2527 (2)	0.64597 (19)	0.0160 (7)
C2B	0.6988 (2)	-0.3504 (2)	0.6444 (2)	0.0191 (7)
H2BA	0.6757	-0.4152	0.6588	0.023*
N1B	0.8223 (2)	-0.26157 (19)	0.59977 (17)	0.0225 (6)
C4B	0.7739 (3)	-0.1680 (2)	0.6022 (2)	0.0232 (7)
H4BA	0.7994	-0.1047	0.5878	0.028*
C5B	0.6885 (2)	-0.1573 (2)	0.62439 (19)	0.0153 (6)
C6B	0.6661 (2)	0.0318 (2)	0.6087 (2)	0.0207 (7)
H6BA	0.7215	0.0325	0.5916	0.025*
C7B	0.6158 (2)	0.1352 (2)	0.61273 (19)	0.0156 (6)
C8B	0.5309 (2)	0.1391 (2)	0.63648 (19)	0.0177 (7)
H8BA	0.5058	0.0751	0.6501	0.021*
C9B	0.4837 (3)	0.2381 (2)	0.6398 (2)	0.0197 (7)
H9BA	0.4278	0.2407	0.6562	0.024*
C10B	0.5212 (2)	0.3334 (2)	0.6183 (2)	0.0183 (7)
C11B	0.6066 (3)	0.3310 (2)	0.5956 (2)	0.0219 (7)
H11B	0.6321	0.3950	0.5823	0.026*
C12B	0.6525 (3)	0.2323 (2)	0.5929 (2)	0.0242 (7)
H12B	0.7094	0.2303	0.5776	0.029*
C13B	0.4722 (3)	0.4381 (3)	0.6207 (2)	0.0222 (7)
H2NA	0.999 (3)	0.077 (3)	0.366 (2)	0.028 (9)*
H1NA	0.996 (3)	0.199 (3)	0.365 (3)	0.048 (12)*
H2NB	0.527 (3)	-0.184 (3)	0.654 (2)	0.044 (11)*
H1NB	0.538 (3)	-0.311 (3)	0.675 (2)	0.039 (10)*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
N1A	0.0241 (14)	0.0159 (14)	0.0297 (13)	-0.0010 (11)	0.0180 (12)	0.0011 (11)
N2A	0.0128 (12)	0.0135 (14)	0.0182 (11)	0.0012 (10)	0.0089 (10)	0.0005 (9)
N3A	0.0293 (16)	0.0154 (17)	0.0494 (17)	0.0009 (13)	0.0290 (14)	-0.0006 (13)
N4A	0.0240 (14)	0.0166 (15)	0.0305 (13)	0.0020 (11)	0.0173 (12)	0.0021 (11)
C1A	0.0197 (15)	0.0196 (18)	0.0157 (13)	-0.0026 (12)	0.0097 (12)	-0.0022 (12)
C2A	0.0239 (16)	0.0113 (17)	0.0263 (15)	-0.0020 (13)	0.0152 (13)	0.0005 (12)
C3A	0.0225 (16)	0.0148 (17)	0.0260 (15)	-0.0034 (13)	0.0140 (13)	0.0002 (12)
C4A	0.0154 (15)	0.0197 (18)	0.0203 (14)	0.0019 (12)	0.0103 (13)	-0.0010 (12)
C5A	0.0153 (14)	0.0134 (16)	0.0165 (13)	0.0020 (12)	0.0082 (12)	-0.0009 (11)

C6A	0.0141 (14)	0.0202 (18)	0.0179 (13)	0.0025 (12)	0.0094 (12)	-0.0004 (12)
C7A	0.0119 (13)	0.0140 (16)	0.0130 (12)	-0.0016 (11)	0.0053 (11)	0.0002 (11)
C8A	0.0216 (15)	0.0155 (17)	0.0191 (14)	0.0006 (12)	0.0113 (13)	-0.0034 (12)
C9A	0.0213 (16)	0.0210 (18)	0.0226 (14)	-0.0014 (13)	0.0149 (13)	-0.0006 (12)
C10A	0.0147 (14)	0.0136 (16)	0.0161 (13)	0.0008 (12)	0.0073 (12)	0.0008 (11)
C11A	0.0171 (15)	0.0163 (17)	0.0209 (14)	0.0033 (12)	0.0101 (13)	-0.0015 (12)
C12A	0.0151 (14)	0.0175 (17)	0.0209 (14)	0.0008 (12)	0.0121 (12)	-0.0004 (12)
C13A	0.0169 (15)	0.0178 (17)	0.0178 (14)	0.0000 (13)	0.0102 (12)	0.0006 (12)
C3B	0.0184 (15)	0.0199 (18)	0.0216 (15)	-0.0007 (12)	0.0109 (13)	-0.0029 (12)
N2B	0.0195 (13)	0.0145 (15)	0.0220 (12)	0.0044 (10)	0.0124 (11)	0.0027 (10)
N3B	0.0266 (15)	0.0137 (17)	0.0392 (15)	0.0009 (13)	0.0239 (13)	0.0023 (12)
N4B	0.0262 (14)	0.0168 (16)	0.0405 (15)	0.0033 (12)	0.0206 (13)	0.0012 (12)
C1B	0.0094 (13)	0.0216 (18)	0.0130 (13)	-0.0003 (12)	0.0040 (11)	0.0003 (11)
C2B	0.0182 (15)	0.0154 (17)	0.0218 (14)	0.0008 (12)	0.0103 (13)	0.0033 (12)
N1B	0.0213 (14)	0.0181 (15)	0.0310 (13)	0.0015 (11)	0.0166 (12)	-0.0030 (11)
C4B	0.0260 (16)	0.0171 (18)	0.0287 (15)	-0.0014 (13)	0.0169 (14)	0.0023 (13)
C5B	0.0126 (14)	0.0164 (17)	0.0142 (13)	0.0026 (11)	0.0060 (11)	-0.0008 (11)
C6B	0.0220 (15)	0.0198 (18)	0.0285 (15)	0.0034 (13)	0.0193 (13)	0.0038 (13)
C7B	0.0137 (14)	0.0169 (17)	0.0163 (13)	-0.0002 (12)	0.0087 (12)	-0.0007 (11)
C8B	0.0206 (16)	0.0141 (17)	0.0195 (14)	-0.0015 (12)	0.0120 (13)	0.0015 (11)
C9B	0.0192 (15)	0.0211 (18)	0.0238 (14)	-0.0015 (13)	0.0152 (13)	0.0001 (12)
C10B	0.0188 (15)	0.0151 (17)	0.0168 (13)	0.0004 (12)	0.0077 (12)	-0.0021 (11)
C11B	0.0264 (16)	0.0172 (18)	0.0264 (15)	-0.0005 (13)	0.0175 (14)	-0.0001 (12)
C12B	0.0269 (17)	0.0200 (18)	0.0345 (17)	0.0018 (13)	0.0227 (15)	0.0034 (13)
C13B	0.0192 (15)	0.0229 (19)	0.0246 (15)	-0.0025 (14)	0.0124 (13)	0.0000 (13)

Geometric parameters (Å, °)

N1A—C3A	1.343 (3)	C3B—N1B	1.354 (3)
N1A—C4A	1.343 (3)	C3B—C2B	1.379 (4)
N2A—C6A	1.281 (3)	C3B—H3BA	0.9300
N2A—C5A	1.413 (3)	N2B—C6B	1.267 (3)
N3A—C1A	1.354 (4)	N2B—C5B	1.411 (3)
N3A—H2NA	0.93 (3)	N3B—C1B	1.366 (4)
N3A—H1NA	0.88 (4)	N3B—H2NB	0.88 (4)
N4A—C13A	1.143 (3)	N3B—H1NB	0.90 (4)
C1A—C2A	1.399 (4)	N4B—C13B	1.147 (4)
C1A—C5A	1.422 (4)	C1B—C2B	1.391 (4)
C2A—C3A	1.381 (4)	C1B—C5B	1.415 (4)
C2A—H2AA	0.9300	C2B—H2BA	0.9300
C3A—H3AA	0.9300	N1B—C4B	1.336 (4)
C4A—C5A	1.392 (4)	C4B—C5B	1.393 (4)
C4A—H4AA	0.9300	C4B—H4BA	0.9300
C6A—C7A	1.466 (4)	C6B—C7B	1.461 (4)
C6A—H6AA	0.9300	C6B—H6BA	0.9300
C7A—C12A	1.389 (4)	C7B—C12B	1.396 (4)
C7A—C8A	1.406 (4)	C7B—C8B	1.399 (4)
C8A—C9A	1.370 (4)	C8B—C9B	1.390 (4)

C8A—H8AA	0.9300	C8B—H8BA	0.9300
C9A—C10A	1.402 (4)	C9B—C10B	1.394 (4)
C9A—H9AA	0.9300	C9B—H9BA	0.9300
C10A—C11A	1.398 (4)	C10B—C11B	1.393 (4)
C10A—C13A	1.437 (4)	C10B—C13B	1.458 (4)
C11A—C12A	1.379 (4)	C11B—C12B	1.375 (4)
C11A—H11A	0.9300	C11B—H11B	0.9300
C12A—H12A	0.9300	C12B—H12B	0.9300
C3A—N1A—C4A	114.9 (2)	N1B—C3B—C2B	124.1 (3)
C6A—N2A—C5A	120.6 (2)	N1B—C3B—H3BA	117.9
C1A—N3A—H2NA	121.2 (19)	C2B—C3B—H3BA	117.9
C1A—N3A—H1NA	123 (2)	C6B—N2B—C5B	121.8 (2)
H2NA—N3A—H1NA	112 (3)	C1B—N3B—H2NB	115 (2)
N3A—C1A—C2A	121.7 (3)	C1B—N3B—H1NB	115 (2)
N3A—C1A—C5A	121.2 (3)	H2NB—N3B—H1NB	125 (3)
C2A—C1A—C5A	117.0 (3)	N3B—C1B—C2B	122.7 (3)
C3A—C2A—C1A	119.9 (3)	N3B—C1B—C5B	120.4 (3)
C3A—C2A—H2AA	120.0	C2B—C1B—C5B	116.8 (3)
C1A—C2A—H2AA	120.0	C3B—C2B—C1B	120.2 (3)
N1A—C3A—C2A	124.6 (3)	C3B—C2B—H2BA	119.9
N1A—C3A—H3AA	117.7	C1B—C2B—H2BA	119.9
C2A—C3A—H3AA	117.7	C4B—N1B—C3B	115.3 (3)
N1A—C4A—C5A	126.3 (3)	N1B—C4B—C5B	125.4 (3)
N1A—C4A—H4AA	116.9	N1B—C4B—H4BA	117.3
C5A—C4A—H4AA	116.9	C5B—C4B—H4BA	117.3
C4A—C5A—N2A	126.5 (3)	C4B—C5B—N2B	125.6 (3)
C4A—C5A—C1A	117.2 (3)	C4B—C5B—C1B	118.2 (3)
N2A—C5A—C1A	116.2 (2)	N2B—C5B—C1B	116.2 (2)
N2A—C6A—C7A	121.0 (2)	N2B—C6B—C7B	122.8 (3)
N2A—C6A—H6AA	119.5	N2B—C6B—H6BA	118.6
C7A—C6A—H6AA	119.5	C7B—C6B—H6BA	118.6
C12A—C7A—C8A	119.3 (3)	C12B—C7B—C8B	118.8 (3)
C12A—C7A—C6A	119.3 (2)	C12B—C7B—C6B	120.0 (2)
C8A—C7A—C6A	121.3 (3)	C8B—C7B—C6B	121.2 (3)
C9A—C8A—C7A	120.2 (3)	C9B—C8B—C7B	120.4 (3)
C9A—C8A—H8AA	119.9	C9B—C8B—H8BA	119.8
C7A—C8A—H8AA	119.9	C7B—C8B—H8BA	119.8
C8A—C9A—C10A	120.3 (3)	C8B—C9B—C10B	119.2 (3)
C8A—C9A—H9AA	119.8	C8B—C9B—H9BA	120.4
C10A—C9A—H9AA	119.8	C10B—C9B—H9BA	120.4
C11A—C10A—C9A	119.6 (3)	C11B—C10B—C9B	121.0 (3)
C11A—C10A—C13A	119.7 (3)	C11B—C10B—C13B	118.9 (3)
C9A—C10A—C13A	120.7 (3)	C9B—C10B—C13B	120.1 (3)
C12A—C11A—C10A	119.8 (3)	C12B—C11B—C10B	118.9 (3)
C12A—C11A—H11A	120.1	C12B—C11B—H11B	120.5
C10A—C11A—H11A	120.1	C10B—C11B—H11B	120.5
C11A—C12A—C7A	120.8 (3)	C11B—C12B—C7B	121.6 (3)

C11A—C12A—H12A	119.6	C11B—C12B—H12B	119.2
C7A—C12A—H12A	119.6	C7B—C12B—H12B	119.2
N4A—C13A—C10A	178.6 (3)	N4B—C13B—C10B	178.8 (3)
N3A—C1A—C2A—C3A	-179.5 (3)	N1B—C3B—C2B—C1B	0.1 (4)
C5A—C1A—C2A—C3A	0.1 (4)	N3B—C1B—C2B—C3B	-178.8 (2)
C4A—N1A—C3A—C2A	0.2 (4)	C5B—C1B—C2B—C3B	-0.7 (4)
C1A—C2A—C3A—N1A	-0.3 (4)	C2B—C3B—N1B—C4B	0.4 (4)
C3A—N1A—C4A—C5A	-0.1 (4)	C3B—N1B—C4B—C5B	-0.3 (4)
N1A—C4A—C5A—N2A	177.7 (2)	N1B—C4B—C5B—N2B	179.8 (2)
N1A—C4A—C5A—C1A	0.0 (4)	N1B—C4B—C5B—C1B	-0.3 (4)
C6A—N2A—C5A—C4A	15.3 (4)	C6B—N2B—C5B—C4B	0.4 (4)
C6A—N2A—C5A—C1A	-167.0 (2)	C6B—N2B—C5B—C1B	-179.5 (2)
N3A—C1A—C5A—C4A	179.6 (2)	N3B—C1B—C5B—C4B	178.9 (2)
C2A—C1A—C5A—C4A	-0.1 (3)	C2B—C1B—C5B—C4B	0.7 (4)
N3A—C1A—C5A—N2A	1.7 (4)	N3B—C1B—C5B—N2B	-1.1 (4)
C2A—C1A—C5A—N2A	-178.0 (2)	C2B—C1B—C5B—N2B	-179.3 (2)
C5A—N2A—C6A—C7A	-179.3 (2)	C5B—N2B—C6B—C7B	-179.1 (2)
N2A—C6A—C7A—C12A	-179.0 (2)	N2B—C6B—C7B—C12B	178.9 (3)
N2A—C6A—C7A—C8A	-1.2 (4)	N2B—C6B—C7B—C8B	-0.9 (4)
C12A—C7A—C8A—C9A	-0.7 (4)	C12B—C7B—C8B—C9B	0.3 (4)
C6A—C7A—C8A—C9A	-178.5 (2)	C6B—C7B—C8B—C9B	-179.9 (2)
C7A—C8A—C9A—C10A	0.5 (4)	C7B—C8B—C9B—C10B	0.6 (4)
C8A—C9A—C10A—C11A	-0.4 (4)	C8B—C9B—C10B—C11B	-1.2 (4)
C8A—C9A—C10A—C13A	-179.9 (2)	C8B—C9B—C10B—C13B	179.5 (2)
C9A—C10A—C11A—C12A	0.4 (4)	C9B—C10B—C11B—C12B	1.0 (4)
C13A—C10A—C11A—C12A	179.9 (2)	C13B—C10B—C11B—C12B	-179.7 (2)
C10A—C11A—C12A—C7A	-0.5 (4)	C10B—C11B—C12B—C7B	-0.1 (4)
C8A—C7A—C12A—C11A	0.7 (4)	C8B—C7B—C12B—C11B	-0.5 (4)
C6A—C7A—C12A—C11A	178.6 (2)	C6B—C7B—C12B—C11B	179.7 (3)

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N3A—H2NA \cdots N4A ⁱ	0.92 (4)	2.26 (4)	3.155 (4)	165 (3)
N3A—H1NA \cdots N1B ⁱⁱ	0.89 (4)	2.33 (5)	3.080 (4)	143 (4)
N3B—H2NB \cdots N1A ⁱⁱⁱ	0.89 (4)	2.42 (4)	3.112 (4)	136 (4)
N3B—H1NB \cdots N4B ⁱ	0.90 (4)	2.36 (4)	3.220 (4)	159 (2)

Symmetry codes: (i) $x, y-1, z$; (ii) $-x+2, -y, -z+1$; (iii) $-x+1, -y, -z+1$.