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4-(4-Nitrobenzenesulfonamido)-pyridinium trichloroacetate

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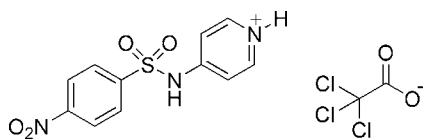
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Key indicators: single-crystal X-ray study; $T = 294$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; R factor = 0.039; wR factor = 0.094; data-to-parameter ratio = 14.6.

In the title compound, $\text{C}_{11}\text{H}_{10}\text{N}_3\text{O}_4\text{S}^+\cdot\text{C}_2\text{Cl}_3\text{O}_2^-$, the benzene ring forms an angle of $85.21(13)^\circ$ with the pyridinium ring. The nitro group is nearly coplanar with its attached benzene ring [dihedral angle = $3.68(12)^\circ$]. In the crystal structure, strong $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds link the ion-pairs. The packing is further consolidated by weak $\text{C}-\text{H}\cdots\text{O}$ interactions.

Related literature

For the synthesis and structure of the unprotonated amine, see: Yu & Li (2007). For reference geometrical data, see: Allen *et al.* (1987).



Experimental

Crystal data

$\text{C}_{11}\text{H}_{10}\text{N}_3\text{O}_4\text{S}^+\cdot\text{C}_2\text{Cl}_3\text{O}_2^-$
 $M_r = 442.65$
 Monoclinic, $P2_1/n$

$a = 5.9929(11)$ Å
 $b = 17.790(3)$ Å
 $c = 16.325(3)$ Å

$\beta = 98.377(3)^\circ$
 $V = 1721.9(6)$ Å³
 $Z = 4$
 Mo $K\alpha$ radiation

$\mu = 0.69$ mm⁻¹
 $T = 294(2)$ K
 $0.24 \times 0.22 \times 0.18$ mm

Data collection

Bruker SMART 1K CCD diffractometer
 Absorption correction: multi-scan (SADABS; Sheldrick, 1996)
 $T_{\min} = 0.852$, $T_{\max} = 0.886$

9742 measured reflections
 3528 independent reflections
 2427 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.038$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.039$
 $wR(F^2) = 0.094$
 $S = 1.01$
 3528 reflections
 241 parameters

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\max} = 0.37$ e Å⁻³
 $\Delta\rho_{\min} = -0.38$ e Å⁻³

Table 1
 Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N1}-\text{H1}\cdots\text{O6}$	0.87 (3)	1.93 (3)	2.803 (3)	172 (3)
$\text{N2}-\text{H2A}\cdots\text{O6}^i$	0.79 (3)	2.00 (3)	2.785 (3)	171 (3)
$\text{C2}-\text{H2}\cdots\text{O5}^i$	0.93	2.53	3.240 (3)	133
$\text{C3}-\text{H3}\cdots\text{O1}^{ii}$	0.93	2.50	3.391 (4)	162
$\text{C8}-\text{H8}\cdots\text{O5}^{iii}$	0.93	2.47	3.149 (3)	130
$\text{C10}-\text{H10}\cdots\text{O3}^{iv}$	0.93	2.50	3.308 (4)	146
$\text{C5}-\text{H5}\cdots\text{O2}$	0.93	2.33	2.978 (4)	126
$\text{C7}-\text{H7}\cdots\text{O2}$	0.93	2.57	2.938 (3)	104

Symmetry codes: (i) $x + \frac{1}{2}, -y + \frac{1}{2}, z + \frac{1}{2}$; (ii) $x + \frac{1}{2}, -y + \frac{1}{2}, z - \frac{1}{2}$; (iii) $-x, -y + 1, -z + 1$; (iv) $-x + 2, -y + 1, -z + 2$.

Data collection: SMART (Bruker, 1997); cell refinement: SAINT (Bruker, 1997); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 1997); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997); molecular graphics: SHELXTL (Bruker, 1997); software used to prepare material for publication: SHELXTL.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB2642).

References

- Allen, F. H., Kennard, O., Watson, D. G., Brammer, L., Orpen, A. G. & Taylor, R. (1987). *J. Chem. Soc. Perkin Trans. 2*, pp. S1–19.
 Bruker (1997). SMART, SAINT and SHELXTL. Bruker AXS Inc., Madison, Wisconsin, USA.
 Sheldrick, G. M. (1996). SADABS. University of Göttingen, Germany.
 Sheldrick, G. M. (1997). SHELXS97 and SHELXL97. University of Göttingen, Germany.
 Yu, H.-J. & Li, J.-S. (2007). *Acta Cryst.* **E63**, o3399.

supporting information

Acta Cryst. (2008). E64, o254 [https://doi.org/10.1107/S160053680706504X]

4-(4-Nitrobenzenesulfonamido)pyridinium trichloroacetate

Bin Zhou, Peng-Wu Zheng, Ke-Yue Liu and Dan Cheng

S1. Comment

The title compound, (I), comprises of a pyridinium cation and a trichloroacetate anion (Fig. 1). In the cation, the short C—N distance [N2—C1 = 1.390 (3) Å] occurs between typical C=N (1.34–1.38 Å) and C—N (1.47–1.50 Å) bond lengths (Allen *et al.*, 1987), indicative of significant double-bond character, despite of a strong electron-withdrawing sulfonyl group. The benzene ring forms an angle of 85.21 (13)° with the pyridinium ring. The nitro group is nearly coplanar and make an acute angle of 3.68 (12)° with the connected benzene ring.

The cation and anion are connected by a strong N—H···O hydrogen bond and weak C—H···O interactions (Table 1) complete the structure. Two short intramolecular C—H···O contacts also arise in the cation.

S2. Experimental

4-Nitro-(*N*-pyridyl)benzenesulfonamide was prepared by the method of Yu & Li (2007). Colourless blocks of (I) were grown by natural evaporation from a trichloroacetic acid solution of the amide.

S3. Refinement

The N-bound H atoms were located in a difference map and their positions were freely refined. The C-bound H atoms were positioned geometrically (C—H = 0.93 Å) and refined as riding atoms. The constraint $U_{iso}(H) = 1.2 U_{eq}(C \text{ and } N)$ was applied.

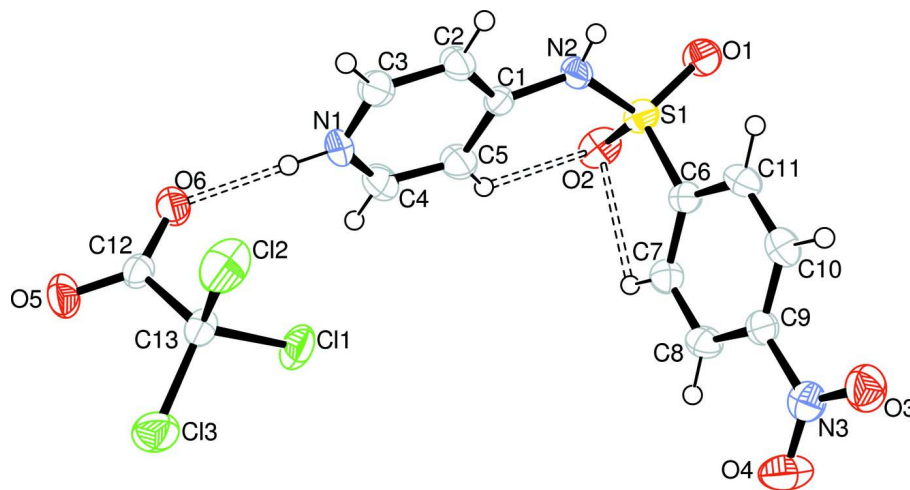


Figure 1

A view of (I) with displacement ellipsoids drawn at the 50% probability level (arbitrary spheres for the H atoms). Hydrogen bonds are indicated by double-dashed lines.

4-(4-Nitrobenzenesulfonamido)pyridinium trichloroacetate

Crystal data

 $C_{11}H_{10}N_3O_4S^+ \cdot C_2Cl_3O_2^-$ $M_r = 442.65$ Monoclinic, $P2_1/n$ Hall symbol: $-P\ 2_1n$ $a = 5.9929$ (11) Å $b = 17.790$ (3) Å $c = 16.325$ (3) Å $\beta = 98.377$ (3)° $V = 1721.9$ (6) Å³ $Z = 4$ $F(000) = 896$ $D_x = 1.707$ Mg m⁻³Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 2690 reflections

 $\theta = 2.5$ – 25.7 ° $\mu = 0.69$ mm⁻¹ $T = 294$ K

Block, colourless

 $0.24 \times 0.22 \times 0.18$ mm

Data collection

Bruker SMART 1K CCD

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

 ω scans

Absorption correction: multi-scan

(SADABS; Sheldrick, 1996)

 $T_{\min} = 0.852$, $T_{\max} = 0.886$

9742 measured reflections

3528 independent reflections

2427 reflections with $I > 2\sigma(I)$ $R_{\text{int}} = 0.038$ $\theta_{\max} = 26.4$ °, $\theta_{\min} = 1.7$ ° $h = -7 \rightarrow 7$ $k = -9 \rightarrow 22$ $l = -20 \rightarrow 20$

Refinement

Refinement on F^2

Least-squares matrix: full

 $R[F^2 > 2\sigma(F^2)] = 0.039$ $wR(F^2) = 0.094$ $S = 1.01$

3528 reflections

241 parameters

0 restraints

Primary atom site location: structure-invariant

direct methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H atoms treated by a mixture of independent and constrained refinement

 $w = 1/[\sigma^2(F_o^2) + (0.0353P)^2 + 0.9782P]$ where $P = (F_o^2 + 2F_c^2)/3$ $(\Delta/\sigma)_{\max} = 0.001$ $\Delta\rho_{\max} = 0.37$ e Å⁻³ $\Delta\rho_{\min} = -0.38$ e Å⁻³

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.20681 (12)	0.30768 (4)	0.84918 (4)	0.03291 (18)
O1	0.2730 (4)	0.26231 (11)	0.92092 (11)	0.0449 (5)
O2	-0.0263 (3)	0.32167 (11)	0.82059 (13)	0.0448 (5)

O3	0.8738 (4)	0.60031 (12)	0.95523 (13)	0.0495 (5)
O4	0.5848 (4)	0.65976 (12)	0.89304 (15)	0.0575 (6)
N1	0.2323 (4)	0.30604 (14)	0.52313 (15)	0.0413 (6)
H1	0.215 (5)	0.3155 (17)	0.470 (2)	0.050*
N2	0.3239 (4)	0.26681 (13)	0.77572 (13)	0.0321 (5)
H2A	0.422 (5)	0.2386 (17)	0.7939 (18)	0.039*
N3	0.6823 (4)	0.60178 (14)	0.91655 (14)	0.0372 (6)
C1	0.2886 (4)	0.28291 (14)	0.69149 (15)	0.0270 (6)
C2	0.4432 (5)	0.25395 (16)	0.64290 (16)	0.0358 (7)
H2	0.5671	0.2263	0.6673	0.043*
C3	0.4103 (5)	0.26665 (16)	0.55900 (17)	0.0416 (7)
H3	0.5130	0.2476	0.5267	0.050*
C4	0.0846 (5)	0.33458 (17)	0.56820 (18)	0.0430 (7)
H4	-0.0383	0.3615	0.5418	0.052*
C5	0.1084 (5)	0.32541 (16)	0.65210 (17)	0.0386 (7)
H5	0.0059	0.3472	0.6826	0.046*
C6	0.3464 (4)	0.39542 (14)	0.86639 (15)	0.0284 (6)
C7	0.2334 (4)	0.46213 (15)	0.84301 (16)	0.0320 (6)
H7	0.0850	0.4610	0.8166	0.038*
C8	0.3429 (4)	0.53001 (15)	0.85921 (16)	0.0337 (6)
H8	0.2698	0.5751	0.8443	0.040*
C9	0.5640 (4)	0.52928 (15)	0.89817 (15)	0.0299 (6)
C10	0.6797 (5)	0.46378 (16)	0.92142 (17)	0.0347 (7)
H10	0.8285	0.4653	0.9474	0.042*
C11	0.5696 (4)	0.39591 (15)	0.90528 (16)	0.0334 (6)
H11	0.6438	0.3510	0.9202	0.040*
O5	0.1076 (4)	0.40196 (11)	0.23515 (11)	0.0457 (5)
O6	0.1360 (3)	0.33610 (11)	0.35309 (11)	0.0412 (5)
C11	0.23049 (13)	0.47867 (4)	0.44855 (4)	0.0447 (2)
C12	0.60598 (13)	0.41336 (5)	0.37668 (6)	0.0579 (3)
C13	0.34546 (14)	0.53626 (4)	0.29679 (5)	0.0486 (2)
C12	0.1730 (4)	0.39111 (15)	0.30852 (16)	0.0293 (6)
C13	0.3314 (4)	0.45299 (15)	0.35494 (16)	0.0305 (6)

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
S1	0.0421 (4)	0.0312 (4)	0.0291 (4)	-0.0018 (3)	0.0176 (3)	-0.0008 (3)
O1	0.0724 (15)	0.0374 (12)	0.0295 (10)	-0.0025 (10)	0.0229 (10)	0.0056 (9)
O2	0.0358 (11)	0.0469 (13)	0.0559 (13)	-0.0052 (10)	0.0205 (10)	-0.0086 (10)
O3	0.0457 (13)	0.0540 (14)	0.0462 (12)	-0.0078 (11)	-0.0020 (10)	0.0001 (11)
O4	0.0529 (14)	0.0316 (12)	0.0862 (18)	0.0019 (11)	0.0046 (13)	0.0025 (12)
N1	0.0639 (18)	0.0385 (14)	0.0203 (11)	0.0062 (13)	0.0024 (12)	0.0031 (11)
N2	0.0416 (14)	0.0330 (14)	0.0230 (11)	0.0110 (11)	0.0088 (10)	0.0022 (10)
N3	0.0369 (14)	0.0409 (15)	0.0351 (13)	-0.0002 (12)	0.0096 (11)	-0.0020 (11)
C1	0.0348 (15)	0.0230 (13)	0.0241 (13)	-0.0013 (11)	0.0072 (11)	-0.0008 (11)
C2	0.0434 (17)	0.0366 (16)	0.0288 (15)	0.0147 (13)	0.0099 (12)	0.0046 (13)
C3	0.060 (2)	0.0372 (17)	0.0307 (16)	0.0131 (15)	0.0165 (15)	0.0015 (13)

C4	0.0457 (18)	0.0436 (18)	0.0369 (17)	0.0128 (15)	-0.0032 (14)	0.0050 (14)
C5	0.0423 (17)	0.0390 (17)	0.0360 (16)	0.0128 (14)	0.0108 (13)	0.0019 (13)
C6	0.0334 (15)	0.0285 (15)	0.0259 (13)	0.0022 (12)	0.0134 (11)	0.0004 (11)
C7	0.0288 (15)	0.0344 (16)	0.0329 (15)	0.0020 (12)	0.0045 (12)	-0.0002 (12)
C8	0.0333 (15)	0.0310 (16)	0.0374 (16)	0.0080 (12)	0.0069 (12)	0.0031 (12)
C9	0.0342 (15)	0.0316 (15)	0.0254 (13)	-0.0006 (12)	0.0089 (12)	0.0008 (11)
C10	0.0301 (15)	0.0429 (18)	0.0312 (15)	0.0028 (13)	0.0046 (12)	0.0045 (13)
C11	0.0339 (15)	0.0328 (16)	0.0344 (15)	0.0089 (13)	0.0085 (12)	0.0084 (12)
O5	0.0601 (14)	0.0476 (13)	0.0255 (10)	-0.0095 (10)	-0.0068 (9)	0.0008 (9)
O6	0.0569 (13)	0.0356 (11)	0.0289 (10)	-0.0117 (10)	-0.0014 (9)	0.0025 (9)
C11	0.0561 (5)	0.0464 (4)	0.0335 (4)	-0.0041 (4)	0.0126 (3)	-0.0148 (3)
C12	0.0328 (4)	0.0677 (6)	0.0683 (6)	0.0159 (4)	-0.0085 (4)	-0.0226 (5)
C13	0.0610 (5)	0.0360 (4)	0.0514 (5)	-0.0091 (4)	0.0168 (4)	0.0002 (3)
C12	0.0289 (14)	0.0295 (15)	0.0279 (14)	0.0011 (12)	-0.0009 (11)	-0.0054 (12)
C13	0.0279 (14)	0.0339 (15)	0.0291 (14)	0.0019 (12)	0.0020 (11)	-0.0071 (11)

Geometric parameters (Å, °)

S1—O2	1.429 (2)	C4—H4	0.9300
S1—O1	1.430 (2)	C5—H5	0.9300
S1—N2	1.645 (2)	C6—C7	1.392 (3)
S1—C6	1.774 (3)	C6—C11	1.395 (4)
O3—N3	1.227 (3)	C7—C8	1.381 (4)
O4—N3	1.220 (3)	C7—H7	0.9300
N1—C4	1.331 (4)	C8—C9	1.384 (4)
N1—C3	1.338 (4)	C8—H8	0.9300
N1—H1	0.87 (3)	C9—C10	1.381 (4)
N2—C1	1.390 (3)	C10—C11	1.383 (4)
N2—H2A	0.79 (3)	C10—H10	0.9300
N3—C9	1.481 (3)	C11—H11	0.9300
C1—C5	1.396 (4)	O5—C12	1.220 (3)
C1—C2	1.402 (3)	O6—C12	1.258 (3)
C2—C3	1.374 (4)	C11—C13	1.783 (3)
C2—H2	0.9300	C12—C13	1.778 (3)
C3—H3	0.9300	C13—C13	1.768 (3)
C4—C5	1.366 (4)	C12—C13	1.574 (4)
O2—S1—O1	120.47 (13)	C1—C5—H5	120.4
O2—S1—N2	109.85 (12)	C7—C6—C11	121.0 (2)
O1—S1—N2	104.53 (12)	C7—C6—S1	120.5 (2)
O2—S1—C6	108.34 (12)	C11—C6—S1	118.5 (2)
O1—S1—C6	107.38 (12)	C8—C7—C6	119.6 (3)
N2—S1—C6	105.25 (12)	C8—C7—H7	120.2
C4—N1—C3	120.6 (2)	C6—C7—H7	120.2
C4—N1—H1	119 (2)	C7—C8—C9	118.4 (2)
C3—N1—H1	120 (2)	C7—C8—H8	120.8
C1—N2—S1	127.63 (19)	C9—C8—H8	120.8
C1—N2—H2A	120 (2)	C10—C9—C8	122.9 (3)

S1—N2—H2A	112 (2)	C10—C9—N3	118.2 (2)
O4—N3—O3	123.3 (3)	C8—C9—N3	118.9 (2)
O4—N3—C9	118.6 (2)	C9—C10—C11	118.5 (3)
O3—N3—C9	118.0 (2)	C9—C10—H10	120.7
N2—C1—C5	124.3 (2)	C11—C10—H10	120.7
N2—C1—C2	117.8 (2)	C10—C11—C6	119.5 (2)
C5—C1—C2	118.0 (2)	C10—C11—H11	120.3
C3—C2—C1	119.5 (3)	C6—C11—H11	120.3
C3—C2—H2	120.2	O5—C12—O6	129.2 (2)
C1—C2—H2	120.2	O5—C12—C13	116.5 (2)
N1—C3—C2	120.8 (3)	O6—C12—C13	114.2 (2)
N1—C3—H3	119.6	C12—C13—C13	113.58 (18)
C2—C3—H3	119.6	C12—C13—C12	107.13 (17)
N1—C4—C5	121.8 (3)	C13—C13—C12	108.93 (14)
N1—C4—H4	119.1	C12—C13—C11	109.61 (18)
C5—C4—H4	119.1	C13—C13—C11	107.11 (14)
C4—C5—C1	119.2 (3)	C12—C13—C11	110.50 (14)
C4—C5—H5	120.4		
O2—S1—N2—C1	40.4 (3)	S1—C6—C7—C8	178.0 (2)
O1—S1—N2—C1	171.0 (2)	C6—C7—C8—C9	0.3 (4)
C6—S1—N2—C1	-76.0 (3)	C7—C8—C9—C10	0.1 (4)
S1—N2—C1—C5	-15.4 (4)	C7—C8—C9—N3	-179.4 (2)
S1—N2—C1—C2	165.2 (2)	O4—N3—C9—C10	176.8 (2)
N2—C1—C2—C3	178.0 (3)	O3—N3—C9—C10	-3.5 (3)
C5—C1—C2—C3	-1.5 (4)	O4—N3—C9—C8	-3.6 (4)
C4—N1—C3—C2	0.8 (5)	O3—N3—C9—C8	176.0 (2)
C1—C2—C3—N1	-0.2 (4)	C8—C9—C10—C11	-0.2 (4)
C3—N1—C4—C5	0.4 (5)	N3—C9—C10—C11	179.3 (2)
N1—C4—C5—C1	-2.1 (5)	C9—C10—C11—C6	-0.1 (4)
N2—C1—C5—C4	-176.8 (3)	C7—C6—C11—C10	0.5 (4)
C2—C1—C5—C4	2.6 (4)	S1—C6—C11—C10	-178.2 (2)
O2—S1—C6—C7	-7.4 (2)	O5—C12—C13—C13	11.0 (3)
O1—S1—C6—C7	-138.9 (2)	O6—C12—C13—C13	-170.43 (19)
N2—S1—C6—C7	110.1 (2)	O5—C12—C13—C12	-109.3 (2)
O2—S1—C6—C11	171.35 (19)	O6—C12—C13—C12	69.2 (3)
O1—S1—C6—C11	39.8 (2)	O5—C12—C13—C11	130.8 (2)
N2—S1—C6—C11	-71.2 (2)	O6—C12—C13—C11	-50.7 (3)
C11—C6—C7—C8	-0.6 (4)		

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N1—H1 \cdots O6	0.87 (3)	1.93 (3)	2.803 (3)	172 (3)
N2—H2A \cdots O6 ⁱ	0.79 (3)	2.00 (3)	2.785 (3)	171 (3)
C2—H2 \cdots O5 ⁱ	0.93	2.53	3.240 (3)	133
C3—H3 \cdots O1 ⁱⁱ	0.93	2.50	3.391 (4)	162
C8—H8 \cdots O5 ⁱⁱⁱ	0.93	2.47	3.149 (3)	130

C10—H10 \cdots O3 ^{iv}	0.93	2.50	3.308 (4)	146
C5—H5 \cdots O2	0.93	2.33	2.978 (4)	126
C7—H7 \cdots O2	0.93	2.57	2.938 (3)	104

Symmetry codes: (i) $x+1/2, -y+1/2, z+1/2$; (ii) $x+1/2, -y+1/2, z-1/2$; (iii) $-x, -y+1, -z+1$; (iv) $-x+2, -y+1, -z+2$.