

2-Amino-4-methylpyridinium 4-methylbenzoate

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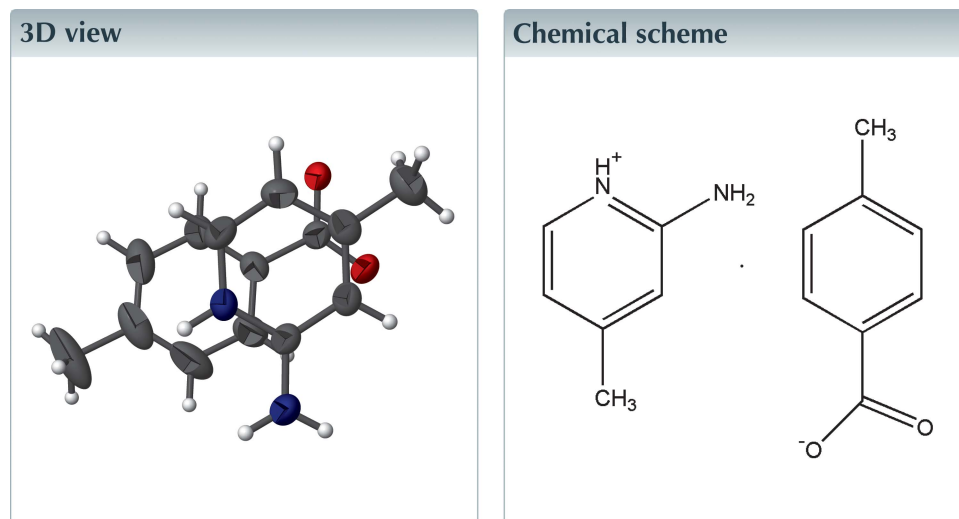
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Keywords: molecular salt; crystal structure; hydrogen bonding..

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Structural data: full structural data are available from iucrdata.iucr.org

In the title molecular salt, $C_6H_9N_2^+ \cdot C_8H_7O_2^-$, the cation is protonated at its pyridine N atom and the dihedral angle between the carboxylate group and its attached benzene ring in the anion is $8.54(17)^\circ$. In the crystal, $N-H \cdots O$ hydrogen bonds link the components into [001] chains. Weak $C-H \cdots O$ and aromatic $\pi-\pi$ stacking [centroid-centroid separation = $3.8503(18) \text{ \AA}$] link the chains into a three-dimensional network.



Structure description

As part of our ongoing studies of hydrogen-bonding patterns in pyridyl molecular salts (Sivakumar *et al.*, 2016*a,b*), we report herein the synthesis and crystal structure of the title molecular salt (Fig. 1).

Weak $\pi-\pi$ [$Cg1 \cdots Cg2 = 3.850(1) \text{ \AA}$; $Cg1$ and $Cg2$ are the centroids of the rings $C7-C12$ and $N1/C1-C5$, respectively] interactions are observed between the benzene and pyridine rings. In the crystal, adjacent anions and cations are connected by $N-H \cdots O$ ($N1-H1A \cdots O1^i$ and $N2-H2B \cdots O2^i$) hydrogen bonds, thereby generating an $R_2^2(8)$ ring-motif; these units are further linked by an $N2-H2A \cdots O2^{ii}$ hydrogen bond and a $C1-H1 \cdots O1^{iii}$ contact (Table 1) into chains propagating along [001] (Fig. 2).

Synthesis and crystallization

The title salt was synthesized by mixing 2-amino-4-methylpyridine (1.081 g) and *p*-toluic acid (1.361 g) in an equimolar ratio in an acetone and methanol (1:1) solvent mixture. After a period of 30 days, colourless blocks of the title compound were obtained.

Table 1
Hydrogen-bond geometry (Å, °).

<i>D</i> –H··· <i>A</i>	<i>D</i> –H	H··· <i>A</i>	<i>D</i> ··· <i>A</i>	<i>D</i> –H··· <i>A</i>
N1–H1A···O1 ⁱ	0.86 (2)	1.80 (2)	2.661 (3)	175 (3)
N2–H2B···O2 ⁱ	0.86 (1)	1.93 (1)	2.792 (4)	176 (3)
N2–H2A···O2 ⁱⁱ	0.85 (2)	1.99 (2)	2.835 (3)	169 (3)
C1–H1···O1 ⁱⁱⁱ	0.93	2.58	3.223 (3)	126

Symmetry codes: (i) *x*, *y*, *z* + 1; (ii) $-x$, $-y + 1$, $z + \frac{1}{2}$; (iii) $-x$, $-y$, $z + \frac{1}{2}$.

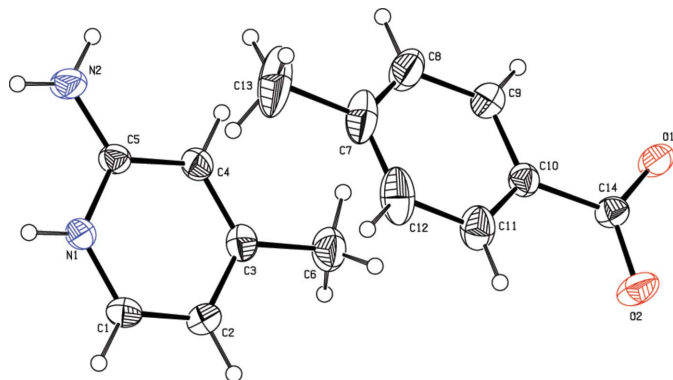


Figure 1
The molecular structure of the title molecular salt, with 30% probability displacement ellipsoids.

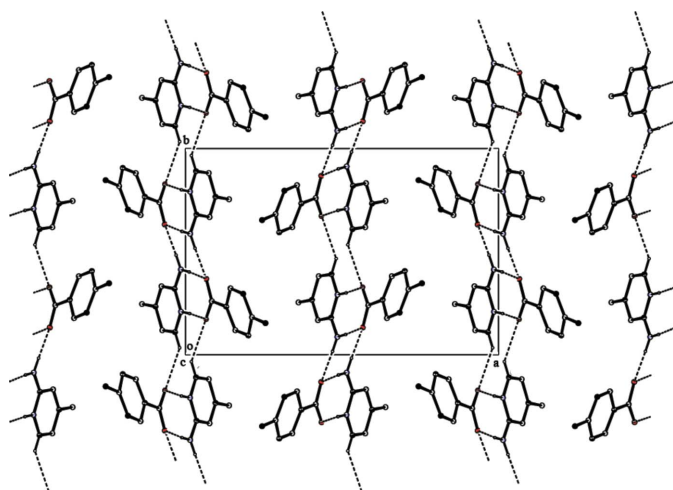


Figure 2
The crystal packing of the title molecular salt viewed along *c* axis. Hydrogen bonds are shown as dashed lines. H atoms not involving in hydrogen bonding have been omitted for clarity.

Table 2
Experimental details.

Crystal data	
Chemical formula	$C_6H_9N_2^+ \cdot C_8H_7O_2^-$
<i>M_r</i>	244.29
Crystal system, space group	Orthorhombic, <i>Pna</i> ₂₁
Temperature (K)	295
<i>a</i> , <i>b</i> , <i>c</i> (Å)	16.7541 (18), 11.0261 (11), 7.2064 (7)
<i>V</i> (Å ³)	1331.3 (2)
<i>Z</i>	4
Radiation type	Mo <i>K</i> α
μ (mm ⁻¹)	0.08
Crystal size (mm)	0.28 × 0.26 × 0.22
Data collection	
Diffractometer	Bruker Kappa APEXII CCD
Absorption correction	Multi-scan (<i>SADABS</i> ; Bruker, 2004)
<i>T_{min}</i> , <i>T_{max}</i>	0.661, 0.745
No. of measured, independent and observed [<i>I</i> > 2σ(<i>I</i>)] reflections	17657, 2372, 1903
<i>R_{int}</i>	0.037
(sin θ/λ) _{max} (Å ⁻¹)	0.626
Refinement	
<i>R</i> [<i>F</i> ² > 2σ(<i>F</i> ²)], <i>wR</i> (<i>F</i> ²), <i>S</i>	0.039, 0.102, 1.11
No. of reflections	2372
No. of parameters	177
No. of restraints	4
H-atom treatment	H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$, $\Delta\rho_{min}$ (e Å ⁻³)	0.12, -0.19

Computer programs: *APEX2* and *SAINT* (Bruker, 2004), *SHELXS97* (Sheldrick, 2008), *SHELXL2016* (Sheldrick, 2015) and *PLATON* (Spek, 2009).

Refinement

Crystal data, data collection and structure refinement details are summarized in Table 2.

Acknowledgements

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full crystallographic data

IUCrData (2016). **1**, x161411 [doi:10.1107/S2414314616014115]

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Crystal data

$C_6H_9N_2^+ \cdot C_8H_7O_2^-$

$M_r = 244.29$

Orthorhombic, $Pna2_1$

$a = 16.7541$ (18) Å

$b = 11.0261$ (11) Å

$c = 7.2064$ (7) Å

$V = 1331.3$ (2) Å³

$Z = 4$

$F(000) = 520$

$D_x = 1.219$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 6146 reflections

$\theta = 2.4$ – 23.5°

$\mu = 0.08$ mm⁻¹

$T = 295$ K

Block, colourless

$0.28 \times 0.26 \times 0.22$ mm

Data collection

Bruker Kappa APEXII CCD
diffractometer

ω and φ scan

Absorption correction: multi-scan
(SADABS; Bruker, 2004)

$T_{\min} = 0.661$, $T_{\max} = 0.745$

17657 measured reflections

2372 independent reflections

1903 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.037$

$\theta_{\max} = 26.4^\circ$, $\theta_{\min} = 2.2^\circ$

$h = -20 \rightarrow 20$

$k = -13 \rightarrow 13$

$l = -9 \rightarrow 6$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.039$

$wR(F^2) = 0.102$

$S = 1.11$

2372 reflections

177 parameters

4 restraints

Hydrogen site location: mixed

H atoms treated by a mixture of independent
and constrained refinement

$w = 1/[\sigma^2(F_o^2) + (0.0394P)^2 + 0.3164P]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.12$ e Å⁻³

$\Delta\rho_{\min} = -0.19$ e Å⁻³

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	-0.03357 (17)	0.1098 (2)	0.5620 (5)	0.0458 (7)

H1	-0.018668	0.033638	0.605372	0.055*
C2	-0.07416 (19)	0.1189 (3)	0.4012 (5)	0.0481 (8)
H2	-0.087627	0.049539	0.334745	0.058*
C3	-0.09593 (17)	0.2338 (3)	0.3350 (4)	0.0442 (7)
C4	-0.07669 (17)	0.3322 (2)	0.4382 (4)	0.0414 (7)
H4	-0.091582	0.408950	0.397487	0.050*
C5	-0.03497 (16)	0.3202 (2)	0.6043 (4)	0.0376 (6)
C6	-0.1390 (2)	0.2460 (3)	0.1553 (5)	0.0639 (9)
H6A	-0.146260	0.330402	0.127061	0.096*
H6B	-0.190096	0.207161	0.164216	0.096*
H6C	-0.108306	0.208399	0.058567	0.096*
C7	0.2124 (2)	0.1820 (5)	0.5775 (5)	0.0713 (11)
C8	0.2006 (2)	0.0940 (4)	0.4461 (6)	0.0705 (11)
H8	0.220675	0.016444	0.465941	0.085*
C9	0.1591 (2)	0.1190 (3)	0.2836 (5)	0.0574 (9)
H9	0.151011	0.057853	0.196778	0.069*
C10	0.12997 (17)	0.2337 (3)	0.2502 (4)	0.0413 (7)
C11	0.1437 (2)	0.3221 (3)	0.3798 (5)	0.0595 (9)
H11	0.125739	0.400643	0.358355	0.071*
C12	0.1841 (2)	0.2957 (4)	0.5422 (6)	0.0786 (12)
H12	0.192168	0.356814	0.629174	0.094*
C13	0.2554 (3)	0.1533 (5)	0.7564 (6)	0.1122 (19)
H13A	0.217202	0.144162	0.854656	0.168*
H13B	0.291374	0.218301	0.785960	0.168*
H13C	0.285005	0.079365	0.742203	0.168*
C14	0.08400 (17)	0.2600 (2)	0.0769 (4)	0.0403 (7)
N1	-0.01429 (14)	0.2087 (2)	0.6605 (3)	0.0397 (6)
N2	-0.01482 (18)	0.4132 (2)	0.7108 (4)	0.0547 (7)
O1	0.06504 (14)	0.17304 (17)	-0.0243 (3)	0.0545 (6)
O2	0.06616 (14)	0.36798 (18)	0.0423 (3)	0.0597 (7)
H2B	0.0117 (17)	0.402 (3)	0.812 (3)	0.056 (10)*
H2A	-0.0303 (17)	0.4832 (16)	0.675 (4)	0.054 (10)*
H1A	0.0133 (15)	0.200 (2)	0.760 (3)	0.039 (8)*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0521 (16)	0.0305 (14)	0.0548 (19)	0.0018 (12)	-0.0027 (16)	0.0029 (15)
C2	0.0507 (17)	0.0412 (16)	0.052 (2)	-0.0013 (13)	-0.0055 (15)	-0.0108 (15)
C3	0.0377 (15)	0.0546 (18)	0.0403 (17)	0.0010 (13)	-0.0007 (13)	0.0005 (15)
C4	0.0438 (15)	0.0376 (14)	0.0426 (19)	0.0020 (12)	-0.0046 (13)	0.0059 (14)
C5	0.0422 (14)	0.0312 (13)	0.0394 (17)	0.0007 (11)	-0.0005 (13)	0.0029 (13)
C6	0.059 (2)	0.081 (2)	0.052 (2)	0.0050 (19)	-0.0138 (17)	-0.0044 (19)
C7	0.0448 (18)	0.125 (3)	0.045 (2)	0.004 (2)	-0.0062 (17)	0.011 (2)
C8	0.058 (2)	0.083 (3)	0.070 (3)	0.0028 (19)	-0.0138 (19)	0.029 (2)
C9	0.061 (2)	0.0524 (19)	0.059 (2)	-0.0040 (15)	-0.0159 (17)	0.0100 (16)
C10	0.0395 (15)	0.0471 (16)	0.0373 (15)	-0.0038 (12)	-0.0023 (13)	0.0028 (14)
C11	0.058 (2)	0.067 (2)	0.053 (2)	0.0074 (17)	-0.0110 (16)	-0.0150 (18)

C12	0.063 (2)	0.118 (3)	0.055 (3)	0.006 (2)	-0.0154 (19)	-0.029 (2)
C13	0.068 (3)	0.217 (6)	0.052 (3)	0.015 (3)	-0.019 (2)	0.019 (3)
C14	0.0480 (16)	0.0367 (14)	0.0362 (16)	0.0021 (12)	0.0008 (14)	0.0005 (14)
N1	0.0468 (14)	0.0346 (13)	0.0378 (14)	0.0034 (10)	-0.0087 (11)	0.0045 (11)
N2	0.081 (2)	0.0319 (14)	0.0516 (18)	0.0052 (13)	-0.0215 (15)	0.0002 (13)
O1	0.0787 (15)	0.0360 (10)	0.0489 (14)	0.0060 (11)	-0.0226 (11)	0.0005 (11)
O2	0.0981 (17)	0.0337 (10)	0.0474 (15)	0.0131 (11)	-0.0163 (13)	-0.0015 (10)

Geometric parameters (Å, °)

C1—N1	1.340 (4)	C8—C9	1.389 (5)
C1—C2	1.347 (4)	C8—H8	0.9300
C1—H1	0.9300	C9—C10	1.377 (4)
C2—C3	1.403 (4)	C9—H9	0.9300
C2—H2	0.9300	C10—C11	1.369 (4)
C3—C4	1.354 (4)	C10—C14	1.495 (4)
C3—C6	1.489 (5)	C11—C12	1.383 (5)
C4—C5	1.392 (4)	C11—H11	0.9300
C4—H4	0.9300	C12—H12	0.9300
C5—N2	1.325 (4)	C13—H13A	0.9600
C5—N1	1.340 (3)	C13—H13B	0.9600
C6—H6A	0.9600	C13—H13C	0.9600
C6—H6B	0.9600	C14—O1	1.246 (3)
C6—H6C	0.9600	C14—O2	1.253 (3)
C7—C12	1.365 (6)	N1—H1A	0.858 (13)
C7—C8	1.370 (6)	N2—H2B	0.860 (13)
C7—C13	1.510 (5)	N2—H2A	0.855 (13)
N1—C1—C2	121.1 (3)	C10—C9—C8	120.5 (3)
N1—C1—H1	119.5	C10—C9—H9	119.8
C2—C1—H1	119.5	C8—C9—H9	119.8
C1—C2—C3	119.4 (3)	C11—C10—C9	118.3 (3)
C1—C2—H2	120.3	C11—C10—C14	121.2 (3)
C3—C2—H2	120.3	C9—C10—C14	120.4 (3)
C4—C3—C2	118.4 (3)	C10—C11—C12	120.7 (3)
C4—C3—C6	121.4 (3)	C10—C11—H11	119.7
C2—C3—C6	120.2 (3)	C12—C11—H11	119.7
C3—C4—C5	121.0 (3)	C7—C12—C11	121.4 (4)
C3—C4—H4	119.5	C7—C12—H12	119.3
C5—C4—H4	119.5	C11—C12—H12	119.3
N2—C5—N1	118.0 (3)	C7—C13—H13A	109.5
N2—C5—C4	123.5 (2)	C7—C13—H13B	109.5
N1—C5—C4	118.5 (2)	H13A—C13—H13B	109.5
C3—C6—H6A	109.5	C7—C13—H13C	109.5
C3—C6—H6B	109.5	H13A—C13—H13C	109.5
H6A—C6—H6B	109.5	H13B—C13—H13C	109.5
C3—C6—H6C	109.5	O1—C14—O2	123.6 (3)
H6A—C6—H6C	109.5	O1—C14—C10	118.1 (2)

H6B—C6—H6C	109.5	O2—C14—C10	118.3 (2)
C12—C7—C8	118.2 (3)	C5—N1—C1	121.6 (2)
C12—C7—C13	121.1 (4)	C5—N1—H1A	119.7 (19)
C8—C7—C13	120.7 (4)	C1—N1—H1A	118.7 (19)
C7—C8—C9	120.9 (4)	C5—N2—H2B	120 (2)
C7—C8—H8	119.5	C5—N2—H2A	116 (2)
C9—C8—H8	119.5	H2B—N2—H2A	123 (3)
N1—C1—C2—C3	0.6 (5)	C9—C10—C11—C12	-1.7 (5)
C1—C2—C3—C4	-1.4 (4)	C14—C10—C11—C12	177.9 (3)
C1—C2—C3—C6	178.4 (3)	C8—C7—C12—C11	0.8 (6)
C2—C3—C4—C5	1.2 (4)	C13—C7—C12—C11	-179.1 (4)
C6—C3—C4—C5	-178.6 (3)	C10—C11—C12—C7	0.9 (6)
C3—C4—C5—N2	-179.8 (3)	C11—C10—C14—O1	-171.3 (3)
C3—C4—C5—N1	-0.2 (4)	C9—C10—C14—O1	8.2 (4)
C12—C7—C8—C9	-1.7 (6)	C11—C10—C14—O2	8.1 (4)
C13—C7—C8—C9	178.2 (4)	C9—C10—C14—O2	-172.3 (3)
C7—C8—C9—C10	1.0 (5)	N2—C5—N1—C1	178.9 (3)
C8—C9—C10—C11	0.7 (5)	C4—C5—N1—C1	-0.7 (4)
C8—C9—C10—C14	-178.8 (3)	C2—C1—N1—C5	0.5 (4)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1A...O1 ⁱ	0.86 (2)	1.80 (2)	2.661 (3)	175 (3)
N2—H2B...O2 ⁱ	0.86 (1)	1.93 (1)	2.792 (4)	176 (3)
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