

Thermal decomposition of the title compounds has been studied. Final products of thermal decomposition are single-phase hexagonal closed-packed metal solid solutions $\text{Rh}_{0.25}\text{Re}_{0.75}$ ($a = 2.749(2)$, $c = 4.395(3)$) and $\text{Ir}_{0.25}\text{Re}_{0.75}$ ($a = 2.758(2)$, $c = 4.393(3)$). $\text{Ru}_{0.25}\text{Re}_{0.75}$ ($a = 2.756(2)$, $c = 4.414(3)$ Å).

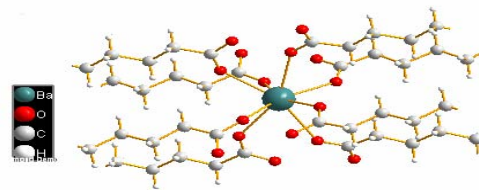
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Synthesis And Structures Of Dicarboxylato-Bridged Earth-Alcaline Polymers ¹Aliouane. K. , ⁽¹⁾Guehria-Laidoudi A., ⁽²⁾Dahaoui S. and ⁽²⁾Lecomte C. ^aLaboratoire de cristallographie appliquée, Institut de Chimie, U.S.T.H.B., B.P. 32, El-Alia 16111, Bab-Ezzouar, Alger, Algérie ⁽²⁾ LCM³B-CNRS UMR 7036, Faculté des Sciences Université Henri Poincaré BP 239, 54506 Vandoeuvre-lès-Nancy, France.

Keywords: polymeric structure, earth-alkaline, dicarboxylate

The binding of earth-alkaline ions to dicarboxylate ligands partially $[\text{CO}_2(\text{CH}_2)_n\text{CO}_2\text{H}]^-$ or completely $[(\text{CH}_2)_n - (\text{CO}_2)_2]^{2-}$ deprotonated often lead to polymeric structures. The resulting model complexes are designed to permit the significance of interactions in such extended structures which are used in several fields, particularly in biology and medicine. An extensive polymeric arrays involving earth-alkaline ion depends on the geometric constraints of the metal ion. It is well known that barium, calcium and strontium exhibit a great flexibility if compared to magnesium, in which a rigid requirement of octahedral geometry is invariably found, preventing aggregation. However, in its hydrated α , ω dicarboxylates, Mg realizes chain structure for small values of n (1 to 3) and layer structure for higher values¹.

The three anhydrous Ba, Ca and Sr sebasate, have been synthesized. Single X-ray diffraction studies, carried out at (293K) and (120K), show that they contain partially deprotonated legands. Their space groups, centrosymmetric for Sr (P4₂/nbc) and Ca (P-1) complexes and non centrosymmetric for Ba (P-4b2) point to the invalidity of n parity criteria in the centrosymmetry. In the two tetragonal symmetries, the values of cell parameters, particularly c parametrer, provide evidence that this later, is not related to the lengthening of carbon backbone. The metals are eight-coordinated, and μ -1, 3 oxo-bridges, lead to a two-dimensional layered structure, exhibiting wide empty channels. The general structural features are very similar to those lanthanum complexes, obtained with a relatively long alkyl chain² ($n \geq 5$), even though, the tetragonal symmetry seems to be, in the series, particular to anhydrous barium complexes whatever n values³.



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The Study of Unprecedented Tubular Channels Formation in Alkyl-Substituted Oxamide Oxime Metal Complexes Yurii Chumakov^{a,b}, Dominique Luneau^a, Fatma Yuksel^{a,c}, Vefa Ahsen^{c,d}, Gulay Gumus^d, Ilke Gurol^d. ^aUniversité Claude Bernard Lyon 1, France, ^bInstitute of Applied Physics of Academy of Science of Moldova, ^cGebze Institute of Technology, ^dTUBITAK-Marmara Research Center, Turkey

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Keywords: oxamide oxime metal complexes, tubular channels, conformational analysis

It was found that nickel(II) and palladium(II) complexes of hexyl- substituted oxamide oxime are self-assembled through hydrogen bonds involving the oximic oxygens and amino nitrogen atoms to afford metal-organic frameworks with tubular channels partly filled by the hexyl chains [1]. But whether the chains cooperate in the formation of these structures, or simply fill the empty space, was unclear. Therefore, the nickel (II) (**1a-3a**) and palladium(II) (**1b-3b**, **5b,6b**) complexes of oxamide oximes substituted with alkyl chains of different length (C3-C8) were synthesized and characterized by X-ray and spectroscopic methods as well as their conformational analysis was carried out in order to understand the role of chains in formation of the tubular channels in the crystals. All studied molecules are centrosymmetric at the metal center forming the perfectly planar central cores containing the metal ion $[\text{M}\{\text{O1-N1-C1(N3)-C2(N4)-N2-O2}\}_2]$ $[\text{M}=\text{Ni(II), Pd(II)}]$. It was found that only palladium(II) (**1b-3b**, **5b,6b**) complexes form the infinite tubular channels filled by the chains. The channels in these ones, as well as in nickel(II) (**4a**) and palladium(II) (**4b**) complexes [1], are walled by the central cores which are self-assembled *via* N-H...O hydrogen bonds involving the oximic oxygens and amino nitrogen atoms. The building blocks which form the crystal packing of **2b-6b** and **4a** have been extracted in order to perform the quantitative analysis of channels topology and they were characterized by the packing indexes. It was shown that the space in the channels which is filled by chains in compounds of **2b**, **3b**, **5b** and **6b** is greater in 2.6-3.6 times then in **4a** and **4b**. On the base of conformational analysis using the module Conformational Search implemented in HyperChem 6.03 it was found that the formation of infinite tubular channels in the crystals of alkyl-substituted oxamide oxime nickel (II) and palladium (II) complexes depends on the orientation of the chains relatively to the