

*Chem. B.*, 2005, **109**, 1304. [3] Urabe H., Sugawara Y., Kasuya T., *Phys. Rev. B*, 1995, **51**, 5666.

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#### A Novel Spin Transition Curve in [tris(2-picolylamine)Fe(II)]Cl<sub>2</sub> Allyl Solvate

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The title compound exhibits a temperature dependent spin crossover with an intermediate plateau in the spin transition curve (STC) around 111K at an unusual high-spin concentration of about 30%. This is well below the plateau with a 50% concentration observed for the isomorphous ethanol solvate [1].

We report here single crystal diffraction data collected at 200, 132, 111 and 94K. The systematic absences are consistent with *B*<sub>2</sub>/*c* space group symmetry at 200 and 132K. At 111K, *i.e.* in the range of the plateau, a superstructure is observed. The consequential reflections violate the original *B*-centering and lead to *P*<sub>2</sub>/*1*/*c* space group symmetry with a doubled unit cell volume. At 94K the original *B*<sub>2</sub>/*c* symmetry is recovered, thus indicating a reentrant sequence of phase transformations [2].

This observation shows that, as in the cases of the ethanol and 2-propanol solvates [1,3], a plateau in the STC reflects the appearance of a new ordered structure. The plateau region is likely a result of the coupling of the spin conversion with the concurrent conformational ordering of the crystal architecture.

[1] Chernyshov D., Hostettler M., Törnroos K. W., Bürgi H.-B., *Angew. Chem. Int. Ed.* 2003, **42**, 3825. [2] Chernyshov D., Bürgi H.-B., Hostettler M., Törnroos K. W., *Phys. Rev. B*, 2004, **70**, 094116. [3] *in preparation*.

**Keywords:** spin crossover, phase transition, order-disorder

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#### Solid State Synthesis and Characterization of Some Novel Sodium Rare Earth Phosphates

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Recently, much attention has been paid to the rare earth phosphates because of their potential applications for optical materials, including laser, phosphors, and more recently, anti-UV materials [1]. MOPO<sub>4</sub> type materials possess properties that make them potentially useful for catalytic, electronic and ion exchange applications[2].

In this work, Ln<sub>2</sub>O<sub>3</sub> (where Ln= La, Nd, Sm, Eu, Gd, Dy, Ho, Er, Yb) were used as a rare earth source, NH<sub>4</sub>H<sub>2</sub>PO<sub>4</sub> was used as a phosphate source and Na<sub>2</sub>CO<sub>3</sub> was used as a sodium source to obtain sodium rare earth oxyphosphates. Reactants were heated at 1100 °C for 20 hours. X-ray diffraction patterns, IR and Raman analysis, SEM pictures and EDX analysis were taken for characterization. According to these results, structures of these products were compared with previous ones[2-3].

[1] Hirai H., Masui T., Imanaka N., Adachi G., *J. Alloys and Compounds*, 2004, **374**, 84-88. [2] Gönen Z.S., Kizilyalli M., Pamuk H.Ö., *J. Alloys and Compounds*, 2000, **303-304**, 416-420. [3] Uztielik- Amour A., Kizilyalli M., *J. Solid State Chem.*, 1995, **120**, 275-278.

**Keywords:** rare-earths, oxyphosphates, solid-state

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#### The Systems Li-Ho-P-O and K-Ho-P-O: A Study in Inert Atmosphere

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The phosphates with open frameworks are materials that are composed by (PO<sub>4</sub>) tetrahedral and by others polyhedral as octahedral (XO<sub>6</sub>) and dodecahedral (XO<sub>8</sub>) units. So, this structural conformation prove several applications of such materials as ionic conductors, ion exchangers, scintillating materials for gamma ray detection [1], catalysts [2], anticorrosive [3], etc. As part of our phosphates with open frameworks research, we study the phases present in the systems Li-Ho-P-O and K-Ho-P-O using different temperatures reaction and atmospheres. In this work, different phosphates compound like rare earth phosphate (Xenotime-type HoPO<sub>4</sub>), alkali metaphosphate (APO<sub>3</sub>, A= Li, K) and an alkali-rare earth pyrophosphate (LiHoP<sub>2</sub>O<sub>7</sub>) [4] were synthesized. These phases were characterized by X ray powder diffraction, differential thermal analysis and microscopy methods. A crystallochemistry study relating the crystal structures was performed.

[1] Natarajan S., Eswaramoorthy M., Cheetham A.K., Rao C.N.R., *Chem. Commun.*, 1998, 1561, and references there in. [2] Clearfield A., Thakurb D.S., *Applied Catalysis.*, 1986, 1. [3] Deyá M.C., Blustein G., Romagnoli R., Del Amo B., *Surf.Coat.Tech.*, 2002, 133. [4] Hamady A., Fauzi Zid M., Jouini T., *J.Solid State Chem.*, 1994, 120.

**Keywords:** phosphates, phase identification, X-ray diffraction

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#### Studies of Spin Crossover Complexes via Solvo-thermal Syntheses and their Thermal Relaxation of Light Excited Kinetic Phenomena

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The unique spin crossover 1D ladder complex [Fe<sub>2</sub>(bpt)<sub>2</sub>(NCS)<sub>2</sub>(bpy)<sub>2</sub>] · MeOH (**1**) have been synthesized successfully by one-step solvo-thermal syntheses or directly transformed from *trans*-[Fe(abpt)<sub>2</sub>(NCS)<sub>2</sub>] complex by hydrothermal process.

The magnetic measurement reveals that the 1D spin-crossover complex **1** has an abrupt spin transition at 130 K and possesses an unusual magnetic behavior. The π-π interactions of **1** do vary slightly with spin transition. The distances along the (110) and (101) planes are 3.65(1) Å and 3.52(5) Å at 295 K but 3.53(3) Å and 3.45(2) Å at 100 K.

The novel dinuclear Double bridging complex [Fe<sub>2</sub>(μ-bpt)<sub>2</sub>(NCS)<sub>2</sub>(CH<sub>3</sub>OH)<sub>2</sub>](**2**) and [Fe<sub>2</sub>(μ-bpt)<sub>2</sub>(NCS)<sub>2</sub>(py)<sub>2</sub>](**3**) have also been synthesized and characterized by x-ray diffraction. The bond lengths of Fe-N are 2.292(3) Å, 2.094(3) Å for complex **2** and 2.110(2) Å, 2.321(2) Å for complex **3** at HS state.

The variable-temperature magnetic susceptibility measurement reveals that **2** and **3** are in high-spin state during 60-400 K. Interestingly, the complex **3** has a spin transition from HS to LS at low temperature under 808 nm radiation and the thermal relaxation behavior of such induced LS state will be discussed.

**Keywords:** LIESST, spin crossover, thermal relaxation

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#### Study of the Mechanism of Cu-Zn Mechanical Alloying by X-ray Powder Diffraction

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